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if two or more different compounds are composed of the same two elements, then the ratio of the masses of the second element combined with a certain mass of the first element is always a ratio of small whole numbers; textbook: if two or more different compounds are composed of the same two elements with a certain mass of the first element is always a ratio of small whole numbers

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see figure 3-3 page 67 . C: Modern Atomic Theory. Some of Dalton's postulates have been proven to be incorrect. His third postulate indicated that atoms cannot be subdivided. His second postulate indicates atoms of the same element have identical mass. Isotopes. His theory has been modified -- not discarded. Homework: Section 3-1

### Chapter 3 Modern Chemistry - Prepchem

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CHAPTER 3 REVIEW Atoms: The Building Blocks of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Explain the difference between the mass number and the atomic number of a nuclide. Mass number is the total number of protons and neutrons in the nucleus of an isotope.

### 3 Atoms: The Building Blocks of Matter - Ms. Sheehan's ...

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### 3 Atoms: The Building Blocks of Matter

CHAPTER 10 REVIEW States of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Match description on the right to the correct crystal type on the left. ... SECTION 5 MODERN CHEMISTRY STATES OF MATTER. STATES OF MATTER MODERN CHEMISTRY.

### 10 States of Matter - Ms. Agostine's Chemistry Page

Modern Chemistry 1 The Periodic Law CHAPTER 5 The Periodic Law SECTION 1 History of the Periodic Table OBJECTIVES 1. Explain the roles of Mendeleev and Moseley in the development of the periodic table 2. Describe the modern periodic table. 3.

### CHAPTER 5 The Periodic Law - mchsapchemistry.com

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MODERN CHEMISTRY STOICHIOMETRY 73 ... CHAPTER 9 REVIEW Stoichiometry SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. b The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. ... mc06se\_cFMsr\_i-vi.qxd  
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CHAPTER 6 REVIEW Chemical Bonding SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. a The notation for sodium chloride, NaCl, stands for one (a) formula unit. (c) crystal. (b) molecule. (d) atom. 2. d In a crystal of an ionic compound, each cation is surrounded by a number of (a) molecules. (c) dipoles. (b) ...

### 6 Chemical Bonding - Effingham County Schools / Overview

Modern Chemistry 1 Solutions CHAPTER 12 REVIEW Solutions Teacher Notes and Answers Chapter 12 SECTION 1 SHORT ANSWER 1. c 2. a 3. b 2. a. alcohol b. water c. the gels 3. The mixture is a colloid. The properties are consistent with those reported in Table 3 on page 404 of the text. The particle size is small, but not too small, and the mixture

### CHAPTER 12 REVIEW Solutions - Weebly

CHAPTER 5 REVIEW The Periodic Law SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. c In the modern periodic table, elements are ordered (a) according to decreasing atomic mass. (b) according to Mendeleev's original design. (c) according to increasing atomic number. (d) based on when they were discovered. 2. d Mendeleev noticed that certain similarities in the ...

### 5 The Periodic Law - Office of Instructional Technology

The Arrangement of Electrons in Atoms chapter of this Holt McDougal Modern Chemistry Companion Course helps students learn the essential lessons associated with how electrons are arranged in atoms.

### Holt McDougal Modern Chemistry Chapter 4: Arrangement of ...

CHAPTER 1 REVIEW Matter and Change MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Classify each of the following as a homogeneous or heterogeneous substance. a. sugar d. plastic wrap b. iron filings e. cement sidewalk c. granola bar 2. For each type of investigation, select the most appropriate branch of chemistry from the following

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