

Chemistry IH - Study Guide

Taylor Blau

January 6, 2015

Contents

1	Introduction	5
1.1	Types of Chemistry	5
1.2	Types of Matter	5
1.3	Separation of a Mixture	6
1.4	General Properties of Substances	6
1.4.1	Chemical vs. Physical	7
1.5	SI Units	7
1.5.1	Base Units	7
1.5.2	Derived Units	8
1.5.3	Prefix System	8
1.5.4	Significant Figures	8
1.5.5	Rules for Arithmetic	9
2	Atoms and Elements	11
2.1	History	11
2.2	Basic Laws	11
2.3	Dalton's Atomic Theory	11
2.4	Modern Atomic Theory	12
2.4.1	Universal Truths	12
2.4.2	Variations on Dalton's Theory	12
2.5	Modern Structure of an Atom	12
2.6	Experiments	13
2.6.1	Cathode Ray Tube	13
2.6.2	Milikan Oil Drop	13
2.6.3	Rutherford's Gold Foil	14
2.7	Elements	15
2.7.1	Isotopes	15
2.8	Nomenclature	15
2.8.1	Common Nomenclature	15
2.8.2	Stock-system Nomenclature	16
2.9	Types of Ions	16
2.9.1	Monatomic Ions	16
2.9.2	Polyatomic Ions	16

3	Atomic Structure and Periodic Trends	19
3.1	Electron-Configuration Notation	19
3.1.1	Noble-gas Notation	19
3.2	Periodic Table and Trends	19
3.2.1	The Periodic Table	19
3.2.2	Modern-day	20
3.3	Periodic Trends	20
3.3.1	Atomic Radii	20
3.3.2	Ionization Energy	20

Chapter 1

Introduction

1.1 Types of Chemistry

Inorganic Chemistry relates to the chemistry of molecules that do not contain a carbon atom

Organic Chemistry the chemistry of molecules which do contain a carbon atom. These molecules may also contain a hydrogen atom, and always contain carbon bonds.

Biochemistry the chemistry of proteins and large bio-molecules

Analytical evolving methods to identify and measure things (relating to chemistry)

Theoretical considered the *opposite* of analytical chemistry. Theoretical chemistry is focused on trying to predict things about chemistry using calculus.

Physical uses the theory developed from the above discipline to compare it against what *actually happened*

1.2 Types of Matter

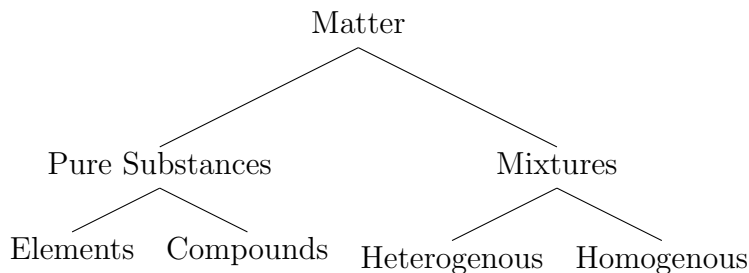


Figure 1.1: The types of matter.

For each type of matter, a definition and several real-world examples are found below.

Elements matter that cannot be broken down into more pure substances *i.e., they have but only one unique type of atom*

- Carbon
- Hydrogen
- Oxygen

Compounds pure substances made of several elements of which are *chemically* joined together

- H₂O
- CO₂

Mixtures impure combinations of two or more substances and are *not* chemically joined

Homogenous uniform content throughout

- Alloys of metal
- Filtered air

Heterogenous non-uniform content throughout

- Non-pasturized apple juice

1.3 Separation of a Mixture

Filtration used on a heterogenous mixture by means of a porous barrier. Exploits the physical size difference *ex., using a coffee filter to make coffee.*

Distillation used on homogenous mixtures and takes advantage of temperature differences. *ex., alcohol boils at temperature A, water boils at temperature B. A temperature between the two is picked to boil off one compound, and then condense it later.*

Chromatography used on both types of mixtures to determine properties of compounds to separate. In practice, a mixture is placed on the end of a piece of paper, and water draws out all the compounds contained within it.

