

## ***Solution Stoichiometry Name Chem Worksheet 15 6***

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The strategy used for solving these solution stoichiometry problems is to set up the problem so that the units cancel. When the volume of a solution is multiplied by the molarity of a solution the resulting units are moles. A balanced equation allows us to convert from moles of a known substance to moles of an unknown.

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Solution Stoichiometry . Name\_\_\_\_\_ CHEMISTRY 110 . last first . 1] How many grams of calcium phosphate can be produced from the reaction of 2.50 L of 0.250 M Calcium chloride with an excess of phosphoric acid?

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### **Solution Chemistry Worksheets - Teacher Worksheets**

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Stoichiometry Worksheet #1 Answers 1. Given the following equation:  $2 \text{C}_4\text{H}_{10} + 13 \text{O}_2 \rightarrow 8 \text{CO}_2 + 10 \text{H}_2\text{O}$ , show what the following molar ratios should be. a.  $\text{C}_4\text{H}_{10} / \text{O}_2$  b.  $\text{O}_2 / \text{CO}_2$  c.  $\text{O}_2 / \text{H}_2\text{O}$  d.  $\text{C}_4\text{H}_{10} / \text{CO}_2$  e.  $\text{C}_4\text{H}_{10} / \text{H}_2\text{O}$  2. Given the following equation:  $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$  a. How many moles of O

### **Stoichiometry Worksheet #1 Answers**

Stoichiometry Problems Name \_\_\_\_\_ Chem Worksheet 12-2 Stoichiometry Strategy Amount moles molar mass (g/mol) 22.4 L/mol KNOWN UNKNOWN Mass grams V olume f gas at STP iters Particles atoms, molecules, formula units  $6.02 \times 10^{23}$  particles/mol Mass grams at STP liters Particles atoms, molecules, formula units Amount B moles molar mass (g/mol)

## **Solution Stoichiometry Name Chem Worksheet 15 6**

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