





Atmospheres as a Window to Rocky Exoplanet Surfaces

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ABSTRACT

As the characterization of exoplanet atmospheres proceeds, providing insights into atmospheric chemistry and composition, a key question is how much deeper into the planet we might be able to see from its atmospheric properties alone. For small planets with modest atmospheres and equilibrium temperatures, the first layer below the atmosphere will be their rocky surface. For such warm rocky planets, broadly Venus-like planets, the high temperatures and moderate pressures at the base of their atmospheres may enable thermochemical equilibrium between rock and gas. This links the composition of the surface to that of the observable atmosphere. Using an equilibrium chemistry code, we find a boundary in surface pressure–temperature space which simultaneously separates distinct mineralogical regimes and atmospheric regimes, potentially enabling inference of surface mineralogy from spectroscopic observations of the atmosphere. Weak constraints on the surface pressure and temperature also emerge. This regime boundary corresponds to conditions under which SO₂ is oxidized and absorbed by calcium-bearing minerals in the crust, thus the two regimes reflect the sulphidation of the crust. The existence of these atmospheric regimes for Venus-like planets is robust to plausible changes in the elemental composition. Our results pave the way to the prospect of characterizing exoplanetary surfaces as new data for short period rocky planet atmospheres emerge.

Key words: planets and satellites: atmospheres – planets and satellites: composition – planets and satellites: surfaces – planets and satellites: terrestrial planets – planets and satellites: individual: Venus

1 INTRODUCTION

Since the discovery of the first exoplanet in orbit around a main sequence star (Mayor & Queloz 1995), over 5 500 exoplanets have been discovered at the time of writing¹. High-resolution transmission spectroscopic facilities are enabling the increasingly routine characterization of exoplanetary atmospheres, with recent high-fidelity detections of CO₂ (Ahler et al. 2022), SO₂ (Alderson et al. 2022), and CH₄ (Madhusudhan et al. 2023) using *JWST*. As we gain greater insight into exoplanet atmospheres a key challenge is going to be seeing further into the planet, using the atmosphere as a window into the geological processes beneath.

One key question to ask from atmospheric data is ‘what is the surface of the exoplanet?’. There has been significant interest in identifying whether a planet has a discrete surface–atmosphere boundary at all (e.g., Hu et al. 2021; Yu et al. 2021; Tsai et al. 2021). This is important where the bulk density and instellation of the planet leaves open multiple possibilities for the nature of the atmosphere–interior connection: whether it is a water ocean underlying a thick atmosphere (e.g., Hycean worlds; Madhusudhan et al. 2021), a magma ocean (e.g., Lichtenberg 2021), or a deep high pressure gas envelope (as with Jupiter and Saturn in our own solar system). However, where the bulk density is constrained to be high, the equilibrium temper-

ature is modest, and observations rule out thick atmospheres (e.g., phase curve observations; Kreidberg et al. 2019; Zieba et al. 2023), then such planets may have atmospheres in contact with rock. This opens up a regime of surface–atmosphere interaction of which we have direct analogues in the solar system’s terrestrial planets.

The mechanism of surface–atmosphere interaction depends primarily on the surface temperature, as depicted in Fig. 1. On cool planets, such as Earth, liquid surface water facilitates geochemical interactions between a planet’s lithospheric, oceanic, and atmospheric chemistry (e.g., Kite & Ford 2018). At intermediate temperatures – like those at the surface of Venus (~ 740 K), the so-called *warm* exoplanet regime – the surface and atmosphere may approach thermochemical equilibrium. In this latter case, the chemical species present at the base of the atmosphere and on the surface may then be prescribed by equilibrium thermodynamics, particularly at the high-temperature end of this regime.

Venus provides valuable insight into the potential nature of surface–atmosphere interaction on warm rocky worlds (e.g., see Feigley 2014, for a review). After the first Venus missions found atmospheric temperatures in excess of 500 K, it was suggested that the planet’s lithosphere is in equilibrium with its atmosphere (Mueller 1963). Subsequent lander missions measured surface temperatures of over 700 K, and pressures of over 90 bar (Marov 1972), consistent with conditions under which minerals metamorphose deep in the Earth’s crust. The findings of these missions fuelled the first theoretical considerations of surface–atmosphere interactions on Venus’

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¹ <https://exoplanets.nasa.gov/>

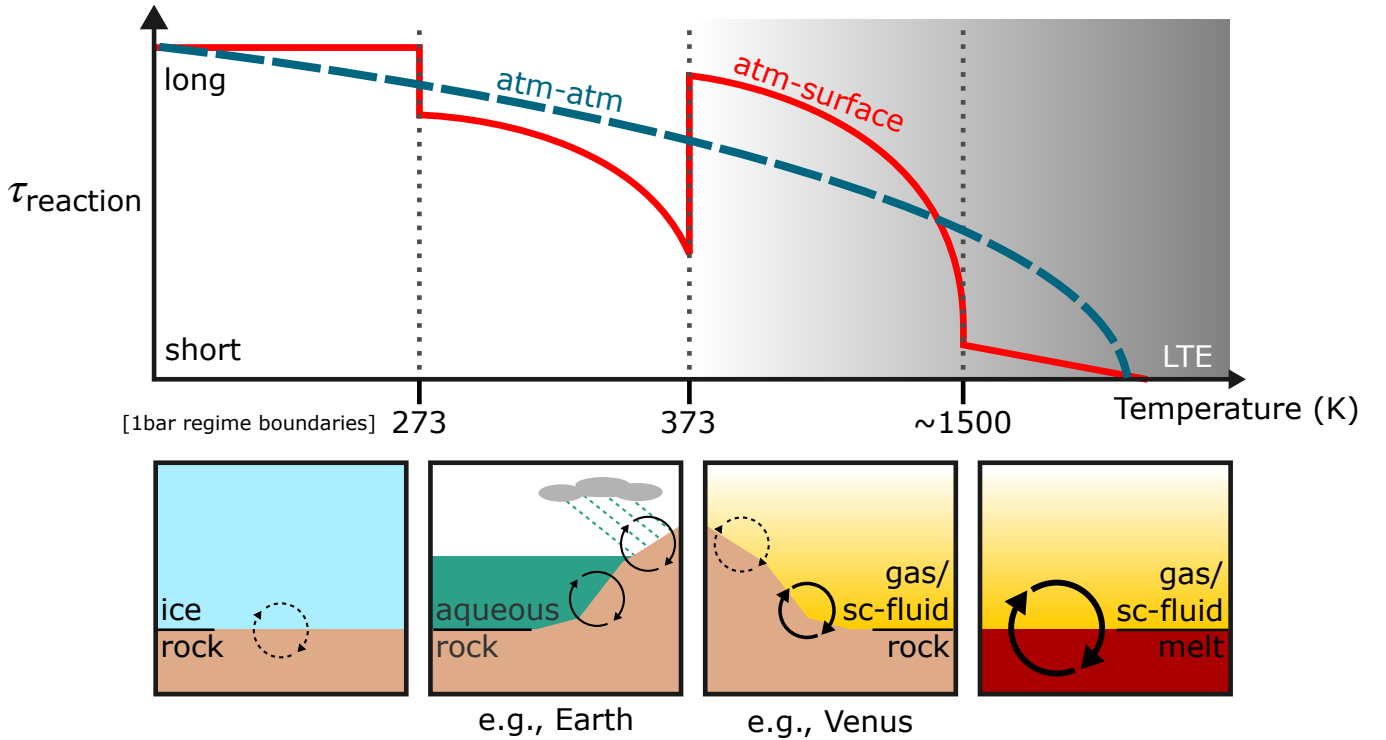


Figure 1. The timescales of surface–atmosphere interaction (red line) and atmospheric equilibration (blue dashed line) across different rocky planet climatic-geodynamic regimes. A reference timescale is shown for pure gas-phase chemical equilibrium to be reached, which is a simple monotonic function of temperature (Arrhenius in nature). Timescales of surface-atmosphere interaction on rocky planets deviate from this as a function of temperature as the climatic/geodynamic phases mediating the reactions change: schematics of these various planetary climatic-geodynamic regimes are shown at top, above their corresponding region of the reaction timescale graph. Cold planets feature slow solid-solid interactions between rock and ice layers (red line) and similarly cold atmospheres (dashed blue line). Where the temperature allows liquid surface water, weathering processes efficiently couple the atmospheric, oceanic, and surface chemistry, albeit in a disequilibrium state. Reaction rates in this regime increase with temperature as water temperature and increased rainfall drive more reaction (Walker et al. 1981; Maher & Chamberlain 2014). In the absence of liquid water, the rate of interaction between the atmosphere and surface may sharply decrease. However, at higher surface temperatures warm rocky planets such as Venus feature increasingly strong surface–atmosphere coupling. In this regime thermochemical equilibrium is achievable potentially both in the lower atmosphere (Liggins et al. 2023) and between the atmosphere and surface. Very hot terrestrial worlds have molten surfaces, and there is likely efficient mass transfer (and therefore equilibration) between the surface and atmosphere.

surface (e.g., Lewis 1970; Volkov et al. 1986), which indicated that a thermochemical equilibrium at the surface may regulate the abundances of major gases such as CO_2 and SO_2 in Venus’ atmosphere.

