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A. 25.00 mL aliquot of concentrated hydrochloric acid (11.7 M) is added to 175.00 mL of 3.25 M hydrochloric acid. a. is this a dilution? b. determine the number of moles of hydrochloric acid from the 175.00 mL of 3.25 M hydrochloric acid. c. determine the number of moles of hydrochloric acid in the solution after the addition of the concentrated hydrochloric acid. assume the volumes are additive.

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Molarity = \_\_\_\_ Problems: Show all work and circle your final answer. 1. To make a 4.00 M solution, how many moles of solute will be needed if 12.0 liters of solution are required?  $4.00 \text{ M} = \frac{\text{moles of solute}}{12.0 \text{ L}}$  moles of solute = 48.0 mol 2. How many moles of sucrose are dissolved in 250 mL of solution if the solution

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Dilutions Worksheet W 329 Everett Community College Student Support Services Program 1) If 45 mL of water are added to 250 mL of a 0.75 M  $\text{K}_2\text{SO}_4$  solution, what will the molarity of the diluted solution be? 2) If water is added to 175 mL of a 0.45 M KOH solution until the volume is 250 mL, what will the molarity of the diluted solution be?

### **Dilutions Worksheet W 329 - Everett Community College**

is used to make a 2.5 L dilution, what is the molarity of the dilution? 3. If 65.5 mL of HCl stock solution is used to make 450.-mL of a 0.675 M HCl dilution, what is the molarity of the stock solution? 4. How do you prepare 500.-mL of a 1.77 M  $\text{H}_2\text{SO}_4$  dilution from an 18.0 M  $\text{H}_2\text{SO}_4$  stock solution? Answers: 1.

### **Molarity Problems Worksheet - Mrs Getson's Blog**

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p. 31 Molarity By Dilution (mixing a concentrated solvent with water) Acids are usually acquired from chemical supply houses in concentrated form. These acids are diluted to the desired concentration by adding water. Since moles of acid before dilution = moles of acid after dilution, and moles of acid : CV then,  $C_1 \times V_1 = C_2 \times V_2$ .

### **Molarity By Dilution - Ms Beaucage - Homework**

Dilutions Worksheet 1) If I add 25 mL of water to 125 mL of a 0.15 M NaOH solution, what will the molarity of the diluted solution be? 2) If I add water to 100 mL of a 0.15 M NaOH solution until the final volume is 150 mL, what will the molarity of the diluted solution be? 3) How much 0.05 M HCl

solution can be made by diluting 250 mL of 10 M HCl?

**Dilutions Worksheet - Awesome Science Teacher Resources**

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