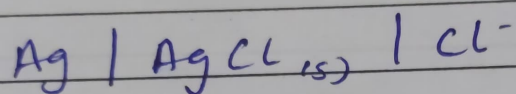
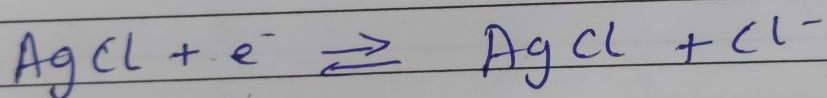


Silver - Silver chloride electrode:

It consists of a Silver electrode coated with a sparingly soluble AgCl and is immersed in a solution containing Cl^- ions (Saturated KCl). The Silver - Silver chloride electrode can be represented as



Electrode reaction is



The electrode potential depends on chloride ions. For 1 Molar solution the electrode potential is 0.223 V .