Subsequent work has shown that Venus’ atmospheric chemistry may mediate a large network of mineral equilibria, involving geochemically important minerals such as calcite (CaCO_3), anhydrite (CaSO_4), pyrite (FeS_2), and haematite (Fe_2O_3) (e.g., Fegley et al. 1997). The intractability of analytically analysing the vast number of important thermochemical equilibria on Venus and elsewhere requires equilibrium chemistry codes, e.g., GGCHEM (Woitke et al. 2018) and PERPLE_X (Connolly & Petrini 2002), to calculate the stable abundances of gas and mineral species at various temperatures and pressures.

Lewis (1970) suggested that a surface–atmosphere equilibrium on Venus meant that the new knowledge of the planet’s atmospheric chemistry could enable inferences on its surface mineralogy. We posit that the same principle applies to exoplanets. Exoplanetary atmospheres can be characterized in increasing detail with modern facilities, and for warm rocky exoplanets this offers as an exciting by-product the possibility of characterizing their surface mineralogy. In turn, this surface mineralogy links to the composition and geodynamic history of planets, and is the key to understanding their wider geological history. Such inferences could complement the use

of emission spectroscopy in identifying surface mineralogy of exoplanets with a thin or absent atmosphere.

In this paper we investigate how the atmospheres of warm rocky exoplanets may provide information on their surface mineralogies. The paper is organized as follows. In Section 2, we give an overview of relevant thermodynamics, and of the model we have used for our investigations. In Section 3, we run the model over a broad 2D grid in surface pressure–temperature (p_0 – T_0) space, identifying a boundary where the atmospheric chemistry and the surface mineralogy change abruptly and simultaneously. Section 4 explores the effects of altering the underlying elemental composition, to assess the generality of our findings. In Section 5, we discuss implications and caveats to our work, and Section 6 provides a summary.

2 METHODS

The model setup is shown schematically in Fig. 2. We consider a system at fixed pressure and temperature that solves the heterogeneous solid–gas phase equilibria. Depending on the temperature and pressure of the system, components (notably SO_2 , as we will see in Section 3) will transition from the gas to solid phase. Given the temperature range we consider in this study, major rock-forming

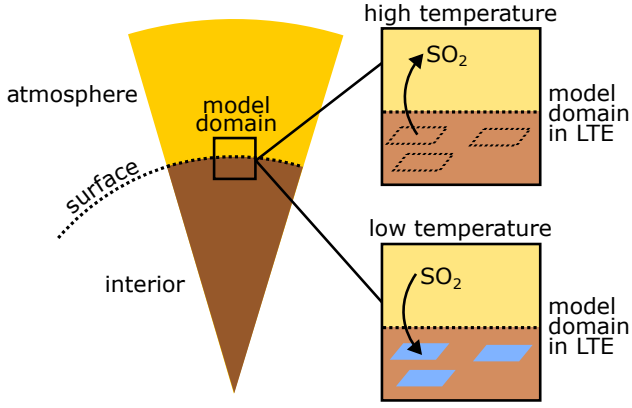


Figure 2. A schematic of the model setup. We simulate conditions of lower atmosphere-surface interaction by calculating heterogeneous phase equilibria (gas-solid) in a single model domain. At high temperatures components that can be present in solid form enter the gas phase, notably SO_2 in our simulations. At low temperature, those components can react back with the solid components into mineral form. LTE = local thermochemical equilibrium.

oxides such as CaO , SiO_2 etc. always remain in the solid phase. Thereby, we can set the extent to which the solids (the ‘surface’ of our model planet) buffer the chemistry of the gas (the ‘atmosphere’) by changing the proportions of dominantly gas-phase to dominantly solid-phase elemental abundances. A description of the approach to solving for thermodynamic equilibrium and the benchmarking of our calculations follows.

2.1 Thermodynamic modelling

We first outline the requisite thermodynamics for calculating equilibrium mineral and gas abundances. In thermochemical equilibrium, the set of volume mixing ratios (VMRs) $\{x_i\}$ of all gas species at a given pressure and temperature is that which minimizes the Gibbs free energy (e.g., White et al. 1958):

$$\frac{G(p, T, \{x_i\})}{RT} = \sum_i x_i \left[\frac{G_{m,i}^\circ(T)}{RT} + \ln\left(\frac{p}{p^\circ}\right) + \ln x_i \right], \quad (1)$$

where $G_{m,i}^\circ(T)$ is the molar Gibbs free energy at the standard pressure $p^\circ = 1$ bar and temperature T , and p is the total pressure. This minimization is subject to the constraints:

$$\sum_i x_i = 1; \quad \sum_i a_{ij} x_i = b_j, \quad (2)$$

where a_{ij} is the number of atoms of element j in species i and b_j is the total number density of atoms j . The only free parameters are the pressure, temperature, and relative elemental abundances. There are several codes which, given these parameters, calculate the equilibrium VMRs of a gas, but few also calculate the abundances of solid/liquid condensates that would be present in equilibrium with such a gas. One code that does is GGChem (Woitke et al. 2018), which we use throughout this work.

It is worth justifying the use of this code to model surface-atmosphere equilibrium. The initial use case of GGChem was to study the condensation sequence and the formation of astrophysical dust (Woitke et al. 2018). To this end the code finds, under user-inputted conditions, the mixture of solids and gases that would be present at thermochemical equilibrium. The solids are in this case envisioned as dust grains floating in the surrounding gas. However, with the assumption of equilibrium, this is essentially equivalent

to a surface in contact with an atmosphere. Although the gas-solid interaction area is lower in the surface-atmosphere case, this is a kinetic barrier, not a thermodynamic one. Provided the timescale for equilibration is significantly shorter than that of non-equilibrium chemical processes – as is the case on Venus-like planets – a surface mineralogy would be able to reach thermochemical equilibrium with an overlying atmosphere. Indeed, it is expected that the surface mineralogy of Venus is in equilibrium with the lower atmosphere (see, e.g. Mueller 1963; Krasnopolsky & Parshev 1981; Volkov et al. 1986; Fegley et al. 1997) though this is to date unconfirmed by in situ measurements. We show in Appendix A that a layer of rock with a thickness of just 15 cm suffices to react with the entire mass of the atmosphere; over this thickness crustal fractures allow effective contact with the atmosphere, facilitating equilibrium between the two. Where the elemental composition of Venus is required, we use the procedure outlined in section 3.2 of Rimmer et al. (2021), which uses Vega 2 measurements of oxide mass fractions on Venus’ surface from Surkov et al. (1986) to estimate elemental abundances. The resulting composition is given in Table B1 in Appendix B. We only consider the elements H, C, N, O, F, S, Cl, Fe, Mn, Si, Mg, Ca, Al, Na, K, and Ti, as well as He, Ne, and Ar. As it is the surface minerals which provide the vast majority of the atoms in the surface-lower-atmosphere system, this surface composition effectively constitutes the overall composition of the equilibrium system. Naturally, the underlying elemental abundance impacts the resulting chemical compositions of the surface and atmosphere; this is explored in Section 4.

2.2 Benchmarking against modern Venus

To verify the accuracy of GGChem’s calculations, we sought to validate the code against observations of Venus’ altitude-dependent mineralogy. The *Pioneer Venus* and *Magellan* missions found abnormally high radar reflectivity in the Venusian highlands ($z \gtrsim 2.6$ km, relative to Venus’ modal radius; Pettengill et al. 1982, 1996). Venus’ surface pressure and temperature both vary with altitude, resulting in the equilibrium abundances of surface minerals and gas species being altitude-dependent. Hence, the highland radar reflectivity problem may be resolved by invoking some mineral that is thermodynamically stable only under Venus’ high-altitude conditions, and not at low altitudes. Various candidates for this mineral have been proposed, the most widely-accepted being pyrite (e.g., Klose et al. 1992; Kohler et al. 2015; Semprich et al. 2020), whose semiconductivity confers high radar reflectivity. The presence of pyrite on the surface of the Venusian highlands is expected to be verified in the early 2030s by the VERITAS (Smrekar et al. 2020), DAVINCI (Garvin 2021) and EnVision (Widemann et al. 2021) missions.

