FBQ1: The nature of the charge on Gamma rays is ---------charge
Answer: Neutral

FBQ2: The type of charge on Beta ray is ----------
Answer: negative

FBQ3: Low ionising power on gases is found on.....ray
Answer: Beta

FBQ4: The nature of the charge on an electron is ----------
Answer: Negative

FBQ5: The magnitude of the charge is the same on a proton and an .....
Answer: Electron

FBQ6: The constituents of the atom are....., electrons and neutrons
Answer: Protons

FBQ7: The relative atomic mass of an element is the mass of one atom of the element compared to (1/12) of the mass of one atom of carbon - 12 .....
Answer: Isotopes

FBQ8: Mass spectrometer is the instrument used to determine fairly accurately the relative atomic masses of .....
Answer: Elements

FBQ9: When two or more elements combine chemically in fixed proportion by mass, it results to .....
Answer: compound

FBQ10: The .....of compounds and the elements from which they are formeddiffer from each other
Answer: Properties

FBQ11: Non metals are .....and gases
Answer: Solids

FBQ12: The process in which two or more light nuclei combine to form a heavier nucleus with a release of energy is known as..... fussion
Answer: Nuclear

FBQ13: Electrovalent bonding involves electron transfer from the valence shell of one .....to the valence shell of the other
Answer: Atom

FBQ14: The process in which the nucleus of a heavy element is spilt into two nuclei of nearly equal mass with a release of energy and radiation is known as nuclear .....
Answer: Fission

FBQ15: The path of deflection of Alpha rays is towards a negative .....
Answer: Pole

FBQ16: The path on which Gamma rays deflected have no effect on .....field
Answer: Electric

FBQ17: The statement that atoms tend to gain or lose electrons until there are eight electrons in their valence shell refers to..... rule_____
Answer: Octet

FBQ18: Michael faraday’s experiment is in……..
Answer: Electricity

FBQ19: Covalent bonding is a form of bonding arises from the .....of electrons

among atoms
Answer: Sharing

FBQ20: Vaander Waal's forces are form of bonding exist even between uncombined atoms and non .....molecules?
Answer: Polar

FBQ21: Radioactivity can be associated with ......nuclei
Answer: Unstable

FBQ22: ....life is a measure of the time taken for half of the radioactive substance to decay
Answer: Half

FBQ23: A lot of energy is often required to split compounds into the ......elements
Answer: Constituent

FBQ24: Physical methods can be used to separate a .....with varying components?
Answer: Mixture

FBQ25: Chemical reactions are explain by combinations and..... of atoms
Answer: Rearrangement

FBQ26: Theory is a ......proposal to explain an observed statement of facts
Answer: Tested

FBQ27: .....postulate was discarded as a result of The knowledge of radioactivity.
Answer: Daltons

FBQ28: Faraday, ...and Millikan conducted experiments that accounted for the evidences about the nature of matter
Answer: Thompson

FBQ29: Chemical symbols and ..... are used to represent elements and compounds
Answer: Formulae

FBQ30: Lustre is a general characteristics of .....
Answer: Metals

FBQ31: The third particle in the nucleus is.............
Answer: Neutron

FBQ32: Element cannot be ......
Answer: Split

FBQ33: Substance that can be broken down into elements is called.....
Answer: Compound

FBQ34: Non metals .....cannot heat
Answer: 8Conduct

FBQ35: Metal is a solid at .....temperature
Answer: Room

MCQ1: What is the nature of the charges on an electron?
Answer: negatively

MCQ2: What is the magnitude of the charge on a proton and an electron?
Answer: the same

MCQ3: What is the constituents of the atom?

Answer: protons, electrons and neutrons

MCQ4: For the atom to be electrically neutral, what should be the magnitude of the charges of the protons and electrons?
Answer: number of protons and electrons must be equal

MCQ5: What is the relative atomic mass of an element?
Answer: the mass of one atom of the element compared to (1/12) of the mass of one atom of carbon - 12 isotopes

MCQ6: What instrument would you use to determine fairly accurately the relative atomic masses of elements
Answer: mass spectrometer

MCQ7: What do Chemical symbols and formulae used to represent ?
Answer: elements and compounds

MCQ8: Which of these laws was discarded as a result of The knowledge of radioactivity?
Answer: Daltons postulate

MCQ9: Which of these can be associated with unstable nuclei?
Answer: Radioactivity

MCQ10: Which of these statements best define the phenomenon of half-life?
Answer: a measure of the time taken for half of the radioactive substance to decay

MCQ11: What radiations does a radioactive element in its decay emit?
Answer: Alpha, beta and gamma rays

MCQ12: What type of charges do Alpha rays bear?
Answer: positively charged

MCQ13: Identify the correct statement about Alpha rays
Answer: high ionising power on gases

MCQ14: What is the nature of the charge on Gamma rays?
Answer: neutrally charged

MCQ15: What type of charges do Beta rays carry?
Answer: negatively charged

MCQ16: Which of these statements hold true about Beta rays?
Answer: low ionising power on gases

MCQ17: Which of these statements hold true about Gamma rays?
Answer: very low ionising power on gases

MCQ18: Which of these statements hold true for an element ?
Answer: cannot be split

MCQ19: What is a pure substance that can be broken down into elements called
Answer: Compound

MCQ20: How many groups can you subdivide an element into?
Answer: 2

MCQ21: What results when two or more elements combine chemically in fixed proportion by mass.
Answer: compound

MCQ22: Which of these can be associated with unstable nuclei?
Answer: Radioactivity

MCQ23: How much energy is often required to split compounds into the constituent elements
Answer: a lot of

MCQ24: What is the numbers of compounds available.
Answer: Limitless

MCQ25: How would you seperate a mixture with varying components?
Answer: physical methods

MCQ26: Which of these methods can be used to separate a mixture
Answer: all of the options

MCQ27: How would you explain chemical reactions of atoms
Answer: combinations and rearrangement

MCQ28: What is a tested proposal to explain an observed statement of fact
Answer: Theory

MCQ29: Which law states that all elements are made up of small, indivisible particles called atoms
Answer: Daltons Atomic theory

MCQ30: which of these scientists conducted experiments that accounted for the evidences about the nature of matter?
Answer: Faraday, Thompson and Millikan

MCQ31: What form of bonding exist even between uncombined atoms and non polar molecules?
Answer: Vaander Waal's forces

MCQ32: Which of these is the nature of Matter?
Answer: Electrical

MCQ33: The process in which two or more light nuclei combine to form a heavier nucleus with a release of energy is known as
Answer: nuclear fussion

MCQ34: Which of these is used for radioactivity?
Answer: all of the options

MCQ35: What does the statement that atoms tend to gain or lose electrons until there are eight electrons in their valence shell refers to?
Answer: Covalent bonding