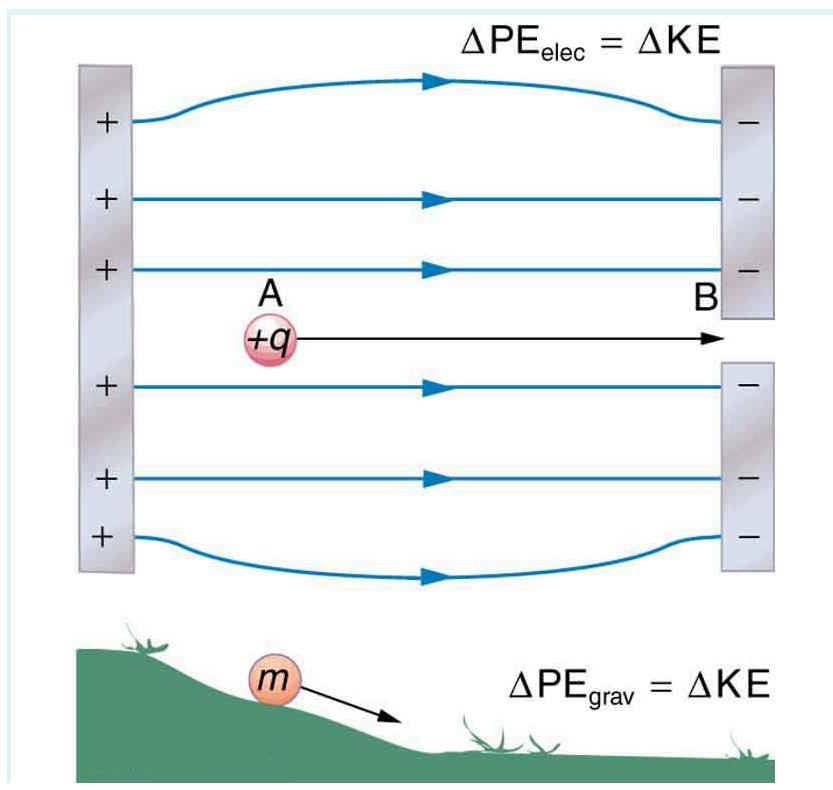


# ELECTRIC POTENTIAL ENERGY: POTENTIAL DIFFERENCE

## LEARNING OUTCOMES:

- Define electric potential and electric potential energy.
- Describe the relationship between potential difference and electrical potential energy.
- Explain electron volt and its usage in submicroscopic process.
- Determine electric potential energy given potential difference and amount of charge.

When a free positive charge  $q$  is accelerated by an electric field, such as shown in [Figure 1], it is given kinetic energy. The process is analogous to an object being accelerated by a gravitational field. It is as if the charge is going down an electrical hill where its electric potential energy is converted to kinetic energy. Let us explore the work done on a charge  $q$  by the electric field in this process, so that we may develop a definition of electric potential energy.



**Figure 1:** A charge accelerated by an electric field is analogous to a mass going down a hill. In both cases potential energy is converted to another form. Work is done by a force, but since this force is conservative, we can write  $W = -\Delta PE$ .

The electrostatic or Coulomb force is conservative, which means that the work done on  $q$  is independent of the path taken. This is exactly analogous to the gravitational force in the absence of dissipative forces such as friction. When a force is conservative, it is possible to define a potential energy associated with the force, and it is usually easier to deal with the potential energy (because it depends only on position) than to calculate the work directly.

We use the letters PE to denote electric potential energy, which has units of joules (J). The change in potential energy,  $\Delta PE$ , is crucial, since the work done by a conservative force is the negative of the change in potential energy; that is,  $W = -\Delta PE$ . For example, work  $W$  done to accelerate a positive charge from rest is positive and results from a loss in PE, or a negative  $\Delta PE$ . There must be a minus sign in front of  $\Delta PE$  to make  $W$  positive. PE can be found at any point by taking one point as a reference and calculating the work needed to move a charge to the other point.

