

Mastering Chemistry Answers Chapter 1

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Mastering Chemistry: Chapter 1 CHEM 14, 11 P05SP17 Chapter 1 Exercise 1.86 Exercise 1.86 Part C Fill in the blanks to complete the table Complete the third column symbol A #p e n Charge Express your answers Si 14 14 14 Cu²⁺ 15 16 A- 28.58.31 Submit M Incorrect: Try Ag Part D Complete the fourth column Express your answers a dip 14,16,29,15 pr Correct Part E

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Best Answer: here's the idea... The empirical formula is the lowest whole number MOLE ratio of the different elements in the compound. The %'s are very often given in MASS %'s.. as they are in this problem.. so to solve this, we will need to (1) convert mass %'s to grams to moles (2) simplify the mole ...

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The two isotopes of chlorine are and .Which isotope is the most abundant? Hint D.1 How to

approach the problem Hint not displayed Hint D.2 Find the atomic mass Hint not displayed Hint D.3 Mass of each isotope Hint not displayed ANSWER: Correct Since the atomic mass of chlorine, 35.45 , is closer to 35 than it is to 37, must be the most abundant isotope. does in fact account for 75.53% of the ...

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Express your answer using two significant figures. ANSWER: Chapter 10 Question 10 - Bimodal Description: (a) A sample of a gas originally at 25 °C and 1.00 atm pressure in a 2.5 L container is subject to a pressure of 0.85 atm and a temperature of 15 °C. The final volume of the gas is ____ L. Part A

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Make an outline of the materials you have been studying the chapter, without having the book open. How close does your chapter outline compare to the chapter summary and key terms, key skills, and key equations at the end of each chapter? Ensure you have the big picture and focus on meeting the chapter objectives. 5.) Practice.

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