GGChem was run on Venus’ surface conditions and elemental composition over the altitude range -2 km to $+12$ km, spanning Venus’ surface hypsometry (Fegley 2014). The full surface composition at $z = 0$ km is given in Table 1. We corroborate the results of Rimmer et al. (2021), finding a solid composition of mainly enstatite (MgSiO_3), anorthite ($\text{CaAl}_2\text{Si}_2\text{O}_8$), and albite ($\text{NaAlSi}_3\text{O}_8$). The only iron- or sulphur-bearing minerals at $z = 0$ km are Fe_2O_3 and CaSO_4 . The p - T profile and the calculated abundances of selected minerals are shown in Figure 3, including all iron- and sulphur-bearing minerals. We see that pyrite becomes stable only when the pressure falls below 75 bar, corresponding to altitudes above roughly 2.8 km. At lower altitudes, all of the iron is present as Fe_2O_3 ; only a small fraction of the iron is transferred from Fe_2O_3 to FeS_2 . The only other significant change in mineralogy with altitude is the conversion of corundum (Al_2O_3) and some quartz (SiO_2) to andalusite (Al_2SiO_5), however this occurs at much higher altitudes ($z \approx 5.6$ km).

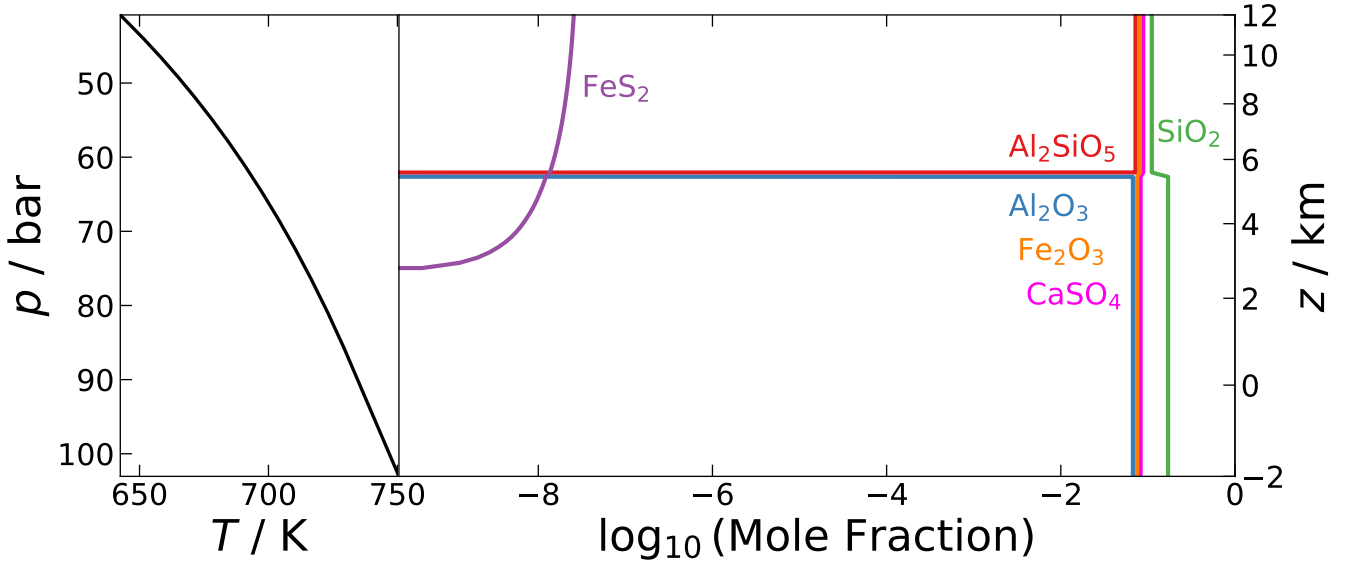


Figure 3. Venus’ surface conditions and equilibrium abundances of selected minerals. The left plot shows the p – T profile on Venus, as provided by GGChem; the profile is not shown for altitudes higher than that of Skaði Mons (Venus’ highest point) and is extrapolated down to the depth of the Diana Chasma (Venus’ lowest point). The right plot shows the equilibrium abundances of selected minerals, in terms of their mole fractions of the solid components. The SiO_2 line represents quartz itself, not the sum of all silicate minerals as is sometimes denoted by this. The only two iron-bearing minerals present at any altitude are Fe_2O_3 and FeS_2 , and the only two sulphur-bearing minerals are CaSO_4 and FeS_2 , consistent with previous thermochemical calculations (e.g., Rimmer et al. 2021). The dark purple line shows that FeS_2 only appears above an altitude of approximately $z = 2.8$ km, corresponding roughly to the altitude above which Venus shows high radar reflectivity (Pettengill et al. 1982, 1996). We see that below an altitude of approximately $z = 5.6$ km, andalusite (Al_2SiO_5 ; red line) decomposes into corundum (Al_2O_3 ; blue) and quartz (green), the latter’s abundance being slightly higher below this altitude.

and andalusite is not a semiconductor, so its appearance is unrelated to the radar reflectivity observations. The agreement between GGChem’s calculations and the likely source of Venus’ highland radar reflectivity is evidence of the accuracy of the thermodynamic model under these conditions, justifying its application to Venus-like exoplanets.

3 COUPLED SURFACE AND ATMOSPHERE CHEMICAL REGIMES

3.1 Gas and mineral chemistry

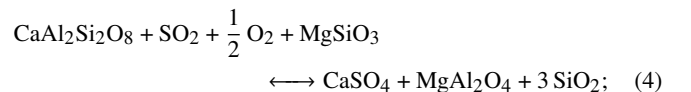
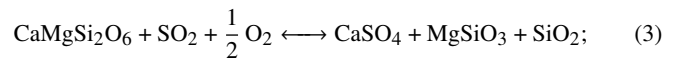
The altitudinal variations in pressure and temperature on Venus’ surface correspond to a 1D track in p_0 – T_0 space. To explore a wider range of conditions that may exist on warm rocky exoplanets, GGChem was run over a grid in p_0 – T_0 space, with p_0 and T_0 distributed logarithmically in the respective ranges $[10^{-1}, 10^4]$ bar and $[650, 1700]$ K: a broad range of conditions at which thermochemical equilibrium between the surface and the lower atmosphere, and in the lower atmosphere itself, may occur. The temperature at the highest surface altitude on Venus is 650 K; Liggins et al. (2023) find that volcanically-derived atmospheres can only achieve thermochemical equilibrium in the gas phase at $T \gtrsim 700$ K. The upper bound of 1700 K is roughly the temperature at which quartz melts (Chase 1986): in GGchem this was found to be the major mineral which melts at the lowest temperature. In fact, melting is likely to become important at lower temperatures than this, due to the stabilisation of eutectic melts. Results above around 1500 K should therefore be treated as illustrative only. We elaborate on this point in Section 5.2.5.

In these calculations we again use the estimated elemental composition of Venus as given in Appendix B.

Abundances of several minerals are shown in the top row of Figure 4. Each plot shows an arc-shaped boundary, from $(p_0 / \text{bar}, T_0 / \text{K}) \approx (10^{-1}, 1200)$ to $(10^3, 2000)$, across which the mineral abundances change sharply. This is particularly stark for CaSO_4 , diopside ($\text{CaMgSi}_2\text{O}_6$), and spinel (MgAl_2O_4), each of which disappears completely on one side of the boundary².

The equilibrium VMRs of CO_2 , O_2 , SO_2 , and SO_3 are shown in the bottom row of Figure 4. There is a clear transition in the atmospheric chemistry, across a boundary which roughly coincides with the boundaries seen in the mineral abundances in the top row. Whereas on one side the atmospheres are CO_2 -dominated, on the other they consist of a mixture of SO_2 and O_2 , with SO_3 constituting a minor component near to the boundary.

Mass balance investigations show that the boundary on the top row of Figure 4 are conceptually due to the two reactions:



each of which involves the oxidation of SO_2 into CaSO_4 at the expense of a calcium-bearing mineral and O_2 . These reactions occur at

² Spinel is also lost isothermally below ~ 1200 K, though this is likely to be an example of what Woitke et al. (2018) refer to as a type-2 transition, where one mineral is suddenly switched off. This behaviour would not be exhibited in a natural system.

Solid Composition (mass fraction)										
MgSiO ₃	CaAl ₂ Si ₂ O ₈	NaAlSi ₃ O ₈	Fe ₂ O ₃	CaSO ₄	SiO ₂	Al ₂ O ₃	KAlSi ₃ O ₈	Mn ₃ Al ₂ Si ₃ O ₁₂	TiO ₂	MgF ₂
29.7%	21.7%	17.6%	8.9%	8.3%	7.6%	5.1%	0.6%	0.3%	0.2%	trace

Table 1. Equilibrium surface composition at $z = 0$ km on Venus. Our results are identical to table 3 of [Rimmer et al. \(2021\)](#) at the precision shown.

