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Let's try a real chemistry problem and use this concept to determine the limiting reagent and see how it determines how much product will be formed. Problem : Nitrogen gas reacts with hydrogen gas to produce gaseous ammonia ( $\text{NH}_3$ ). If we react 25.0 g of nitrogen with 25.0 g of hydrogen, how much ammonia will be produced.

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#### **mc06se cFMsr i-vi**

Modern Chemistry 1 Solutions CHAPTER 12 REVIEW Solutions Teacher Notes and Answers Chapter 12 SECTION 1 SHORT ANSWER 1. c 2. a 3. b 2. a. alcohol b. water c. the gels 3. The mixture is a colloid. The properties are consistent with those reported in Table 3 on page 404 of the text. The particle size is small, but not too small, and the mixture

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