Specific Heat Of Water Answer

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Specific Heat Of Water Answer
Best Answer: Mass X Change in Temperature (ΔT) X Specific Heat (SH) = Heat This is what we know: 45.0 g X (? final - 20.7 °C initial) X 4.184 J/1 cal = 3560 J Divide the heat (3560J) by the mass (45 grams) times the amount of joules in one gram of water (4.184J/g H20)

Specific Heat of Water? | Yahoo Answers
Best Answer: shorter experience. you're right. substance with smaller specific heat capacity(SHC) experiences a greater temperature drop when energy is released. assume the heat released for two substances, one of high SHC and one of lower SHC. relate to the equation Q=mc(dT). 'c' is inversely proportional ...

specific heat of water? | Yahoo Answers
Answer . Soil does not have a higher specific heat than water. . Specific heat capacity , also known simply as specific heat (Symbol: C or c) is the measure of the heat energy required to raise ...

Specific heat of water - answers.com
1) Calculate the quantity of heat required to raise the temperature of 60.0 g of water from 19.9 °C to 39.4 °C. 2) Calculate the final temperature, in degrees Celcius, when 75.0 g of water, initially at 22.1 °C, absorbs 4580 J of heat. Follow. 1 answer 1. Report Abuse.

Specific Heat of Water? | Yahoo Answers
Views · View Upvoters. The specific heat of water is 1 calorie/gram °C = 4.186 joule/gram °C which is higher than any other common substance. As a result, water plays a very important role in temperature regulation.

What is the specific heat of water ? - Quora
Water has a large specific heat. That means that it takes more heat energy to change the temperature of water than it does to change the temp of land; thus places near large bodies of water are ...

Specific heat of pure water - answers.com
The specific heat capacity of water is high, 4.184 J/gºC. The presence of impurities in a sample of water lowers its specific heat capacity. What is the specific heat capacity of the sample if 100 grams of it now requires 200 Joules of heat for a 1.8ºC temperature increase? What are the effects of the low specific heat capacity of the sample of water?

specific heat capacity help? | Yahoo Answers
Specific Heat Answer Key. 1. According to Joule's Law, the internal energy of a gas is a function of the kinetic energy of its molecules. 2. When working gas law problems, all temperatures must be converted to the. Celsius scale. Fahrenheit scale. Boyle scale.

Specific Heat Answer Key - HelpTeaching.com
The specific heat of calcium metal is 0.653 j/g*degrees C. If 75.0 g of calcium metal at 35.0 degrees celcius are placed into 100.0 g of water at 20.0 degrees C, what will be the final temperature of the water? A rich student wishes to determine the specific heat of gold. He places a 52.6g piece of gold in boiling water until it reaches 100.0 degrees C.

Specific heat questions? | Yahoo Answers
Specific Heat Problems 1) How much heat must be absorbed by 375 grams of water to raise its temperature by 25° C? 2) What mass of water can be heated from 25.0° C to 50.0° C by the addition of 2825 J? 3) What is the final temperature when 625 grams of water at 75.0° C loses 7.96 x 104 J?

Specific Heat Problems - mmsphyschem.com
The specific heat capacity of a solid or liquid is defined as the quantity of heat required to change

the temperature of a unit mass of a substance through a unit change in temperature. Our result from this experiment was somewhat close to the specific heat capacity for water, but still off the mark.

Determining Heat Capacity of Water Lab Answers ...
Worksheet- Introduction to Specific Heat Capacities Heating substances in the sun: The following table shows the temperature after 10.0 g of 4 different substances have been in direct sunlight for up to 60 minutes.

Name: Per: Worksheet- Introduction to Specific Heat Capacities
Given the specific heat of copper is 0.385J/gC and that of water is 4.184J/gC? what is the final temperature if 985g of copper at 95C is placed into 1102.2g of water at 8.0C Follow

Given the specific heat of copper ... - au.answers.yahoo.com
The heat of the water is transferred to the coastal areas by wind - thus warm air promotes growth of vegetation. Cold air blowing from land to sea - forms fog Cold currents adhere fog when wind blows the cold air inland Coastal areas will generally have more moderate temperatures than inland areas because of the heat capacity of the ocean.

How are the climates of coastal ... - uk.answers.yahoo.com
Initial temperature of hot chocolate liquid: 25 °C. Final temperature of the hot chocolate liquid: 40 °C. Hot Chocolate Calculations Answers will vary, based on collected data. See example answers below. 1. Find the specific heat (SH) of the hot chocolate using the equation: Q = mcΔT. a. Find ΔT for the hot chocolate.

Specific Heat Capacity Handout Answer Key
To calculate specific heat, start by reading the problem carefully, then write down each known and unknown variable to get a better sense of what you're working with. Next, plug the known factors into the specific heat equation, then solve the equation as you normally would to get your answer!

How to Calculate Specific Heat (with Calculator) - wikiHow
The specific heat of water is 4.18 J/g•°C. How much heat does 225.0 g of water release when it cools from 85.5°C to 50.0°C? Enter your answer below, in whole numbers, without any commas.

The specific heat of water is 4.18 J/g•°C. How much heat ...
Water has a very high specific heat capacity of 4.1814 J/(g·K) at 25 °C – the second highest among all the heteroatomic species (after ammonia), as well as a high heat of vaporization (40.65 kJ/mol or 2257 kJ/kg at the normal boiling point), both of which are a result of the extensive hydrogen bonding between its molecules.

Properties of water - Wikipedia
This specific heat calculator is a tool that determines the heat capacity of a heated or a cooled sample. Specific heat is just the amount of thermal energy you need to supply to a sample weighing 1 kg to increase its temperature by 1 K. Read on to learn how to apply the heat capacity formula correctly to obtain a valid result.

Specific Heat Calculator - Omni
Specific heat capacity questions and equation Whenever two objects with different initial temperatures are put in contact with each other, the warmer one will cool down, and the cooler one will warm up, until they reach the same temperature.

Specific Heat Of Water Answer

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