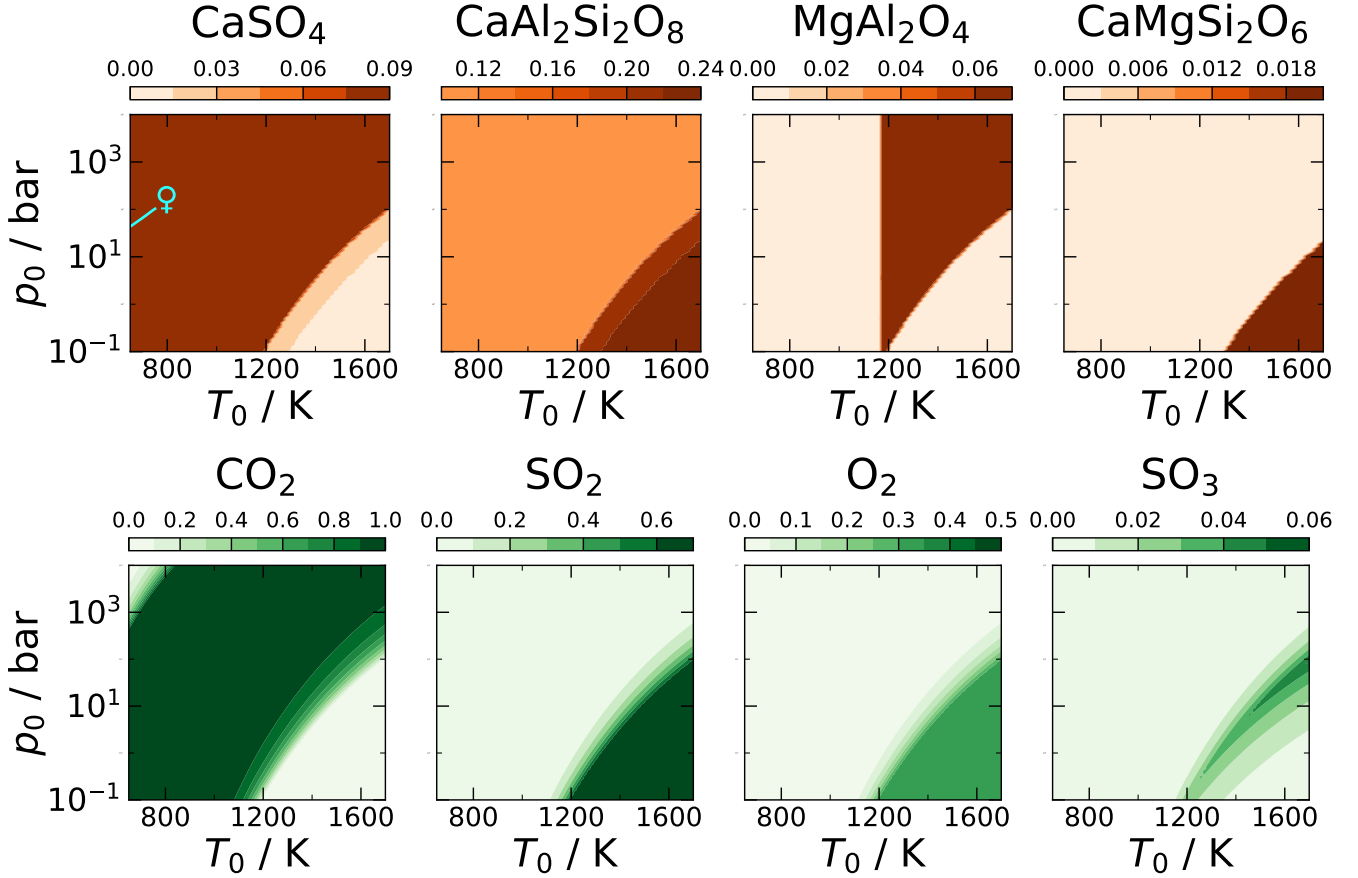


Figure 4. Equilibrium abundances of selected minerals and gases on Venus-composition exoplanets. Equilibrium abundances of four major minerals are shown, in terms of mole fraction of the condensates, in the top row. The surface conditions on Venus are shown by the short blue track in the first panel. There is a clear boundary in the abundance patterns of these minerals, across which their abundances change abruptly. The second row, shows the abundances of four major gases presented as VMRs, and demonstrates that the mineralogical boundary also demarcates regimes in the atmospheric chemistry. The high- T_0 , low- p_0 corner is dominated by a mixture of SO₂ and O₂, with SO₃ a minor component. The rest of the atmospheres are dominated by CO₂.

slightly different pressure–temperature conditions, causing the formation of CaSO₄ to occur in two stages as the pressure is raised (or as the temperature is lowered).

The change in the VMR of CO₂ across the boundary is not related to the stability of any carbon-bearing mineral³, unlike with SO₂ and O₂: at least 99.4 per cent of the carbon atoms are contained within CO₂ on both sides of the boundary. CO₂ is therefore present across almost the entire grid. On the lower-pressure side of the boundary it is vastly outnumbered by SO₂ and O₂, which dominate the atmosphere under these conditions. At higher pressures, reactions 3 and

4 sequester SO₂ and O₂ into CaSO₄, leaving CO₂ as the dominant component.

One important note about these simulations is that as GGchem is run at fixed p and T . The consequence of this is that reactions sequestering gas phase species lead to a multiplying up of the actual number density of every species in proportion so that the atmosphere maintains the requisite pressure. Therefore, although the element ratios are preserved between different points on the grid, each point has a different system mass. We therefore perform mass balance analyses by looking at the *fraction* of the atoms of a given element that are contained within different species. For example, on the higher-pressure side of the boundary, almost all of the sulphur atoms are in CaSO₄; on the lower-pressure side almost all are in SO₂ and SO₃.

³ There is in fact another boundary at much higher pressure and lower temperature at which CO₂ is released by the decomposition of MgCO₃, but we do not consider it of further relevance to the Venus-like planet scenarios.

3.2 Transmission spectra across mineral-gas regimes

We have seen above that the compositions of the high and low temperature (low and high pressure, respectively) atmospheric regimes are starkly different. It follows that transmission observations of the atmosphere of a Venus-composition exoplanet could identify whether the atmosphere is CO₂- or SO₂-O₂-dominated, and hence whether the surface pressure–temperature conditions are above or below the boundary. As the mineralogy is also, to first order, dependent on which side of the boundary the surface conditions are, such observations would then also provide unprecedented information about the planet’s mineralogy.

Using the radiative transfer package PETITRADTRANS (Mollière et al. 2019), and the comprehensive line lists provided by HITRAN (Gordon et al. 2022) and ExoMol (Tennyson et al. 2016), we generated idealized transmission spectra for a series of hypothetical Venus-like exoplanets whose surface conditions lie on a transect straddling the identified gas–mineral boundary (Figure 5). These spectra were calculated for a planet orbiting a TRAPPIST-1-like star. We emphasize that these transmission spectra are intended only to be illustrative, as they have been generated under the unrealistic assumption of a homogeneous and isothermal atmosphere.

However, these spectra do serve to illustrate that atmospheres on either side of this boundary are likely to be distinguishable by transmission spectroscopy, and that observations could therefore identify the surface pressure–temperature regime – and hence mineralogical regime – of a warm rocky exoplanet. For example, spectral features due to SO₂ at 2.5 μ m, 4.0 μ m, and 5.4 μ m are prominent in one regime but muted in the other, where the CO₂ feature at 4.4 μ m is more prominent.

Atmospheric inhomogeneity, lapse rate, photochemistry and other non-equilibrium processes all undoubtedly influence the chemistry of the upper atmosphere (to which infrared transmission spectra are primarily sensitive; e.g. Kitzmann et al. 2018) and further work would be necessary to determine the best ways to distinguish these atmospheric regimes in practice. None the less, we suggest that the chemistries of these atmospheres are sufficiently distinct that accounting for these complications would not erase the identification of the two regimes in real observations. This is discussed further in Section 5.2.

4 ROLE OF A PLANET’S ELEMENTAL COMPOSITION IN SETTING ATMOSPHERIC REGIME

The correlation described above between atmospheric and mineralogical regimes has been calculated based on the composition of the crust being equal to that inferred for Venus (Rimmer et al. 2021). While short-period (warm) exoplanets may have surface pressure–temperature conditions conducive to achieving gas–mineral thermochemical equilibrium, the existence of a large population of exoplanets with identical crustal compositions to Venus is unlikely.

Estimating the bulk composition of an exoplanet is challenging in practice (Bond et al. 2010; Elser et al. 2012), though constraints may be placed based on the refractory composition of the host star, as obtainable by spectroscopy (e.g., Hinkel et al. 2014; Thiabaud et al. 2015). The composition of the crust can then be estimated from that of the bulk planetary mantle using thermodynamics codes which simulate magmatic processes, such as MELTS (Ghiorso & Sack 1995). Although the link between the stellar and planetary crustal compositions is contingent on unknown properties of the system, previous work may none the less enable a rough estimation

of the relative elemental abundances of refractory elements in an exoplanet’s crust. From this variation, we can then evaluate the robustness of the mineral–atmosphere regimes of Figure 4 and therefore understand how well-constrained the composition of a planet needs to be for its surface mineralogy to be inferable from atmospheric observations.

