



Cambridge International AS & A Level

CHEMISTRY

9701/22

Paper 2 AS Structured Questions

October/November 2024

MARK SCHEME

Maximum Mark: 60

<p>Published</p>

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge International will not enter into discussions about these mark schemes.

Cambridge International is publishing the mark schemes for the October/November 2024 series for most Cambridge IGCSE, Cambridge International A and AS Level components, and some Cambridge O Level components.

This document consists of **11** printed pages.

Generic Marking Principles

These general marking principles must be applied by all examiners when marking candidate answers. They should be applied alongside the specific content of the mark scheme or generic level descriptions for a question. Each question paper and mark scheme will also comply with these marking principles.

GENERIC MARKING PRINCIPLE 1:

Marks must be awarded in line with:

- the specific content of the mark scheme or the generic level descriptors for the question
- the specific skills defined in the mark scheme or in the generic level descriptors for the question
- the standard of response required by a candidate as exemplified by the standardisation scripts.

GENERIC MARKING PRINCIPLE 2:

Marks awarded are always **whole marks** (not half marks, or other fractions).

GENERIC MARKING PRINCIPLE 3:

Marks must be awarded **positively**:

- marks are awarded for correct/valid answers, as defined in the mark scheme. However, credit is given for valid answers which go beyond the scope of the syllabus and mark scheme, referring to your Team Leader as appropriate
- marks are awarded when candidates clearly demonstrate what they know and can do
- marks are not deducted for errors
- marks are not deducted for omissions
- answers should only be judged on the quality of spelling, punctuation and grammar when these features are specifically assessed by the question as indicated by the mark scheme. The meaning, however, should be unambiguous.

GENERIC MARKING PRINCIPLE 4:

Rules must be applied consistently, e.g. in situations where candidates have not followed instructions or in the application of generic level descriptors.

GENERIC MARKING PRINCIPLE 5:

Marks should be awarded using the full range of marks defined in the mark scheme for the question (however; the use of the full mark range may be limited according to the quality of the candidate responses seen).

GENERIC MARKING PRINCIPLE 6:

Marks awarded are based solely on the requirements as defined in the mark scheme. Marks should not be awarded with grade thresholds or grade descriptors in mind.

Science-Specific Marking Principles

1 Examiners should consider the context and scientific use of any keywords when awarding marks. Although keywords may be present, marks should not be awarded if the keywords are used incorrectly.

2 The examiner should not choose between contradictory statements given in the same question part, and credit should not be awarded for any correct statement that is contradicted within the same question part. Wrong science that is irrelevant to the question should be ignored.

3 Although spellings do not have to be correct, spellings of syllabus terms must allow for clear and unambiguous separation from other syllabus terms with which they may be confused (e.g. ethane / ethene, glucagon / glycogen, refraction / reflection).

4 The error carried forward (ecf) principle should be applied, where appropriate. If an incorrect answer is subsequently used in a scientifically correct way, the candidate should be awarded these subsequent marking points. Further guidance will be included in the mark scheme where necessary and any exceptions to this general principle will be noted.

5 'List rule' guidance

For questions that require ***n*** responses (e.g. State **two** reasons ...):

- The response should be read as continuous prose, even when numbered answer spaces are provided.
- Any response marked *ignore* in the mark scheme should not count towards ***n***.
- Incorrect responses should not be awarded credit but will still count towards ***n***.
- Read the entire response to check for any responses that contradict those that would otherwise be credited. Credit should **not** be awarded for any responses that are contradicted within the rest of the response. Where two responses contradict one another, this should be treated as a single incorrect response.
- Non-contradictory responses after the first ***n*** responses may be ignored even if they include incorrect science.

6 Calculation specific guidance

Correct answers to calculations should be given full credit even if there is no working or incorrect working, **unless** the question states 'show your working'.

For questions in which the number of significant figures required is not stated, credit should be awarded for correct answers when rounded by the examiner to the number of significant figures given in the mark scheme. This may not apply to measured values.

For answers given in standard form (e.g. $a \times 10^n$) in which the convention of restricting the value of the coefficient (a) to a value between 1 and 10 is not followed, credit may still be awarded if the answer can be converted to the answer given in the mark scheme.

Unless a separate mark is given for a unit, a missing or incorrect unit will normally mean that the final calculation mark is not awarded. Exceptions to this general principle will be noted in the mark scheme.

7 Guidance for chemical equations


Multiples / fractions of coefficients used in chemical equations are acceptable unless stated otherwise in the mark scheme.

State symbols given in an equation should be ignored unless asked for in the question or stated otherwise in the mark scheme.


