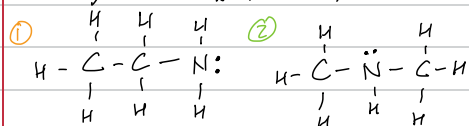


# Ch1 HW - 1-34 c, 1-39, 1-41i, d, c, k, l, 1-42 b, h, i, 1-44, 1-46, h, j, 1-52 a, f, 1-54d, 1-55d, e, 1-58 a, b, f, 1-59 c, f

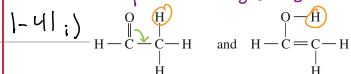
1-34 c) two Lewis for  $C_2H_7N$ 

$$C: R.C = 4 - \frac{1}{2}(6) = 0 \quad N: R.C = 5 - \frac{1}{2}(6) = 2 = 0$$

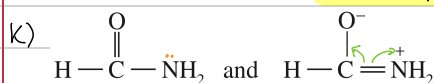
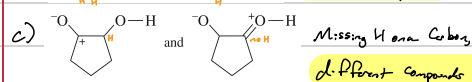
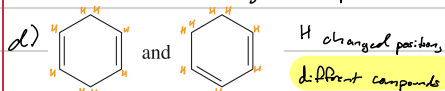
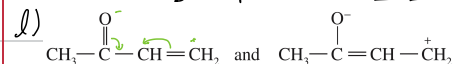
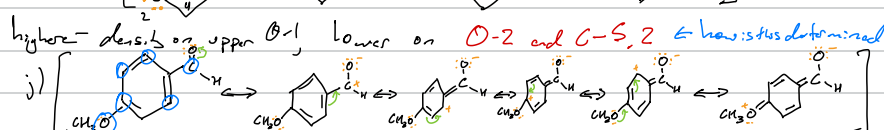
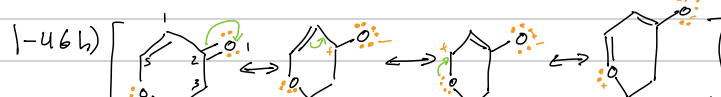
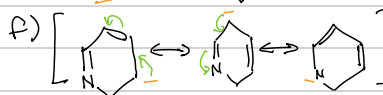
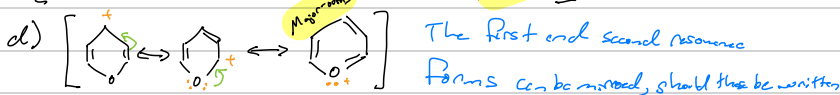
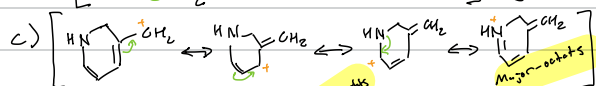
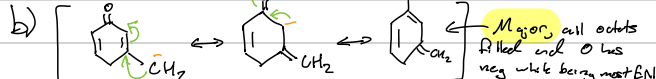
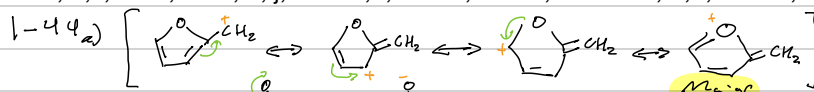
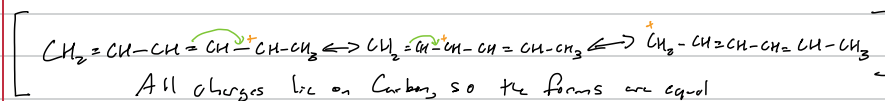
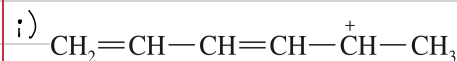
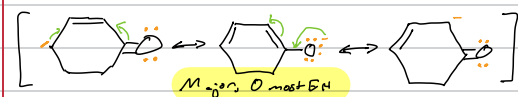
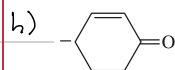
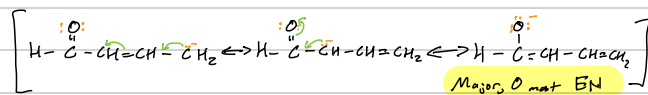
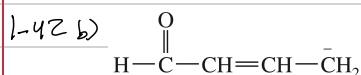
Is there a reliable way to predict BH differences?

- 1-39 a)  $C-Cl$  large (7.5) b)  $C-H$  small  
 c)  $C-Li$  large d)  $C-N$  small  
 e)  $C-O$  large f)  $C-B$  large  
 g)  $C-Mg$  large h)  $N-H$  large  
 i)  $O-H$  large j)  $C-Br$  small

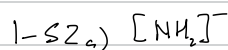
if atom position changes, they are different compounds!



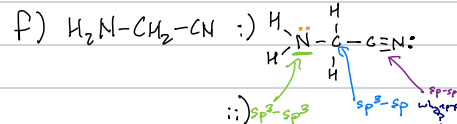
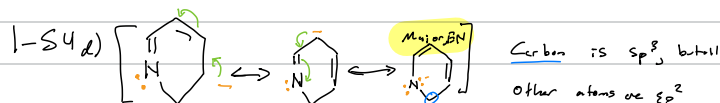
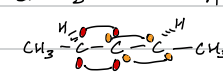
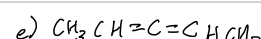
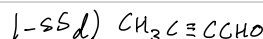
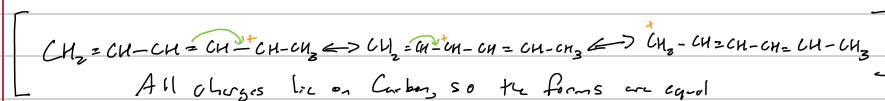
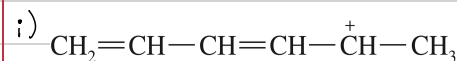
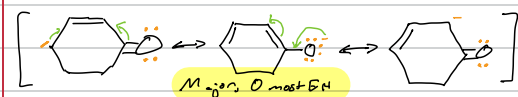
D: P Forst Compounds, H moved position

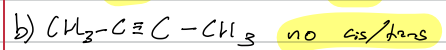
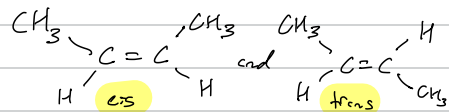
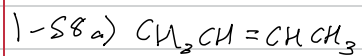
Resonance,  $e^-$  position changed onlyResonance,  $e^-$  position changed only

blue circles indicate areas of lower electron density

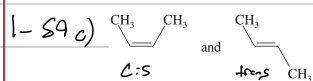
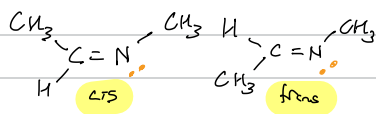


ii) how?

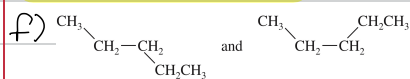
iii)  $< 109^\circ$ N end  $C: 109^\circ$  C end  $N: 180^\circ$ Carbon is  $sp^2$ , but H other atoms are  $sp^3$ Pi bond  $sp^2$  bonded orbitals are color coded



cannot occur around a triple bond!



These are cis/trans isomers



Same exact compound, different