1.4 General Properties of Substances

Intensive properties of substances which are not related to the amount of that substance which is present

- Boiling point
- Specific heat

- Density

Extensive properties of substances which are functions of the amount of that substance which is present

- Mass
- Volume
- Quantity of atoms

1.4.1 Chemical vs. Physical

Chemical observed when something changes into a new substance

- Bond
- Toxicity
- Reactivity
- Flamability

Physical observed without changing chemical properties

- Color
- Density
- Temperature

1.5 SI Units

The SI-system is the unit of measurements most widely used by scientists. It defines a prefix-base system which is useful in representing both large and small numbers.

1.5.1 Base Units

Length meter *m*

Mass kilogram *kg*

Time second *sec*

Temperature kelvin *k*

Amount of Substance mole *mol*

1.5.2 Derived Units

Name	Symbol
Area	m ²
Volume	m ³
Density	kg/m ³
Molar Mass	kg/mol
Concentration	mol/L
Molar Volume	L/mol
Velocity	m/s
Force	N

1.5.3 Prefix System

	Symbol	Modifier
mega	<i>M</i>	10e6
kilo	<i>K</i>	10e3
centi	<i>c</i>	10e−2
milli	<i>m</i>	10e−3
micro	<i>μ</i>	10e−6
nano	<i>n</i>	10e−9

1.5.4 Significant Figures

Significant Figures define a set of rules for maintaining precision while doing math. This system ensures that the answer to a problem set accurately reflects the degree of precision of all sub-compounds.

1. Zeros between non-zero digits are considered significant
2. Zeros in front of a number are not significant
3. Zeros to the end of a number and to the right of a decimal place are considered significant
4. Zeros to the right of non-zeros *may* be considered significant. 1000 has one digit of significance, while 1000. has four. Similarly, 1.00e3 has three digits of significance.

1.5.5 Rules for Arithmetic

When using digits of significance in arithmetic problems, it is important to maintain a correct number of significant digits in the answer. When adding and subtracting numbers, the result has the minimum of the digits of significance between the two products. The same rules apply to both multiplication and division.

Chapter 2

Atoms and Elements

2.1 History

Democritus thought that all things are made of indivisible bits

Aristotle disagreed with the idea of the atom and muted discussion

2.2 Basic Laws

The Law of Conservation of Mass matter is neither created nor destroyed during ordinary chemical and physical changes.

The Law of Definite Proportions chemical compounds have the same elements in the same proportions by mass regardless of the size of the sample

The Law of Multiple Proportions when chemical elements combine they will do so in small, whole-number ratios. If two elements can form multiple compounds, the ratio of the second element to a fixed mass of the 1st element will always be a small, fixed number.

2.3 Dalton's Atomic Theory

1. All matter is composed of atoms
2. Atoms of an element are identical, atoms of different elements are dissimilar in proportions
3. Atoms cannot be subdivided, created, or changed
4. Atoms of different elements combine in simple whole number ratios to form chemical compounds

5. In chemical reactions, atoms are combined, separated or rearranged

2.4 Modern Atomic Theory

2.4.1 Universal Truths

- Conservation of mass is obeyed during chemical reactions
- All matter is composed solely of atoms
- Atoms of any one element differ in proportions from others

2.4.2 Variations on Dalton's Theory

- Atoms of a type can have different masses (isotopes)
- Atoms have an internal structure composed of protons, neutrons, and electrons
- Atoms can be created and destroyed, but only during nuclear reactions

2.5 Modern Structure of an Atom

An atom is the smallest unit of an element which still retains the properties of that element.