We address this surface–atmosphere link in the context of changing crustal composition in this section. Starting from Venus’ composition, we vary the abundances of several important elements in turn, and assess the sensitivity of the correlation between the mineralogy and the atmospheric chemistry to the underlying crustal composition. We focus on the same isothermal transect as in Figure 5 (1400 K, running from 10^{−1} bar to approximately 10² bar) for simplicity and ease of visualisation. The compositional ranges we have investigated for refractory elements are motivated by the variation among stars in the Hypatia catalog (Hinkel et al. 2014) and solar system geology. These ranges are likely representative of diverse planetary crusts.

4.1 Calcium

In our calculations the transformation from an SO₂-O₂-dominated atmosphere to a CO₂-dominated atmosphere is driven by the incorporation of SO₂ and O₂ into CaSO₄. The availability of calcium is thus likely to impact the nature of the boundary between the two regimes. Naïvely increasing the abundance of Ca on its own would severely disrupt the redox state of the system, and lead to unrealistic highly-oxidized or -reduced systems: any added Ca would consume available oxygen to produce CaO, and Ca removed would leave behind O. We avoid this issue by adding/subtracting stoichiometric CaO, effectively altering the CaO weight percentage (wt%) of the crust. The relative abundances of all other elements are kept constant.

Figure 6 shows the equilibrium atmospheric compositions and the mole fraction of CaSO₄ for four calcium abundances, expressed as the CaO wt% of the solids. We see that in each panel – independent of the available calcium – the atmosphere is SO₂-O₂-dominated at pressures lower than about $p_0 = 0.4$ bar. At higher pressures, SO₂ and O₂ are incorporated into CaSO₄, leaving CO₂ to prevail as a significant component of the atmosphere. However, the pressure range over which this occurs does depend on the CaO wt%, as does the eventual contribution of CO₂ to the atmospheres at high pressure.

For similar calcium abundances to Venus, mass balance investigations show that reaction 3 occurs at $p_0 = 0.5$ bar, then reaction 4 at $p_0 = 4$ bar. (These pressures given for the above reactions only apply at $T_0 = 1400$ K; Figure 4 shows that at higher temperatures these reactions occur at higher pressures.)

The second panel of Fig. 6 shows that for Venus’ composition, reaction 3 (where diopside acts as a sink for atmospheric sulphur) has little effect on the atmospheric chemistry, with the VMR of CO₂ increasing only slightly from 0.011 to 0.013. There is then a range of pressures for which a small amount of CaSO₄ is present yet atmospheric CO₂ is still very low. For higher CaO wt%, however, reaction 3 desulphidates the atmosphere to such an extent that CO₂ attains a significant VMR at $p_0 = 0.5$ bar and above. The more calcium is available, the more CaMgSi₂O₆ is present in the low-pressure mineralogy compared to CaAl₂Si₂O₈, enabling more SO₂ to be incorporated into the crust at this lower pressure, rather than the higher pressure at which CaSO₄ is formed from CaAl₂Si₂O₈.

By 11.9 wt% CaO, there is enough CaMgSi₂O₆ to react with all of the atmospheric SO₂ at low pressure. This proceeds through a different reaction at a slightly lower pressure of $p_0 = 0.4$ bar, facilitated by the conversion of Mg₂SiO₄ (forsterite, not present for

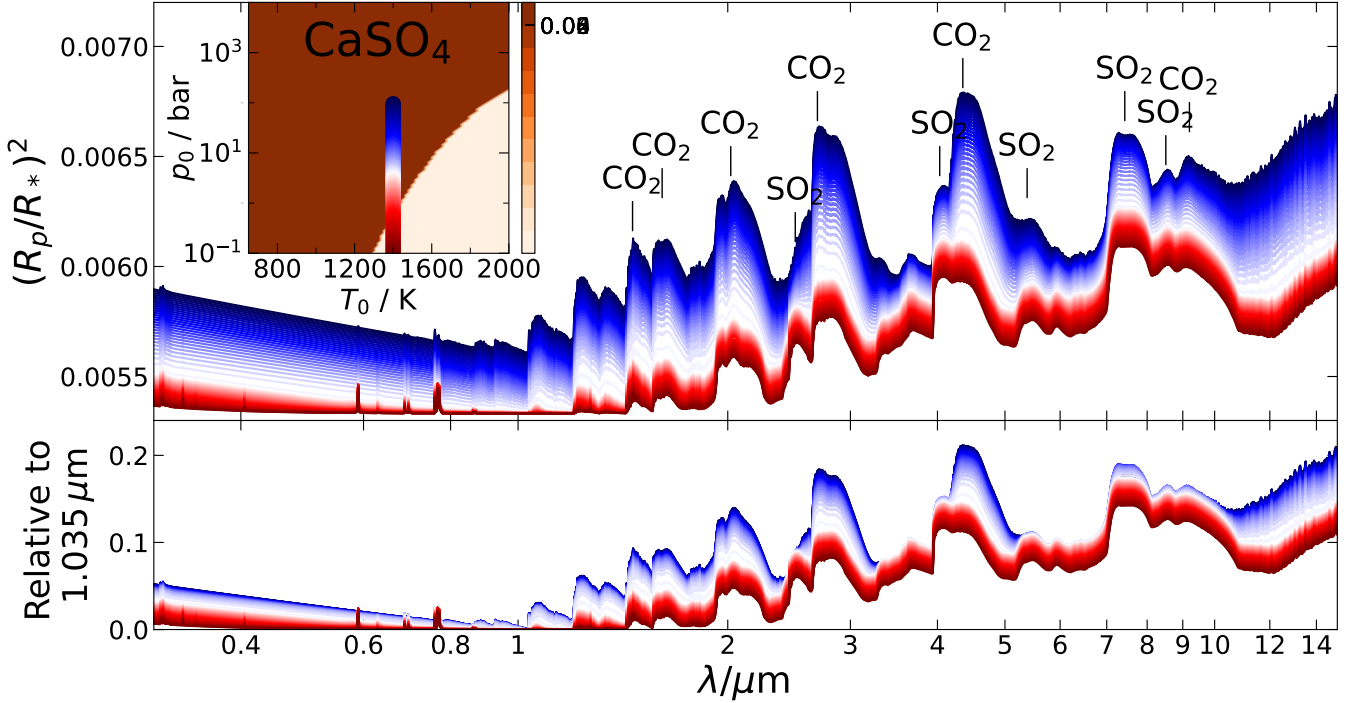
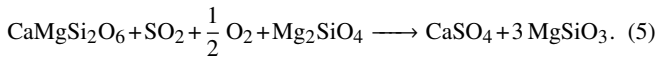


Figure 5. Idealized transmission spectra across the boundary between atmospheric regimes. The inset shows an isothermal transect through p_0 – T_0 space, at $T_0 = 1400$ K, and crossing between the two surface/atmospheric regimes. Assuming planets of Venus’ mass and radius, orbiting a TRAPPIST-1-like star, transmission spectra are shown for homogeneous isothermal atmospheres, whose surface conditions are colour-coded to the inset. The transit radius increases almost monotonically as the surface pressure is increased, due to the reduction in the mean molecular weight on transitioning from an atmosphere dominated by SO_2 ($\mu \approx 64$) to one dominated by CO_2 ($\mu \approx 44$). The lower figure shows the relative deviation in the transit radius from that at $1.035 \mu\text{m}$, and emphasizes the very different spectral shapes, particularly between $2 \mu\text{m}$ and $6 \mu\text{m}$.

lower calcium abundances) to MgSiO_3 :



For Ca-poor crusts, there is little or no $\text{CaMgSi}_2\text{O}_6$ present at low pressure, and the SO_2 can only be absorbed by $\text{CaAl}_2\text{Si}_2\text{O}_8$ at 4 bar. At 3.3 wt% CaO, there is not even sufficient $\text{CaAl}_2\text{Si}_2\text{O}_8$ to absorb all of the SO_2 , and some sulphur remains in the atmosphere even at high pressure. Enough is sequestered in the crust, however, that CO_2 is able to constitute a mixing ratio of order 0.1 in the high-pressure atmospheres, as shown in the top panel of Figure 6.

Although the correlation between the mineralogy and the atmospheric chemistry is more nuanced when the amount of calcium is uncertain, some results remain valid. For the vast majority of compositions and conditions, the presence of CaSO_4 is associated with CO_2 being present in the atmosphere; the only exceptions being where the surface pressure is between 0.5 and 4 bar and the CaO wt% is between about 4 and 9 per cent. Being a highly radiatively active gas, it is likely that CO_2 would be detectable by transmission spectroscopy even if its VMR is small, enabling the inference of surface CaSO_4 for a range of calcium abundances.

4.2 Aluminium

For Venus’ elemental composition, the mineral which is responsible for most of the high-pressure sulphidation of the crust is $\text{CaAl}_2\text{Si}_2\text{O}_8$. The crust’s aluminium content is therefore also likely to be an important factor in determining its effect at thermochemical equilibrium

on atmospheric chemistry. Again, to preserve the oxidation state we add 1.5 O atoms for every Al atom, effectively adding Al_2O_3 to the composition.

