

Kinetic theory

Wenhao

October 2, 2021

1 Kinetic theory of ideal gas

We describe here the problem of motion of molecules in a gas, ignoring its rotational and vibration degrees of freedom, with energy:

$$E_k = \frac{1}{2}mv^2 = \frac{1}{2}m(v_x^2 + v_y^2 + v_z^2) \quad (1)$$

we consider some assumption: 1) the molecules do not interact with each other most of the time (only through collision) and 2) the collision is not frequent. Through the collision, molecules exchange energy with each other, but every thing remain in equilibrium. We also assume the motion before and after the collision is independent, thus their kinetic energies are independent variables. With the above assumption, we can consider each molecules as an independent system connected to the heat reservoir at temperature T , where heat reservoir is "all other molecules in the gas"¹ and energy is exchanged with collisions².

Velocity distribution

The velocity distribution $g(\mathbf{v})$ is defined to the fraction of molecules with velocity \mathbf{v} (\mathbf{v} is the vector velocity while v is the absolute value).

$$g(\mathbf{v}) \propto e^{-mv^2/2k_B T} dv_x dv_y dv_z \quad (2)$$

$$\propto e^{-m(v_x^2 + v_y^2 + v_z^2)/2k_B T} dv_x dv_y dv_z \quad (3)$$

$$= g(v_x)g(v_y)g(v_z) \quad (4)$$

we find the normalization factor for $g(v_x)$:

$$\int_{-\infty}^{\infty} e^{-mv_x^2/2k_B T} dv_x = \sqrt{\frac{2\pi k_B T}{m}} \quad (5)$$

$$g(v_x) = \sqrt{\frac{m}{2\pi k_B T}} e^{-mv_x^2/2k_B T} \quad (6)$$

with $g(v_i)$ normalized to 1, $g(v)$ from Eq.4 is then also normalized to 1. we can calculate the following values:

$$\begin{aligned} \langle v_x \rangle &= \int_{-\infty}^{\infty} v_x g(v_x) dv_x = 0 \\ \langle v_x^2 \rangle &= \int_{-\infty}^{\infty} v_x^2 g(v_x) dv_x = \frac{k_B T}{m} \end{aligned}$$

To find the distribution with respect to speed v instead of velocity \mathbf{v} , we consider the volume in velocity phase space corresponding to velocity v :

$$dV = 4\pi v^2 dv \quad (7)$$

which correspond to a thin spherical shell at radius v . The speed distribution function is then:

$$f(v)dv \propto v^2 dv e^{-mv^2/2k_B T} \quad (8)$$

¹Blundell, p48

²This point is not stated very clear in the book, but collision is the only way a gas molecule can interact with other molecules. The motion of the molecule after the collision is independent of their motion before the collision, therefore, the motion after the collision is entirely determined by the heat bath.

normalize $\int_0^\infty f(v)dv = 1$, we find the distribution function:

$$f(v) = \frac{4}{\sqrt{\pi}} \left(\frac{m}{2k_B T} \right)^{3/2} v^2 e^{-mv^2/2k_B T} \quad (9)$$

which is known as the **Maxwell-Boltzmann distribution**. We can find the following expectation values:

$$\begin{aligned} \langle v \rangle &= \int_0^\infty v f(v) dv = \sqrt{\frac{8k_B T}{\pi m}} \\ \langle v^2 \rangle &= \int_0^\infty v^2 g(v) dv_x = \frac{3k_B T}{m} \\ v_{rms} &= \sqrt{\langle v^2 \rangle} = \sqrt{\frac{3k_B T}{m}} \end{aligned} \quad (10)$$

where v_{rms} is the "root mean square" value. The mean kinetic energy of a gas molecule is then:

$$\langle E_k \rangle = \frac{1}{2} m \langle v^2 \rangle = \frac{3}{2} k_B T \quad (11)$$

