

Stoichiometry Instructional Fair Answers

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Stoichiometry Instructional Fair Answers

STOICHIOMETRY: MIXED PROBLEMS Name What volume of NH_3 at STP is produced if 25.0 g of N_2 is reacted with an excess of H_2 ? 2. $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$ If 5.0 g of KClO_3 is decomposed, what volume of O_2 is produced at STP? 3. How many grams of KCl are produced in Problem 2? ...
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Stoichiometry: Mole-Mole Problems 1. $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ How many moles of hydrogen are needed to completely react with 2.00 moles of nitrogen? 2. ... Chemistry IF8766 Page 62 Instructional Fair, Inc. Title: Microsoft Word - pg 62 - Stoichiometry - mole-mole problems.doc

Stoichiometry: Mole-Mole Problems

Stoichiometry: Mixed Problems (KEY) 1) $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ What volume of NH_3 at STP is produced if 25.0 g of N_2 is reacted with an excess of H_2 ? 3 3 3 2 3 2 2 40.0L NH_3 1mol NH_3 22.4L NH_3 1mol N_2 2mol NH_3 28.0g N_2 25.0g N_2 1mol N_2 $\times \times \times = 2$ 2) $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$ If 5.0g of KClO_3 is decomposed, what volume of O_2 is produced at STP? 2

Stoichiometry: Mixed Problems (KEY)

Stoichiometry: Mole-Mole Problems 1. $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ How many moles of hydrogen are needed to completely react with 2.00 moles of nitrogen? 2. $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$ How many moles of oxygen are produced by the decomposition of 6.00 moles of potassium chlorate? 3. $\text{Zn} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$

Stoichiometry: Mole-Mole Problems

STOICHIOMETRY: MASS-MASS PROBLEMS Name Kao RcllJs= 79 Rci + O_2 How many grams of potassium chloride are produced if 25 g of potassium chlorate decompose? 05 101 How many grams of hydrogen are necessary to react completely with 50.0 g of nitrogen in the above reaction? 50e0 Z 3. How many grams of ammonia are produced in the reaction in Problem 2?

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Stoichiometry: Volume-Volume Problems 1. $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ What volume of hydrogen is necessary to react with 5.00 L of nitrogen to produce ammonia? 2. What volume of ammonia is produced in the reaction in Problem 1? 3. $\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$ If 20.0 L of oxygen are consumed in the above reaction, how many liters of carbon ...

Stoichiometry: Volume-Volume Problems

ANSWER: 0.89 mole Chemistry IF8766 ... 8-14a,b Mixed Problems--Mole-Mole and Mole-Mass wkst-Key. KEY. Answer each of the following ... Gram Formula Mass Practice Worksheet - TMA-CHEMISTRY - home

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