Increasing the amount of Al has a very similar effect to *decreasing* the amount of Ca. Both changes to the crustal composition increase the amount of $\text{CaAl}_2\text{Si}_2\text{O}_8$ relative to $\text{CaMgSi}_2\text{O}_6$ at low pressure, which causes more of the SO_2 to be absorbed at 4 bar and less at 0.5 bar. Surfaces which have a high Al_2O_3 weight percentage therefore show an abrupt boundary at around 4 bar where $\text{CaAl}_2\text{Si}_2\text{O}_8$ absorbs all the atmospheric sulphur to form CaSO_4 (by reaction 4); this occurs when the Al_2O_3 is > 18 wt% (cf. Venus: (16 ± 1.8) wt%, [Surkov et al. 1986](#)). Conversely, surfaces with very *low* Al_2O_3 wt% ($\lesssim 2.5$ wt%) have enough of their calcium present as $\text{CaMgSi}_2\text{O}_6$ to absorb all the atmospheric sulphur by reaction 3, giving instead a sharp boundary at 0.5 bar.

4.3 Magnesium

Figure 7 shows the atmospheric chemistry and CaSO_4 abundance at the boundary when the amount of MgO is increased from Venus’ level. Reducing the amount of MgO compared to Venus has very little effect. This could be due to magnesium’s only influence on the chemistry being through $\text{CaMgSi}_2\text{O}_6$ and Mg_2SiO_4 , the latter of which is absent for magnesium contents equal to or less than Venus’.

With a higher MgO content, some magnesium *is* found as Mg_2SiO_4 at low pressure. This enables some SO_2 to be absorbed at 0.4 bar, by reaction 5, as was the case for very CaO-rich crusts. If MgO is $\gtrsim 13$ wt%, there will still be some Mg_2SiO_4 remaining after

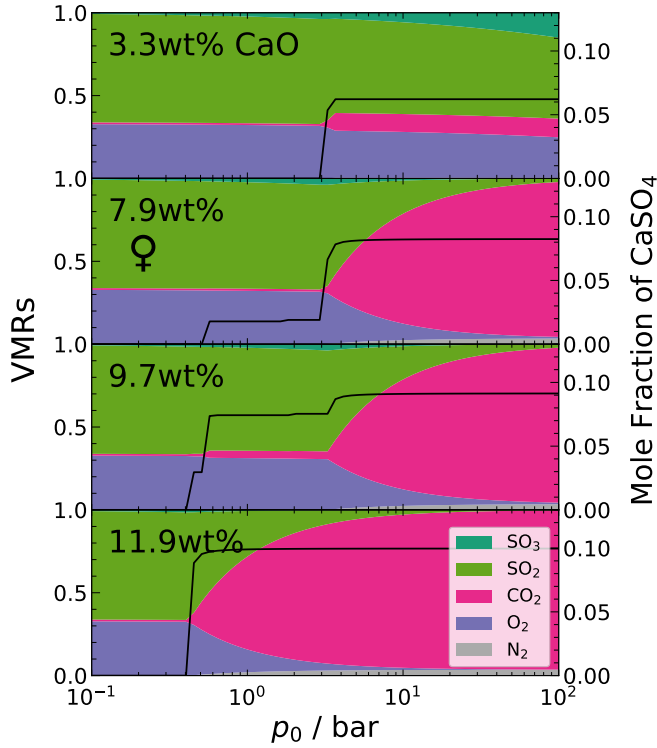
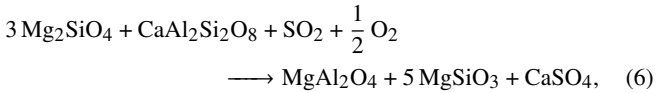


Figure 6. Equilibrium atmospheric compositions for different calcium abundances at 1400 K. The black lines show the abundances of CaSO_4 as a function of pressure. The second panel used Venus’ elemental composition. Depending on the CaO wt%, the sulphidation of the calcium bearing minerals – and hence the change in atmospheric composition – occurs at different pressures, in different stages, and to different extents.

this reaction. This then becomes unstable at around 1.5 bar, forming MgAl_2O_4 and MgSiO_3 by the reaction:



leading to SO_2 being absorbed from the atmosphere at another intermediate pressure. For ≥ 17.9 wt% MgO, the overall picture is complicated by the presence of more exotic phases such as MgFe_2O_4 (magnesioferrite) and $\text{Mn}_3\text{Al}_2\text{Si}_3\text{O}_{12}$ (spessartine), and their auxiliary interactions with the sulphur chemistry. These species cause the emergence of CO_2 in the atmosphere to proceed in several steps, likewise the appearance of CaSO_4 in the crust.

The interference of these additional Mg-bearing minerals with the sulphur chemistry would complicate mineralogical inferences in practice, particularly if an exoplanet’s surface conditions and composition are poorly-constrained initially. However, the subtleties which emerge are confined within a relatively narrow region of p_0 – T_0 space around the boundary, and if other observations suggest that a planet’s surface conditions are far away from the boundary then such inferences would remain feasible.

4.4 Oxygen

As the most abundant metal in the Universe – and the largest constituent of terrestrial planets – variation in the amount of available oxygen is likely. It has long been suggested that Venus itself may

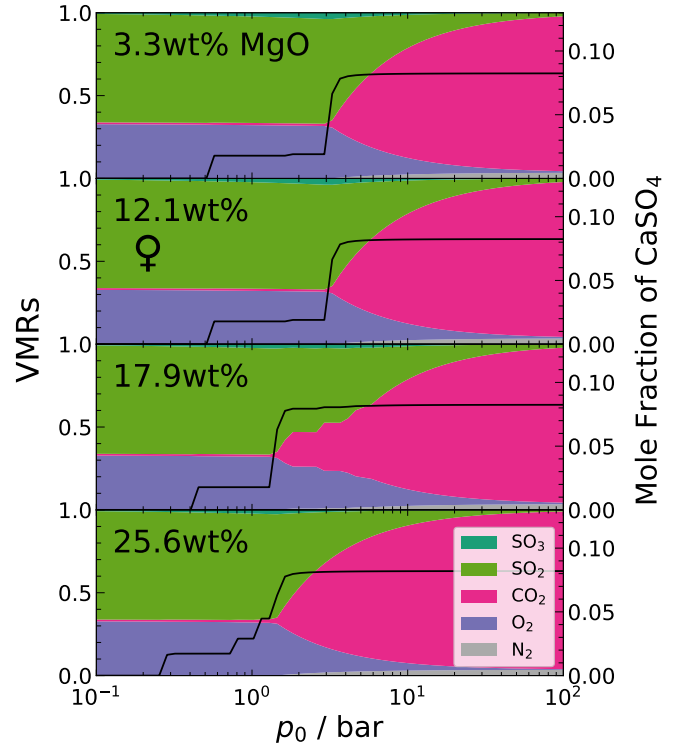


Figure 7. Equilibrium atmospheric compositions for different magnesium abundances at 1400 K. Increasing the amount of MgO in the crust significantly complicates the sulphidation of the crust, which proceeds in several steps. Decreasing the amount of MgO, even very significantly, has little effect.

have once had abundant surface water that was subsequently lost, with a runaway greenhouse effect followed by photolysis of H_2O and hydrogen escape (e.g., [Kasting et al. 1984](#)), resulting in a significant release of O_2 into the atmosphere. O_2 may also be released due to photolysis of CO_2 ([Wong et al. 2019](#)). Whatever its source, liberated oxygen may oxidize surface minerals before accumulating in the atmosphere, leading to an abiotic buildup of O_2 (e.g., [Luger & Barnes 2015](#)). It is therefore prudent to explore the effect of altering the amount of available oxygen to the system.

For each oxygen abundance, we again run GGchem along the same isothermal track in p_0 – T_0 space. The oxygen abundance for a given run is quantified by the fugacity of O_2 at the pressure of 95 bar, $f_{\text{O}_2}^{95}$. f_{O_2} is a useful quantifier of how oxidising the system is and the oxidation state of the surface ([Fegley et al. 1997](#)), and a pressure of 95 bar is chosen to aid comparison with Venus. At 1400 K, this pressure is in the regime where the equilibrium atmospheres are CO_2 -dominated and O_2 -poor, as with Venus. Venus’ surface conditions are however at much lower temperatures, far away from the boundary across which the atmosphere becomes SO_2 -dominated. The $f_{\text{O}_2}^{95}$ values given below, being measured at surface conditions much closer to the boundary, are much higher than the f_{O_2} on Venus’ surface; [Fegley et al. \(1997\)](#) give $\log f_{\text{O}_2}$ between -21.7 and -20.0 at 0 km.

The atmospheric composition and abundance of CaSO_4 for varying oxygen abundances are shown in Figure 8.