### POTENTIAL ENERGY

$W = -\Delta PE$ . For example, work  $W$  done to accelerate a positive charge from rest is positive and results from a loss in PE, or a negative  $\Delta PE$ . There must be a minus sign in front of  $\Delta PE$  to make  $W$  positive. PE can be found at any point by taking one point as a reference and calculating the work needed to move a charge to the other point.

Gravitational potential energy and electric potential energy are quite analogous. Potential energy accounts for work done by a conservative force and gives added insight regarding energy and energy transformation without the necessity of dealing with the force directly. It is much more common, for example, to use the concept of voltage (related to electric potential energy) than to deal with the Coulomb force directly.

Calculating the work directly is generally difficult, since  $W = Fd \cos \theta$  and the direction and magnitude of  $F$  can be complex for multiple charges, for odd-shaped objects, and along arbitrary paths. But we do know that, since  $F = qE$ , the work, and hence  $\Delta PE$ , is proportional to the test charge  $q$ . To have a physical quantity that is independent of test charge, we define **electric potential**  $V$  (or simply potential, since electric is understood) to be the potential energy per unit charge:

$$V = \frac{PE}{q}. \quad \text{Eq.(1)}$$

### ELECTRIC POTENTIAL

This is the electric potential energy per unit charge.

$$V = \frac{PE}{q} \quad \text{Eq.(2)}$$

Since PE is proportional to  $q$ , the dependence on  $q$  cancels. Thus  $V$  does not depend on  $q$ . The change in potential energy  $\Delta PE$

is crucial, and we are concerned with the difference in potential or potential difference  $\Delta V$  between two points, where

$$\Delta V = V_B - V_A = \frac{\Delta PE}{q}. \quad \text{Eq.(3)}$$

The **potential difference** between points A and B,  $V_B - V_A$ , is thus defined to be the change in potential energy of a charge  $q$  moved from A to B, divided by the charge. Units of potential difference are joules per coulomb, given the name volt (V) after Alessandro Volta.

$$1V = 1 \frac{J}{C} \quad \text{Eq.(4)}$$

### POTENTIAL DIFFERENCE

The potential difference between points A and B,  $V_B - V_A$ , is defined to be the change in potential energy of a charge  $q$  moved from A to B, divided by the charge. Units of potential difference are joules per coulomb, given the name volt (V) after Alessandro Volta.

$$1V = 1 \frac{J}{C} \quad \text{Eq.(5)}$$

The familiar term **voltage** is the common name for potential difference. Keep in mind that whenever a voltage is quoted, it is understood to be the potential difference between two points. For example, every battery has two terminals, and its voltage is the potential difference between them. More fundamentally, the point you choose to be zero volts is arbitrary. This is analogous to the fact that gravitational potential energy has an arbitrary zero, such as sea level or perhaps a lecture hall floor.

In summary, the relationship between potential difference (or voltage) and electrical potential energy is given by

$$\Delta V = \frac{\Delta PE}{q} \text{ and } \Delta PE = q\Delta V. \quad \text{Eq.(6)}$$

### POTENTIAL DIFFERENCE AND ELECTRICAL POTENTIAL ENERGY

The relationship between potential difference (or voltage) and electrical potential energy is given by

$$\Delta V = \frac{\Delta PE}{q} \text{ and } \Delta PE = q\Delta V. \quad \text{Eq.(7)}$$

The second equation is equivalent to the first.

Voltage is not the same as energy. Voltage is the energy per unit charge. Thus a motorcycle battery and a car battery can both have the same voltage (more precisely, the same potential difference between battery terminals), yet one stores much more energy than the other since  $\Delta PE = q\Delta V$ . The car battery can move more charge than the motorcycle battery, although both are 12 V batteries.

**Example 1: Calculating Energy**

Suppose you have a 12.0 V motorcycle battery that can move 5000 C of charge, and a 12.0 V car battery that can move 60 000 C of charge. How much energy does each deliver? (Assume that the numerical value of each charge is accurate to three significant figures.)

**Strategy**

To say we have a 12.0 V battery means that its terminals have a 12.0 V potential difference. When such a battery moves charge, it puts the charge through a potential difference of 12.0 V, and the charge is given a change in potential energy equal to  $\Delta PE = q\Delta V$ .

