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The Laws of Gases

By S.J. Robert Boyle

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Excerpt: . the value of a is deduced from p , p , v , v , given by experiment. Hence one obtains the atomic volume of the mass of gas subjected to the experiments from which the curves were constructed. This mass is defined by the normal pressure and the temperature at which the manometers were charged with gas. I have preferred to refer the atomic volume to the unit of volume of the gas at 0 0. and 76 centimetres of mercury, obtaining the following values: Hydrogen 0.00078 Carbon dioxide 0.00170 Ethylene 0.00232 Finally, I may remark that the interpretation given for the value of a is independent of the hypothesis on which the theory of the gaseous state is founded. It will readily be seen that it is enough that the curves of compressibility should tend to become straight lines in the limit; for under...


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