Example Problem

Try to calculate the thermodynamic quantities using the following equation. The coefficients are the red digits. The reactants and products are shown.

$$1\ C_2H_4 + 3\ O_2 \rightarrow 2\ CO_2 + 2\ H_2O$$

Reactants: 1 C2H4, 3 O2
Products: 2 CO2 , 2 H2O

On the Heroku App:
You can enter the formulas exactly as they are written:
C2H4
O2
CO2
H2O

You will be prompted for the coefficients. This is what the output should look like.

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The reactants are: ['1 of C2H4 ethene g', '3 of O2 oxygen g']
The products are: ['2 of CO2 carbon dioxide g', '2 of H2O water l']
The enthalpy change (dH) is: -1408.0 kJ per mol
The free energy change (dG) is: -1330.0 kJ per mol
The entropy change (dS) is: -0.2668000000000001 kJ per mol per Kelvin
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This is the output on the Textual version.