

# Term Paper Introduction

## General Notes

- (i) Decide Vertical vs. Horizontal Approach
- (ii) Published, Peer reviewed articles are expectation
- (iii) No date limit on citations but you must read it & a review article, or book chapters books not strict published but they should not dominate
- (i) Do not cite web addresses as references
- (ii) Try to avoid writing a review article based on summarizing a review article, or book chapters
- (iii) Any more than 50 consecutive word match violates copyright law

## Possible Term Paper Subjects

- (i) **Vitrimers**: A chemical group that can be broken apart under certain conditions, while having thermoset characteristics at room temp. Can be considered a dynamic bond
- (ii) **Green Flame Retardants**: Traditionally Highly Toxic or produce toxic chemicals when heated. Cancer Causing usually. This focuses on developing one that isn't as toxic.
- (iii) **Bio-crystals**: Making mm x nm crystals with no defects. This small size warranted this time.  
**Example:** Carbon Nanotubes  
Needle like crystals which can be prepared from cellulose, hemicellulose, or others
- (iv) **MXene**: 'Maxenes'. Layers can be separated. Such as graphite being separated into individual layers of graphene. **Incredible Aspect Ratio**
- (v) **Glass Transition Temperature**: Super easy to measure. Popular topic of polymer studies but not really known. Nobel Prize worthy.  
How can we predict glass transition temperature without any assumptions?
- (vi) **Collection of Rare Metals**: Natural Oil Reserves running out. Flaking helps this, but horrible to environment.  
↳ Contaminates with excess natural disasters. Metals much harder to replace. Using by-products and will exhaust before oil. Found in low concentrations in seawater.
- (vii) **Porous Materials**: Used to assist in topic of (vi). Analogous used on space shuttles. Very brittle/fragile but remarkably flame resistant. Carbonaceous porous  
Can create non-metal catalysts with porous nts!
- (viii) **Extreme Temperature Polymers**  
↳ Being Scramjet can fly in very thin air at fast speed. Falcon HTV-2 also incredibly hot, compared to current record of cmauth 3. Used for weapons, but can turn into passenger travel, reduces trips to other continents to months of hours.  
Polymers need to be developed to survive 750-1000°C. Breaks temps but is less sensitive to a slow, safe, stable travel.
- (ix) **Non-metallocene Catalyst**: Carbon's surface has defects  $\rightarrow$  acts as catalytic centers. Can add other atoms to defects  $\rightarrow$  more active.
- (x) **H-bonding (Supramolecular synthesis)**: Technique to let monomers contact in non-covalent manner and lets you selectively design polymerization (usually random). Catalyst allows polymers to obtain in a desired way  $\rightarrow$  crystalline properties. Can also through H-bonding control this you can apply desired polymerization

**Pages:** 1 page to 10,000 pages. 1 page is harder to write. Subject chosen, date due, length.

**Review Draft:** Acts as a check point along with other milestones outlined in syllabus

Ask Questions!

Ishida knows his shit, don't plagiarize or copy references word-for-word!

# **Chapter 1**

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**Goal: Refresh information from ENGR145**

# Concepts and Nomenclature

Polymer = poly + mer 3 Repetition Units Oligomer = oligo + mer  
 ↑ multi + unit requires 2-4 units to be counted as poly mer  
 Sub Category a few + unit

Plus ics ← πλαστικού (Greek for plastic)  
 means to 'mold into'

Macromolecules: Contain both polymeric and non-polymeric large molecules

How many units to be polymers?

Rubber (AKA elastomers): polymers that deform elastically and to very large strains  
 ↳ become glass at low temperatures → become brittle (glass transition temperature)

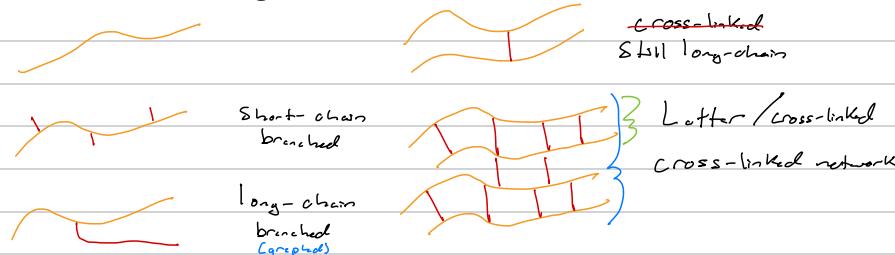
Not repeating

00 "

