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29.100 cm<sup>3</sup> of 1 N HCl is added to 100 cm<sup>3</sup> of 1 N NaOH solution. The rise in temperature is t<sub>1</sub>. 200 cm<sup>3</sup> of 2 N HCl is added to 200 cm<sup>3</sup> of 2 N NaOH solution. The rise in temperature is t<sub>2</sub>. Then  
1] 2 t<sub>1</sub> = t<sub>2</sub> 2] t<sub>1</sub> = 2 t<sub>2</sub> 3] t<sub>2</sub> = t<sub>1</sub> 4] t<sub>2</sub> = 4 t<sub>1</sub>  
30. For a gaseous reaction K<sub>c</sub> > K<sub>p</sub>, the rate of the reaction is favored by  
1] low pressure 2] low temperature 3] high pressure 4] high temperature

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