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31 Chapter 2 Section 2 Units of Measurements pages 33-43 Density What is the density of a block of marble that occupies 310 cm3 and has a mass of 853 g? Diamond has a density of 3.26g/cm3. What is the mass of a diamond that has a volume of cm3? What is the volume of a sample of liquid mercury that has a mass of 76.2 g, given the density of mercury is 13.6 g/mL? p. 40 g/cm g mL Chapter 2 ...

Modern Chemistry Chapter 2 Measurements and Calculations ...

Modern Chemistry • CHAPTER 7 HOMEWORK 7-2 ... CHAPTER 7 • HOMEWORK 7-2 VOCABULARY 1. d 2. c 3. e 4. a 5. g 6. h 7. b 8. f GRAPHIC ORGANIZER Name of Compound or Ion Formula of Compound Copper(II) ion Cu2+ Ammonium ion NH 4 + Bismuth(III) bromide BiBr 3 Stannous fluoride SnF 2 Magnesium fluoride MgF 2 Iron(III) cyanide Fe(CN) 3

HOMEWORK 7-2 - Kettering City School District

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2. Which statement is true? a. In a number with no zeros, all the digits are significant. b. In chemistry, the term significant means "certain." c. Both significant and insignificant digits are always reported in the results of an experiment. d. Zeros at the end of a number are not significant digits.

CHAPTER 2 HOMEWORK 2-5 - apcscience.com

Accelerated Chemistry Anderson - MCHS Modern Chemistry 1 Chemical Equations & Reactions CHAPTER 8 Chemical Equations and Reactions SECTION 1 Describing Chemical Reactions OBJECTIVES 1. List three observations that suggest that a chemical reaction has taken place. 2.

CHAPTER 8 Chemical Equations and Reactions

Holt Modern Chemistry Review CHAPTER 2: MEASUREMENTS AND CALCULATIONS Include graphic organizer(s) for this chapter The following pages contain the bulk (but not all) of the information for

the chapter 2 test. Focus on this content, but make sure to review class notes, activities, handouts, questions, etc. If you study this document and NOTHING else, you should at least be able to PASS the test.

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3.1 discusses some of the most fundamental assertions of chemistry, including the law of conservation of mass, the law of definite proportions, and the law of multiple proportions. It also gives a ...

3.1 The Atom

CHAPTER 10 REVIEW States of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Match description on the right to the correct crystal type on the left. b ionic crystal (a) has mobile electrons in the crystal c covalent molecular crystal (b) is hard, brittle, and nonconducting a metallic crystal (c) typically has the lowest melting point of the four

10 States of Matter - Ms. Agostine's Chemistry Page

CHAPTER 1 REVIEW Matter and Change MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Classify each of the following as a homogeneous or heterogeneous substance. a. sugar d. plastic wrap b. iron filings e. cement sidewalk c. granola bar 2. For each type of investigation, select the most appropriate branch of chemistry from the following

mc06se cFMsr i-vi

CHAPTER 3 REVIEW Atoms: The Building Blocks of Matter MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. The element boron, B, has an atomic mass of 10.81 amu according to the periodic table.

3 Atoms: The Building Blocks of Matter - Ms. Sheehan's ...

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