

## 10 Sc Chemical reactions and equations

10.0072

10.0072

10.0072. What is the type of the reaction between sodium sulphate and barium chloride?

(a) Decomposition (b) Direct combination (c) Single displacement (d) Double displacement

## 10 Sc Chemical reactions and equations

*Answer*

10.0072

D

## 10 Sc Chemical reactions and equations

10.0075

10.0075

10.0075. Name the precipitate formed when aqueous solutions of sodium sulphate and barium chloride are mixed?

(a) Barium sulphate (b) Sodium sulphate (c) Barium hydroxide (d) Barium sulphate

## 10 Sc Chemical reactions and equations

*Answer*

10.0075

D

## 10 Sc Chemical reactions and equations

10.0077

10.0077

10.0077. What is the colour of the aqueous solution of Barium chloride?

(a) Blue (b) Green (c) Colourless (d) Purple

## 10 Sc Chemical reactions and equations

*Answer*

10.0077

C

## 10 Sc Chemical reactions and equations

10.0084

10.0084

10.0084. The heating of ferrous sulphate is an example of \_\_\_\_\_ reaction and the reaction between iron and copper sulphate is an example of \_\_\_\_\_ reaction.

- (a) displacement, decomposition (b) decomposition, displacement  
(c) combination, displacement (d) combination, decomposition

## 10 Sc Chemical reactions and equations

*Answer*

10.0084

B



## 10 Sc Chemical reactions and equations

10.0088

10.0088

10.0088. What is the colour of the precipitate which is formed when aqueous solutions of sodium sulphate and barium chloride are mixed?

(a) Yellow (b) White (c) Green (d) Orange

## 10 Sc Chemical reactions and equations

*Answer*

10.0088

B

## 10 Sc Chemical reactions and equations

10.009

10.009

10.0090. What is the color of the solution of copper sulphate?

(a) Green (b) Light green (c) Blue (d) Pink

## 10 Sc Chemical reactions and equations

*Answer*

10.009

C

## 10 Sc Chemical reactions and equations

10.0118

10.0118

10.0118. What is the colour of the aqueous solution of sodium sulphate?

(a) Blue (b) Green (c) Colourless (d) Purple

## 10 Sc Chemical reactions and equations

*Answer*

10.0118

C

## 10 Sc Chemical reactions and equations

10.0165

10.0165

10.0165. Heating ferrous sulphate is an example of which type of reaction?

- (a) Thermal decomposition (b) Thermal analysis  
(c) Double decomposition (d) Catalytic decomposition

## 10 Sc Chemical reactions and equations

*Answer*

10.0165

A



## 10 Sc Chemical reactions and equations

10.0291

10.0291

10.0291. What is the color of the residue left in the test tube after thermal decomposition of ferrous sulphate?

(a) White (b) Black (c) Green (d) Yellowish brown

## 10 Sc Chemical reactions and equations

*Answer*

10.0291

D

## 10 Sc Chemical reactions and equations

10.0292

10.0292

10.0292. What is the odor of the gases produced on thermal decomposition of ferrous sulphate?

(a) Like rotten eggs (b) Like burning sulphur (c) Sweet smell (d) Fruity smell

## 10 Sc Chemical reactions and equations

*Answer*

10.0292

B

## 10 Sc Chemical reactions and equations

10.0293

10.0293

10.0293. How the colour changes when the gases after thermal decomposition of ferrous sulphate come in contact with an acidified solution of potassium dichromate?

- (a) Green to orange (b) Orange to green  
(c) Blue to green (d) Red to colourless

## 10 Sc Chemical reactions and equations

*Answer*

10.0293

B

## 10 Sc Chemical reactions and equations

10.0294

10.0294

10: What is the color of solid ferrous sulphate?

(a) Blue (b) Red (c) Light green (d) Light yellow

## 10 Sc Chemical reactions and equations

*Answer*

10.0294

C



## 10 Sc Chemical reactions and equations

10.0295

10.0295

10.0295. What is the nature of the new product which is formed by the action of water on quick lime?

(a) Neutral (b) Acidic (c) Amphoteric (d) Basic

## 10 Sc Chemical reactions and equations

*Answer*

10.0295

C

## 10 Sc Chemical reactions and equations

10.0296

10.0296

10.0296. The action of water on quick lime is an example of which type of reaction?

(a) Combination (b) Displacement (c)  
Decomposition (d) Redox

## 10 Sc Chemical reactions and equations

*Answer*

10.0296

A

## 10 Sc Chemical reactions and equations

10.0297

10.0297

10.0297. What becomes the colour of iron nails after they have been kept in the blue solution of copper sulphate for about 20 minutes?