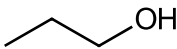
Question	Answer	Marks												
1(a)(i)	lowest energy state owtte examples <ul style="list-style-type: none">• (with) lowest energy• (with) no external energy• (with) no added energy• (where) electrons are not promoted to a higher energy level• most energetically stable	1												
1(a)(ii)	[Ar] <table><tr><td>↑</td><td>↑</td><td>↑</td><td></td><td></td><td>↑↓</td></tr></table> <i>The 3 × d electrons may go in any of the d orbitals</i>	↑	↑	↑			↑↓	1						
↑	↑	↑			↑↓									
1(a)(iii)	12	1												
1(b)(i)	<table><tr><td>isotope</td><td>RIM</td><td># protons</td><td># neutrons</td></tr><tr><td>⁹³₄₁Nb</td><td>92.91</td><td>• 41</td><td>• 52</td></tr><tr><td>¹⁸¹₇₃Ta</td><td>180.95</td><td>• 73</td><td>• 108</td></tr></table>	isotope	RIM	# protons	# neutrons	⁹³ ₄₁ Nb	92.91	• 41	• 52	¹⁸¹ ₇₃ Ta	180.95	• 73	• 108	2
isotope	RIM	# protons	# neutrons											
⁹³ ₄₁ Nb	92.91	• 41	• 52											
¹⁸¹ ₇₃ Ta	180.95	• 73	• 108											

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Question	Answer	Marks
1(b)(ii)	M1 mass of an (atom of an) isotope	1
	M2 relative/compared to (the mass of) the unified atomic mass unit OR on a scale in which a carbon-12 atom / isotope has a mass of exactly 12 units OR divided by / compared to $\frac{1}{12}$ of the mass of a carbon-12 atom / isotope Alternative route using mass of 1 mol throughout M1 mass of one mol of an (atom of an) isotope M2 relative/compared to $\frac{1}{12}$ of the mass of 1 mol of C-12 atom / isotope OR when the mass of one mol C-12 atom / isotope is exactly 12(.000) g	1
1(b)(iii)	$92.91 \times 0.909 + 180.95 \times 0.091$	1
	$= 100.92$	1

Question	Answer	Marks
2(a)(i)	M1 $4\text{Na} + \text{O}_2 \rightarrow 2\text{Na}_2\text{O}$	1
	M2 $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$	1
2(a)(ii)	Al^{3+} 	1
2(a)(iii)	it increases	1
	as number of valence electrons (which can be used in bonding / lost / shared / donated) increases	1

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Question	Answer	Marks									
2(b)(i)	<table border="1"> <tr> <td>Period 3 oxide</td><td>product of rxn with water</td><td>pH of solution formed</td></tr> <tr> <td>• MgO</td><td>Mg(OH)₂</td><td>• 8 ≤ pH ≤ 12</td></tr> <tr> <td>P₄O₁₀</td><td>• H₃PO₄</td><td>• 1 ≤ pH ≤ 4</td></tr> </table> <p style="text-align: right;">•✓•✓</p>	Period 3 oxide	product of rxn with water	pH of solution formed	• MgO	Mg(OH) ₂	• 8 ≤ pH ≤ 12	P ₄ O ₁₀	• H ₃ PO ₄	• 1 ≤ pH ≤ 4	2
Period 3 oxide	product of rxn with water	pH of solution formed									
• MgO	Mg(OH) ₂	• 8 ≤ pH ≤ 12									
P ₄ O ₁₀	• H ₃ PO ₄	• 1 ≤ pH ≤ 4									
2(b)(ii)	CH ₃ CN + 2H ₂ O + (1)H ⁺ → CH ₃ COOH + NH ₄ ⁺	1									
2(b)(iii)	 	2									
2(b)(iv)	alkyl groups are electron donating / have positive inductive effect.	1									
	strengthens the O—H bond / makes H ⁺ less likely to be donated	1									
2(c)(i)	(molecules) H ₂ S to H ₂ Te / they have greater number of electrons	1									
	stronger instantaneous dipole–induced dipole / London dispersion forces OR more energy required to overcome the instantaneous dipole–induced dipole / London dispersion forces	1									
2(c)(ii)	(only) H ₂ O has hydrogen bonding AND hydrogen bonding (much) stronger than the other intermolecular forces OR H ₂ O has hydrogen bonding AND because of higher electronegativity of O compared to S OR stronger van der waals' forces because (only) H ₂ O has hydrogen bonding OR hydrogen bond AND increased strength of permanent dipoles in H ₂ O outweighs the increase in strength of id-id in the others	1									

