

## *Modern Chemistry Chapter Test B Answers*

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### **Modern Chemistry Chapter Test B**

Modern Chemistry 34 Chapter Test Name Class Date Chapter Test B, continued PART IV Write the answers to the following questions in the space provided. 27. What two early twentieth-century observations involving the interaction of light and matter could not be explained solely by the wave theory of light? 28.

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Modern Chemistry 88 Chapter Test Chapter: States of Matter PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

\_\_\_\_ 1. According to the kinetic-molecular theory, particles of an ideal gas a. attract each other but do not collide. b. repel each other and collide.

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Modern Chemistry 141 Chapter Test Name Class Date Chapter Test B, continued 18. When a strong acid is titrated with a weak base, the pH of the solution at the equivalence point is \_\_\_\_\_. 19. A is a highly purified solid used to check the concentration of a standard solution. 20. A 1 M solution of NaOH will have a pH that is \_\_\_\_.

### **Assessment Chapter Test B - Baumapedia**

Holt McDougal Modern Chemistry 3 Chapter Test Chapter Test B, continued 16. The measure of the ability of an atom in a chemical compound to attract electrons from another atom in the compound is called \_\_\_\_\_. 17. The energy required to remove one electron from an atom is called its \_\_\_\_\_. 18.

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### **Modern Chemistry: Chapter Tests with Answer Key, 2006 ...**

Modern Chemistry 136 Chapter Test Name Class Date Chapter Test A, continued \_\_\_\_ 14. A pH meter measures the pH of a solution by measuring the a. amount of charge on the hydronium ion. b. color of the solution. c. concentration of the solution. d. voltage difference between two electrodes in the meter's probe.

### **Assessment Chapter Test A - Baumapedia**

Holt McDougal Modern Chemistry 2 Chapter Test Chapter Test A, continued \_\_\_\_ 3. The periodic law states that a. no two electrons with the same spin can be found in the same place in an atom. b. the physical and chemical properties of the elements are functions of their atomic number. c. wave patterns repeat at regular intervals.

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CHAPTER 10 REVIEW States of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Match description on the right to the correct crystal type on the left. b ionic crystal (a) has mobile electrons in the crystal c covalent molecular crystal (b) is hard, brittle, and

nonconducting a metallic crystal (c) typically has the lowest melting point of the four

### 10 States of Matter - Ms. Agostine's Chemistry Page

Modern Chemistry 47 Chapter Test Name Class Date Chapter Test A, continued \_\_\_\_ 7. The melting points of ionic compounds are higher than the melting points of molecular compounds because a. ionic substances tend to vaporize at room temperature. b. ionic substances are brittle. c. attractive forces between ions are greater than the attractive ...

### Assessment Chapter Test A - Kettering City School District

Modern Chemistry 56 Chapter Test Assessment Chapter Test B Chapter: Chemical Equations and Reactions PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. \_\_\_\_ 1. The production of a slightly soluble solid compound in a double-

### Assessment Chapter Test B - Hatboro

CHAPTER 3 TEST continued Date Class FILL IN THE BLANK Write the correct term (or terms) in the space provided. 9, If a particular compound is composed of elements A and B, the ratio of the mass of B to the mass of A will always be the same. This is a statement of the law of exactly 12 g of carbon-12 is referred to as a(n) 11.

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(b) It will be more difficult to break the bonds in compound A than those in compound B. (c) Compound B has larger crystals than compound A. (d) Compound A has larger crystals than compound B. 5. b The forces of attraction between molecules in a molecular compound are generally (a) stronger than the attractive forces among formula units in ...

### 6 Chemical Bonding - Effingham County Schools / Overview

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CHAPTER 1 REVIEW Matter and Change MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Classify each of the following as a homogeneous or heterogeneous substance. a. sugar d. plastic wrap b. iron filings e. cement sidewalk c. granola bar 2. For each type of investigation, select the most appropriate branch of chemistry from the following

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CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Write formulas for the following compounds: CuCO<sub>3</sub> a. copper(II) carbonate Na<sub>2</sub>SO<sub>3</sub> b. sodium sulfite (NH<sub>4</sub>)<sub>3</sub>PO<sub>4</sub> c. ammonium phosphate SnS<sub>2</sub> d. tin(IV) sulfide HNO<sub>2</sub> e. nitrous acid 2.

### 7 Chemical Formulas and Chemical Compounds

Holt Chemistry 40 Matter and Energy Section: Measurements and Calculations in Chemistry In the space provided, write the letter of the term or phrase that best answers the question. \_\_\_\_ 1. A

measurement is said to have good precision if it a. agrees closely with an accepted standard. b. agrees closely with other measurements of the same quantity.

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