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Chapter 20 Oxidation–Reduction Reactions221 SECTION 20.1 THE MEANING OF OXIDATION AND REDUCTION (pages 631–638) This section explains oxidation and reduction in terms of the loss or gain of electrons, and describes the characteristics of a redox reaction. It also explains how to identify oxidizing and reducing agents.

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Chapter 20 Oxidation-Reduction Reactions 523 Practice Problems In your notebook, solve the following problems. SECTION 20.1 THE MEANING OF OXIDATION AND REDUCTION Determine what is oxidized and what is reduced in each reaction. Identify the oxidizing agent and the reducing agent. 1.  $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$  2.  $2\text{Li} + \text{S} \rightarrow 2\text{Li}_2\text{S}$  3.  $2\text{Cs} + \text{Br}_2 \rightarrow 2\text{CsBr}$  4.  $3\text{Mg} + \text{N}_2 \rightarrow \dots$

### SECTION 20.1 THE MEANING OF OXIDATION AND REDUCTION

b. butanoic acid from 1-butanol 5. Classify each of the reactions in problem 4 as an oxidation-reduction, substitution, addition, or esterification reaction. SECTION 23.4 POLYMERIZATION 1. Draw the structures of propene (propylene) and tetrafluoroethene. Draw the basic repeating units when each of these compounds polymerizes. Describe

### SECTION 23.1 INTRODUCTION TO FUNCTIONAL GROUPS

SECTION 20.2 OXIDATION NUMBERS 1. Give the oxidation number of each kind of atom or ion. a. Sn c.  $\text{S}^{2-}$  e. Se g.  $\text{Sn}^{4+}$  b. K d.  $\text{Fe}^{3+}$  f.  $\text{Mg}^{2+}$  h. Br 2. Calculate the oxidation number of chromium in each of the following formulas. a.  $\text{Cr}_2\text{O}_3$  b.  $\text{H}_2\text{Cr}_2\text{O}_7$  c.  $\text{CrSO}_4$  d.  $\text{CrO}_4^{2-}$  3. Use the changes in oxidation number to determine which elements are oxidized ...

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