So to find the energy output, we multiply the charge moved by the potential difference.

**Solution**

For the motorcycle battery,  $q = 5000\text{C}$  and  $\Delta V = 12.0\text{V}$ . The total energy delivered by the motorcycle battery is

$$\begin{aligned}\Delta PE_{\text{cycle}} &= (5000\text{C})(12.0\text{V}) \\ &= (5000\text{C})(12.0\text{J/C}) \\ &= 6.00 \times 10^4\text{J}.\end{aligned}\quad \text{Eq.(8)}$$

Similarly, for the car battery,  $q = 60000\text{C}$  and

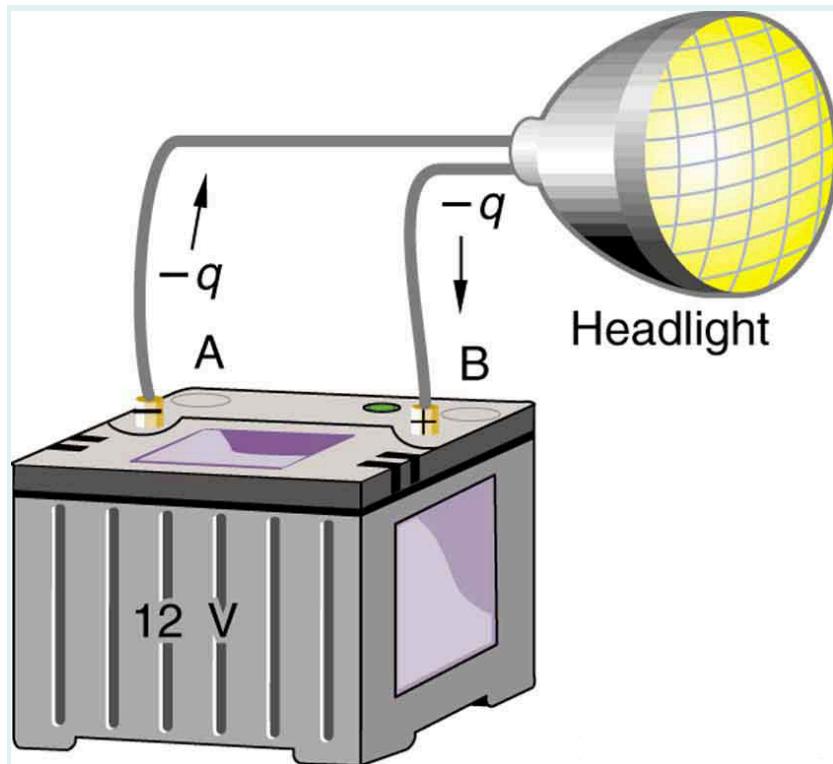
$$\begin{aligned}\Delta PE_{\text{car}} &= (60000\text{C})(12.0\text{V}) \\ &= 7.20 \times 10^5\text{J}.\end{aligned}\quad \text{Eq.(9)}$$

**Discussion**

While voltage and energy are related, they are not the same thing. The voltages of the batteries are identical, but the energy supplied by each is quite different. Note also that as a battery is discharged, some of its energy is used internally and its terminal voltage drops, such as when headlights dim because of a low car battery. The energy supplied by the battery is still calculated as in this example, but not all of the energy is available for external use.

Note that the energies calculated in the previous example are absolute values. The change in potential energy for the battery is negative, since it loses energy. These batteries, like many electrical systems, actually move negative charge—electrons in particular. The batteries repel electrons from their negative terminals (A) through

whatever circuitry is involved and attract them to their positive terminals (B) as shown in [Figure 2]. The change in potential is  $\Delta V = V_B - V_A = +12V$  and the charge  $q$  is negative, so that  $\Delta PE = q\Delta V$  is negative, meaning the potential energy of the battery has decreased when  $q$  has moved from A to B.



**Figure 2:** A battery moves negative charge from its negative terminal through a headlight to its positive terminal. Appropriate combinations of chemicals in the battery separate charges so that the negative terminal has an excess of negative charge, which is repelled by it and attracted to the excess positive charge on the other terminal. In terms of potential, the positive terminal is at a higher voltage than the negative. Inside the battery, both positive and negative charges move.