Any oxygen which is added relative to Venus’ composition has no noticeable effect on the sulphur chemistry, instead simply increasing the amount of O_2 in an already fully-oxidized system. When enough oxygen is added that $\log f_{\text{O}_2}^{95} > 1.5$, there is so much atmospheric O_2 that when SO_2 and O_2 are incorporated into crustal CaSO_4 , the loss

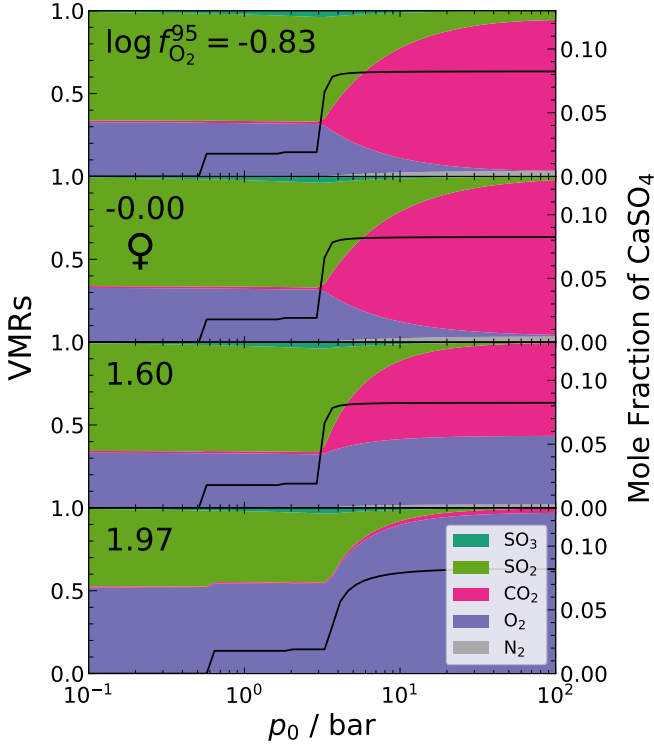


Figure 8. Equilibrium atmospheric compositions for different oxygen abundances at 1400 K. By coincidence, Venus’ elemental composition gives an equilibrium f_{O_2} very close to 1 bar at $p_0 = 95$ bar and $T_0 = 1400$ K, hence $\log f_{\text{O}_2}^{95} \approx 0$ for this composition. Changing the amount of available oxygen has little effect on the formation of CaSO_4 , any excess oxygen simply going into atmospheric O_2 .

of atmospheric SO_2 actually causes the VMR of O_2 to increase. In such oxygen-rich cases, the high-pressure atmospheres are dominated by an increasingly O_2 -rich mixture of CO_2 and O_2 .

Although these atmospheres are very different from those emerging from Venus’ composition, the observational distinguishability of the two regimes is likely to remain. As a homonuclear diatomic molecule, O_2 has very low opacity at infrared wavelengths, and hence would only have a direct effect on exoplanet transmission spectra at visible and UV wavelengths. The more IR-active species CO_2 and SO_2 would have a reduced impact on the spectrum compared with the low-oxygen cases, as their VMRs are drowned out by the dominant O_2 , but with much higher opacities they would likely still be diagnostic of the surface state. Even atmospheres flooded with O_2 are therefore expected to retain the surface–atmosphere correlation we have described.

4.5 Sulphur

Sulphur’s transference from atmospheric SO_2 to CaSO_4 forms the dividing line between the two mineral–atmosphere regimes identified above. It is therefore important to understand the effect of changing its abundance. By analogy with the previous subsection, we measure the amount of available sulphur through the parameter $f_{\text{S}_2}^{95}$, which on the surface of Venus is thought to be between $10^{-4.2}$ and 10^{-6} bar (Zolotov 2018). Again, Venus’ surface conditions are very different to those near the regime boundary we are investigating, so here we find very different fugacities.

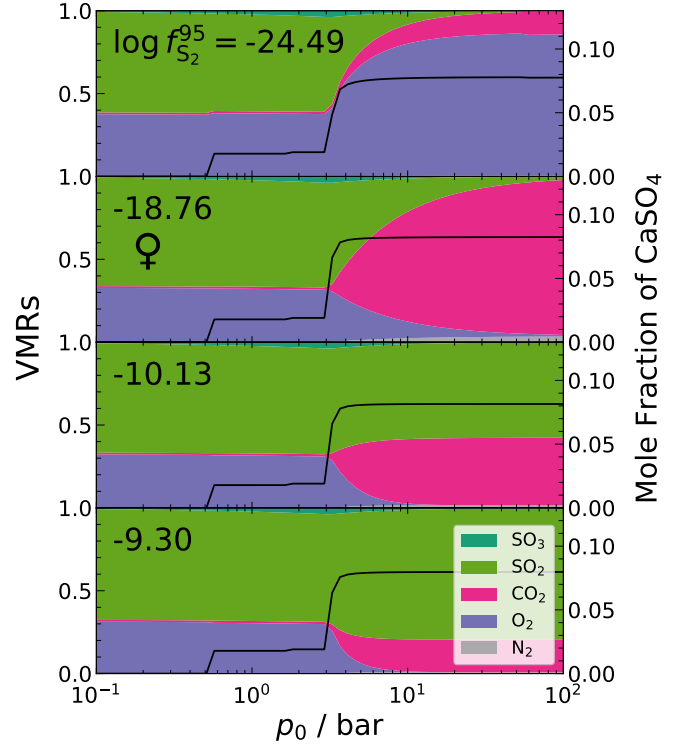


Figure 9. Equilibrium atmospheric compositions for different sulphur abundances at 1400 K. Altering the amount of sulphur has no noticeable effect on the sulphidation process, simply changing the balance between SO_2 and O_2 in the high-pressure atmospheres.

The effects of altering the amount of available sulphur are shown in Figure 9. As with oxygen, changing the sulphur has little effect on the absorption of sulphur into the crust, and the VMR of CO_2 increases when CaSO_4 sequesters the atmospheric SO_2 and O_2 . The only major effect is the balance between SO_2 and O_2 in the atmosphere. Reducing the amount of sulphur has a similar effect to increasing the amount of oxygen. The atmosphere becomes dominated by CO_2 and O_2 when SO_2 is removed into the crust, as shown in the top panel of Figure 9. Increasing the amount of sulphur simply increases the VMR of SO_2 in the high-pressure atmospheres. When $\log f_{\text{S}_2}^{95} > -10$, the atmosphere is so SO_2 -rich that when SO_2 and O_2 form CaSO_4 the loss of O_2 actually causes the SO_2 VMR to increase.

Increasing the amount of sulphur leads to the buildup of a radiatively-active gas (SO_2) superceding the CO_2 at high pressure. In fact, the bottom panel of Figure 9 shows that there is actually a higher VMR of SO_2 at high pressure than at low pressure. When observing such atmospheres, the presence of SO_2 is no longer indicative of a lack of surface CaSO_4 , or the presence of $\text{CaAl}_2\text{Si}_2\text{O}_8$ or $\text{CaMgSi}_2\text{O}_6$. CO_2 maintains its association with the sulphidated mineralogy, though its reduced VMR may hinder its detectability if the exoplanet being observed is more sulphur-rich.

5 DISCUSSION

5.1 Implications

We have demonstrated that the atmospheres and surfaces of warm rocky exoplanets, when coupled sufficiently strongly by thermochemical equilibrium, may enable observational inferences on the surface

mineralogy of such worlds. Facilities already in operation are more than capable of distinguishing between atmospheres that are dominated by CO₂ and those that are dominated by SO₂. Observations using such facilities could therefore determine whether a planet's exterior sulphur reserves are present in its atmosphere or in its crust; we have shown that the intermediate region of surface conditions – where some sulphur is in both – is rather narrow for most compositions.

Knowledge of where the sulphur is also gives a loose constraint on the surface pressure and temperature, information which is difficult to otherwise constrain. Furthermore, knowing whether or not the crust is sulphidated informs the mineralogy: whether calcium is present as CaSO₄, or as CaAl₂Si₂O₈; whether magnesium is present as MgAl₂O₄, or as CaMgSi₂O₆. Inferences on exoplanet mineralogies are as yet unprecedented in astronomy, and represent the natural next step in our ever-more-detailed understanding of the exoplanet population.

Future work would employ the principle of surface–atmosphere equilibrium as a useful lower boundary condition, fixing the atmospheric composition at low altitudes. A natural next step would be to couple the equilibrium to a 1D atmospheric model.

5.2 Caveats and future directions to modelling surface–atmosphere interaction

5.2.1 Chemical Gradients and Non-Equilibrium Processes

The theoretical study of surface–atmosphere equilibrium can of course only yield information about the chemistry at the base of the atmosphere. Observations of exoplanetary atmospheres however probe only the uppermost layers of the atmosphere ($p \lesssim 10^{-2}$ bar, e.g., [Kitzmann et al. 2018](#)). As such, to be able to practically exploit the correlation that we have identified between the surface and atmospheric regimes, we require the chemical distinctiveness of the atmospheric regimes to carry through to higher altitudes/lower pressures. This does not necessarily require the exact chemistry to be homogeneous throughout the atmosphere, but rather that surface regimes none the less produce distinct chemistry at higher altitudes. For example, atmospheres which are SO₂-rich at the surface may instead contain SO₃, SO, COS, or H₂S at higher altitudes, simply due to the relevant sulphur equilibria shifting at lower temperatures and pressures. Though not the direct product of the reactions occurring at the surface, the presence/absence of these species is still capable of indicating whether sulphur is absent/present from the surface mineralogy, and would likely be detectable by transmission spectroscopy if present at high enough VMRs.

