

## *Modern Chemistry Chapter 14 Section 3 Review Answers*

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Modern Chemistry 120 Acids and Bases SECTION 2 continued 4. a. Write the equation for the first ionization of  $\text{H}_2\text{CO}_3$  in aqueous solution. Assume that water serves as the reactant that attaches to the hydrogen ion released from the  $\text{H}_2\text{CO}_3$ . Which of the reactants is the Brønsted-Lowry ... CHAPTER 14 REVIEW Acids and Bases SECTION 3

### CHAPTER 14 REVIEW Acids and Bases - Manasquan Public Schools

CHAPTER 14 REVIEW Ions in Aqueous Solutions and Colligative Properties SECTION 14-1 SHORT ANSWER Answer the following questions in the space provided. 1. Use the guidelines in Table 14-1 on page 427 of the text to predict the solubility of the following

### 14 Ions in Aqueous Solutions and Colligative Properties

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SECTION 3 continued 4. Table 6 on page 485 of the text lists several amphoteric species, but only one other than water is neutral.  $\text{NH}_3$  a. Identify that neutral compound. b. Write two equations that demonstrate this compound's amphoteric properties. (Answers will vary, but one equation should form  $\text{NH}_4^+$ )

### 14 Acids and Bases - mchsapchemistry.com

CHAPTER 1 REVIEW Matter and Change MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Classify each of the following as a homogeneous or heterogeneous substance. a. sugar d. plastic wrap b. iron filings e. cement sidewalk c. granola bar 2. For each type of investigation, select the most appropriate branch of chemistry from the following

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### **New Page 1 [srvhs.org]**

CHAPTER 14 REVIEW Ions in Aqueous Solutions and Colligative Properties MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Match the four compounds on the right to their descriptions on the left.

### **14 Ions in Aqueous Solutions and Colligative Properties**

Modern Chemistry Chapter 15 Test Review 25 multiple choice definition of self-ionization  $[H_3O^+] [OH^-] = 1 \times 10^{-14}$  definition and formula for calculating  $pH = -\log[H_3O^+]$   $pH > 7$  is a base;  $pH < 7$  is an acid;  $pH = 7$  is neutral  $pH$  range is normally 0 to 14 calculate  $pH$  from  $[H_3O^+]$  calculate  $[H_3O^+]$  from  $pH$  definitions of indicators, transition ...

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SECTION SUMMARY. 14.3 Ideal Gases Summary: The ideal gas law permits you to solve for the number of moles of a contained gas when the pressure, volume, and ... Chapter 7 Acids, Bases, and Solutions Section 3 Summary ...

### **Section Summary Chapter 10 Answer Key**

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