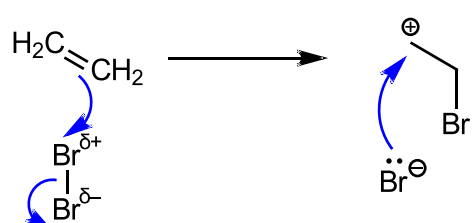
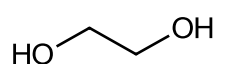
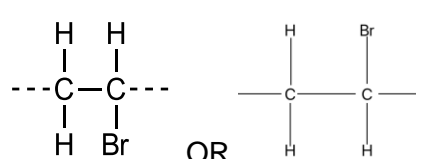
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Question	Answer	Marks
3(a)(i)	(natural =)lightning	1
	(man-made =)internal combustion engines	1
3(a)(ii)	$2\text{NO}_2 + \text{H}_2\text{O} \rightarrow \text{HNO}_2 + \text{HNO}_3$ OR $4\text{NO}_2 + \text{O}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{HNO}_3$	1
3(a)(iii)	It / NO_2 reacts with (unburned) hydrocarbons / VOCs ALLOW reaction of unburned hydrocarbons / VOCs in presence of NO_2	1
3(a)(iv)	$2\text{HNO}_3 + \text{CaO} \rightarrow \text{Ca}(\text{NO}_3)_2 + \text{H}_2\text{O}$	1
3(a)(v)	brown fumes given off	1
3(b)(i)	(+)5 / V	1
3(b)(ii)	aluminium / Al	1
3(b)(iii)	(NH_3 is an) H^+ acceptor	1
3(b)(iv)	tetrahedral	1
3(c)(i)	$\text{P}^{2+}(\text{g}) \rightarrow \text{P}^{3+}(\text{g}) + \text{e}^{(-)}$	1
3(c)(ii)	general increase	1
	increase in IE between IE5 and IE6 is noticeably the largest increase between IEs shown	1

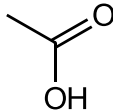
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Question	Answer			Marks
3(d)	element	nitrogen	phosphorus	2
	state and appearance	colourless gas	white solid	
	electrical conductivity	• poor	poor	
	type of bonding	• covalent	• covalent	
	type of structure	simple	• simple	
3(e)	simple molecular (lattice structure)			1
3(f)(i)	M1 one sigma / σ bond AND head-on (overlap of) p / sp (orbitals)			1
	M2 two pi / π bond(s) AND side-on (overlap of / involving) p (orbitals)			1
3(f)(ii)	$\text{P}\equiv\text{P}$ is much weaker so P_2 is more reactive (than N_2)			1

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Question	Answer	Marks
4(a)	 <p>pt 1 dipole on Br₂</p> <p>pt 2 curly arrow from C=C bond to a Br^{δ+}</p> <p>pt 3 curly arrow from Br–Br bond to the other Br</p> <p>pt 4 correct intermediate</p> <p>pt 5 curly arrow from lone pair on Br[–] to C⁺</p> <p>5• = 3 marks. 4• = 2 marks. 3• = 2 marks. 2• = 1 marks. 1• = 0 marks.</p>	3
4(b)	= –37.8 (kJ mol ^{–1})	1
4(c)(i)		1
	bromoethene	1
4(c)(ii)		1
4(c)(iii)	C ₂ H ₂	1
4(d)(i)	structural / positional	1
4(d)(ii)	aldehyde	1

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Question	Answer	Marks								
4(d)(iii)	<table><tr><th>reagent</th><th>observation with F</th></tr><tr><td>2,4-DNPH</td><td>red / orange / yellow ppt</td></tr><tr><td>Tollens' reagent</td><td>silver mirror OR grey ppt OR black ppt</td></tr><tr><td>alkaline I₂(aq)</td><td>yellow ppt</td></tr></table>	reagent	observation with F	2,4-DNPH	red / orange / yellow ppt	Tollens' reagent	silver mirror OR grey ppt OR black ppt	alkaline I ₂ (aq)	yellow ppt	3
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2,4-DNPH	red / orange / yellow ppt									
Tollens' reagent	silver mirror OR grey ppt OR black ppt									
alkaline I ₂ (aq)	yellow ppt									
4(e)(i)	<div><div></div><p>M1 identify H =</p><p>M2 (broad) absorption within the range 3600–2500 cm⁻¹ so O-H (bond) O-H (bond) is equivalent to OH <u>bond</u></p><p>M3 pt 1 and pt 2 •✓ OR pt 1 and pt 3 •✓</p><p>•pt 1 absorption within the range 1670–1750 cm⁻¹ so C=O (bond)</p><p>•pt 2 absorption within the range 1040–1300 cm⁻¹ so C-O (bond)</p><p>•pt 3 (M⁺ at <i>m/ e</i> = 60 so it has) molecular mass / Mr = 60</p></div>									
4(e)(ii)	oxidising agent	1								