



AMERICAN INTERNATIONAL UNIVERSITY – BANGLADESH

Department of Natural Science (Chemistry)

Faculty of Science & Technology

Programs: B.Sc. Eng'g (EEE/CSE/IPE)

CHEM 1101: CHEMISTRY

Chemistry Lab Report

Semester: Spring

Session: 2022-2023

NO EXPERIMENT, NO REPORT

Experiment No: 8

Name of the Experiment: Determination of strength of a weak acid (CH_3COOH) against a strong alkali (NaOH) solution by measuring conductance.

Date of Performance: 5-4-23

, Date of Submission: 12-4-23

Course-Teacher:

Instructions:

1. A lab report consists of three parts: a cover page, body of the report and a data and results sheet (lab-sheet).
2. This is the cover page of a report and students will collect and preserve the lab-sheet of a particular experiment to be performed.
3. Body of the report includes-(1) Objective of the Experiment, (2) Theory, (3) Name of the Chemicals, (4) Name of the Apparatus, (5) Percentage of Error (if necessary) and (6) Discussion (I. Precautions taken, II. Possible errors).
4. Use A4-size off-set paper, write on one side of the paper by hand keeping suitable margin.
5. Staple the lab-sheet at the end of the report and cover page on the top.
6. Submit the report in time to avoid deduction of marks.
7. Students working in a group will write and submit the report individually.
8. Copying of the report from others is strictly prohibited.

Name of the Student: TRIDIB SARKAR
ID No: 22-46444-1, Section: F, Group: 03

FOR FACULTY USE ONLY

Faculty comments: , Signature:
Date:

Objective:

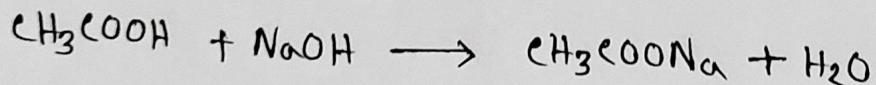
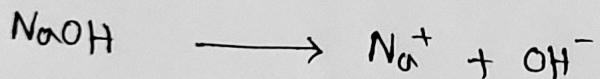
- * To draw the titration curves by measuring the conductance.
- * To find out the end-point of an acid-base reaction.
- * To know the strength of supplied sample solution (acid or base).

Theory:

(i) Methods: The conductance of an acid solution varies with the amount of alkali solution added to it because of the change in the number and nature of the ions in the solution. The curve showing this variation is called the conductance titration curve. After complete neutralization, the conductance increases due to the addition of the conducting ions of the alkali. At the end-point there is a sharp change in the conductivity. Therefore, if the conductance titration curve is drawn graphically, the end-point of the titration can easily be determined. The conductance of a dilute acetic acid solution is due to the small amounts of H^+ ions and acetate ions resulting from the dissociation of the weak acid. When small amount of alkali is added and the H^+ ions are neutralized, an equal amount of H^+ ions are not generated by further dissociation of the weak acid because the acetate ions suppress the dissociation due to common ion effect. Therefore the conductance decreases. Upon further addition of alkali the conductance increases because of the addition of Na^+ ions.

and formation of acetate ions. After the end-point, the conductance increases at a sharper rate due to addition of the fast moving OH^- ions, in addition to the Na^+ ions.

(ii) Reactions:

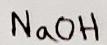


Required chemicals:

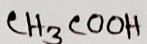
Name of the chemicals

1. Supplied 0.1 N
Sodium hydroxide solution

Chemical formula



2. Supplied d.m.
Acetic acid solution



Apparatus:

- (i) Conductivity meter.
- (ii) Burette (50 ml).
- (iii) Pipette (10 ml)
- (iv) Conical flask (250 ml).
- (v) Volumetric flask (100 ml)

- (vi) Plastic beaker .
- (vii) Watch glass .
- (viii) Pipette filler .
- (ix) Dropper .
- (x) Glass rod .
- (xi) Stand and clamp .

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Attention: Please bring one graph paper

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Experiment 8

CHEM 1101: CHEMISTRY (EEE/CoE/CSE/IPE)

EXPERIMENT NO. 8: DETERMINATION OF STRENGTH OF A WEAK ACID (CH_3COOH) AGAINST A STRONG ALKALI (NaOH) SOLUTION BY MEASURING CONDUCTANCE.