0△0 "

0△△0 Beginning of structure

## Structure of Polymer Chains



## Copolymers

- **AAAAAA** Mono polymer
- **AAAAAA BBBB** A,B-d block polymer
- **AAAAAA BBBB** graft polymers
- **A B A B A B A** alternating copolymer
- **A B B B A B A** random copolymer

Given Properties of homopolymers • Can you predict  
 A and B • Copolymer properties?  
 properties of A on one side, B on other like soap

even "mix" of properties: AB can be considered a monomer  
 unbalanced mix of properties: true random mixture  
 ↳ can be determined roughly with % of each

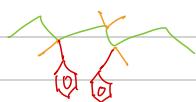
## Conformation vs. Configuration

Conformation: Structural change caused by chain rotation. Can be changed easily (reversibly) by changing temp.

trans: 'zig-zag' chain on a plane gauche: chain rotates  $\pm 60^\circ$  or  $-60^\circ$

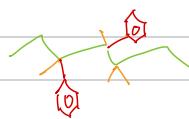
Configuration: Permanent structural difference caused during synthesis. Mild temp change does not affect this

a) Head-to-Head



Less likely to occur  
as this is energetically demanding

b) Head-to-Tail



Much less steric  
hindrance → thermodynamically more stable

## Stereoregularity (stereo isomers)

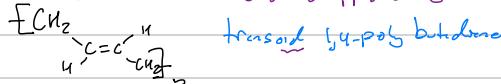
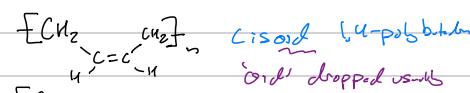
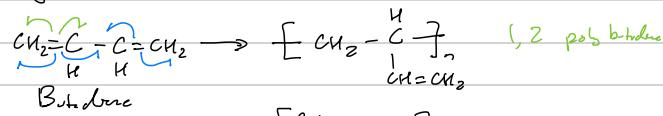
Isotactic

Can crystallize  
↳ Material looks shiny

amorphous A tactic (Random/Not polarized)  
 ↳ Doesn't crystallize → looks very glasslike

Most cost effective

## Regiosomers



## IUPAC Naming

- Add 'poly' to start of monomer name
- Not accurate! Polybenzoxazide is not wrong!

## Molecular Weight

Provide MW in Paper

Tensile Strength  
Cut 1 =  $10^8$  MW  
In half, more  
 weaker  
 $1/200$  strength

$$\text{Number Avg MW: } M_n = \frac{\sum N_x M_x}{\sum N_x} \quad \text{Related to tensile strength}$$

Basic average idea

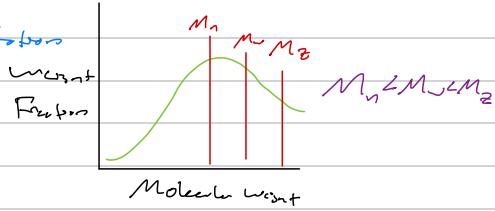
$$\text{Weight Avg MW: } M_w = \frac{\sum W_x M_x}{\sum W_x} \quad \text{How, } W_x = N_x M_x \rightarrow M_w = \frac{\sum N_x M_x^2}{\sum N_x M_x}$$

Average from weight perspective

$$\text{Z-average MW: } M_z = \frac{\sum N_x M_x^3}{\sum N_x M_x^2} \quad \text{Related to precipitators}$$

$$\text{Poly Dispersal Index (PDI): } PDI = \frac{M_w}{M_n}$$

$\hookrightarrow$  Higher PDI  $\rightarrow$  broader MW



# Molecular Weight & Macromolecules

## Various MW Measurement Methods

Method	Type	M.W. type	MW range
Colligative properties		Mn	
Boiling point elevation	Absolute	Mn	< $10^4$
Freezing point depression	Absolute	Mn	< $10^4$
Vapor pressure lowering	Absolute	Mn	< $10^4$
Osmotic pressure	Absolute	Mn	$2 \times 10^4 - 10^6$
Terminal group concentration	Absolute	Mn	< $10^4$
Light scattering	Absolute	Mw	$10^4 - 10^8$
Sedimentation	Absolute	Mw, Mz	$10^2 - 5 \times 10^6$
Intrinsic viscosity	Relative	Mv	$10^3 - 10^8$
Size exclusion chromatography(SEC)	Relative	Mn, Mw, Mz	$10^2 - 10^7$

