

Student Exploration Chemical Equations Answer Key

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Student Exploration Chemical Equations Answer

Chemical Equations. Practice balancing chemical equations by changing the coefficients of reactants and products. As the equation is manipulated, the amount of each element is shown as individual atoms, histograms, or numerically. Molar masses of reactants and products can also be calculated and balanced to demonstrate conservation of mass. Use for 5 minutes a day.

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Activity A: Balancing equations Get the Gizmo ready: • Check that the Combination reaction is selected and that all coefficients are set to one. (The coefficients are the numbers in the boxes.) Introduction: The equation $H_2 + O_2 \rightarrow H_2O$ is unbalanced because there are two oxygen atoms on the reactants side of the equation, and only one on the products side of the equation.

Student Exploration: Balancing Chemical Equations

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Introduction: In a chemical reaction, the reactants are the substances that enter into the reaction, and the products are the substances that are made in the reaction. A chemical reaction is balanced if the numbers of reactant atoms match the numbers of product atoms. Goal: Learn to balance any chemical equation. 1.

Student Exploration: Chemical Equations

Balancing Chemical Equations Answer Key Vocabulary: coefficient, combination, compound, decomposition, double replacement, element, molecule, product, reactant, single replacement,

subscript Prior Knowledge Questions (Do these BEFORE using the Gizmo.) [Note: The purpose of these questions is to activate prior knowledge and get students thinking.

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Student Exploration: Balancing Chemical Equations Learning Objectives Students will... Adjust coefficients to balance a chemical equation. Classify chemical equations as combination (synthesis), decomposition, single replacement, or double replacement. Vocabulary: coefficient, combination, compound, decomposition, double replacement, element, molecule, product, reactant, single replacement ...

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