

Modern Chemistry Test B Answer Key Chapter10

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Modern Chemistry Test B Answer

Modern Chemistry 88 Chapter Test Chapter: States of Matter PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

____ 1. According to the kinetic-molecular theory, particles of an ideal gas a. attract each other but do not collide. b. repel each other and collide.

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Modern Chemistry 71 Chapter Test Name Class Date Chapter Test B, continued 17. A substance combines with oxygen, releasing a large amount of energy as heat and light, in a(n) . 18. The decomposition of a substance by an electric current is called . 19. A(n) orders the elements by the ease with which they undergo certain chemical reactions. 20.

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Holt McDougal Modern Chemistry 1 Chapter Test Assessment Chapter Test B Chapter: The Periodic Law PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. ____ 1. In his periodic table, Mendeleev did not list all of the elements in order

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Modern Chemistry 46 Chapter Test Chapter: Chemical Bonding In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. ____ 1. The charge on an ion is a. always positive. b. always negative. c. either positive or negative. d. zero. ____ 2.

Assessment Chapter Test A - Kettering City School District

Modern Chemistry 143 Chapter Test Name Class Date Chapter Test B, continued PART V Write the answers to the following problems on the line to the left, and show your work in the space provided. ____ 39. What is the hydronium ion concentration of an aqueous solution that has a pH of 5.0? ____ 40.

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SHORT ANSWER Answer the following questions in the space provided. 1. b In metals, the valence electrons are considered to be (a) attached to particular positive ions. (c) immobile. (b) shared by all surrounding atoms. (d) involved in covalent bonds. 2. a The fact that metals are malleable and ionic crystals are brittle is best explained in ...

6 Chemical Bonding - Effingham County Schools / Overview

CHAPTER 10 REVIEW States of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Match description on the right to the correct crystal type on the left. b ionic crystal (a) has mobile electrons in the crystal c covalent molecular crystal (b) is hard, brittle, and nonconducting a metallic crystal (c) typically has the lowest melting point of the four

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Modern Chemistry 56 Chapter Test Assessment Chapter Test B Chapter: Chemical Equations and Reactions PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. ____ 1. The production of a slightly soluble solid compound in a double-

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Holt McDougal Modern Chemistry 3 Chapter Test Chapter Test A, continued ____ 9. Which of the following elements is a transition metal? a. calcium b. ron i c. sodium d. sulfur ____ 10. All the alkali metal elements are found on the periodic table in a. Group 1. b. Group 2. c. Period 1. d. Perod i 2. ____ 11.

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Modern Chemistry 12 Chapter Test Name Class Date Chapter Test A, continued ____ 20. In division and multiplication, the answer should have the same number of significant figures as the a. number in the calculation with the fewest significant figures. b. number in the calculation with the most significant figures.

Assessment Chapter Test A - Kettering City School District

CHAPTER 1 REVIEW Matter and Change MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Classify each of the following as a homogeneous or heterogeneous substance. a. sugar d. plastic wrap b. iron filings e. cement sidewalk c. granola bar 2. For each type of investigation, select the most appropriate branch of chemistry from the following

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