

## *Subatomic Heavyweights Worksheet Answers*

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SUBATOMIC PARTICLES and ISOTOPES WORKSHEET ... Answer each of the following using your knowledge of chemistry and the Periodic Table. 4) An atom contains 55 protons. ... What is the atomic number of bromine? 35 7) What is the number of total subatomic particles in an atom of B-11? 16 8) What is the atomic number of Zn-65? 30

**SUBATOMIC PARTICLES and ISOTOPES WORKSHEET**

Lesson 13 - Subatomic Heavyweights. jena23chemistry ♦ September 6, 2012 ♦ 1 Comment. Today I learned about isotopes and how to calculate the average atomic mass of an element. Average atomic mass (AAM) is used on the periodic table and can be calculated by averaging the masses of all the isotopes found for that element.

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Best Answer: I'm going to try and walk you through this as best I can. In this exercise, you are looking at ten boron atoms. Imagine that these 10 B atoms were taken at random from a sample of boron (which is a gray powder). You can tell that all the atoms shown on the picture are boron. How? The protons ...

**Help with tenth grade chemistry PLEASE! Subatomic ...**

50 Unit 1 Alchemy Living By Chemistry Teaching and Classroom Resources Lesson 13 • Worksheet © 2012 W. H. Freeman and Company/BFW Subatomic Heavyweights Isotopes ...

### **Purpose Part 1: The Average Boron Atom - Weebly**

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### **Subatomic Particle Calculations - Lesson Worksheets**

An isotope is an atom of the same element with a different number of neutrons. Most elements have at least 2 isotopes. The average atomic mass is the weighted average of the mass of the isotopes in an element.

### **Lesson 13: Subatomic Heavyweights- Isotopes - Unit 1: Alchemy**

- Isotopes of an element have the same number of protons and electrons but different numbers of neutrons.
- The average atomic mass of an element listed in the periodic table is the weighted average mass of the naturally occurring isotopes of that element.
- Isotopes are referred to by their mass numbers, as in carbon-12.

### **Isotopes - NPS WWW2**

Subatomic Heavyweights HW Read and outline (Cornell style) Section 4.3 in your chemistry textbook. Then answer the following assessment questions: 1. Which subatomic particle identifies an atom as that of a particular element? How is this particle related to the atom's atomic number? 2. What is an isotope? Give an example of an element with ...

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