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Modern Chemistry 91 Chapter Test Name Class Date Chapter Test B, continued 27. Distinguish between ionic crystals and metallic crystals. PART V Write the answers to the following questions on the line to the left, and show your work in the space provided. The molar enthalpy of fusion for water is 6.008 kJ/mol. 28.

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Modern Chemistry 35 Chapter Test Name Class Date Chapter Test B, continued Write the orbital notation for the following elements in the space provided. 37. lithium, atomic number 3 38. carbon, atomic number 6 39. neon, atomic number 10 PART VI Write the answers to the questions on the line to the left, and show your work in the space provided.

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Modern Chemistry 84 Chapter Test Chapter: States of Matter In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. ____ 1. The kinetic-molecular theory of gases assumes that the particles of an ideal gas are separated by

great distances. This implies that the gas

Assessment Chapter Test A - Kettering City School District

SHORT ANSWER Answer the following questions in the space provided. 1. Use the concept of potential energy to describe how a covalent bond forms between two atoms. As the atoms involved in the formation of a covalent bond approach each other, the electron-proton attraction is stronger than the electron-electron and proton-proton repulsions.

6 Chemical Bonding - Effingham County Schools / Overview

CHAPTER 10 REVIEW States of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Match description on the right to the correct crystal type on the left. b ionic crystal (a) has mobile electrons in the crystal c covalent molecular crystal (b) is hard, brittle, and nonconducting a metallic crystal (c) typically has the lowest melting point of the four

10 States of Matter - Ms. Agostine's Chemistry Page

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Modern Chemistry 143 Chapter Test Name Class Date Chapter Test B, continued PART V Write the answers to the following problems on the line to the left, and show your work in the space provided.
____ 39. What is the hydronium ion concentration of an aqueous solution that has a pH of 5.0?
____ 40.

Assessment Chapter Test B - Mr. Marshall's Classroom

SHORT ANSWER Answer the following questions in the space provided. 1. b A balanced chemical equation represents all the following except (a) experimentally established facts.

8 Chemical Equations and Reactions - David Brearley High ...

MODERN CHEMISTRY HR W material copyrighted under notice in this work. Name CHAPTER 3 TEST continued Date Class SHORT ANSWER Write the answers to the following questions in the space provided. 22. Give three of the main concepts in Dalton's atomic theory. 23. What is molar mass? How is it related to atomic mass?

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HONORS CHEMISTRY -- CHAPTER 8-9 TEST - MULTIPLE CHOICE NAME: ____ 1. Which of the following has bonds which are the least ionic? a) CCl₄ b) SiO₂ c) KCl d) NH₃ e) H₂ 2. What type of bonding occurs between phosphorus and hydrogen within a molecule of PH₃?

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