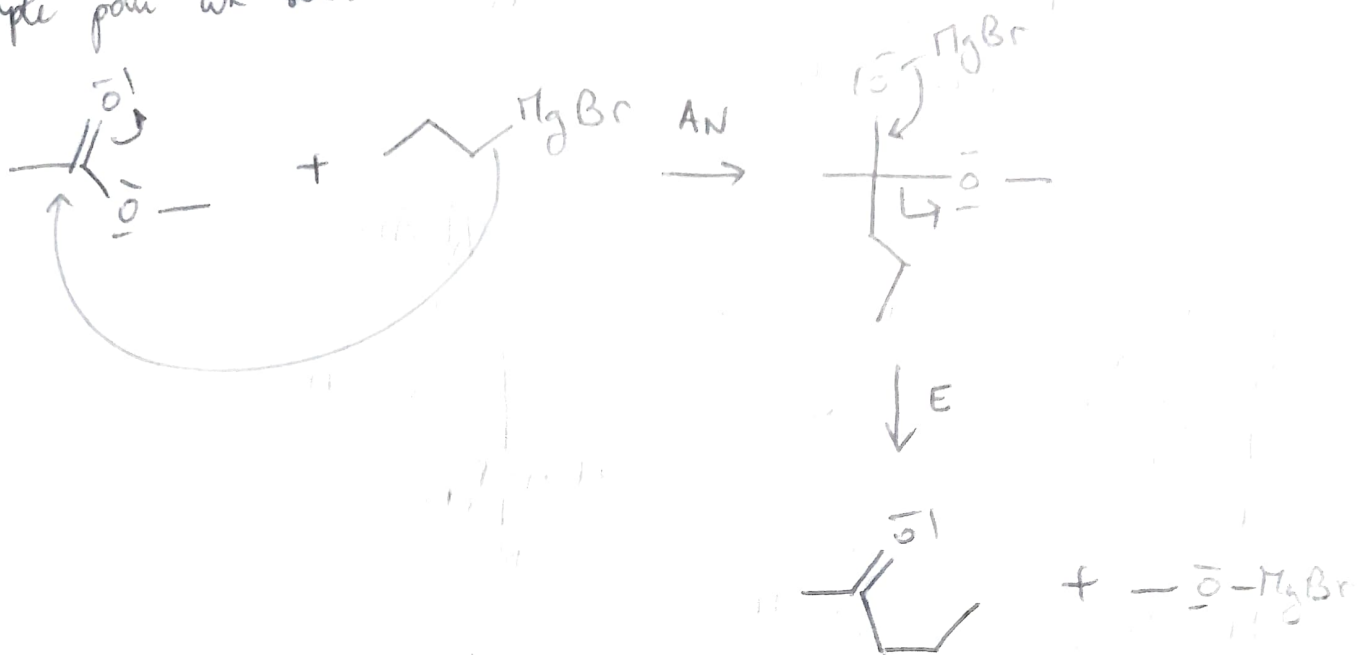


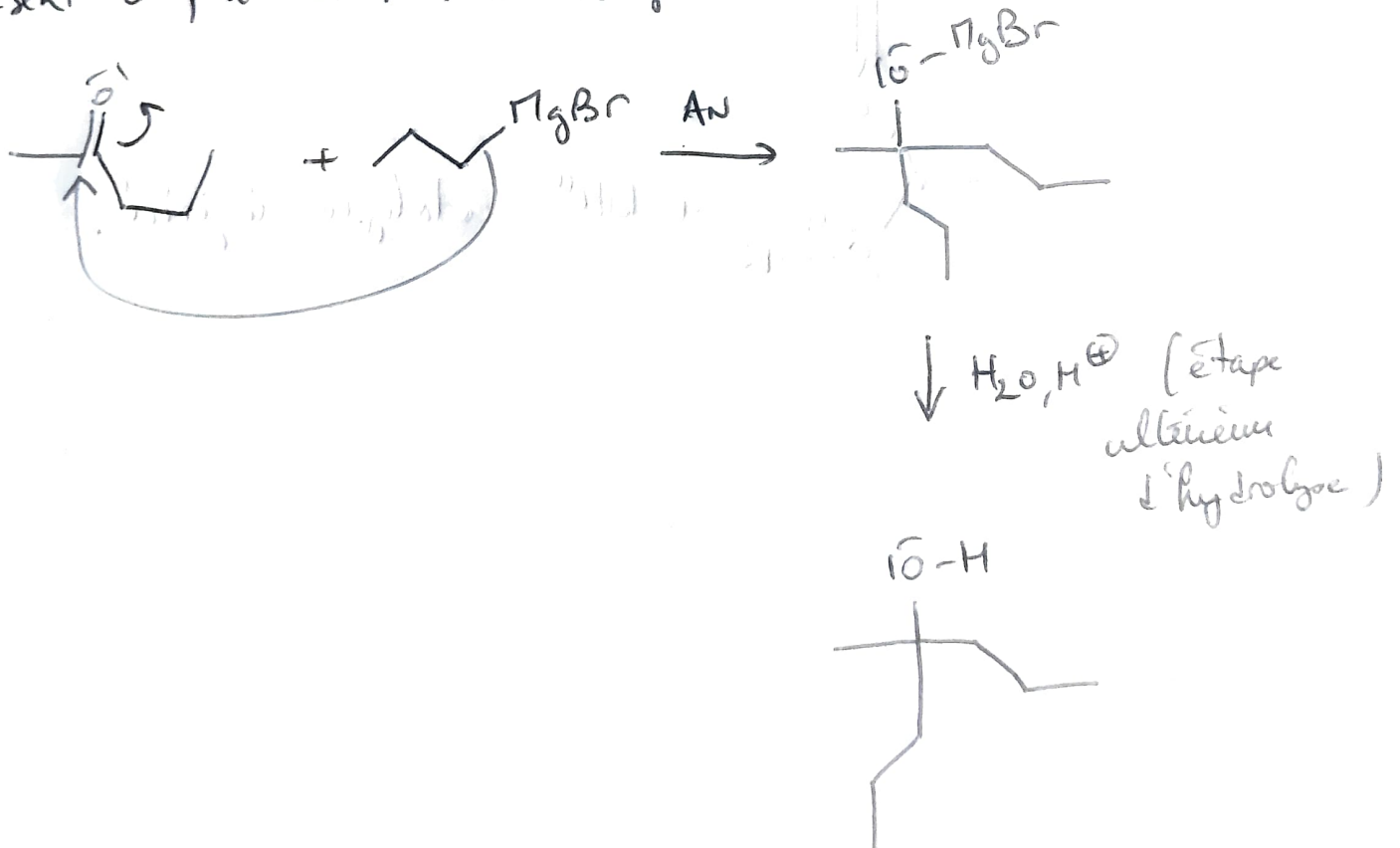
## Annexe Cours 8

2.1.2

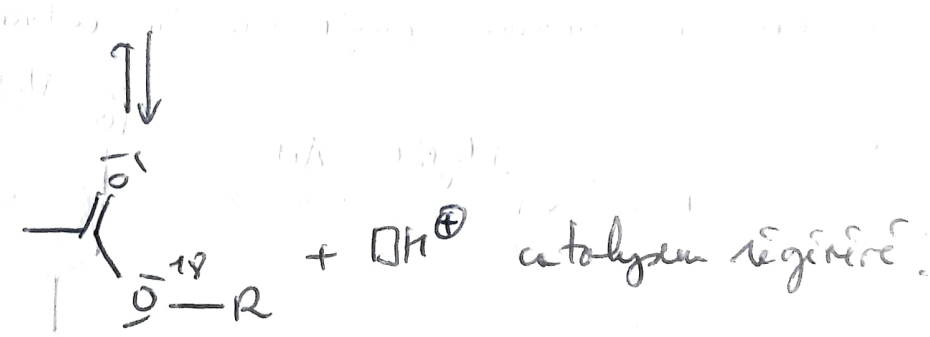