Pressure

Since all molecules are equally likely to travel in any direction, the fraction of these whose velocity lie in solid angle $d\Omega$ is therefore $d\Omega/4\pi$. We consider the molecules traveling at angle between θ and $\theta + d\theta$ with some specific direction (z), the solid angle is given by:

$$d\Omega = 2\pi \sin \theta d\theta \quad (12)$$

So the fraction of molecules having speed between v and $v+dv$ and angel between θ and $\theta + d\theta$ with direction z is then:

$$f(v)dv \frac{2\pi \sin \theta d\theta}{4\pi} = \frac{1}{2} f(v)dv \sin \theta d\theta \quad (13)$$

In a time interval dt , a molecules with speed v travelling at an angle θ to the normal of the wall will hit the container wall if it is $vdt \cos \theta$ away from the wall. the number of such molecules (speed v at an angle θ) hitting the wall of area A is thus

$$A v dt \cos \theta n \cdot \frac{1}{2} f(v)dv \sin \theta d\theta \quad (14)$$

where the first part gives the total number of molecules in the volume $A v dt \cos \theta$ and the second part is the fraction of the molecule in the velocity and angle range.

Pressure of a gas on its container can be calculated by $dP = Fdt$ where dP is the momentum change of the particle reflecting from the wall: $dP = 2mv \cos \theta$ (perpendicular to wall). Pressure is defined to be $p = F/A$, so that We integrate over all possible velocity and angle θ to obtain the expression of p :

$$p = \int_0^\infty dv \int_0^{\pi/2} d\theta (2mv \cos \theta) \left(v \cos \theta n \frac{1}{2} f(v) \sin \theta \right) \quad (15)$$

$$= \frac{1}{3} n m \langle v^2 \rangle \quad (16)$$

using the relationship $n = N/V$ and $\langle v^2 \rangle = 3k_B T/m$, we can obtain the ideal gas law:

$$pV = Nk_B T \quad (17)$$

Partial pressure

If we have a mixture of gases in equilibrium, the total pressure will be the sum of pressures of each component:

$$n = \sum_i n_i \quad (18)$$

$$p = \left(\sum_i n_i \right) k_B T = \sum_i p_i \quad (19)$$

where $n_i = N_i/V$

Effusion

Effusion is the process in which gas escape from a small hole³. we define the flux Φ as the number of molecules going through unit area per unit time, therefore, Φ has the unit of $m^{-2}s^{-1}$. we use the above result of particles hitting the wall to find flux:

$$\Phi = \int_0^\infty dv \int_0^{\pi/2} d\theta \left(v \cos \theta n \frac{1}{2} f(v) \sin \theta \right) \quad (20)$$

$$= \frac{1}{4} n \langle v \rangle \quad (21)$$

with the result:

$$\begin{aligned} n &= p/k_B T \\ \langle v \rangle &= \sqrt{\frac{8k_B T}{\pi m}} \\ \Phi &= \frac{p}{\sqrt{2\pi m k_B T}} \end{aligned} \quad (22)$$

Now, consider a we have a small hole (small enough that the gas escaping does not change the equilibrium distribution of gas near the hole), the number of molecules escaping through the hole is given as:

$$\Phi A = \frac{pA}{\sqrt{2\pi m k_B T}} \quad (23)$$

which is linear to pressure, inversely proportion to T and m .

It should be noted that the speed distribution of the gas through the hole is no longer the Maxwell-Boltzmann distribution, since the amount of gas that effuse through the hole depend on velocity v (in Eq.20), therefore the speed distribution of those that go through the hole as an extra factor v , compare to the speed distribution of gas molecules inside the box.

