PARTE 2 VOLUMEN ESTEAUIOMETRIA		
PARTE 2 VOLUMEN ESTERUIOMETALA [m] Bull -> Ball + 1 OL 0=16 Bull -> Ball + 2 OL Ba = 137. 105 Ball x Irol Ball 2 0,5 mol 0 = 0,03 mol 0	1 P ZNH3+	0c -> NO + 2H20
105 Balle x Irol Balle 2 0,5 mol 0 = 0,03 mdes	2NH3+50	2-> 2ND + 3H20
169,38 1mol Bale	30°C= 303k	650 mm 18 = 0,855 atm
27°C=300K PV=MRT P= MRT = 0,01:0,072:300	1 n= 1/2 = 0	1.20 = 0,805 md
P=0,738 afa 1	10,805m-1 H20	x 2 mol NHs = 0,537 and NHs
rul 2KNO3+3C+5 -> K2S+CUL + N2		0/822 = 72/11
$2kNO_3 + 3C + S \longrightarrow kiS + 3\omega_L + N_L$	1	01822
50g KNU3x 1md KNU3 x 3md Q, = 0,742 mol		
C.N - T=237K PV=NRT		
C.n = T=237x PV=NRT P=1atin V= nRT = 0,742.0,087.21		
V=16,612		
[1] 6002+6H20 > 6H120,+602		
n= N - 1.10 = 0,447 mol. CO2		
0,447 mol Ocx 6 mol Oz x 328 - 14,304g		
1C) H2SOY + NECl -> HCl+ NE HSOY		
n= PV - 1.2 = 0,089 moles HC1		
0,089 mol HCI x 1mol Nacl x 58,55 - 5,25		
[i] H2SOy + Nacl -> HCl + Na HSOY		
26 HCl = 0,089 moles		
0,089 moles HC/x InvolH259 x Trs I and H2504		
= 8,746		