Notes accurate but most convenient

With be discussed  
with Macromolecules

## Colligative Properties

- Add NaCl to water to lower Freezing Point
- Properties related to numbers

## Staudinger's finger

Hypothesis: stronger propylene molecule consists of aggregates of molecules  $\rightarrow$  size of aggregates should change as electronic structure of molecules is altered

## Approach: Hydrogenation of regular Benzene into a polycyclic aromatic hydrocarbon

Saturated cyclohexenyl ring

Materials Used: Polymers small oligomers of polycyclic aromatic hydrocarbons

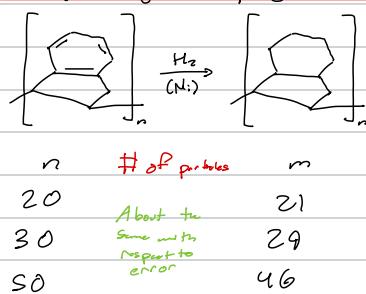
## Meas Method: Colligative property Measurement



Dilute Degree Polymerization as the product was aggregate of molecules

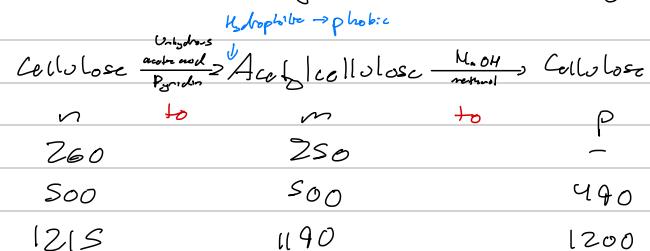
## Staudinger's Experiments

### Initial: Colligative Property Measurement



He was criticized for using small molecules

### Final: Acetylation and regeneration of cellulose; viscosity measurement



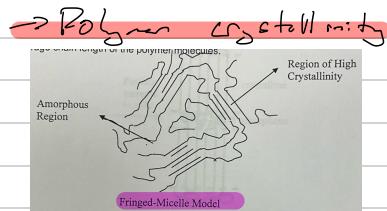
Critics stuck as they rebutted all they were arguing  
Covalent Bonds proved as structure constancy

Directly from his paper, so many value probably measurement error

Viscosity is more sensitive to larger MW molecules than smaller ones (like very cheap go to make 'pamkin' lacquer)

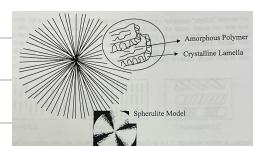
## Morphology

Def: Study of forms of polymer chains, such as crystal structure



## $\rightarrow$ Spherulites

Polarized Light reveals a molten cross

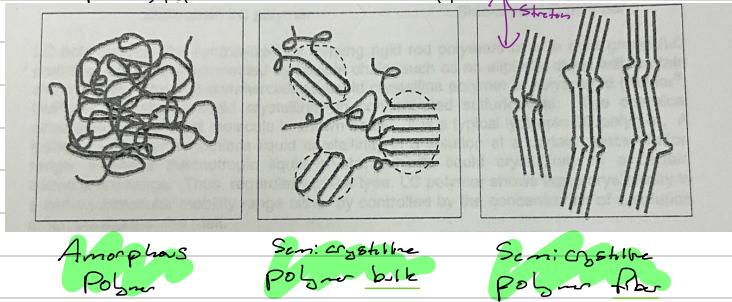


## $\rightarrow$ Lamella Structure

Adjacent rotary model  
 $\rightarrow$  thought to occur due to energy preference

