Stoichiometry Using Copper Lab 31 Answers

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In this lab copper was suppose to be recovered to initial values. As the data shows, this did not happen. The final mass of copper (49.5g Cu) ended up significantly more than the original value (1.962 Cu). The final moles of copper (.77 moles Cu) ended up being significantly more than the intital moles of copper (.03 moles Cu).

Stoichiometry Using Copper Lab - AP Chemistry Labs

Lab 31 Stoichiometry Using Copper.pdf Free Download Here MakingCopper Stoichiometry I - Mt. San Antonio College ... Chemistry 30S/PGL Stoichiometry Lab ... copper ion product in the reaction between copper and silver nitrate ... 0.31 mol/L x 0.025 L ... Stoichiometry - Wikispaces

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Stoichiometry Using Copper Lab 31 Answer Key | Higher ...

Stoichiometry Using Copper Lab. October 23, 2012. ... The goal of the lab is to recover the copper precipitated and compare it to the initial copper mass to verify the Law of Conservation of Mass. The final mass of copper is 13.103 grams, clarifying there was an 11.073 grams increase in mass from the initial. ... Stoichiometry was used in this ...

Stoichiometry Using Copper Lab - Stephanie's Wonderful ...

Stoichiometry Using Copper. Purpose: The purpose is to see how the amount of copper (and copper itself) is altered after a series of reactions. ... The main theory used in the lab was the conservation of mass. While copper was subjected to many different types of reactions the mass of copper was projected to stay the same. ... The types of ...

Stoichiometry Using Copper - Alexia's Ap Chemistry Lab ...

General Chemistry I (FC, 09 - 10) Lab #4: Stoichiometry: The Reaction of Iron with Copper(II) Sulfate Revised 8/19/2009 1 Introduction In this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper(II) sulfate. This reaction produces

General Chemistry I (FC, 09 - 10) Lab #4: Stoichiometry ...

The reaction is also successful. At his point the copper has been retrived and at the end all the compounds and elements should be decanted or evaporated and the only thing left should be the solid copper. Stoichiometry problems were used in each reaction in order to balance equations.

Stoichiometry Using Copper Lab - Yamilet's AP Chemistry Labs

there is significant pre-lab preparation that must be done by the instructor. First, I hand out the LAB - Stoichiometry (aluminum to copper). I ask students to skim through the lab and to name the different sections that they will be completing (Procedure, Observations, Data, Analysis, and

Conclusion).

LAB - Stoichiometry (aluminum to copper).doc - BetterLesson

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Stoichiometry Using Copper Lab; Butane Lab; Antifreeze Lab; Le Chatelier's Principle Lab ... but the same amount should be still be present. As each of the reactions in this lab occurred, the mass of copper should have been equal to the initial amount of around two grams. ... for the Law of Conservation of Mass states it so. Stoichiometry can ...

Stoichiometry Using Copper Lab - AP Chemistry Krebs 2012-2013

Chapter 3 Stoichiometry 3-3 3.1a Avogadro's Number The mole (abbreviated mol) is the unit chemists use when counting numbers of atoms or molecules in a sample. The number of particles (atoms, molecules, or other objects) in one mole is equal to the number of atoms in exactly 12 g of carbon-12.

Chapter 3 Stoichiometry - Oneonta

Stoichiometry is not only used to balance chemical equations but also used in conversions, i.e., converting from grams to moles using molar mass as the conversion factor, or from grams to milliliters using density. For example, to find the amount of NaCl (sodium chloride) in 2.00 g, one would do the following:

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