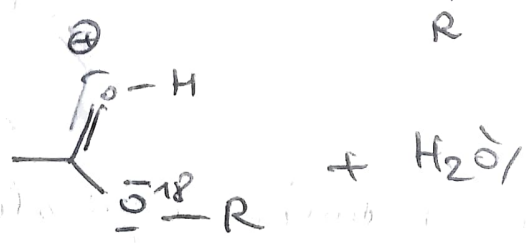
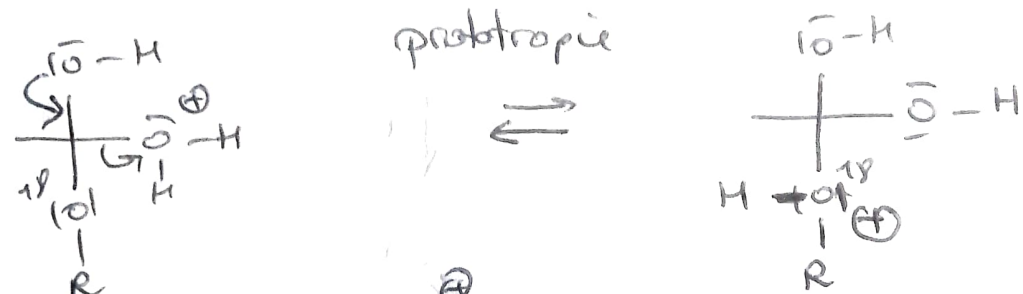
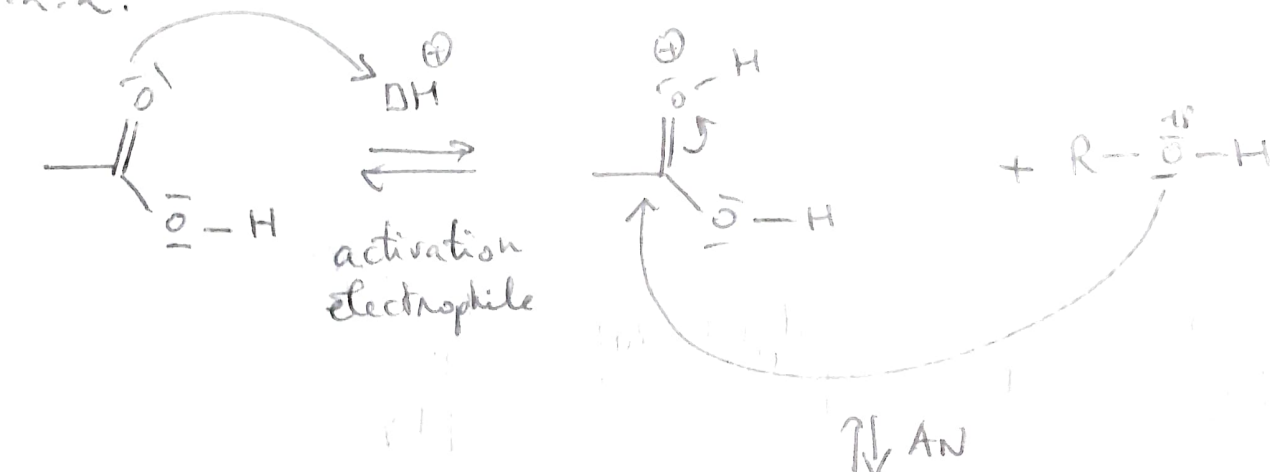
exemple pour un ester:



Dans le cas où il y a un deuxième équivalent d'organomagnésien présent il peut de nouveau réagir sur la cétone:

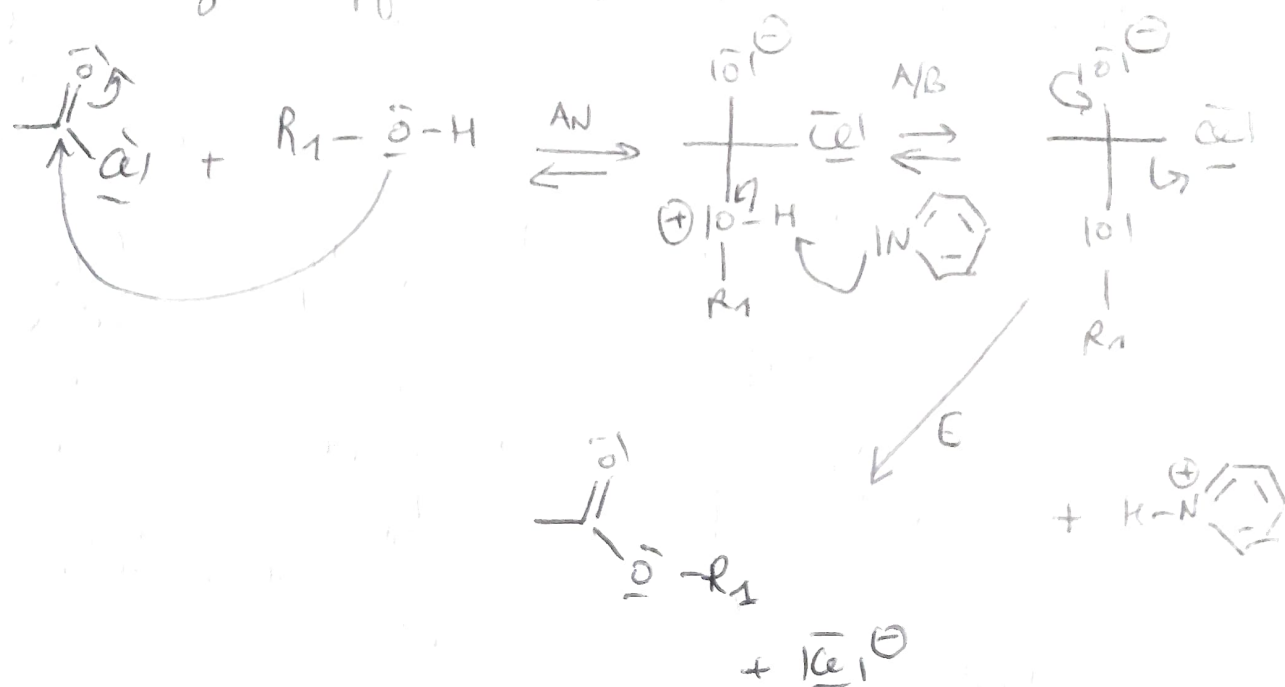


2.2.2.

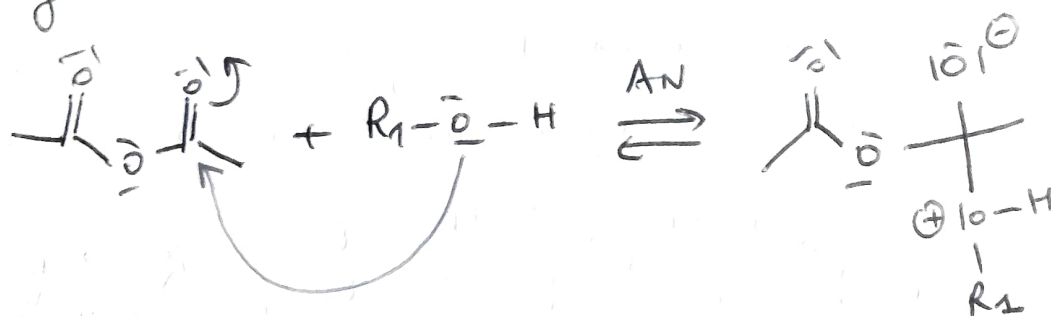


2.2.3.

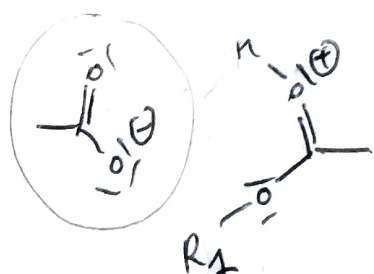
Chlorure d'acyle + pyridine:



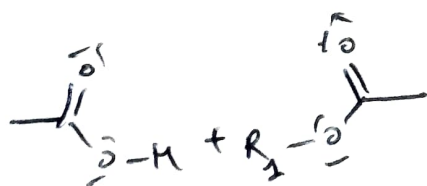
Anhydride d'acide:



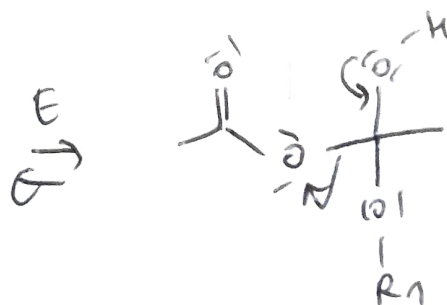
Bon groupe partant



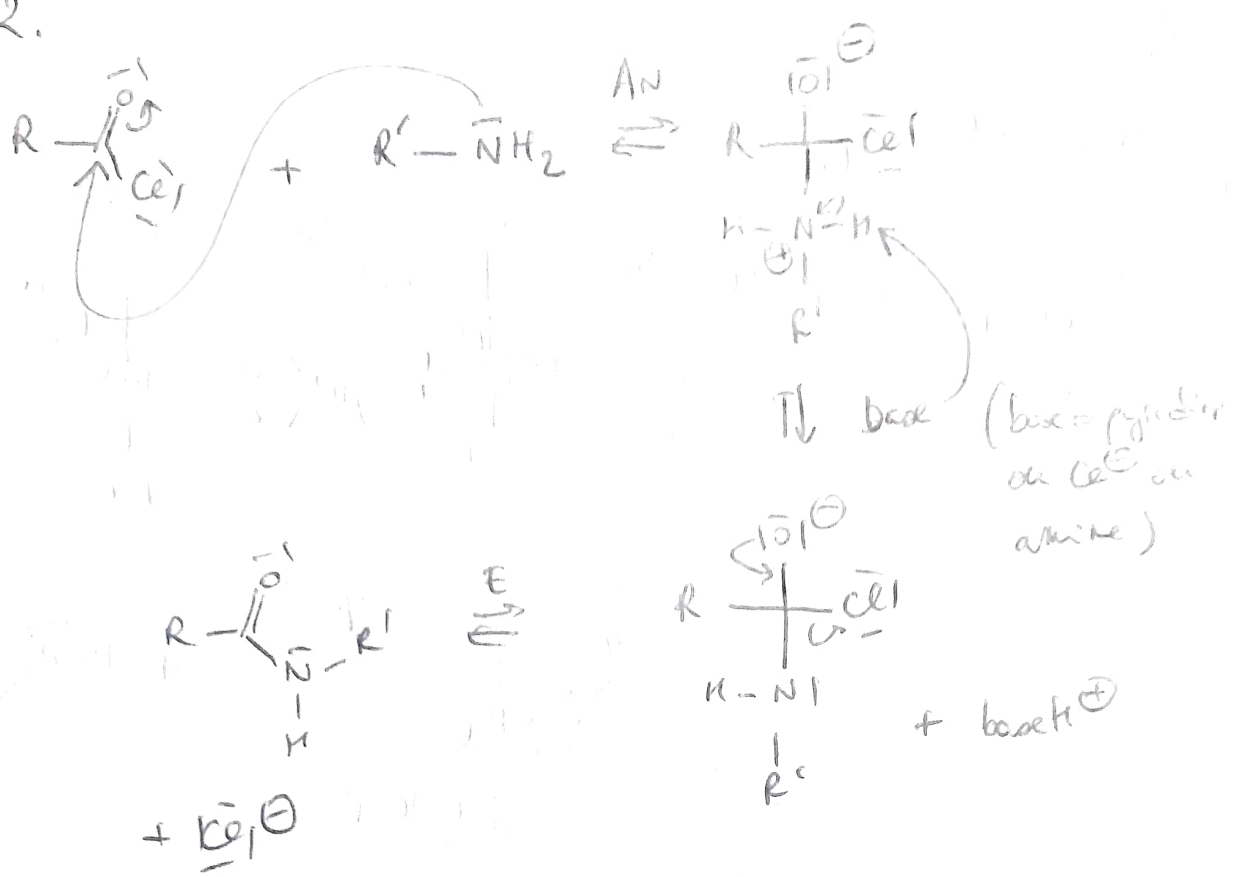
↓ NB



↕ prototropie

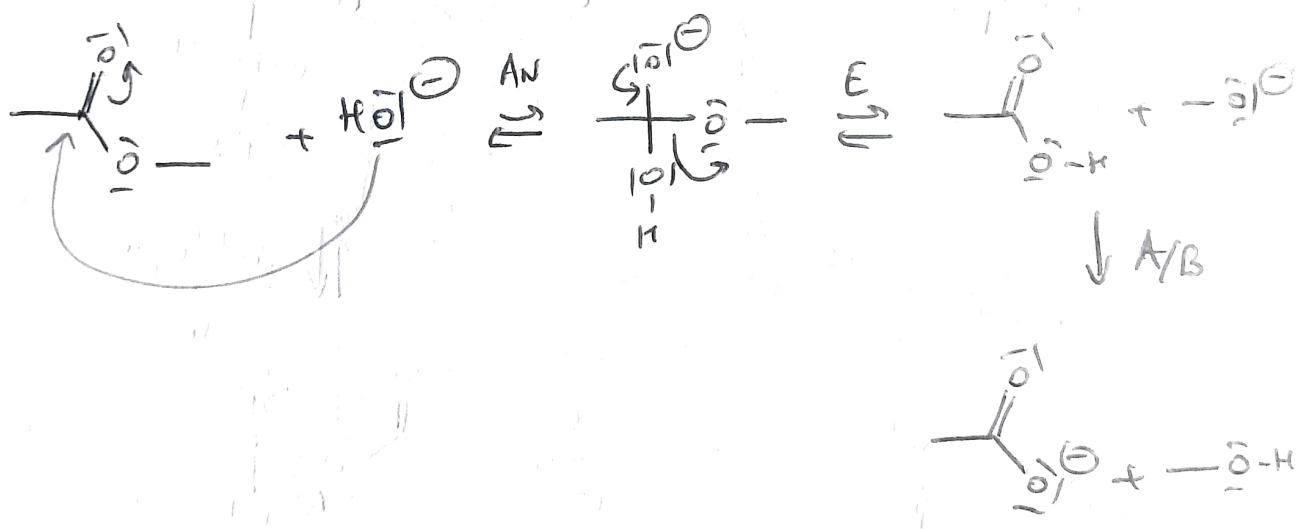


2.3.2.



2.4.2

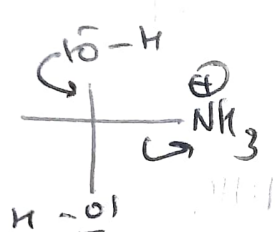