Today Adjacent is a more favorable explanation

## Morphology Controls Stretching

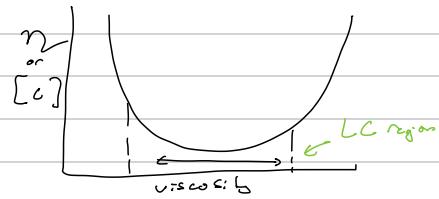


## Stretching

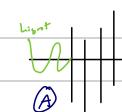
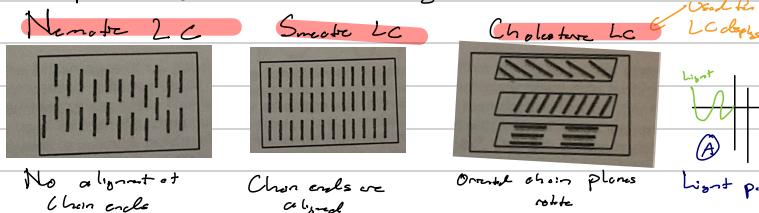
- Molecules realign along stretching direction.
- $\uparrow$  stretching  $\rightarrow$   $\downarrow$  intermolecular distance
  - $\rightarrow$  leads to crystalline region formation
- Cracks are also observed along stretching direction
  - $\rightarrow$  Thus the fibers or structural films are stronger along stretching directions

# Liquid Crystalline (LC) Polymers

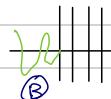
liquid = amorphous  
crystal = perfectly ordered } Two opposing words used to express unusual material  
that is neither completely amorphous nor ordered



## Types of LC Polymers



vs.



Light passes through ① as it needs transverse alignment

Allows microwaveable

supperware

only dissolving conc. sulfuric acid

Sustaining on oriented tie layers so lighter rotate up and out

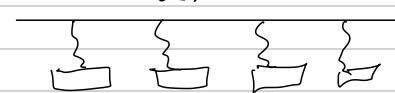
**Thermotropic LC:** formed between certain temperature range  
**Liquid-crystalline LC:** formed between certain concentration range

ex: Aromatic Polyester

ex: Poly(p-phenylene phthalimide), polyaramid Kevlar



Main-chain type LC



Side-chain type LC

□: LC chromophore

## Advantages

- (i) Lower viscosity in LC state than random melt/solvent
- (ii) Molecular alignment even in liquid state → strong fibers

even bullet proof

## **Chapter 2**

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**Goal: Learn what and when to apply certain analytical techniques**

# Qualitative Analysis

You should always start by determining if it's a mixture or pure substance

"If you come into other states; purify, purify, purify, and most importantly purify" - which

(i) Bring to high purity, then crystallize

(ii) If the material cannot be crystallized you must find a better way to purify

## Chromatography

(i) HPLC (High Perf. Liquid Chrom) ↓ (eluent) machines, change type  
↓ based on column and mixture used!

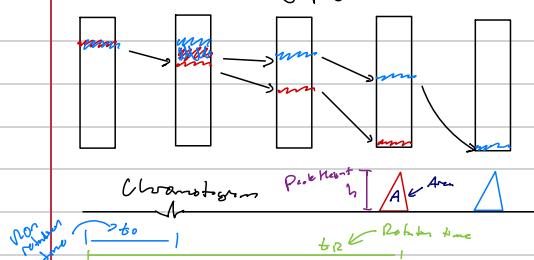
(ii) SEC (Size-Exclusion Chrom)

Chromo = Color

Graphy = Technique

Thin Layer Chromatography (TLC) indicates if compound is a mixture when exposed to UV light

## Column Chromatography



Columns use packing material in a separation column.

HPLC: use silica gel

SEC: use cross linked poly styrene gel

Thus a Mobile & Stationary Phase

## High Resolution Fractionation

Gel Filtration: used in Fract. mode using porous particles where small molecules can get trapped in pores to travel and large ones flow around them.

## Analyses

Collected Fractions examined by spectroscopic techniques to determine concentration of substances eluted. ex spectroscopic techniques: Refractive Index (RI)  
Ultra violet (UV)

(i) Refractive Index: Most Popular ass't is different for molecules with diff. structure  
↳ not as sensitive to absolute mixture results

(ii) Ultraviolet: Order of Magnitude more sensitive

↳ not all molecules can absorb/react to UV

Combining these techniques allows you to check if mixture is homogenous or a copolymer

Mobile Phase = liquid → mobile phase called "eluent"

↳ Stationary Phase: is usually a solid or liquid

↙ (Derived from Liquid Chromatography)

## HPLC

• High Degree of separation

→ Refinement of packing material  
(3 to 10 μm)

• Reduction of Analysis Time

→ Delays of about by pump

Used in tons of industrial studies!

