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Modern Chemistry • CHAPTER 7 HOMEWORK 7-3 (pp. 211–213) VOCABULARY Write the prefix that corresponds to each number. Try not to use the book. Number Prefix Number Prefix 16 27 38 49 510 GRAPHIC ORGANIZER Complete the chart by supplying the missing information. Name of Compound Formula of Compound CO 2 Nitrogen dioxide S 2 Cl 2 Carbon ...

HOMEWORK 7-3 - Kettering City School District

3(s) is strongly heated, it produces solid MgO as gaseous CO 2 is driven off. a. What is the percentage loss in mass as this reaction occurs? b. Assign the oxidation number to each atom in MgCO 3. c. Does the oxidation number of carbon change upon the formation of CO 2? 64 CHEMICAL FORMULAS AND CHEMICAL COMPOUNDS MODERN CHEMISTRY

7 Chemical Formulas and Chemical Compounds

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CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Label each of the following statements as True or False: True a. If the formula mass of one molecule is x amu, the molar mass is x g/mol. False b. Samples of equal numbers of moles of two different chemicals

7 Chemical Formulas and Chemical Compounds

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Modern Chemistry • CHAPTER 7 HOMEWORK 7-5 VOCABULARY Complete each sentence. 1.
Oxidation numbers are assigned to the atoms composing a compound or ion in order to
2. The atoms in a pure element have an oxidation number of 3. The more electronegative
element in a binary molecular compound is assigned the number that

HOMEWORK 7-5 - apcscience.com

see figure 3-3 page 67. C: Modern Atomic Theory. Some of Dalton's postulates have been proven to be incorrect. His third postulate indicated that atoms cannot be subdivided. His second postulate indicates atoms of the same element have identical mass. Isotopes. His theory has been modified -not discarded. Homework: Section 3-1

Chapter 3 Modern Chemistry - Prepchem

Modern Chemistry 57 Chemical Bonding CHAPTER 7 REVIEW Chemical Formulas and Chemical
Compounds SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1
Label each of the following statements as True or False: a. If the formula mass of one
molecule is x u, the molar mass is x g/mol. b.

CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds

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SECTION SUMMARY. 14.3 Ideal Gases Summary: The ideal gas law permits you to solve for the number of moles of a contained gas when the pressure, volume, and ... Chapter 7 Acids, Bases, and Solutions Section 3 Summary ...

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CHAPTER 1 REVIEW Matter and Change MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Classify each of the following as a homogeneous or heterogeneous substance. a. sugar d. plastic wrap b. iron filings e. cement sidewalk c. granola bar 2. For each type of investigation, select the most appropriate branch of chemistry from the following

mc06se cFMsr i-vi

CHAPTER 10 REVIEW States of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Match description on the right to the correct crystal type on the left. b ionic crystal (a) has mobile electrons in the crystal c covalent molecular crystal (b) is hard, brittle, and nonconducting a metallic crystal (c) typically has the lowest melting point of the four

10 States of Matter - Ms. Agostine's Chemistry Page

Modern Chemistry 88 Chapter Test Chapter: States of Matter PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. 1. According to the kinetic-molecular theory, particles of an ideal gas

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