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Modern Chemistry 77 Stoichiometry CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. ____ The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of H₂

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CHAPTER 9 REVIEW Stoichiometry SECTION 9-3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% If the actual yield of a reaction is 22 g and the theoretical yield is 25 g, calculate the percent yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of H₂ according to the following equation: $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$ N₂; 2.0 mol a.

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14 Ions in Aqueous Solutions and Colligative Properties

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Let's try a real chemistry problem and use this concept to determine the limiting reagent and see how it determines how much product will be formed. Problem : Nitrogen gas reacts with hydrogen gas to produce gaseous ammonia (NH₃). If we react 25.0 g of nitrogen with 25.0 g of hydrogen, how much ammonia will be produced.

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Modern Chemistry • CHAPTER 9 HOMEWORK 9-1 (pp. 275-279) VOCABULARY Define. ... Modern Chemistry • CHAPTER 9 HOMEWORK 9-2 (pp. 280-282) VOCABULARY Complete each sentence. 1. The mole ratio is obtained from _____. 2. Theoretical stoichiometric calculations show _____. ...

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