## Advantages

- High Capacity, enabling batch analysis
- Superior quantitative capability and reproducibility
- Moderate conditions (i.e. Sample vaporization unneeded)
- Generally High Sensitivity
- Low sample consumption

Cannot use on solids which don't dissolve or evaporate

## Gradient Systems

- Isoconcentric System: Constant Eluent composition
- Gradient System: Varying eluent composition
  - (i) HPGC - high pressure gradient
  - (ii) LPGB - low pressure gradient

Typical one used

## Reversed Phase Chromatography

- Stationary Phase: Low Polarity  
→ Octadecyl group-bonded silica gel (ODS)
- Mobile Phase: High Polarity  
→ Water, methanol, acetone, etc & salt solutions

Hydrophobic → Hydrophobic surface charge

en: Phenols are hydrophobic

so it avoids hydrophobic surfaces.

Untreated silica gel is usually used (Acidic)

Acid/Base interactions > H-bonding ones

**Detectors:** UV-VIS Detectors, Refractive Index detector, electrochemical detector, mass spectrometer, etc

## Sizing Exclusion Chromatography

- Molecules separated based on size (based on bulkiness)
- Usually applied to large macromolecules
- Does not care about hydrophobic/hydrophilic

Think of it like trying to explore a tight cage

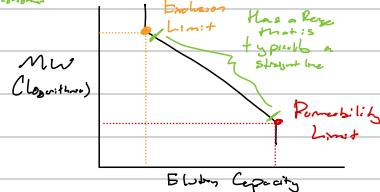
- Huge molecules come out first as small ones need to go through a lot of pores
- Uses cross-linked polystyrene → Created with vinyl aromatic monomers  
↳ This means most polymers will interact with the substrate

## Separation Principle

→ Molecules  $>$  Pore size cannot enter!

For molecules that can enter, there is a inverse logarithmic relationship

P.I. (or monodispersed PDA)  
UV (single/wavelength/multiple wavelengths), Fluorescence, Electrokinetic



You can determine  $M_n$ ,  $M_w$ ,  $M_z$  using Chromatographic techniques (with software)

# Infrared Spectroscopy

Photon in → Photon out

↳ Invisible radiation outside of the red

Used because it has exactly the right energy to see molecules moving at a certain frequency (which requires energy)  
 ↳ As long as molecular structure is different, IR will be different too!  
 Gives you group formations usually

Detects Vibrational Motion

## Uses of Vibrational Spec

- Recognition of functional groups
- Ident. of spectra of individual molecules
- Determination of molecular structure

## Advantages

- Sensitive, Rapid
- Spectra easy to see

## Disadvantages

- Chemical process
- Hard to interpret.

## Spectral Analysis

A spectrum completely defined when following spectral parameters are specified

- The frequencies
- # of constituent bands
- Band shapes
- Band widths
- Amplitudes or intensities

Peak Height is usually called intensity, but  
 you should really use integrated intensity

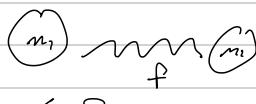
## Model For Vibrational Spectroscopy

↳ Point masses bonded with a Hookean Spring

Allows us to see bonding, stretching, bending, wagging, etc.



$Z_n - 6$  is number of bands,  $n$  is # of atoms/molecules



Hookean spring  
resting force:  $f$

← Displacement →  
(z)

From 2nd order DE  $m_1 r^2 = \frac{m_1 m_2}{m_1 + m_2}$   $\mu$  constant

$V = \frac{1}{2\pi c} \left[ \frac{f}{m_1} \right]^{1/2}$

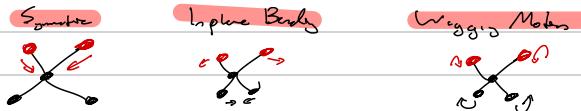
Frequency



- Double Bonds
- Have much higher Frequencies than single bonds.

• Frequency Higher with smaller masses!

