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Modern Chemistry • CHAPTER 7 HOMEWORK 7-1 ... Answer the following questions in the space provided. 1. ... MODERN CHEMISTRY SECTION 7-1 REVIEW. 53.

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The Chemical Formulas and Chemical Compounds chapter of this Holt McDougal Modern Chemistry Companion Course helps students learn the essential lessons associated with chemical compounds and ...

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Modern Chemistry 92 Chapter Test Name Class Date Chapter Test B, continued 31. Calculate the energy released by freezing 13.8 g of a liquid. Its molar mass is 82.9 g/mol, and its molar enthalpy of fusion is 4.60 kJ/mol. PART VI Define the term equilibrium vapor pressure. Use it to explain the following two phenomena.

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Modern Chemistry 71 Chapter Test Name Class Date Chapter Test B, continued 17. A substance combines with oxygen, releasing a large amount of energy as heat and light, in a(n) . 18. The decomposition of a substance by an electric current is called . 19. A(n) orders the elements by the ease with which they undergo certain chemical reactions. 20.

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Holt McDougal Modern Chemistry 1 Chapter Test Assessment Chapter Test A Chapter: The Periodic Law Use the periodic table below to answer the questions in this Chapter Test. In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. ____ 1.

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Holt McDougal Modern Chemistry 3 Chapter Test Chapter Test B, continued 16. The measure of the ability of an atom in a chemical compound to attract electrons from another atom in the compound is called _____. 17. The energy required to remove one electron from an atom is called its _____. 18.

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Modern Chemistry 47 Chapter Test Name _____ Class _____ Date _____ Chapter Test A, continued _____ 7. The melting points of ionic compounds are higher than the melting points of molecular compounds because a. ionic substances tend to vaporize at room temperature. b. ionic substances are brittle. c. attractive forces between ions are greater than the attractive ...

Assessment Chapter Test A - Kettering City School District

SECTION 1 continued 6. Some binary compounds are ionic, others are covalent. The type of bond favored partially depends on the position of the elements in the periodic table.

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