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$$E_{cell} = E_{cathode} - E_{anode}$$

$$E_{cell} = 1.10 - 0.76$$

$$E_{cell} = 0.34 V$$

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Q Calculate the reduction potential & oxidation potential of  $Cr/Cr^{+3}$  at  $25^\circ C$   
 $Cr^{+3} + 3e^- \rightleftharpoons Cr$   
 $E^\circ_{Cr^{+3}/Cr} = -0.75V$

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Q Calculate the reduction potential & oxidation potential of  $Cr/Cr^{+3}$  at  $25^\circ C$

$$E^\circ_{Cr^{+3}/Cr} = -0.75V$$

Q Calculate the emf of the cell ( $E_{cell}$ )



$$T = 25^\circ C (298K)$$

$$E^\circ_{Zn^{+2}/Zn} = -0.76V$$

$$E^\circ_{Sn^{+2}/Sn} = -0.14V$$