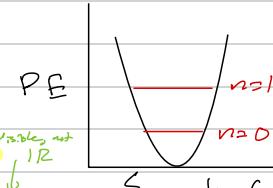
## Analysis of Methylene Models



## Potential Energy Diagrams

$$E_v = (n + \frac{1}{2}) \left( \frac{\hbar}{2\pi c} \right) \left( \frac{f}{m_r} \right)^{1/2}, \text{ for integral values of } n \text{ (vibrational quantum number)}$$

$$\Delta E_{\text{fundamental}} = h\nu_0 = \text{Optical Frequency}$$



Only #1 is visible, can only see n=0 to n=1 and back

## Selective Rules for Vibrational Spectroscopy

### (i) Infrared

- Polar bonds are strong
- Dipole moments present
- Charge is permanent dipole moment upon vibration

- Charge is polarization of molecule
- Non-polar or homonuclear bonds strong in the Raman effect

### (ii) Raman

## Characteristic Group Frequencies

Use characteristic Frequency table to determine groups!

- Func groups have same frequencies regardless of nature of remaining molecule
- Observation of certain frequencies indicates which groups present
- With frequency not present, then func. group is not either

## Molecular 1D-Basis

- UV-vis spectrum is a 'fingerprint' of the molecule
- Data must be from pure compounds!

## CWIR Evolution

- Albert Michelson: Michelson & Morley → Nobel Prize

↳ Prof. of Physics at Cesc

- William W. Coblenz: Established IR spectroscopy as an analytical technique  
↳ Attended Cesc as an OG student

- Jack Koening: First Raman study of carbon

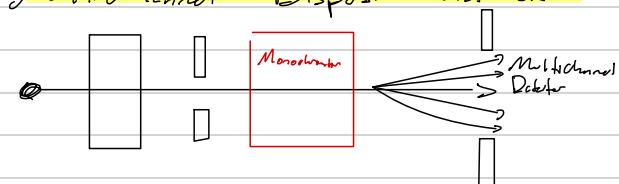
- H. Isidor: FT-IR analysis. Optical Techniques

## Dispersive Infrared Instrument



Solution

## Multichannel Dispersive Instrument

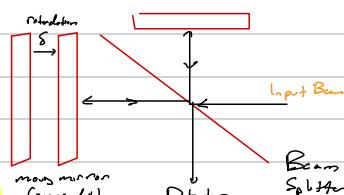


## Disadvantages

- Slow Scanning Process
- Limited energy throughput  
↳ Dispersion loses energy
- Digitalizing difficult

## Michelson Interferometer

So, energy lost!



## Fourier Transform (FTIR) Spectrometer

- Encodes the info in an interferogram using Michelson interferometer
- Performing Inverse Fourier Transform to change Distance → Frequency  
↳ FFT used to transform Interferogram to Spectrum
- Computation:  $I(x) = \int_{-\infty}^{\infty} B(u) e^{2\pi i ux} du$   
 $B(u) = \int_{-\infty}^{\infty} I(x) e^{-2\pi i ux} dx$

## Advantages

- Jacquinot: Energy Throughput ↑
- Fellgett: Multiplex
- Conrad: Frequency Accuracy

FTIR Originally Studied by Frank Sombret

No Disadvantages  
to using FTIR

# Nuclear Magnetic Resonance Spectroscopy

## Nuclear Magnetic Resonance (NMR) Spectroscopy

- Obtains spectra from unique magnetic structures of the nuclei
- Details nuclei Magnetic Property

- (i) Liquid
- (ii) Solid State
- (iii) MRI Imaging

3 types, but same principle. Diff conditions

### Advantages

- Highly Specific
- Simple Interpretation

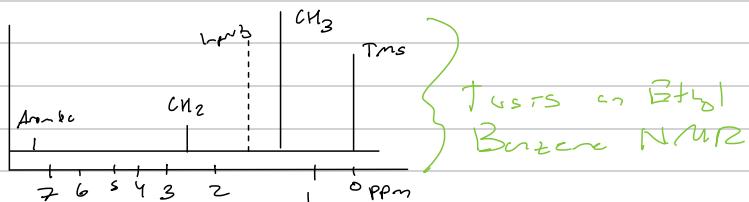
### Disadvantages

- Low Sensitivity
- Measurement time Long
- Difficult to work with complex mixtures
- Very Expensive to operate & maintain
- Solubility limitations of polymers

500-600+ MHz preferred  
1.1 GHz on campus!

