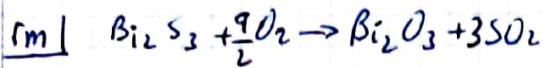


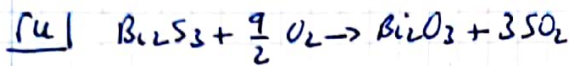
TENAS ESTEQUIOMETRIA

PARTE I MASA-MASA



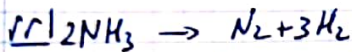
$\text{Bi} = 209 \quad \text{S} = 32 \quad \text{O} = 16$

$$1000\text{g Bi}_2\text{S}_3 \times \frac{1\text{mol Bi}_2\text{S}_3}{514\text{g}} \times \frac{3\text{mol SO}_2}{1\text{mol Bi}_2\text{S}_3} \times \frac{64\text{g}}{1\text{mol SO}_2} = 373,54\text{g}$$

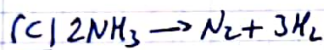


$$5\text{mol Bi}_2\text{S}_3 \times \frac{4,5\text{mol O}_2}{1\text{mol Bi}_2\text{S}_3} \times \frac{32\text{g}}{1\text{mol O}_2} =$$

720g

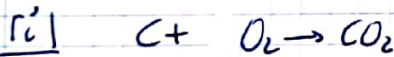


$$68\text{g NH}_3 \times \frac{1\text{mol NH}_3}{17\text{g}} \times \frac{3\text{mol H}_2}{2\text{mol NH}_3} \times \frac{2\text{g}}{1\text{mol H}_2} = 12\text{g}$$



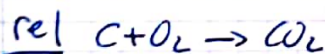
$$68\text{g NH}_3 \times \frac{1\text{mol NH}_3}{17\text{g}} \times \frac{3\text{mol H}_2}{2\text{mol NH}_3} \times \frac{6,022 \times 10^{23} \text{ molecules}}{1\text{mol H}_2} =$$

$3,61 \cdot 10^{24}$ molecules



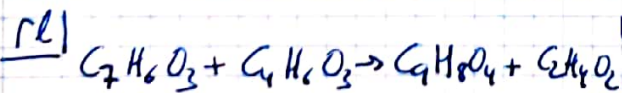
$$10\text{g C} \times \frac{1\text{mol C}}{12\text{g}} \times \frac{1\text{mol O}_2}{1\text{mol C}} \times \frac{32\text{g}}{1\text{mol O}_2} =$$

26,67g



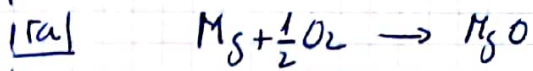
$$10\text{g C} \times \frac{1\text{mol C}}{12\text{g}} \times \frac{1\text{mol CO}_2}{1\text{mol C}} \times \frac{44\text{g}}{1\text{mol CO}_2} =$$

36,67g



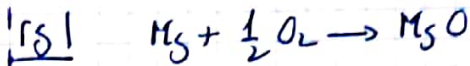
$$50\text{g C}_9\text{H}_8\text{O}_4 \times \frac{1\text{mol C}_9\text{H}_8\text{O}_4}{180\text{g}} \times \frac{1\text{mol C}_7\text{H}_6\text{O}_3}{1\text{mol C}_9\text{H}_8\text{O}_4} \times$$

$$\times \frac{138\text{g}}{1\text{mol C}_7\text{H}_6\text{O}_3} = \underline{\underline{38,33\text{g}}}$$



$$200\text{g Mg} \times \frac{1\text{mol Mg}}{24,3\text{g}} \times \frac{1\text{mol MgO}}{1\text{mol Mg}} \times \frac{40,3\text{g}}{1\text{mol MgO}} =$$

331,69g MgO



$$200\text{g Mg} \times \frac{1\text{mol Mg}}{24,3\text{g}} \times \frac{0,5\text{mol O}_2}{1\text{mol Mg}} \times \frac{32\text{g}}{1\text{mol O}_2} =$$

131,68g