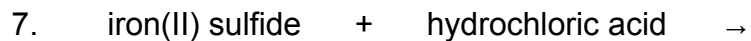
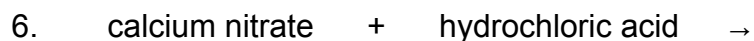
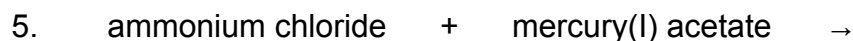
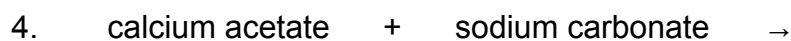
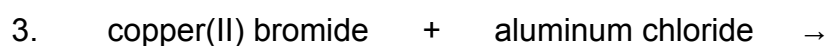
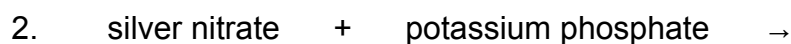
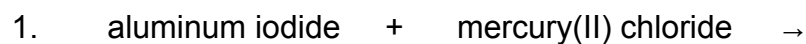
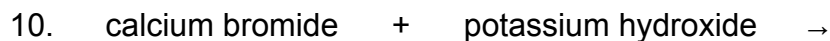


Worksheet #5: Double-Replacement Reactions

In these reactions, all you do is look at the names of the reactants, and "switch partners". Just be sure that the new pairs come out with the positive ion named first, and paired with a negative ion.

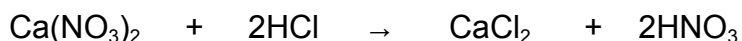
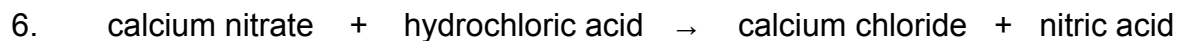
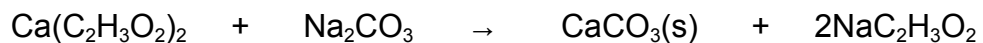
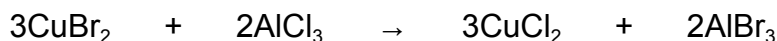
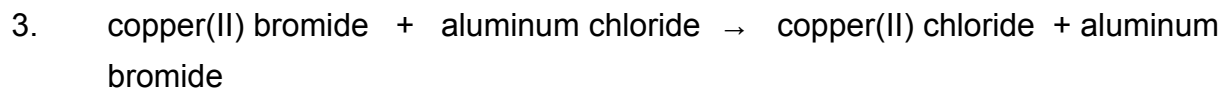
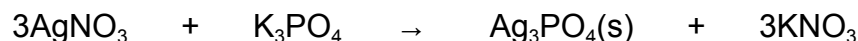
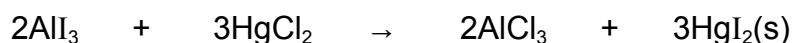
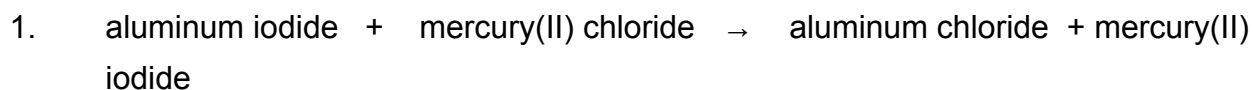




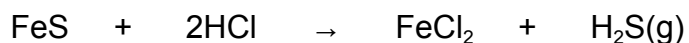
Examine the products of the reactions on this page, and determine in each whether a gas, water, or a precipitate is formed. If there is no gas, water, or precipitate produced, put an "X" through the yield sign, because no reaction occurs.

Worksheet #5: Double-Replacement Reactions

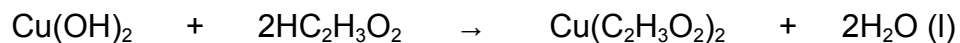
In these reactions, all you do is look at the names of the reactants, and "switch partners". Just be sure that the new pairs come out with the positive ion named first, and paired with a negative ion.



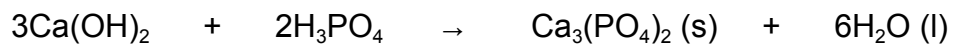
7. iron(II) sulfide + hydrochloric acid → iron(II) chloride + hydrogen sulfide (g)



8. copper(II) hydroxide + acetic acid → copper(II) acetate + water



9. calcium hydroxide + phosphoric acid → calcium phosphate + water



10. calcium bromide + potassium hydroxide → calcium hydroxide + potassium bromide