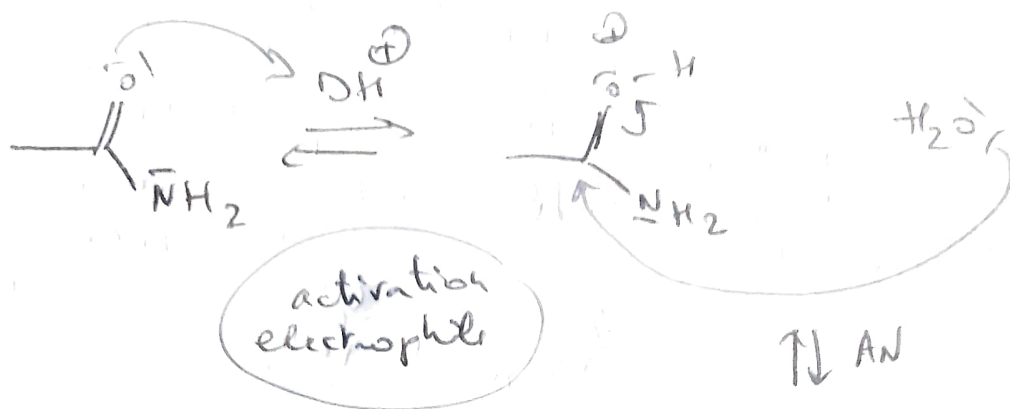
Saponification:



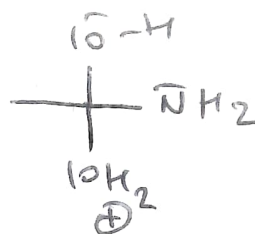
C'est la réaction A/B totale de la fin qui déplace les équilibres de la saponification.

