CHM191
The Bunsen burner regulator can be adjusted to produce a ---------- flame.

Non-luminous
Electric oven can also be referred to as -------- oven.

Drying
If corrosive fumes may evolve during evaporation, the process must be carried out in a ---------------.

Fume hood
The most conventional size for volumetric work is the 250ml ------------------

Conical flask
Liquid reagent should be taken with the help of -----------------.

Droppers
The --------------- is the basic unit for the quantity of a substance.

Mole
The three different techniques in titrimetric analysis are direct, indirect and ---------- titration.

Back
Aluminium is --------------- reactive than calcium.

Less
A substance that lose electron is said to be ----------------

Oxidized
The more ----------------- the potential, the easier the oxidation will be.

Positive
Heat of -------------- is the amount of heat released or absorbed for a given amount of reactant or products.

Reaction
The -------------- is used to measure the heat of reaction.

Calorimeter
-------------- are volumetric analysis involving iodine.

Iodimetry
KMnO4 is characteristically -------------- in both acid and base media.

Coloured
In --------------- reactions , oxidation and reduction occurs simultaneously

Redox
The two balanced ---------------- equations should add up to give redox equations.

Half
Phenolphthalein can be used in the titration of HCl and NaOH since the titration is between -------------- acid and strong base.

strong
When standard solution directly reacts with the substance being determined, the titration is said to be ------------------- titration.

Direct
Each types of titrimetric analysis is based on certain types of ------------.

Reaction

Titration results are best recorded in ---------- decimal places.

Two
---------- of a substance is achieved by titrating the approximate concentration with standard solution of known concentration.

Standardization
Is it practically possible to prepare a solution of high concentration from a solution of lower concentration?

NO
A solution contains 2.65g of anhydrous sodium carbonate in 125ml of solution, what is its concentration in mole per dm cube?

0.2
A standard solution is prepared by weighing a ------- solute and dissolving it in a suitable solvent.

Pure
The equivalent weight of solute per volume of solution in dm3 is termed ----------

Normality
Sometimes precipitates are heated to a high temperature to convert it into a ---------- of constant composition.

Compound
Ash less filter paper leaves little residue on ----------.

Ignition
For laboratory purpose, filter papers of ---------- grades are generally made.

Three
Proper ---------- of filter paper can increase the rate of filteration.

Folding
An antidote should be given only to a person that swallow ---------- poisons.

Non-corrosive
When you are very careful and follow all the laboratory rules of safety ,all accidents in the laboratory will be avoided. True or False

False
A ------- notebook should be used for laboratory record.

Bound
Ordinary filtration can be speeded up by the use of ---------- filtration.

Suction
Care must be taken so that the liquid does not ---------- violently while evaporating in a boiling tube.

Bump
Conical flask is the container in which the ---------- meet and the products is/are formed.

Reactants
One of the following causes mental confusion

CHCl3

Which of the following is most dangerous to the skin?

HF

Bunsen burner can be adjusted to a moderately high temperature of about _____.

600 oC

Which of these operations is performed in wet chemical analysis?

Precipitation

The number of moles of solute in one kilogram of solution is called _____

Molality

All pure substance have the following characteristics except _____.

Readily soluble in water

There are _____ types of titrimetric analysis.

Four

Which indicator is suitable for the titration of ethanoic acid and ammonia.

None

Which of the following metals has the least oxidation potential?

Ag

A reaction involving iron and copper ions, which of the ions will be oxidized?

Fe

The reaction is _____ when the reaction vessel feels warmer.

exothermic

The amount of heat required to raise the temperature of one mole of substances through one degree Celsius is called _____.

Molar heat capacity

Displacement reactions are usually _____.

oxidized

The most commonly used external indicator for iodine titration is _____.

starch

It is very difficult to read the meniscus of potassium permanganate solution in burette because _____.

Of its intense colouration

Which of these acid is suitable to catalyze permanganate reaction?

Sulphuric acid

Oxidation is defined as the _____

Loss of electron

Given the equation, $C + H_2SO_4 \rightarrow CO_2 + SO_2 + H_2O$, the oxidation number of sulphur decrease by _____.

2

The oxidation number of sulphur in $H_2SO_4^{2-}$ is _____.

4

H_2O_2 is oxidized by MnO_2 to give _____.

O_2

Volumetric method of analysis measures _____.

Volume of a gas or volume of solution of unknown concentration

The concept of _____ method is used in the preparation of a solution of lower concentration from higher concentration

Dilution

Which of the following is true of molar and molal concentration of sodium carbonate?

Molal concentration is higher

Precipitation is common to _____ analysis.

Gravimetric

What should be given to a person that swallowed corrosive poison as soon as possible?

Calcium hydroxide solution

Burns caused by bromine can be treated with _____

Ammonia

Burns are mostly caused by _____.

Hot equipment and reagents

What is the molar concentration of a solution containing 2.5g of potassium hydroxide in 200cm3 of solution. (K = 39, H=1,O=16)

0.22mol/dm3

In a titration, 25cm3of an impure anhydrous sodium trioxocarbonate (iv) containing 5.0g in 1dm3 of solution was neutralized by 22.20cm3of 0.10mol/dm3 of HCl, find the mass concentration of the trioxocarbonate (iv).

4.71g/dm3

A molar solution of caustic soda is prepared by dissolving _____.

40g NaOH in 1000g of water

What volume of 0.5M sulphuric acid will exactly neutralize 20cm3 of 0.1M NaOH solution?

2.0 cm3

In which type of titration is two standard solution needed?

Back

The volume of acid used is usually _____

Average titre value

All electron donors are normally _____.

Reducing agent

At what reaction condition will the absolute value of enthalpy equals the absolute value of heat of reaction?

Constant pressure