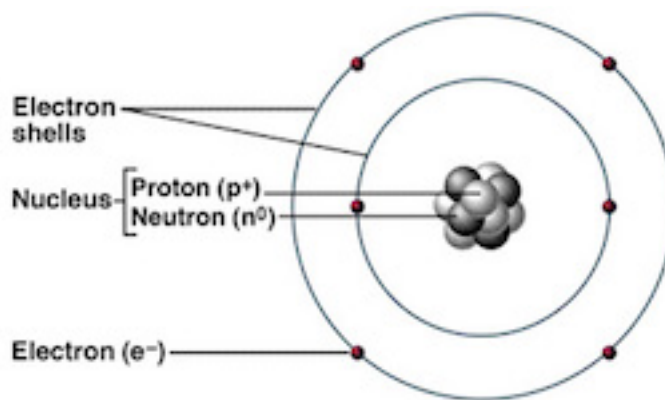


Figure 2.1: Modern-day structure of an atom (basic)

There are three components that make up the important parts of an atom. They are described below:

Nucleus The center of the atom. Very small, and very dense. Contains protons and neutrons.

Protons positively charged sub-atomic particles

Neutrons neutrally charged sub-atomic particles

Electrons Small, negatively-charged particles. They orbit the nucleus in *electron clouds*. That is to say that we do not know the definite position of an electron, but we can reason about where it is likely to be.

2.6 Experiments

2.6.1 Cathode Ray Tube

Experiment performed by J.J. Thompson. Contributed first to the discovery of a sub-atomic particle (the electron). Showed that a cathode-ray had negatively charged particles because it responded to magnetic fields.

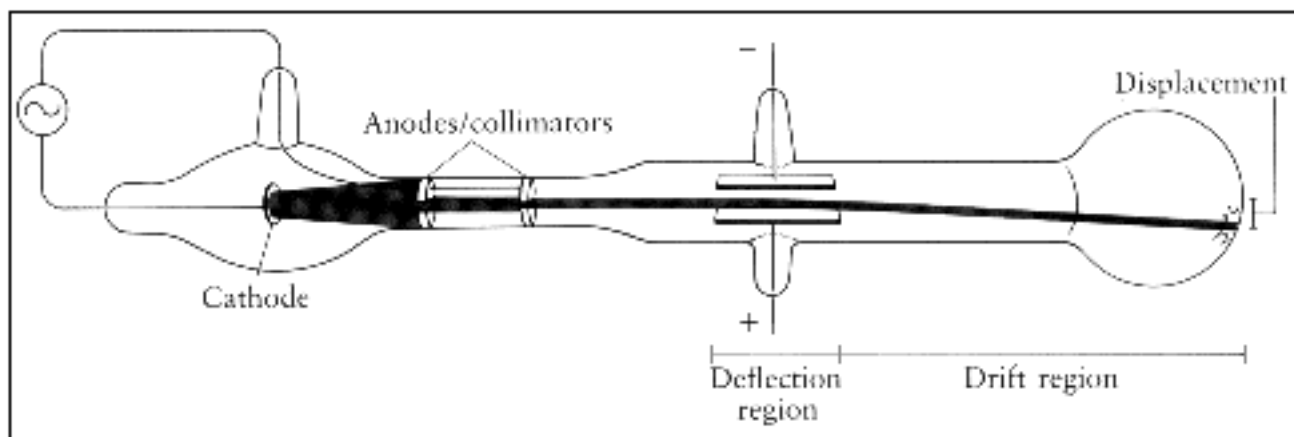


Figure 2.2: Diagram of the cathode-ray tube experiment.

2.6.2 Milikan Oil Drop

Determined the more specific charge of the electron. Showed that small particles were able to be suspended when they picked up negative charge. Using the charge of the plates involved, the charge was able to be determined.

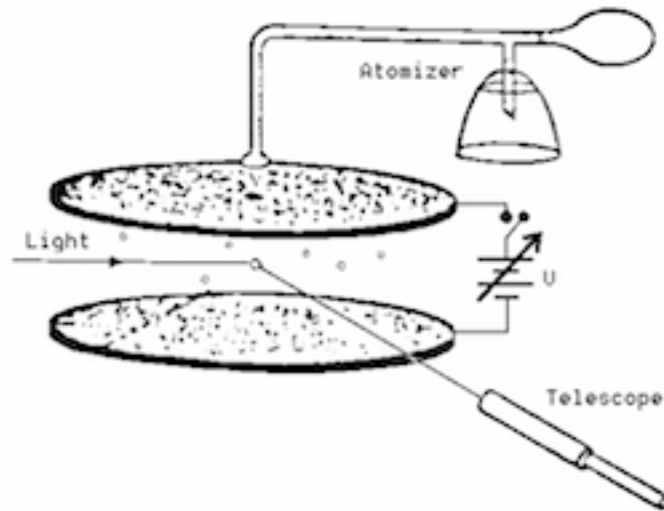


Figure 2.3: Diagram of the oil-drop experiment.

