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A first-level text stressing chemistry of natural and polluted water and its application to waste-water treatment. Discusses principles of chemical kinetics, dilute solution equilibria, effects of temperature and ionic strength, and thermodynamics in relation to water chemistry. Strong emphasis given to graphical procedures.

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Reference Texts: Water Chemistry, Brezonik and Arnold (Oxford, 2011) Aquatic Chemistry, Stumm and Morgan (Wiley, 1996) Water Chemistry, Snoeyink and Jenkins (Wiley, 1980) Aqueous Environmental Geochemistry, Langmuir (Prentice-Hall, 1997) Principles & Applications of Aquatic Chemistry, Morel and Hering (Wiley, 1993)

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3. Ability to solve problems in groups and present the solutions orally 4. Ability to use current software for solving water chemistry problems 5. Ability to understand and communicate new developments in water chemistry . Outcome Measures 1. Weekly homework problems on principles

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in fresh water, $I=0.002$, $\gamma_{O_2} = 1.0006$ Molar vs. Molal in principle, activity predictions are based on molal concentrations (mole/kg solvent), but since we are often most concerned with dilute solutions, we frequently use molar concentrations. David Reckhow CEE 680 #4 24

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