### Example 2: How Many Electrons Move through a Headlight Each Second?

When a 12.0 V car battery runs a single 30.0 W headlight, how many electrons pass through it each second?

#### Strategy

To find the number of electrons, we must first find the charge that moved in 1.00 s. The charge moved is related to voltage and energy through the equation  $\Delta PE = q\Delta V$ . A 30.0 W lamp uses 30.0 joules per second. Since the battery loses energy, we have  $\Delta PE = -30.0 \text{ J}$  and, since the electrons are going from the negative terminal to the positive, we see that  $\Delta V = +12.0 \text{ V}$ .

#### Solution

To find the charge  $q$  moved, we solve the equation  $\Delta PE = q\Delta V$ :

$$q = \frac{\Delta PE}{\Delta V}. \quad \text{Eq.(10)}$$

Entering the values for  $\Delta PE$  and  $\Delta V$ , we get

$$q = \frac{-30.0\text{J}}{+12.0\text{V}} = \frac{-30.0\text{J}}{+12.0\text{J/C}} = -2.50\text{C}. \quad \text{Eq.(11)}$$

The number of electrons  $n_e$  is the total charge divided by the charge per electron. That is,

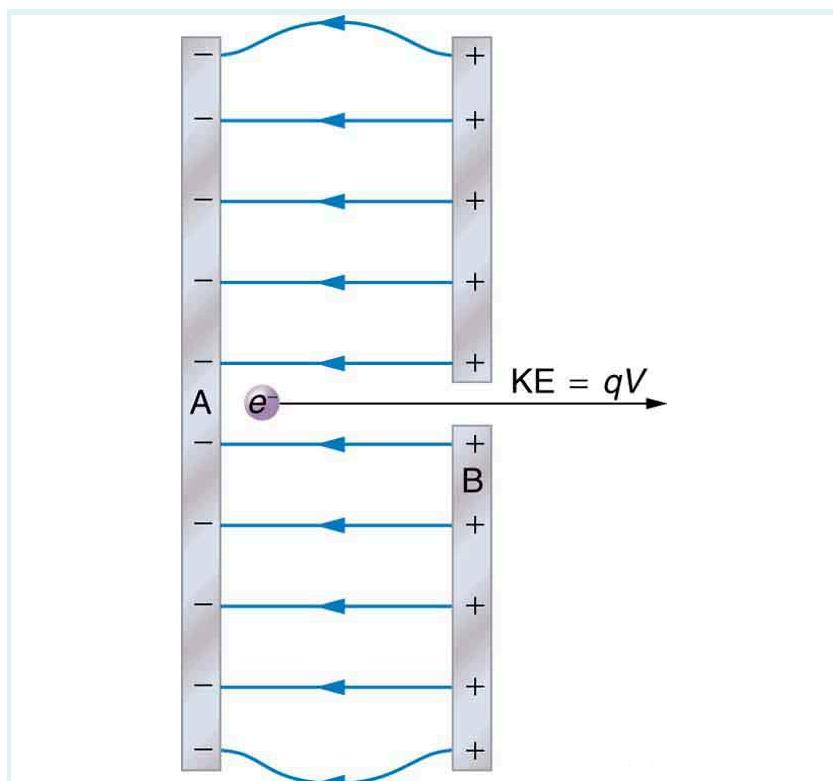
$$n_e = \frac{-2.50\text{C}}{-1.60 \times 10^{-19}\text{C/e}^-} = 1.56 \times 10^{19}\text{electrons.} \quad \text{Eq.(12)}$$

### Discussion

This is a very large number. It is no wonder that we do not ordinarily observe individual electrons with so many being present in ordinary systems. In fact, electricity had been in use for many decades before it was determined that the moving charges in many circumstances were negative. Positive charge moving in the opposite direction of negative charge often produces identical effects; this makes it difficult to determine which is moving or whether both are moving.