2.6.3 Rutherford's Gold Foil

Experiment performed by Rutherford and Geiger. Together, the two of them shot *alpha* particles at a ultra-thin sheet of gold foil. These particles exhibited one of three behaviors, each indicating a different outcome.

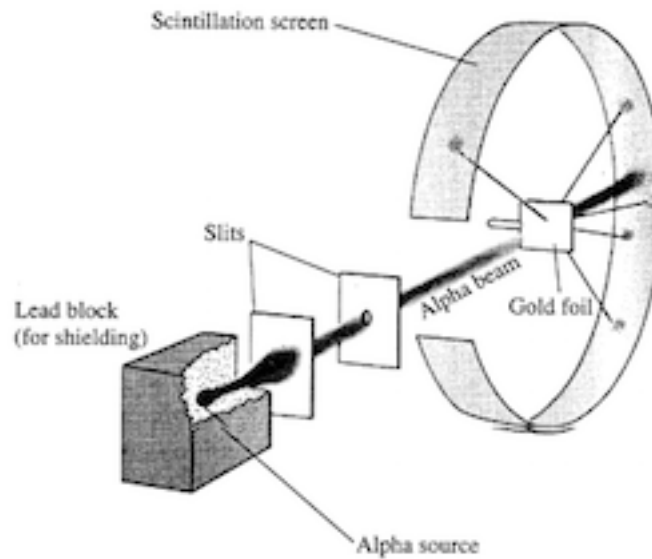


Figure 2.4: Diagram of Rutherford's Gold Foil experiment.

Through the foil the alpha particles went through empty space within an atom, and thus traveled uninterrupted through the atom in a straight path

Deflected through the foil the alpha particles were attracted to something while traveling through an empty space within the atom. As such, their path was changed slightly, and they angled as they travelled.

Bounced off the foil the alpha particles collided with the positively charged nucleus to deflect off of the foil, back at the particle gun.

2.7 Elements

Elements are defined by the number of protons in the nucleus of one of their atoms. No two elements will ever have the same number of protons in their nucleus under normal circumstances. *As such, the number of protons in an element can be considered as its unique key.*

The atomic number of an element has the same value as the number of protons in its nucleus.

2.7.1 Isotopes

Isotopes have the same number of protons, but a different number of neutrons. Isotopes are represented as the following $^{35}_{17}\text{Cl}$. This is Chlorine-35. The superscript represents the atomic mass of an isotope, and the subscript represents the atomic number.

Since the atomic mass is the sum of the number of protons and the number of neutrons in an atom, the number of neutrons in an isotope can easily be determined.

2.8 Nomenclature

2.8.1 Common Nomenclature

Molecular Compounds compounds which contain only non-metal components

Binary Compounds compounds of two different elements. In non-ionic binary compounds, simply change the last component to have an *-ide* suffix, for example CaCl_2 - *Calcium Chloride*. Naming occurs in the order *mono-*, *di-*, *tri-*, *tetra-*, *penta-*, *hexa-*, *hepta-*, *octa-*, *nona-*, *deca-*.

Binary Ionic Compounds binary-compounds (see above) which contain both a metal and a non-metal.

1. Positive charge first
2. The positive element undergoes no change in the name of its element
3. Negative charge gets the *-ide* treatment

2.8.2 Stock-system Nomenclature

When dealing with ions, it is frequently the case where there are more than one charge state available to work with. The stock-system nomenclature makes it obvious which one we are dealing with.

Simply put, you take the charge of an ion, and place it in parenthesis in roman-numerals to the right of the chemical symbol. For example, Fe^{3+} becomes Fe(III).

