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The Beer-Lambert law, also known as Beer's law, the Lambert-Beer law, or the Beer-Lambert-Bouguer law relates the attenuation of light to the properties of the material through which the light is travelling. The law is commonly applied to chemical analysis measurements and used in understanding attenuation in physical optics, for photons, neutrons, or rarefied gases.

Beer-Lambert law - Wikipedia

Beer's Law is an equation that relates the attenuation of light to properties of a material. The law states the concentration of a chemical is directly proportional to the absorbance of a solution. The relation may be used to determine the concentration of a chemical species in a solution using a colorimeter or spectrophotometer. The relation is most often used in UV-visible absorption spectroscopy.

Beer's Law Definition and Equation - ThoughtCo

Here $I(0)$ is the intensity of monochromatic light entering the solution perpendicularly to one face and $I(X)$ is the intensity of light exiting the solution through the opposite face. With an SI unit of $\text{m}^2 \text{mol}^{-1}$, the constant ϵ goes by a variety of names, one of which is molar absorptivity. The dimensionless quantity A , termed the absorbance, is defined by Eq.

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