Average lifetime

We first considering a molecule moving at speed v with other molecules in the gas stationary. Within a time dt , the molecule sweep through an area $\sigma v dt$, where σ is the collision cross-section of the molecule. We denote $P(t)$ as the probability of this molecule travel without any collision up to time t . We have:

$$P(t + dt) = P(t) + \frac{dP}{dt} dt \quad (24)$$

$$P(t + dt) = P(t)(1 - n\sigma v dt) \quad (25)$$

with n the number density of gas and therefore $(1 - n\sigma v dt)$ gives the probability of having no other gas molecule in the volume the moving molecule is going to sweep over from time t to $t + dt$. Withe the above two equation, we obtain the result:

$$P(t) = e^{-n\sigma v t} \quad (26)$$

We have the boundary condition $P(0) = 1$ and $P(\infty) = 0$

The probability that the molecule collide in time interval from t to $t + dt$ is given by the probability that molecule does not collide at time t but do collide at time $t + dt$: $P(t) - P(t + dt)$ (since $P(t)$ decay with time, $dP(t) < 0$), we have:

$$-dP(t) = n\sigma v e^{-n\sigma v t} dt \quad (27)$$

We can calculate the mean scattering time:

$$\begin{aligned} \tau &= \int_0^\infty t n\sigma v e^{-n\sigma v t} dt \\ &= \frac{1}{n\sigma v} \int_0^\infty x e^{-x} dx \\ &= \frac{1}{n\sigma v} \end{aligned} \quad (28)$$

³Blundell, p64

where $x = n\sigma vt$ represent a change of variable.

In the case of hard sphere approximation in a gas consist of molecules with equal radius, the collision cross-section is given as:

$$\sigma = \pi(r_1 + r_2)^2 = 4\pi a^2 \quad (29)$$

Mean free path of gas

With the previous result, we have, for the mean free path

$$\lambda = \langle v \rangle \tau = \langle v \rangle \frac{1}{n\sigma \langle v_r \rangle} \quad (30)$$

$\langle v \rangle$ is the average speed of gas molecules, where for the lifetime, instead of using v , we recognize that we should use relative speed v_r in a real system when all particles are in motion. It is not attempted to derive the value of $\langle v_r \rangle$ here (refer to Blundell, p73). But the final result is:

$$\lambda = \frac{1}{\sqrt{2}n\sigma} \quad (31)$$

2 Transport properties in Ideal gas

Viscosity

Viscosity η measures the resistance of a liquid (gas) to deform by shear stress. Suppose we sandwich fluid between two infinitely large plates. We fix the bottom plate and apply some force F to the top plate. The fraction between the fluid and the top plate will accelerate the fluid near it, then the momentum is passed down inside the liquid through internal interaction (collision).

Now, we wait for long enough time until the whole system is in equilibrium. The top plate will reach some final speed u and the fluid near the top plate will have the same macroscopic speed (condition of equilibrium). The bottom plate and the fluid near the bottom will be stationary (An external force is necessary to keep the bottom plate stationary). Therefore, a velocity gradient will be set up in the fluid between the two plates (in z direction). We define the viscosity by equation:

$$\tau_{shear} = \frac{F}{A} = \eta \frac{d\langle u_x \rangle}{dz} \quad (32)$$

With force F constantly applied on the top plate, we are inputting momentum Fdt per unit area in time dt . In equilibrium, this momentum will be transported completely to the bottom plate. Therefore, we have momentum flux through z direction (note that the momentum itself is along the plate, x direction, but the flux is along z):

$$\Pi_z = -F/A = -\eta \frac{d\langle u_x \rangle}{dz} \quad (33)$$

We have a negative sign because the momentum flow from high velocity area to low velocity area, which is opposite the velocity gradient.

