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CHAPTER 9: STUDY QUESTIONS

Solutions & Their Behavior. Solution concentration can often be expressed in units of molarity, molality, weight percent, or as ppm. Polar solutions are miscible with other polar solutions, and nonpolar solutions are miscible with other nonpolar solutions. Polar and nonpolar solutions do not mix and form heterogeneous mixtures.

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Solutions and . Their Behavior. By. Dr. Andrea Wallace. CGCC . Edited by. John Taylor. FCCJ-North

Campus . Chapter 14 : Solutions and Their Behavior . 14.1 Units of Concentration, p. 656 . A solution consists of a _____ and a _____. Examples: Concentration is a way to express the amount of solute present. Equations:

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71. A solution of toluene (molecular weight 92.1) in benzene (molecular weight 78.1) is prepared. The mole fraction of toluene in the solution is 0.100. ... ideal behavior " argument AP Chemistry - Solutions, Chapter 11/4b FAQ. Q: What are hydrophobic and hydrophilic? ... AP Chemistry study guide for Solutions (Chapter 11) ...

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Chapter 14 Solutions and Their Behavior Study Guide. 1. Define solute, solvent, and solution. 2. Define unsaturated, saturated, and supersaturated. How would you make a supersaturated solution? 3. Define the concentration units molarity, molality, mole fraction, mass percent, pph, ppt and ppm. 4. Define the terms miscible and immiscible. 5.

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