

## *Sodium Thiosulfate Supersaturated Solution Lab*

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### **Sodium Thiosulfate Supersaturated Solution Lab**

If a hot saturated solution is prepared and allowed to cool without any solute undissolving, the solution is supersaturated. Most solutes will undissolve out of a solution as the temperature is lowered, but a few substances remain in solution during the cooling process. One such solid is sodium thiosulfate.

### **Name Date Hour - Chippewa Falls Area Unified School District**

A supersaturated solution of sodium acetate trihydrate (the clear solution) will recrystallize by the addition of a single crystal of the solid. A supersaturated solution is a solution which contains a greater amount of dissolved substance than is present in a saturated solution at the same temperature.

### **Supersaturated Sodium Acetate Solution SCIENTIFIC**

View Lab Report - lab\_supersaturation\_of\_sodium\_thiosulfate from CHEM 1312 at Texas A&M University, Corpus Christi. PAP Chemistry Name \_ Partner\_ Period: \_ 18 / 20 Solutions Solutions & Solubility

### **lab\_supersaturation\_of\_sodium\_thiosulfate - PAP Chemistry Name Partner Period 18 20 Solutions Solutions Solubility Lab Projected Time 25 minutes - coursehero.com**

Solutions and Solubility. Rinse the test tube several times with tap water and then two or three times with distilled water. Fill the test tube about one-half full of solid sodium acetate and then add about 2 mL of distilled water (no more). Warm the tube as much as necessary to dissolve all of the solid.

### **Solutions and Solubility - MiraCosta College**

crystal of sodium thiosulfate into the test tube and record all of your observations. Hold the test tube in your hand. 6. Return your test tube with the sodium thiosulfate to the front of the room to be used by the next class. Do not pour out any of the water or crystals. Supersaturated Solution Investigation 2 supersaturated\_solutions.odt

### **Chemistry II Investigation: "Supersaturated Solutions"**

Crystallization of supersaturated sodium thiosulfate solution is initiated by the addition of the seed crystal.

### **Crystallization of supersaturated sodium thiosulfate**

The sodium thiosulfate pentahydrate is then heated in a test tube and forms a supersaturated solution. The solution is then cooled in an ice bath.

### **Supersaturated Solution Demo**

Also, the reason I'm asking is we're doing a lab with sodium thiosulfate. We heat it up, adding all but one crystal, and then put it in a ice bath and add the last crystal and observe. Is it supersaturated when it is heated? and then the one last crystal causes it to fall out of solution and crystallize?

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