## How does NMR Work?

- Some atomic nuclei behave as tiny bar magnets when placed in field  $\rightarrow$  alignment
- $\hookrightarrow$  They magnetize about the magnet at its Larmor frequency
- Hit Nuclei with pulse to knock its spin, so it will flip spin up back to original state
- ppm reported usually
- Software used for reporting, cause inverse FT to go from func of time to frequency  
 $\hookrightarrow$  By Def, NMR is Fourier transform, so it's not called FT-NMR
- Present molecules show - $\nu$  ps ~ one spot or. You can see impurities as well
- Farthest left are most  $\nu$  ps, meaning they furthest from Larm Frq.  $\hookrightarrow$  7 6 5 4 3 2 1 0 ppm
- TMS is  $\text{Si}(\text{CH}_3)_4$  and is one of the most positive compounds you can find, used as baseline
- NMR has a superconductor magnet - needs to be very cold



## NMR Signal to Noise Ratio

$$\frac{S}{N} = \gamma N I (I+1) \left[ \frac{B_0}{T} \right]^{3/2} F \left( \frac{\Omega V_0}{b} \right)^{1/2}$$

- $\Rightarrow \gamma$  is magnetogyric ratio of nucleus
- $\Rightarrow B_0$  is external magnetic field
- $\Rightarrow N$  is # of magnetically active nuclei
- $\Rightarrow T$  is sample temp

- $\Rightarrow Q$  is quality factor of resonant circuit
- $\Rightarrow F$  is filter factor
- $\Rightarrow b$  is bandwidth of detector
- $\Rightarrow V_0$  is volume of sample

## General Notes on NMR

- Increased chemical sensitivity w/ applied magnetic field  
 $\hookrightarrow$  Can study more neighbors w/ lower frequency
- NMR Spectrum revealed that polyethylene can be short-chain branched  
 $\hookrightarrow$  Butyl and ethyl groups are most frequent
- Intensity Ratio comes from Pascal's triangle with # nub.. neighboring interactions

### Iso tactic Peaks

### Syndiotactic Peaks

### A tactic Peaks



Combination of R other two

# Thermal Property Analysis (DSC & TGA)

## Fundamentals of DSC (Differential Scanning Calorimetry)

- Measures the flux of heat

↳ Temperatures & heat flows associated with transitions in materials

- Calorimeter:** Measures heat into and out of a sample

- Differential Calorimeter:** Measures heat of a sample relative to a reference

- A DSC does both of the above and heat sample w/ linear temp. ramp

(Endotherm: heat flows into system)

(Exotherm: heat flows out of system)

- Crystalline Compounds have a melting endotherm

- Makes Instrumentation consists of usually 4 parts

↳ Usually Use Aluminum unless sample is reactive w/ Al

- T<sub>g</sub> readings are about 100x weaker than the other technique commonly used today

**Exotherm Events:** Crystallization, Solid-state Transitions

- Lack of Peak Thermograms are 2<sup>nd</sup> order transitions & Thermograms

↳ With peaks, these are 1<sup>st</sup> order transitions

- Glass Transitions (second-order transition)

→ 2<sup>nd</sup> order transition

→ characterized by change in heat capacity

→ appears as a step with the midpoint being the reported T<sub>g</sub>

→ Small Endo peak may appear due to shippers / purity issues

- Enthalpy of Fusion given as area under melting curve

- Run by calorimeters based on Unit Heat Equations

$$T_m = T_o - \frac{RT_o^2 \chi}{\partial H_o} \cdot \frac{1}{f}$$

Look for  
Thermograms  
in research

ThermoGram

Enthalpy relaxation  
phenomenon

## Reversing and Non-Reversing Contributions to Total DSC Heat Flow

$$\text{Total Heat Flow} \rightarrow \frac{dQ}{dt} = C_p \cdot \frac{dT}{dt} + f(t, T)$$

Non-reversing heat component  
Reversing signal

Fourier Transform Technique

- Time domain to Frequency domain, but DSC doesn't contain an inverse Fourier

**Temperature Modulated DSC:** Decreases heat capacity and separates heat flow into reversible and non-reversible events

↳ Heating rates: 0-8 °C, Period: 60 sec, Amplitude, +/- 1 °C