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3. What is the molarity of a solution of HNO<sub>3</sub> that contains 12.6 grams HNO<sub>3</sub> in 1.0 L of solution? ?  
 $\text{mol HNO}_3 = \frac{12.6 \text{ g HNO}_3}{63.0 \text{ g HNO}_3} \times 1 \text{ mol HNO}_3 = 0.200 \text{ mol HNO}_3$   
 $M = \frac{0.200 \text{ mol HNO}_3}{1.0 \text{ L}} = 0.200 \text{ M}$   
4. How many grams of potassium nitrate are required to prepare 0.250 L of a 0.700 M solution?  $0.700 \text{ M} = \frac{\text{moles of solute}}{0.250 \text{ L}}$  moles of ...

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