2.4.3

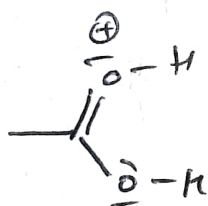
Hydrolyze acid



protonate

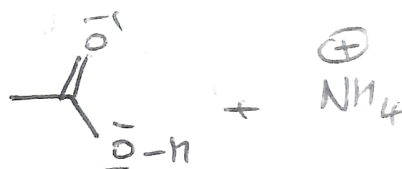


$\downarrow E$

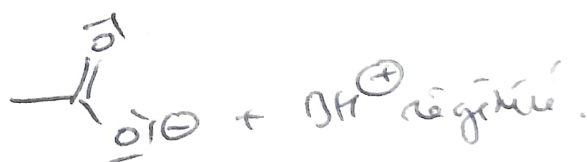


+ / NH<sub>3</sub>

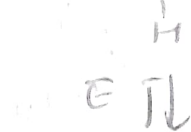
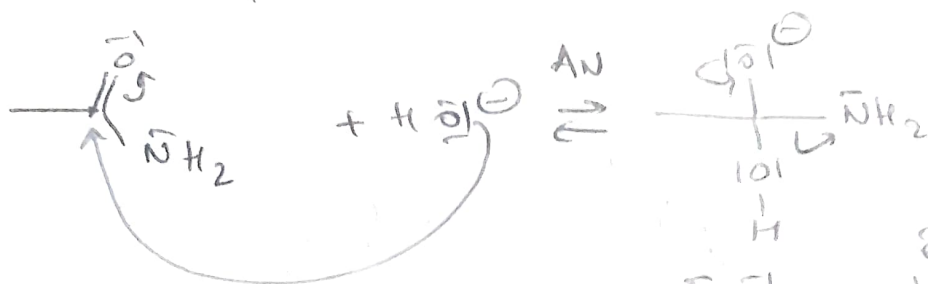
A/B



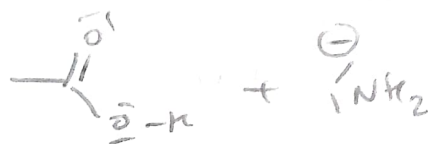
$\downarrow$



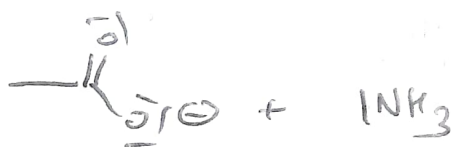
# Hydrolyse basique



étape très  
difficile nécessite  
de chauffer le milieu  
réactionnel.

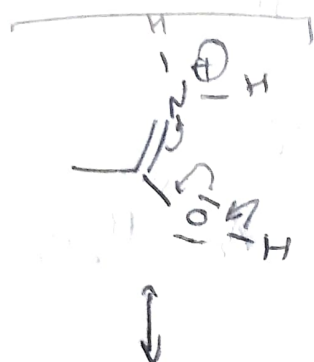
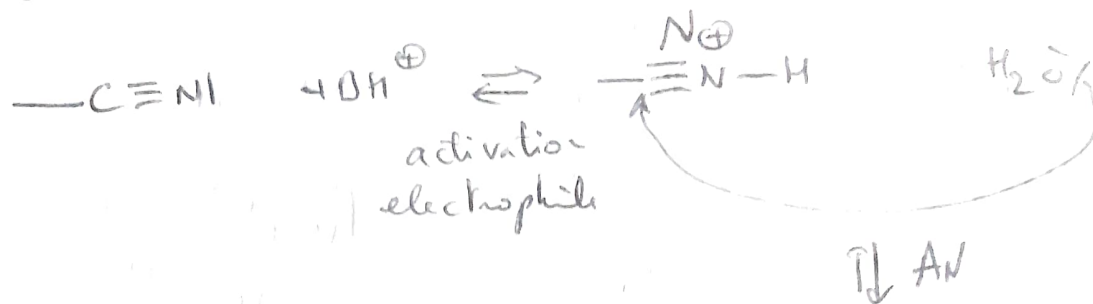


↓ A/B

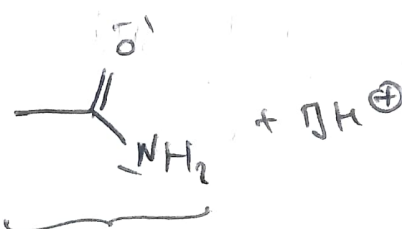
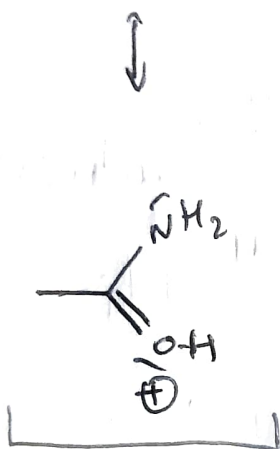
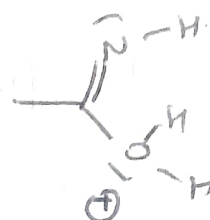