2.9 Types of Ions

Ions take shape in several different types, but come in only two charges. These charges are named *cation* and *anion*. A cation is a positively charged ion, whereas an anion is a negatively charged ion. (*Remember: a-negative-ion... an-ion.*)

2.9.1 Monatomic Ions

Monatomic ions are formed of one ore more of the same kind of element. These ions gain and loose electrons to obtain a full valence-level of electrons for wahtever level they might be at. Examples include H^+ and O^{2-} .

2.9.2 Polyatomic Ions

Polyatomic ions are charged groups of covalently bonded atoms. They carry charge as a group, such that they behave like single ions, even though they are composed of more than one atom.

All relevant polyatomic ions are listed below, grouped by charge.

- Cation¹⁺

Ammonium NH_4^+

Hydronium H_3O^+

- Anion¹⁻

Cyanide CN^-

Bicarbonate HCO_3^-

Hydroxide OH^-

Nitrate NO_3^-

Nitrite NO_2^-

Permanganate MnO_4^-

Acetate CH_3COO^- or $\text{C}_2\text{H}_3\text{O}_2^-$

Dihydrogen Phosphate H_2PO_4^-

- Anion²⁻

Carbonate CO_3^{2-}

Sulfate SO_4^{2-}

Sulfite SO_3^{2-}

- Anion³⁻

Phosphate PO_4^{3-}

Chapter 3

Atomic Structure and Periodic Trends

3.1 Electron-Configuration Notation

A system of numbers and letters is used to designate electron configuration. For example: $1s^2 2s^2 2p^6 3s^2 3p^2$.

1. the level number is used
2. the level designation for the shape of the orbital
3. a superscript is used to represent the number of electrons in that specific orbital.

3.1.1 Noble-gas Notation

To simplify writing out all of these numbers, you can include the noble-gas most-closely following behind of the element you are trying to describe, and then only describe the differing electrons using the notation above. For example: $\text{Ne} 3s^2 3p^2$.

Highest occupied level the electron containing main energy level with the highest quantum number

Inner-shell electrons the electrons that are contained in levels with a lower quantum state

Noble-gas configuration an outer main-energy level that is fully occupied, usually by eight electrons.

3.2 Periodic Table and Trends

3.2.1 The Periodic Table

Mendeleev's Table

In 1860, the first International Conference of Chemists was held and at that time, the first periodic table was discussed. In 1869, 9 years later, that table was published in a textbook

for college students.

Mendeleev was the first to generate a periodic table. He used atomic mass as the basis for his table, arranging elements by increasing mass. Some elements, however, were out of place when compared against the periodic tables of today. He was able to, despite these faults, see periodic trends, which allowed him to predict with a high-degree of accuracy, the qualities of yet undiscovered elements.

Improvements

Moseley observed that elements were better fit as ordered by their nuclear charge, not atomic mass. The nuclear charge is due to the number of protons, which led to the usage of atomic number for description.

Periodic Law states that the physical and chemical properties of the elements are periodic functions of their atomic numbers.

3.2.2 Modern-day

Today, the *Periodic Table* is just an arrangement of elements in the order of their atomic numbers, such that elements of similar properties fall in the family column or group. However, there are but a few differences from Mendeleev's original table:

1. There are more elements now as compared to the first introduction of the periodic table
2. The discovery of Noble Gases and the synthesis of the Lanthanides and Actinides
3. The general arrangement has changed over time as well

3.3 Periodic Trends

3.3.1 Atomic Radii

Atomic radius is defined as one half of the distance between the nuclei of identical atoms that are bonded together. Atomic radii decreases as we move across a period (increase in positive charge) and increases as we descend a family (increasing the main energy level).

3.3.2 Ionization Energy

Ionization energy is the energy required to remove one electron from a neutral atom of an element. It is defined by the following chemical equation: $A + \text{energy} \longrightarrow A^+ + e^-$. Ionization energy increases as you move right across a period, and decreases as you move down a family.

Multiple Ionization Energies

There are multiple ionization energies, each which involves removing another electron from the atom. These are referred to as IE_2 , IE_3 , and so on.