## The Electron Volt

The energy per electron is very small in macroscopic situations like that in the previous example—a tiny fraction of a joule. But on a submicroscopic scale, such energy per particle (electron, proton, or ion) can be of great importance. For example, even a tiny fraction of a joule can be great enough for these particles to destroy organic molecules and harm living tissue. The particle may do its damage by direct collision, or it may create harmful X-rays, which can also inflict damage. It is useful to have an energy unit related to submicroscopic effects. [Figure 3] shows a situation related to the definition of such an energy unit. An electron is accelerated between two charged metal plates as it might be in an old-model television tube or oscilloscope. The electron is given kinetic energy that is later converted to another form—light in the television tube, for example. (Note that downhill for the electron is uphill for a positive charge.) Since energy is related to voltage by  $\Delta PE = q\Delta V$ , we can think of the joule as a coulomb-volt.



**Figure 3:** A typical electron gun accelerates electrons using a potential difference between two metal plates. The energy of the electron in electron volts is numerically the same as the voltage between the plates. For example, a 5000 V potential difference produces 5000 eV electrons.

On the submicroscopic scale, it is more convenient to define an energy unit called the **electron volt** (eV), which is the energy given to a fundamental charge accelerated through a potential difference of 1 V. In equation form,

$$\begin{aligned} 1\text{eV} &= (1.60 \times 10^{-19}\text{C})(1\text{V}) = (1.60 \times 10^{-19}\text{C})(1\text{J/C}) \\ &= 1.60 \times 10^{-19}\text{J}. \end{aligned} \quad \text{Eq.(13)}$$

**ELECTRON VOLT**

On the submicroscopic scale, it is more convenient to define an energy unit called the electron volt (eV), which is the energy given to a fundamental charge accelerated through a potential difference of 1 V. In equation form,

$$\begin{aligned} 1\text{eV} &= (1.60 \times 10^{-19}\text{C}) (1\text{V}) = (1.60 \times 10^{-19}\text{C}) (1\text{J/C}) \\ &= 1.60 \times 10^{-19}\text{J}. \end{aligned} \quad \text{Eq.(14)}$$

An electron accelerated through a potential difference of 1 V is given an energy of 1 eV. It follows that an electron accelerated through 50 V is given 50 eV. A potential difference of 100 000 V (100 kV) will give an electron an energy of 100 000 eV (100 keV), and so on. Similarly, an ion with a double positive charge accelerated through 100 V will be given 200 eV of energy. These simple relationships between accelerating voltage and particle charges make the electron volt a simple and convenient energy unit in such circumstances.

**CONNECTIONS: ENERGY UNITS**

The electron volt (eV) is the most common energy unit for submicroscopic processes. This will be particularly noticeable in the chapters on modern physics. Energy is so important to so many subjects that there is a tendency to define a special energy unit for each major topic. There are, for example, calories for food energy, kilowatt-hours for electrical energy, and therms for natural gas energy.

The electron volt is commonly employed in submicroscopic processes—chemical valence energies and molecular and nuclear binding energies are among the quantities often expressed in electron volts. For example, about 5 eV of energy is required to break up certain organic molecules. If a proton is accelerated from rest through a potential difference of 30 kV, it is given an energy of 30 keV (30 000 eV) and it can break up as many as 6000 of these molecules ( $30000\text{eV} \div 5\text{eV per molecule} = 6000\text{molecules}$ ). Nuclear decay energies are on the order of 1 MeV (1 000 000 eV) per event and can, thus, produce significant biological damage.

**Conservation of Energy**

The total energy of a system is conserved if there is no net addition (or subtraction) of work or heat transfer. For conservative forces, such as the electrostatic force, conservation of energy states that mechanical energy is a constant.

**Mechanical energy** is the sum of the kinetic energy and potential energy of a system; that is,  $KE+PE = \text{constant}$ . A loss of PE of a charged particle becomes an increase in its KE. Here PE is the electric potential energy. Conservation of energy is stated in equation form as

Eq.(15)

$$KE+PE = \text{constant}$$

or

Eq.(16)

$$KE_i + PE_i = KE_f + PE_f,$$

where i and f stand for initial and final conditions. As we have found many times before, considering energy can give us insights and facilitate problem-solving.