We can calculate the momentum flux by considering the microscopic motion of gas molecules. We consider that the motion of free ideal gas superimposed with the collective drifting motion with velocity $\langle u_z \rangle$. A molecule travelling along z direction will change their momentum by collision with other molecules in the final position z_2 ($z_1 \rightarrow z_2$). The number of particles with velocity v travelling at an angle θ with z direction is $v \cos \theta n \cdot \frac{1}{2} f(v) dv \sin \theta d\theta$ per unit area A and time dt . How far they travel will be given by the mean free path λ . The momentum difference of a molecule just after collision at z_1 and just after collision at z_2 is given by:

$$-m \left(\frac{\partial \langle u_x \rangle}{\partial z} \right) \lambda \cos \theta \quad (34)$$

which is the momentum flux created by this single molecule. Summing over all molecules with different speed and angle, we have:

$$\Pi_z = \int_0^\infty dv \int_0^{\pi/2} d\theta v \cos \theta n \cdot \frac{1}{2} f(v) \sin \theta \cdot m \left(\frac{\partial \langle u_x \rangle}{\partial z} \right) \lambda \cos \theta \quad (35)$$

$$= -\frac{1}{3} nm \lambda \langle v \rangle \frac{\partial \langle u_x \rangle}{\partial z} \quad (36)$$

We obtain the viscosity:

$$\eta = \frac{1}{3}nm\lambda\langle v \rangle \quad (37)$$

Using the previous result $\lambda = (\sqrt{2}n\sigma)^{-1}$ and $\langle v \rangle = \left(\frac{8k_B T}{\pi m}\right)^{1/2}$, we can also write:

$$\eta = \frac{2}{3\sigma} \left(\frac{mk_B T}{\pi} \right)^{1/2} \quad (38)$$

We have the following observations:

1. η is independent of pressure.
2. $\eta \propto T^{1/2}$.
3. For the above approximation of momentum transfer to be correct, we require $L \ll \lambda \ll d$, where d is the size of the molecule and L is the size scale of the container.

Thermal conductivity

We define the heat flux in the z direction

$$J_z = -\kappa \left(\frac{\partial T}{\partial z} \right) \quad (39)$$

$$\mathbf{J} = -\kappa \nabla T \quad (40)$$

The gas molecules carrier heat through their kinetic energy, which depend on temperature T via $\langle E_k \rangle = 3k_B T/2$. Defining heat capacity of a molecule as C , we can calculate the total heat flux:

$$J_z = \int_0^\infty dv \int_0^{\pi/2} d\theta v \cos \theta n \frac{1}{2} f(v) \sin \theta \cdot -C \left(\frac{\partial T}{\partial z} \right) \lambda \cos \theta \quad (41)$$

$$= -\frac{1}{3}nC\lambda\langle v \rangle \frac{\partial T}{\partial z} \quad (42)$$

Using $C_V = nC$, the thermal conductivity of gas is therefore

$$\kappa = \frac{1}{3}C_V\lambda\langle v \rangle \quad (43)$$

We observe that Eq.42 is very similar to Eq.36, and we have:

$$\frac{\kappa}{\eta} = \frac{C}{m} \quad (44)$$

Particle Diffusion

We consider a gas of molecules in which some of them is labelled, if those labelled molecules are initially confined in certain area and the confinement is removed, they will start to diffuse (Self-diffuse). Suppose that the diffusion is along z direction and we use $n^*(z)$ to denote the density of those labelled particles, we can define the diffusion coefficient:

$$\Phi_z = -D \left(\frac{\partial n^*}{\partial z} \right) \quad (45)$$

Following the above microscopic picture, we have:

$$\Phi_z = \int_0^\infty dv \int_0^{\pi/2} d\theta v \cos \theta \frac{1}{2} f(v) \sin \theta \cdot - \left(\frac{\partial n^*}{\partial z} \right) \lambda \cos \theta \quad (46)$$

$$= -\frac{1}{3}\lambda\langle v \rangle \frac{\partial T}{\partial z} \quad (47)$$

giving the self-diffusion coefficient

$$D = \frac{1}{3}\langle v \rangle \quad (48)$$

We have the following relationship:

1. $D \propto T^{3/2}$
2. $D\rho = \eta$, where ρ is the density $\rho = nm$

Heat diffusion equation

Appendix A