(a) White (b) Reddish brown (c) Green (d) Blue

## 10 Sc Chemical reactions and equations

*Answer*

10.0297

B

## 10 Sc Chemical reactions and equations

10.0298

10.0298

10.0298. What is the best way to describe the reaction that takes place during the process of respiration?.

(a) Oxidation (b) Endothermic (c) Exothermic  
(d) Displacement

## 10 Sc Chemical reactions and equations

*Answer*

10.0298

A



## 10 Sc Chemical reactions and equations

10.0299

10.0299

10.0299. Which one of the following processes involve chemical reactions?

- (a) Storing of oxygen gas under pressure in a gas cylinder
- (b) Liquefaction of air
- (c) Keeping petrol in a china dish in the open
- (d) Heating copper wire in presence of air at high temperature

## 10 Sc Chemical reactions and equations

*Answer*

10.0299

D

## 10 Sc Chemical reactions and equations

10.03

10.03

10.0300. Which of the following gases can be used for storage of fresh sample of an oil for a long time?

- (a) Carbon dioxide or oxygen
- (b) Nitrogen or oxygen
- (c) Carbon dioxide or helium
- (d) Helium or nitrogen

## 10 Sc Chemical reactions and equations

*Answer*

10.03

D

## 10 Sc Chemical reactions and equations

10.0301

10.0301

10.0301. Solid calcium oxide reacts vigorously with water to form calcium hydroxide accompanied by liberation of heat. This process is called slaking of lime. Calcium hydroxide dissolves in water to form its solution called lime water. Which among the following is (are) true about slaking of lime and the solution formed?

- (i) It is an endothermic reaction
- (ii) It is an exothermic reaction
- (iii) The pH of the resulting solution will be more than seven
- (iv) The pH of the resulting solution will be less than seven

(a) (i) and (ii) (b) (ii) and (iii)

## 10 Sc Chemical reactions and equations

*Answer*

10.0301

B

## 10 Sc Chemical reactions and equations

10.0302

10.0302

10.0302. Which among the following statement(s) is (are) true? Exposure of silver chloride to sunlight for a long duration turns grey due to

- (i) the formation of silver by decomposition of silver chloride
- (ii) sublimation of silver chloride
- (iii) decomposition of chlorine gas from silver chloride
- (iv) oxidation of silver chloride

- (a) (i) only (b) (i) and (iii)
- (c) (ii) and (iii) (d) (iv) only

## 10 Sc Chemical reactions and equations

*Answer*

10.0302

A



## 10 Sc Chemical reactions and equations

10.0303

10.0303

10.0303. Which of the following is not a physical change?

- (a) Boiling of water to give water vapour
- (b) Melting of ice to give water
- (c) Dissolution of salt in water
- (d) Combustion of Liquefied Petroleum Gas (LPG)

## 10 Sc Chemical reactions and equations

*Answer*

10.0303

D

## 10 Sc Chemical reactions and equations

10.0304

10.0304

10.0304. Which of the following are exothermic processes?

- (i) Reaction of water with quick lime
- (ii) Dilution of an acid
- (iii) Evaporation of water
- (iv) Sublimation of camphor (crystals)

- (a) (i) and (ii) (b) (ii) and (iii)
- (c) (i) and (iv) (d) (iii) and (iv)

## 10 Sc Chemical reactions and equations

*Answer*

10.0304

A

## 10 Sc Chemical reactions and equations

10.0305

10.0305

10.0305. Barium chloride on reacting with ammonium sulphate forms barium sulphate and ammonium chloride. Which of the following correctly represents the type of the reaction involved?

- (i) Displacement reaction
- (ii) Precipitation reaction
- (iii) Combination reaction
- (iv) Double displacement reaction

- (a) (i) only (b) (ii) only
- (c) (iv) only (d) (ii) and (iv)

## 10 Sc Chemical reactions and equations

*Answer*

10.0305

D

## 10 Sc Chemical reactions and equations

10.0306

10.0306

10.0306. In the double displacement reaction between aqueous potassium iodide and aqueous lead nitrate, a yellow precipitate of lead iodide is formed. While performing the activity if lead nitrate is not available, which of the following can be used in place of lead nitrate?

- (a) Lead sulphate (insoluble)
- (b) Lead acetate
- (c) Ammonium nitrate
- (d) Potassium sulphate

## 10 Sc Chemical reactions and equations

*Answer*

10.0306

B



## 10 Sc Chemical reactions and equations

10.0307

10.0307

10.0307. Which of the following is(are) an endothermic process(es)?

- (i) Dilution of sulphuric acid
- (ii) Sublimation of dry ice
- (iii) Condensation of water vapours
- (iv) Evaporation of water

(a) (i) and (iii) (b) (ii) only  
(c) (iii) only (d) (ii) and (iv) 12. Which of the following is(are) an endothermic process(es)?

- (i) Dilution of sulphuric acid
- (ii) Sublimation of dry ice
- (iii) Condensation of water vapours
- (iv) Evaporation of water

## 10 Sc Chemical reactions and equations

*Answer*

10.0307

D

## 10 Sc Chemical reactions and equations

10.0308

10.0308

10.0308. Electrolysis of water is a decomposition reaction. The mole ratio of hydrogen and oxygen gases liberated during electrolysis of water is

- (a) 1:1
- (b) 2:1
- (c) 4:1
- (d) 1:2

## 10 Sc Chemical reactions and equations

*Answer*

10.0308

B

## 10 Sc Chemical reactions and equations

10.0912

10.0912

10.0912 Which type of reaction takes place when an iron nail is dipped in a solution of copper sulphate?

(a) Combination (b) Displacement (c) Double displacement (d) Decomposition

## 10 Sc Chemical reactions and equations

*Answer*

10.0912

B