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____ 5. Mendeleev is credited with developing the first successful. a. periodic table. b. method for determining atomic number. c. test for radioactivity. d. use of X rays. ____ 6. Mendeleev did not always list elements in his periodic table in order of increasing atomic mass because he grouped together elements with similar. a. properties. c ...

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Modern Chemistry 33 Chapter Test Name Class Date Chapter Test B, continued 15. The energy state of an atom is called its ground state. 16. The number of waves that pass a point in one second is called. 17. When an electron drops from a higher-energy state to a lower-energy state, a(n) spectrum is produced. 18.

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Unformatted text preview: Chemistry Chapter 5 the Periodic Table In 1860, Stanislao Cannizzaro presented a method for accurately measuring the relative atomic masses of the elements. Chemistry Chapter 5 The Periodic Law In 1869, Dmitri Mendeleev published his first "periodic table". - He started by placing the known elements in order using their atomic masses.

Modern Chemistry Chapter 5 - Chemistry Chapter 5 the ...

Chapter 5 Pre-Test Chapter: The Periodic Law ... Modern Chemistry 2 The Periodic Law Chapter 5 Pre-Test, continued ____ 3. The periodic law states that a. no two electrons with the same spin can be found in the same place in an atom. b. the physical and chemical properties of the elements are functions of ...

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Holt McDougal Modern Chemistry 1 Chapter Test Assessment Chapter Test B Chapter: The Periodic Law PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. ____ 1. In his periodic table, Mendeleev did not list all of the elements in order

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Modern Chemistry 91 Chapter Test Name Class Date Chapter Test B, continued 27. Distinguish between ionic crystals and metallic crystals. PART V Write the answers to the following questions on the line to the left, and show your work in the space provided. The molar enthalpy of fusion for water is 6.008 kJ/mol. 28.

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Modern Chemistry 2 The Periodic Law CHAPTER 5 REVIEW The Periodic Law SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. ____ In the modern periodic table, elements are ordered (a) according to decreasing atomic mass.

CHAPTER 5 The Periodic Law - mchsapchemistry.com

CHAPTER 1 REVIEW Matter and Change MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Classify each of the following as a homogeneous or heterogeneous substance. a. sugar d. plastic wrap b. iron filings e. cement sidewalk c. granola bar 2. For each type of investigation, select the most appropriate branch of chemistry from the following

mc06se cFMs r i-vi

Holt McDougal Modern Chemistry 1 Chapter Test Assessment Chapter Test A Chapter: The Periodic Law Use the periodic table below to answer the questions in this Chapter Test. In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. ____ 1.

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CHAPTER 5 REVIEW The Periodic Law SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. c In the modern periodic table, elements are ordered (a) according to decreasing atomic mass. (b) according to Mendeleev's original design. (c) according to increasing atomic number. (d) based on when they were discovered. 2. d Mendeleev noticed that certain similarities in the ...

5 The Periodic Law - Jefferson Township Public Schools

CHAPTER 10 REVIEW States of Matter SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. Identify whether the descriptions below describe an ideal gas or a real gas. ideal gas a. ... SECTION 5 MODERN CHEMISTRY STATES OF MATTER. STATES OF MATTER MODERN CHEMISTRY.

10 States of Matter - Ms. Agostine's Chemistry Page

CHAPTER 6 REVIEW Chemical Bonding SECTION 5 SHORT ANSWER Answer the following questions in the space provided. 1. Identify the major assumption of the VSEPR theory, which is used to predict the shape of atoms. Pairs of valence electrons repel one another. 2. In water, two hydrogen atoms are bonded to one oxygen atom. Why isn't water a linear ...

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