**Example 3: Electrical Potential Energy Converted to Kinetic Energy**

Calculate the final speed of a free electron accelerated from rest through a potential difference of 100 V. (Assume that this numerical value is accurate to three significant figures.)

**Strategy**

We have a system with only conservative forces. Assuming the electron is accelerated in a vacuum, and neglecting the gravitational force (we will check on this assumption later), all of the electrical potential energy is converted into kinetic energy. We can identify the initial and final forms of energy to be  $\text{KE}_i = 0$ ,  $\text{KE}_f = \frac{1}{2}mv^2$ ,  $\text{PE}_i = qV$ , and  $\text{PE}_f = 0$ .

**Solution**

Conservation of energy states that

$$\text{KE}_i + \text{PE}_i = \text{KE}_f + \text{PE}_f. \quad \text{Eq.(17)}$$

Entering the forms identified above, we obtain

$$qV = \frac{mv^2}{2}. \quad \text{Eq.(18)}$$

We solve this for  $v$ :

$$v = \sqrt{\frac{2qV}{m}}. \quad \text{Eq.(19)}$$

Entering values for  $q$ ,  $V$ , and  $m$  gives

$$\begin{aligned} v &= \sqrt{\frac{2(-1.60 \times 10^{-19} \text{C})(-100 \text{J/C})}{9.11 \times 10^{-31} \text{kg}}} \\ v &= 5.93 \times 10^6 \text{ m/s.} \end{aligned} \quad \text{Eq.(20)}$$

**Discussion**

Note that both the charge and the initial voltage are negative, as in [Figure 3]. From the discussions in [Electric Charge and Electric Field](#), we know that electrostatic forces on small particles are generally very large compared with the gravitational force. The large final speed confirms that the gravitational force is indeed negligible here. The large speed also indicates how easy it is to

accelerate electrons with small voltages because of their very small mass. Voltages much higher than the 100 V in this problem are typically used in electron guns. Those higher voltages produce electron speeds so great that relativistic effects must be taken into account. That is why a low voltage is considered (accurately) in this example.

## Section Summary

- Electric potential is potential energy per unit charge.
- The potential difference between points A and B,  $V_B - V_A$ , defined to be the change in potential energy of a charge  $q$  moved from A to B, is equal to the change in potential energy divided by the charge, Potential difference is commonly called voltage, represented by the symbol  $\Delta V$ .

$$\Delta V = \frac{\Delta PE}{q} \text{ and } \Delta PE = q\Delta V. \quad \text{Eq.(21)}$$

- An electron volt is the energy given to a fundamental charge accelerated through a potential difference of 1 V. In equation form,

$$\begin{aligned} 1\text{eV} &= (1.60 \times 10^{-19}\text{C})(1\text{V}) = (1.60 \times 10^{-19}\text{C})(1\text{J/C}) \\ 1\text{eV} &= 1.60 \times 10^{-19}\text{J}. \end{aligned} \quad \text{Eq.(22)}$$

- Mechanical energy is the sum of the kinetic energy and potential energy of a system, that is,  $KE + PE$ . This sum is a constant.

## Conceptual Questions

### Exercise 1

Voltage is the common word for potential difference. Which term is more descriptive, voltage or potential difference?

**Exercise 2**

If the voltage between two points is zero, can a test charge be moved between them with zero net work being done? Can this necessarily be done without exerting a force? Explain.

**Exercise 3**

What is the relationship between voltage and energy? More precisely, what is the relationship between potential difference and electric potential energy?

**Exercise 4**

Voltages are always measured between two points. Why?

**Exercise 5**

How are units of volts and electron volts related? How do they differ?

## Problems & Exercises

**Exercise 6**

Find the ratio of speeds of an electron and a negative hydrogen ion (one having an extra electron) accelerated through the same voltage, assuming non-relativistic final speeds. Take the mass of the hydrogen ion to be  $1.67 \times 10^{-27} \text{ kg}$ .

