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Modern Chemistry 1 Solutions CHAPTER 12 REVIEW Solutions Teacher Notes and Answers Chapter 12 SECTION 1 SHORT ANSWER 1. c 2. a 3. b 2. a. alcohol b. water c. the gels 3. The mixture is a colloid. The properties are consistent with those reported in Table 3 on page 404 of the text. The particle size is small, but not too small, and the mixture

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Modern Chemistry • CHAPTER 12 HOMEWORK 12-5 (pp. 379–381) ... MODERN CHEMISTRY SECTION 12-3 REVIEW 103 HRW material copyrighted under notice appearing earlier in this work. Name Date Class . SECTION 12-3 continued b. Explain the difference between an open system and a closed system. 5.

HOMEWORK 12-3 - Los Banos Unified School District

Modern Chemistry 88 Chapter Test Chapter: States of Matter PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

____ 1. According to the kinetic-molecular theory, particles of an ideal gas a. attract each other but do not collide. b. repel each other and collide.

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Modern Chemistry 109 Chapter Test Chapter: Solutions PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

____ 1. Molecules whose water solutions can carry electric current a. are nonpolar. b. ionize in water. c. do not dissolve in water. d. decompose in water ...

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CHAPTER 1 REVIEW Matter and Change MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Classify each of the following as a homogeneous or heterogeneous substance. a. sugar d. plastic wrap b. iron filings e. cement sidewalk c. granola bar 2. For each type of investigation, select the most appropriate branch of chemistry from the following

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Modern Chemistry 84 Chapter Test Chapter: States of Matter In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. ____ 1. The kinetic-molecular theory of gases assumes that the particles of an ideal gas are separated by great distances. This implies that the gas

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