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Modern Chemistry 111 Chapter Test Name Class Date Chapter Test B, continued 18. Dissolution processes with negative enthalpies of solution are processes. 19. is the solution process with water as the solvent. 20. Liquid solutes and solvents that are not soluble in each other are. 21. A solute molecule that is surrounded by solvent molecules is ...

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Modern Chemistry 84 Chapter Test Chapter: States of Matter In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. \_\_\_\_ 1. The kinetic-molecular theory of gases assumes that the particles of an ideal gas are separated by great distances. This implies that the gas

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Modern Chemistry 1 Solutions CHAPTER 12 REVIEW Solutions Teacher Notes and Answers Chapter 12 SECTION 1 SHORT ANSWER 1. c 2. a 3. b 2. a. alcohol b. water c. the gels 3. The mixture is a colloid. The properties are consistent with those reported in Table 3 on page 404 of the text. The particle size is small, but not too small, and the mixture

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each question. \_\_\_\_ 1.

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Modern Chemistry 91 Chapter Test Name Class Date Chapter Test B, continued 27. Distinguish between ionic crystals and metallic crystals. PART V Write the answers to the following questions on the line to the left, and show your work in the space provided. The molar enthalpy of fusion for water is 6.008 kJ/mol. 28.

**Assessment Chapter Test B - clarkchargers.org**

CHAPTER 1 REVIEW Matter and Change MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Classify each of the following as a homogeneous or heterogeneous substance. a. sugar d. plastic wrap b. iron filings e. cement sidewalk c. granola bar 2. For each type of investigation, select the most appropriate branch of chemistry from the following

**mc06se cFMsr i-vi**

CHAPTER 10 REVIEW States of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Match description on the right to the correct crystal type on the left. b ionic crystal (a) has mobile electrons in the crystal c covalent molecular crystal (b) is hard, brittle, and nonconducting a metallic crystal (c) typically has the lowest melting point of the four

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12. An unknown compound is a solid that has a melting point of approximately 300°C and is: F. an ionic compound. H. a network solid. G. a covalent compound. J. There is not enough information to determine the type of compound. 13. Which of the following statements explains why ionic compounds generally conduct an electrical current? A.

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