

Section 163 Colligative Properties Of Solutions

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Section Resources Connecting to Your World Section 16.3 Colligative Properties of Solutions 487 a b Pure solvent Higher vapor pressure Solution containing nonvolatile solute Lower vapor pressure Solvent particle Solute particle 16.3 Colligative Properties of Solutions The wood frog is a remarkable creature because it can survive being frozen.

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SECTION 16.3 COLLIGATIVE PROPERTIES OF SOLUTIONS (pages 487–490) This section explains why a solution has a lower vapor pressure, an elevated boiling point, and a depressed freezing point compared with the pure solvent of that solution. Vapor Pressure Lowering (pages 487–488) 1.

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Chapter 16: Colligative Properties of Solutions 45 16-4. The mole fraction of $(\text{NH}_4)_2\text{SO}_4(\text{aq})$ is given by $x_{(\text{NH}_4)_2\text{SO}_4} = \frac{n_{(\text{NH}_4)_2\text{SO}_4}}{n_{(\text{NH}_4)_2\text{SO}_4} + n_{\text{H}_2\text{O}}}$ Because $(\text{NH}_4)_2\text{SO}_4(\text{aq})$ is a strong electrolyte, it dissociates completely into $\text{NH}_4^+(\text{aq})$ and $\text{SO}_4^{2-}(\text{aq})$ ions. Assume a one kilogram solution. The number of moles of ions in one ...

CHAPTER 16. Colligative Properties of Solutions

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Section Resources Connecting to Your World Section 16.4 Calculations Involving Colligative Properties 491 16.4 Calculations Involving Colligative Properties Cooking instructions for a wide variety of foods, from dried pasta to packaged beans to frozen fruits to fresh vegetables, often call for the addition of a small amount of salt to the

16.4 Calculations Involving Colligative Properties 16

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