

WebAssign**Lab #14: Heat Capacity (Homework)**

Yinglai Wang

PHYS 172-SPRING 2012, Spring 2012

Instructor: Virendra Saxena

Current Score : 4 / 4 **Due** : Thursday, April 26 2012 11:59 PM EDT1. 1/1 points | [Previous Answers](#)

Thermodynamics describes the effect of heating an object. An important thermal parameter for a material is its heat capacity (a measure of how much its temperature will rise for a given input of energy).

In this lab, you predicted the heat capacity as a function of temperature for Aluminum. In your model you made use of a number of things that you have learned throughout this course, e.g. the quantized energy of a quantum oscillator, the interatomic spring constant, etc.

Which of the following concepts/principles did you use to predict the heat capacity (Include concepts/principles you used as you studied the underlying phenomena, e.g. the interatomic spring constant)?

- ☐ The angular momentum principle.
- ☒ Energy.
- ☒ Statistical mechanics (entropy, temperature, ...).
- ☒ Momentum.

2. 3/3 points | [Previous Answers](#)

MI2 11.P.57

A nanoparticle containing 6 atoms can be modeled approximately as an Einstein solid of 18 independent oscillators. The evenly spaced energy levels of each oscillator are $3e-21$ J apart. Use $k = 1.4e-23$ J/K.

(a) When the nanoparticle's energy is in the range $5(3e-21)$ J to $6(3e-21)$ J, what is the approximate temperature? (In order to keep precision for calculating the heat capacity, give the result to the nearest tenth of a degree.)

 K

(b) When the nanoparticle's energy is in the range $8(3e-21)$ J to $9(3e-21)$ J, what is the approximate temperature? (In order to keep precision for calculating the heat capacity, give the result to the nearest tenth of a degree.)

 K

(c) When the nanoparticle's energy is in the range $5(3e-21)$ J to $9(3e-21)$ J, what is the approximate

heat capacity per atom?

 J/K

Note that between parts (a) and (b) the average energy increased from "5.5 quanta" to "8.5 quanta".

As a check, compare your result with the high temperature limit of $3k$, where $k = 1.4\text{e-}23$ J/K.