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Modern Chemistry 2 Stoichiometry CHAPTER 9 REVIEW Stoichiometry MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Given the following equation: C ... Modern Chemistry 3 Stoichiometry MIXED REVIEW continued c. If 0.1 mol of N<sub>2</sub> combine with H<sub>2</sub>, what must be true about the quantity of H

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### **14 Ions in Aqueous Solutions and Colligative Properties**

Modern Chemistry • CHAPTER 9 HOMEWORK 9-1 (pp. 275–279) VOCABULARY Define. ... Modern Chemistry • CHAPTER 9 HOMEWORK 9-2 (pp. 280–282) VOCABULARY Complete each sentence. 1. The mole ratio is obtained from \_\_\_\_\_. 2. Theoretical stoichiometric calculations show \_\_\_\_\_. ...

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Let's try a real chemistry problem and use this concept to determine the limiting reagent and see how it determines how much product will be formed. Problem : Nitrogen gas reacts with hydrogen gas to produce gaseous ammonia ( $\text{NH}_3$ ). If we react 25.0 g of nitrogen with 25.0 g of hydrogen, how much ammonia will be produced.

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