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Modern Chemistry 77 Stoichiometry CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. ____ The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of H₂

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14 Ions in Aqueous Solutions and Colligative Properties

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Let's try a real chemistry problem and use this concept to determine the limiting reagent and see how it determines how much product will be formed. Problem : Nitrogen gas reacts with hydrogen gas to produce gaseous ammonia (NH₃). If we react 25.0 g of nitrogen with 25.0 g of hydrogen, how much ammonia will be produced.

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