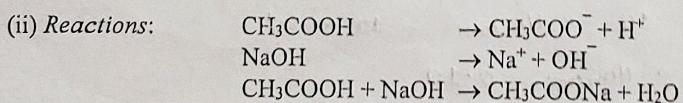
OBJECTIVE:

- To draw the titration curves by measuring the conductance
- To find out the end-point of an acid-base reaction
- To know the strength of supplied sample solution (acid or base)

THEORY:

(i) *Methods:* Conductometric titration,

The conductance of an acid solution varies with the amount of alkali solution added to it because of the change in the number and nature of the ions in the solution. The curve showing this variation is called the conductance titration curve. After complete neutralization, the conductance increases due to the addition of the conducting ions of the alkali. At the end-point there is a sharp change in the conductivity. Therefore, if the conductance titration curve is drawn graphically, the end-point of the titration can easily be determined. The conductance of a dilute acetic acid solution is due to the small amounts of H^+ ions and acetate ions resulting from the dissociation of the weak acid. When small amount of alkali is added and the H^+ ions are neutralized, an equal amount of H^+ ions are not generated by further dissociation of the weak acid because the acetate ions suppress the dissociation due to common ion effect. Therefore the conductance decreases. Upon further addition of alkali the conductance increases because of the addition of Na^+ ions and formation of acetate ions. After the end-point, the conductance increases at a sharper rate due to addition of the fast moving OH^- ions, in addition to the Na^+ ions.



APPARATUS:

Conductivity meter, Burette (50mL), pipette (10mL), conical flask (250mL), volumetric flask (100mL), plastic beaker, watch glass, pipette filler, dropper, glass rod, stand and clamp etc.

REQUIRED CHEMICALS:

(1) Supplied 0.1N NaOH solution, (2) Supplied dil. CH_3COOH solution

Name: TRIDIB SARKAR ID No: 22-46444-1 Section (Group): F-(63)

(Expt. 8 contd.)

PROCEDURE:

Take 10 mL of the supplied CH_3COOH solution in a beaker. Place the previously washed (with hot water) conductance cell in it and add sufficient water (~ 200 mL) to keep the electrodes of the cell immersed. Measure the conductance of the acid solution (1st reading). Fill a burette with the supplied ~ 0.1 N NaOH solution. Add 2/1 mL NaOH solution from the burette, stir the solution and measure the conductance (2nd to 12th reading) after each addition. Get a graph paper ready for the plotting conductance data. Plot conductance data (in μs) against the final volume (in mL) of NaOH solution on graph paper and find the end-point. The end-point gives the required volume of NaOH equivalent to 10 ml of supplied dil. CH_3COOH solution. Now calculate the strength of CH_3COOH solution.

EXPERIMENTAL DATA:

Table: Conductance-measurement of CH_3COOH and NaOH solutions using conductivity meter.

No. of reading	Vol. of CH_3COOH (in mL)	Vol. of NaOH (burette reading) (in mL.)				Conductance (μs), Y
		Initial	Final	Difference	Total X	
1	10	0	0	0	0	098
2		0	2	2	2	092
3		2	4	2	4	128
4		4	6	2	6	191
5		6	8	2	8	243
6		8	10	2	10	289
7		10	11	1	11	336
8		11	12	1	12	375
9		12	13	1	13	480
10		13	14	1	14	525
11		14	15	1	15	575
12		15	16	1	16	620

CALCULATIONS:

Strength of supplied CH_3COOH solution:

$$V_{\text{CH}_3\text{COOH}} \times N_{\text{CH}_3\text{COOH}} = V_{\text{NaOH}} \times N_{\text{NaOH}}$$

$$\Rightarrow 10 \times N_{\text{CH}_3\text{COOH}} = 10 \times 0.1$$

$$\Rightarrow N_{\text{CH}_3\text{COOH}} = 0.10$$

RESULTS:

The strength of supplied CH_3COOH solution 0.10 N

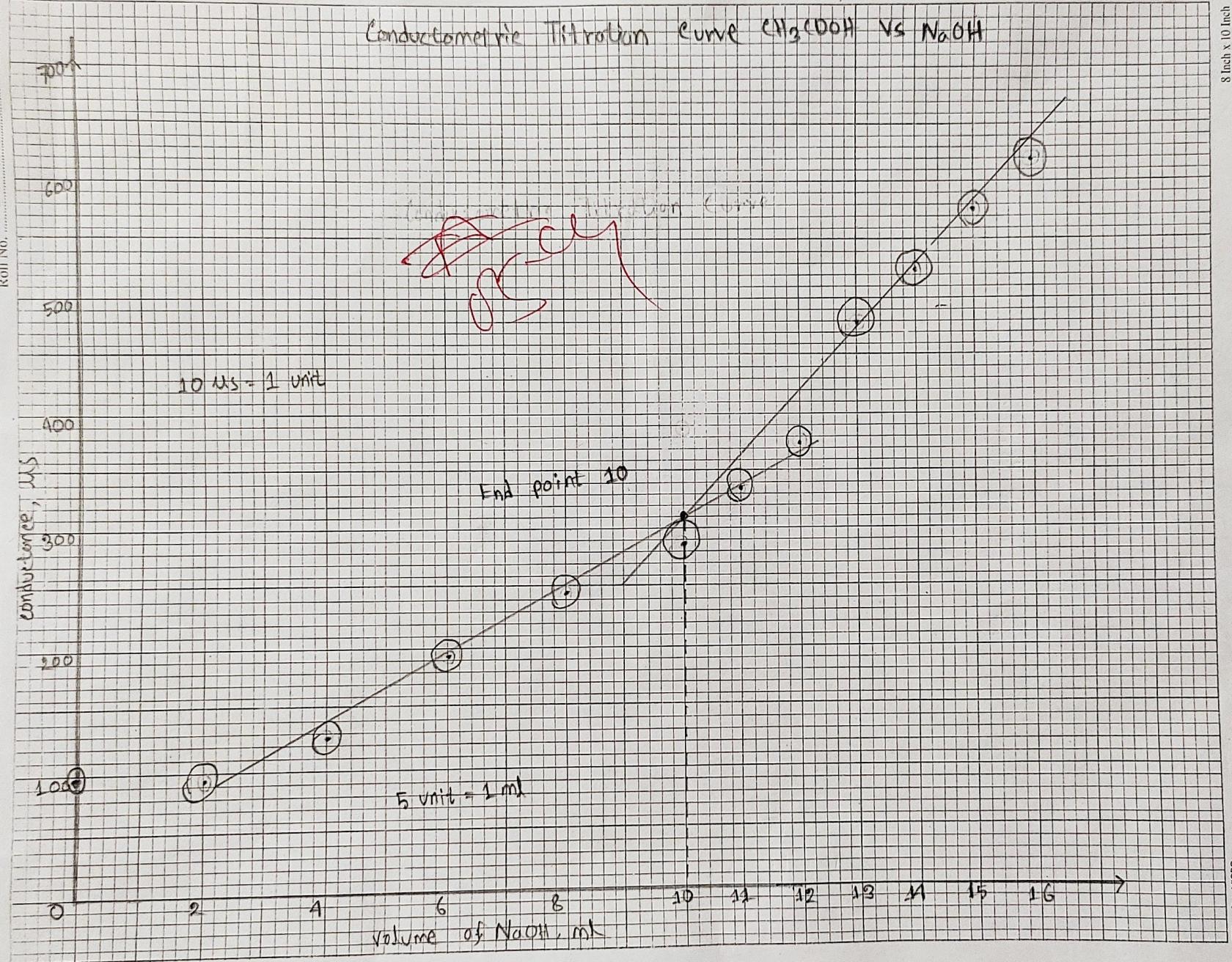
(Attach the graph paper with this lab-sheet)

Students should know

- Define (a) electrolyte, (b) specific conductance, (c) molar conductance
- What is conductance cell?
- How does the molar conductance of a strong electrolyte vary with concentration?
- How is the molar conductance at infinite dilution determined for (a) a strong electrolyte and (b) a weak electrolyte?

Conductometric Titration Curve CH₃COOH vs NaOH

Roll No. 22-4644-1



8 inch x 10 Inch

Discussion:

(i) Precaution:

- ① The apparatus were washed properly as instructed.
- ② The temperature was fixed constant through out the experiment.
- ③ The titrant is 10 times stronger so that the volume change is as little as possible.

(ii) Possible error:

- ① Error might be occurred when took the burette reading.
- ② Error might be occurred while calculating concentration.
- ③ Mistakenly added few drops of NaOH solution might change the output or outcome.