**Exercise 7**

An evacuated tube uses an accelerating voltage of 40 kV to accelerate electrons to hit a copper plate and produce X-rays. Non-relativistically, what would be the maximum speed of these electrons?

**Exercise 8**

A bare helium nucleus has two positive charges and a mass of  $6.64 \times 10^{-27}$  kg. (a) Calculate its kinetic energy in joules at 2.00% of the speed of light. (b) What is this in electron volts? (c) What voltage would be needed to obtain this energy?

**Exercise 9****Integrated Concepts**

Singly charged gas ions are accelerated from rest through a voltage of 13.0 V. At what temperature will the average kinetic energy of gas molecules be the same as that given these ions?

**Exercise 10****Integrated Concepts**

The temperature near the center of the Sun is thought to be 15 million degrees Celsius ( $1.5 \times 10^7$  °C). Through what voltage must a singly charged ion be accelerated to have the same energy as the average kinetic energy of ions at this temperature?

**Exercise 11****Integrated Concepts**

(a) What is the average power output of a heart defibrillator that dissipates 400 J of energy in 10.0 ms? (b) Considering the high-power output, why doesn't the defibrillator produce serious burns?

**Exercise 12****Integrated Concepts**

A lightning bolt strikes a tree, moving 20.0 C of charge through a potential difference of  $1.00 \times 10^2$  MV. (a) What energy was dissipated? (b) What mass of water could be raised from  $15^\circ\text{C}$  to the boiling point and then boiled by this energy? (c) Discuss the damage that could be caused to the tree by the expansion of the boiling steam.

**Exercise 13****Integrated Concepts**

A 12.0 V battery-operated bottle warmer heats 50.0 g of glass,  $2.50 \times 10^2$  g of baby formula, and  $2.00 \times 10^2$  g of aluminum from  $20.0^\circ\text{C}$  to  $90.0^\circ\text{C}$ . (a) How much charge is moved by the battery? (b) How many electrons per second flow if it takes 5.00 min to warm the formula? (Hint: Assume that the specific heat of baby formula is about the same as the specific heat of water.)

**Exercise 14****Integrated Concepts**

A battery-operated car utilizes a 12.0 V system. Find the charge the batteries must be able to move in order to accelerate the 750 kg car from rest to 25.0 m/s, make it climb a  $2.00 \times 10^2$ m high hill, and then cause it to travel at a constant 25.0 m/s by exerting a  $5.00 \times 10^2$ N force for an hour.

**Exercise 15****Integrated Concepts**

Fusion probability is greatly enhanced when appropriate nuclei are brought close together, but mutual Coulomb repulsion must be overcome. This can be done using the kinetic energy of high-temperature gas ions or by accelerating the nuclei toward one another. (a) Calculate the potential energy of two singly charged nuclei separated by  $1.00 \times 10^{-12}$ m by finding the voltage of one at that distance and multiplying by the charge of the other. (b) At what temperature will atoms of a gas have an average kinetic energy equal to this needed electrical potential energy?

**Exercise 16****Unreasonable Results**

- (a) Find the voltage near a 10.0 cm diameter metal sphere that has 8.00 C of excess positive charge on it. (b) What is unreasonable about this result? (c) Which assumptions are responsible?

**Exercise 17****Construct Your Own Problem**

Consider a battery used to supply energy to a cellular phone. Construct a problem in which you determine the energy that must be supplied by the battery, and then calculate the amount of charge it must be able to move in order to supply this energy. Among the things to be considered are the energy needs and battery voltage. You may need to look ahead to interpret manufacturer's battery ratings in ampere-hours as energy in joules.

**GLOSSARY*****electric potential***

potential energy per unit charge

***potential difference (or voltage)***

change in potential energy of a charge moved from one point to another, divided by the charge; units of potential difference are joules per coulomb, known as volt

***electron volt***

the energy given to a fundamental charge accelerated through a potential difference of one volt

***mechanical energy***

sum of the kinetic energy and potential energy of a system; this sum is a constant



(<https://creativecommons.org/licenses/by/4.0/>)

This work is licensed under a Creative Commons Attribution 4.0 International License (<https://creativecommons.org/licenses/by/4.0/>).