

## *Net Ionic Equation Yahoo Answers*

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### Net Ionic Equation Yahoo Answers

What's the net ionic equation for these 3 questions? How do you know what to write? Thank you. 1. Sodium hydroxide and iron (III) chloride react to form sodium chloride and iron (III) hydroxide. 2. Potassium sulfide and copper (II) nitrate react to form potassium nitrate and copper (II) sulfide. 3. Ammonium sulfate and barium chloride react to form ammonium chloride and barium sulfate.

### Net ionic equations help? | Yahoo Answers

what is the net ionic equation including phases for  $\text{CoPO}_4(\text{s})$  dissolving in  $\text{H}_3\text{O}^+(\text{aq})$ ? help please!

### Net Ionic Equation Help? | Yahoo Answers

Which of the following statements are true for the reaction  $\text{Ca}(\text{OH})_2(\text{aq}) + \text{HCl}(\text{aq}) \rightarrow$  a)  $\text{CaCl}(\text{aq})$  is a product b) When balanced the coefficient of  $\text{HCl}(\text{aq})$  is 3 c) Ions cannot appear in the net ionic equation because they always cancel out on both sides of the equation d)  $\text{Ca}(\text{OH})_2(\text{aq})$  appears in the net ionic equation e)  $\text{H}_2\text{O}(\text{l})$  appears in the net ionic equation

### Net Ionic equation? | Yahoo Answers

Yahoo UK & Ireland Answers ... The net ionic equation will be comprised of species that are different on both sides of the arrow. Net ionic equation:  $\text{Fe}^{3+}(\text{aq}) + 3 \text{OH}^-(\text{aq}) \rightarrow \text{Fe}(\text{OH})_3(\text{s})$  Notice that  $\text{FeOH}_3(\text{s})$  is incorrect. You need to include the parentheses for multinuclear ions like  $\text{OH}^-$ .

### Ionic and net ionic equations? | Yahoo Answers

Write balanced complete ionic equation & net ionic equation for  $\text{HC}_2\text{H}_3\text{O}_2(\text{aq}) + \text{K}_2\text{CO}_3(\text{aq}) \Rightarrow \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g}) + \text{KC}_2\text{H}_3\text{O}_2(\text{aq})$  please help me in writing a balanced ionic equation and net ionic equation, identify all of the phases.

### Ionic and Net equations? | Yahoo Answers

Write the balanced net ionic equation for each of these precipitation reactions. Enter the cation in the first box and the anion in the second box, with coefficients (if needed, leave 1s off as they are implied). Put charges in parentheses (charge multiple, if any, followed by sign) and phase in parentheses, e.g.,  $\text{Al}^{3+}(\text{aq})$  should be entered as  $\text{Al}(3+)(\text{aq})$ .

### Write the balanced net ionic equation? | Yahoo Answers

Yahoo New Zealand Answers ... Complete the reactions and write the net ionic equation. identify the spectator ion or ions in each reaction:  $\text{CuBr}_2(\text{aq}) + \text{NaOH}(\text{aq}) \rightarrow$  Follow . 2 answers 2. Report Abuse. Are you sure that you want to delete this answer? Yes No.

### complete the reactions and write the net ionic equation ...

Best Answer: the possible products are nickel nitrate and magnesium bromide and neither of these will precipitate, neither is a gas and neither is a molecule so I'm going with "No Reaction"

### Net ionic equations HELP! Chemistry!? | Yahoo Answers

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### Net Ionic Equation..? | Yahoo Answers

Use the solubility rules to write net ionic equations for the reactions that occur when solutions of the following are mixed: a)  $(\text{NH}_4)_2\text{CO}_3 + \text{Ca}(\text{ClO}_4)_2$  b)  $\text{FeSO}_4 + \text{K}_3\text{PO}_4$  c)  $\text{FeCl}_2 + \text{AgNO}_3$  Cation + Anion  $\rightarrow$  precipitate format (and if there is not a reaction just put none in the precipitate spot.) Thanks!

### Help with Net Ionic equations ... - au.answers.yahoo.com

Net Ionic Equation Worksheet Answers Write balanced molecular, ionic, and net ionic equations (NIE) for each of the following reactions. Assume all reactions occur in aqueous solution. 1.  $2\text{NaCl}(\text{aq}) + \text{Pb}(\text{NO}_3)_2(\text{aq}) \rightarrow \text{PbCl}_2(\text{s}) + 2\text{NaNO}_3(\text{aq})$

### Net Ionic Equation Worksheet Answers

What is the net ionic equation of a dissolution of salt in water? Basically we added salt in water and we want to know the net ionic equation for this.

### Chemistry Problem: What is the net ionic equation of a ...

Write a balanced molecular equation for the reaction that occurs when aqueous solutions of potassium hydroxide, KOH, and nitric acid, HNO<sub>3</sub>, are combined. ----- Write a net ionic equation for the reaction that occurs when aqueous solutions of nitric acid, HNO<sub>3</sub>, and potassium hydroxide, KOH, are combined. Represent the hydronium ion as H<sub>3</sub>O<sup>+</sup>.

### Questions | Yahoo Answers

Answers.com is the place to go to get the answers you need and to ask the questions you want ... A Net Ionic Equation is a chemical equation for a reaction which lists only those species ...

### How do you write a net ionic equation - answers.com

PRACTICE PROBLEMS ON NET IONIC EQUATIONS page 1 of 3 Show the complete ionic and net ionic forms of the following equations. If all species are spectator ions, please indicate that no reaction takes place. Note: you need to make sure the original equation is balanced before proceeding!

### PRACTICE PROBLEMS ON NET IONIC EQUATIONS

It says state if there is a reaction, net ionic equation, total ionic equation and write a balanced chemical equation. Iron (II) Sulphate + Potassium Hydroxide ----> The products should be Iron Hydroxide and Potassium Sulphate, but I don't get how to the net ionic equation because when I do it everything just crosses out

### Net ionic equations help? | Yahoo Answers

Yahoo India Answers ... Answers. Best Answer: Net ionic equations leave out spectator ions (aqueous ions that are unchanged during the reaction). Here the NaOH is a provider of hydroxide ions - the sodium ions are spectator ions - and in the first three cases the metal hydroxide is the product. In the fourth, two compounds are formed, gaseous ...

### How do i write the net ionic ... - in.answers.yahoo.com

Yahoo India Answers ... write the net ionic equation for the reaction of an aqueous solution of potassium hydroxide reacting with aqueous solution of nickel(2) sulfate? I got NiSO<sub>4</sub> + K --- Ni + K<sub>2</sub>SO<sub>4</sub> but I don't think it's right? Follow . 2 answers 2.

### Write the net ionic equation for ... - in.answers.yahoo.com

3. net reaction OH<sup>-</sup> plus H<sup>+</sup> = HOH or H<sub>2</sub>O water does not precipitate but does actually form from its ions producing a non ionized product. 4. Zn<sup>++</sup>(aq) plus S<sup>=</sup>(aq) = ZnS(s) sodium bromide is soluble so its ions remain separate and are left out of the net ionic equation 5.

### Net Ionic Equations.....? | Yahoo Answers

Best Answer: The presence (or lack) of a driving force for reactions done in water can be evaluated using a net ionic equation. The total and net ionic equations show the real physical state of every component of the reaction. The compound is a strong electrolyte (a soluble ionic compound or a strong acid), then the compound is separated into the appropriate aqueous ions.

## Net Ionic Equation Yahoo Answers

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