2.4.4.

mécanisme en condition acide :



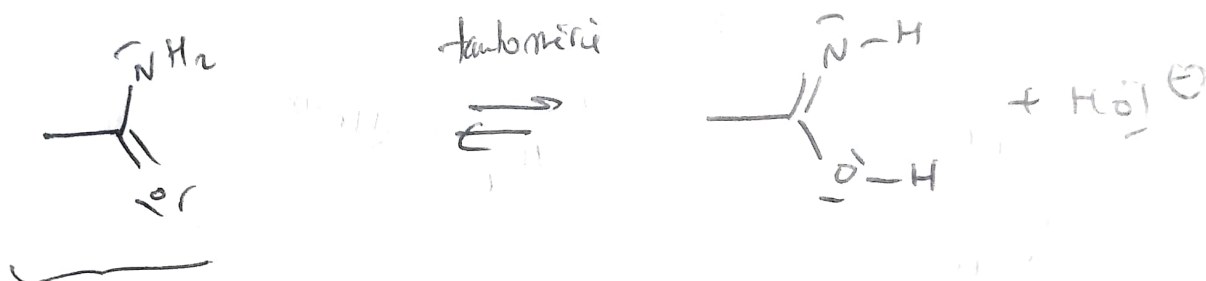
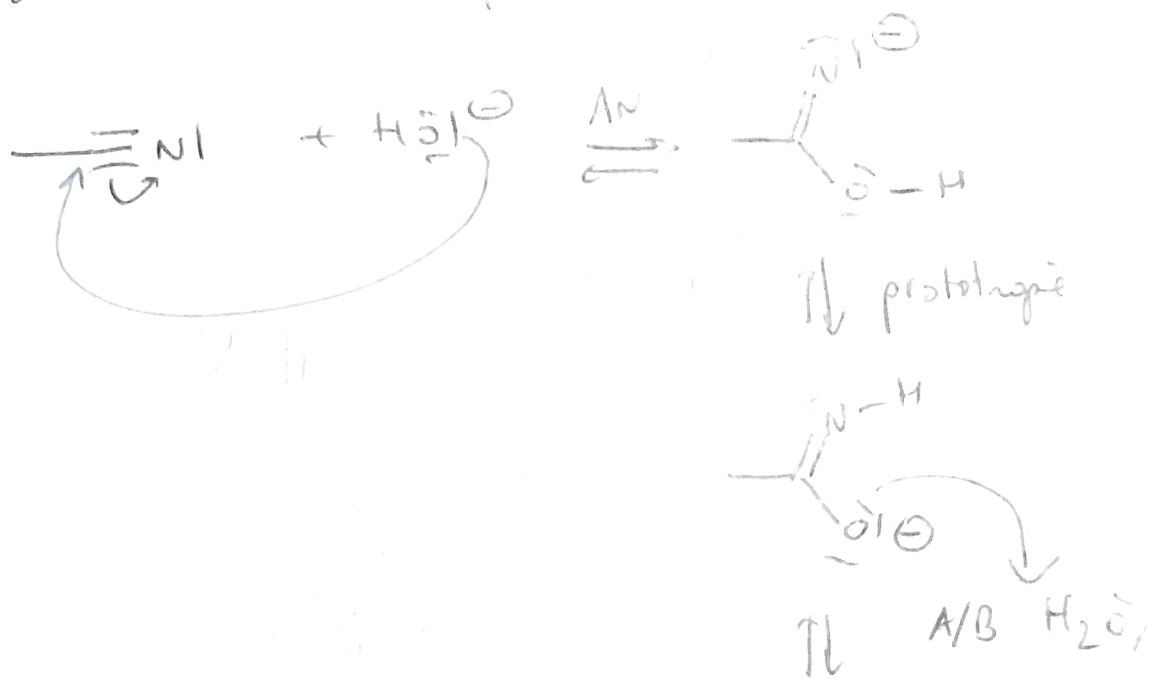
prototropie



amide

↳ la suite de la réaction  
est décrite au § 2.4.3.

mécanisme en conditions basiques:



amide: suite de

l'hydrolyse décrite au § 2.4.3.