Photochemical processes may facilitate the above conversions between sulphur species, or alternatively may oxidize SO₂ to H₂SO₄ droplets, as occurs on Venus (e.g., [Krasnopolsky & Pollack 1994](#)). This process removes SO₂ from Venus' atmosphere efficiently, with the VMR of this species of order 10⁴ lower above the clouds compared to below them ([Mills et al. 2007](#)). This may limit the distinguishability of the two regimes: if all of the sulphur in the SO₂-rich base atmospheres is converted to species which all condense out into clouds and precipitation at higher altitudes, the atmosphere would appear to be sulphur-poor, and one might erroneously infer the presence of CaSO₄ on the surface, where in fact the sulphur is all located above the surface but below the clouds. However, at least for M-dwarf host stars, [Jordan et al. \(2021\)](#) find that their lower UV-flux ensures that SO₂ and other sulphur species survive above the cloud layer of Venus-like planets. For these objects, surface-atmosphere coupling may still be visible in the chemistry of the high-altitude atmosphere.

Vertical mixing can also strongly impact the observable atmospheric chemistry (e.g., [Prinn & Barshay 1977](#)). Although the detailed effects of this phenomenon are beyond the scope of this work, substantial vertical mixing would advect the equilibrium chemistry of the surface up to higher altitudes where it is observable. This tendency to homogenize the atmosphere would act to preserve the distinguishability of the two atmospheric classes (S-rich and S-poor) into the upper atmospheres. Vertical mixing therefore enables transmission spectroscopic observations to more readily identify the regime of the surface mineralogy.

5.2.2 3D planetary effects

The models we have considered are essentially 0-dimensional and do not account for 3D effects that could alter the atmospheric structure and composition. Many known exoplanets are tidally locked, with drastic differences in the conditions on the dayside and nightside of the planet. Different regions of a given planet may therefore occupy different parts of p_0 – T_0 space, perhaps even crossing the boundary we have found. Indeed, perhaps the day–night temperature difference would be so large that surface–atmosphere equilibrium would only be in effect at certain regions of the surface. Counteracting this is atmospheric superrotation, as occurs on Venus (e.g., [Read & Lebonnois 2018](#)), which would mitigate the effects of tidal locking: strong zonal winds would longitudinally homogenize the surface conditions, preventing the strong temperature differential from being communicated to the surface.

Planet hypsometry could have a similar effect: variations in surface altitude (and hence pressure and temperature) may lead to a different mineral assemblages and atmospheric compositions at different altitudes. Indeed, we showed in Section 2 that this is the case on Venus: above a certain altitude the mineralogy may change slightly as FeS₂ becomes stable. However, this is also likely to be a second-order effect: the equilibrium *atmosphere* does not drastically change at the altitude where FeS₂ becomes stable; indeed, only a small amount of FeS₂ is formed. Conversely, this means that FeS₂ would not have sufficient impact on the atmospheric composition for its presence to be inferable from transmission observations of Venus. We therefore stress that surface–atmosphere equilibrium would not permit the inference of altitudinal variations in the mineralogy on a given exoplanet.

5.2.3 Planetary compositional variability

Section 4 investigated the robustness of the mineralogy–atmosphere correlation to variations in the underlying elemental composition, finding that the correlation was maintained for the majority of cases, with certain caveats. We have not, however, explored the full extent of the vast multi-dimensional space of elemental compositions; we have merely looked at the effect of varying certain elements individually. It is possible that crusts which are both, say, more Ca- and Mg-rich than Venus' would exhibit different behaviour. The purpose of this work is simply to demonstrate the principle of relating atmospheric and mineralogical compositions via thermochemical equilibrium, rather than to provide an accurate road map of how the two are related in all possible cases.

The closer an exoplanet's elemental composition is to Venus', the more directly applicable our results are, but it would not be difficult to apply the principles outlined here to planets of differing composition. Our results here demonstrate that, for roughly Venus-composition planets, a boundary exists in p_0 – T_0 space between both atmospheric

and mineralogical regimes. Perhaps in other cases it would be different minerals, or different atmospheric species, that would relate to one another to an observably distinguishable extent. We see no reason why Venus' composition would be special in exhibiting such a boundary between atmospheric and mineralogical regimes.

5.2.4 Sharpness of the regime boundary

The boundary between the two atmospheric/mineralogical regimes is not completely sharp. For most compositions, there is an intermediate regime of surface conditions for which the crust is only partly sulphidated, with some sulphur present in the crust as CaSO_4 and some in the atmosphere as SO_2 or SO_3 . At 1400 K, as examined above, this corresponds to surface pressures between around 0.4 and 4 bar. If the surface conditions happen to lie in this regime, inferences on the mineralogy based on the atmospheric chemistry would be more difficult, as discussed at various points in the previous section. None the less, we are reassured by the narrow width of this intermediate region as shown in Figure 4 for Venus' composition. Although the surface conditions of rocky exoplanets are unlikely to be uniformly distributed across the ranges examined, we see no reason why a large proportion should lie within this strip.

5.2.5 Equilibrium chemistry code limitations

Although GGChem's enormous chemical network accounts for a large number of mineral species, it does not account for mineral solid solution, which for certain phases is strongly favoured in their thermodynamic stability over pure-phase endmember minerals. For example, olivine (any member of the solid solution $(\text{Mg},\text{Fe})_2\text{SiO}_4$) is represented within GGChem only as either pure forsterite (Mg_2SiO_4) or pure fayalite (Fe_2SiO_4). As such, the mineralogies that we have inferred here will not be precisely those that would be stable under the respective equilibrium conditions. This is however likely to be a second-order effect, with our results at least providing an accurate approximation to the true equilibrium mineralogies.

This also highlights consequences for the range of conditions over which surface–atmosphere equilibrium will dominate over other processes, in particular the melting of the crust. At high temperatures, GGChem accounts for the melting of pure solid minerals into their liquid form, at temperatures in excess of 1700 K. However, eutectic melting means that melts will form at lower temperatures than is the case for pure phases. This enables planets to transition to the lava world scenario of Figure 1 at lower temperatures than predicted here, allowing mass transfer and chemical reaction between reservoirs to occur much more rapidly.

As such, although we have limited the scope of this work to surface temperatures below about 1700 K, the assumption of solid–gas equilibrium being the dominant force in setting the atmospheric and surface compositions becomes less valid at the upper end of this temperature range.

6 CONCLUSIONS

We have investigated the thermochemical equilibrium between the surface and the atmosphere of warm rocky exoplanets, finding a boundary in surface pressure–temperature space which delineates the stability regimes of the major gases in the atmosphere, concurrent with the stability of several minerals. This potentially enables observational inferences of the presence or absence of specific minerals on the surfaces of warm rocky exoplanets and of their

surface pressure temperature conditions, using contemporary techniques such as transmission spectroscopy for analysing exoplanetary atmospheres. The inferences we have outlined could provide a complementary method to identifying exoplanet mineralogy, alongside emission spectroscopy.

The surface–atmosphere regime boundary we identify is not entirely sharp, but rather a narrow intermediate region of pressure–temperature space where the mineralogical and atmospheric chemistries transition from one regime to the other. Although this somewhat blurs the picture for a small range of surface conditions, we argue that the correlation that we have found between the two atmospheric/mineralogical regimes is sufficiently stark to enable unprecedented inferences of the surface mineralogy of warm terrestrial exoplanets.

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DATA AVAILABILITY

The GGChem input and abundance files, the resulting output files, and data analysis code will be made available upon acceptance of the manuscript at <https://github.com/xbyrne/sai>.

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APPENDIX A: DEPTH OF ROCK IN EQUILIBRIUM WITH ATMOSPHERE

Whilst on rocky planets the rock–reservoir, by definition, vastly exceeds the atmospheric reservoir in mass, the majority of this is buried and not able to participate in atmospheric chemistry. In the context of surface–atmosphere interaction on Earth, this question is central to the operation of the carbon cycle, and whether climate regulation is rock-limited or not (e.g., [Graham & Pierrehumbert 2020](#)). For the Venus-like case investigated here, the key question is whether there is sufficient unsulfidated calcium in the upper crust to absorb all of the atmospheric SO₂ present in the atmosphere at high temperatures; this paper shows this to be the primary reaction at play.

One could envisage a scenario where there would only be enough

Table B1. Estimated elemental composition of Venus’ crust. This table uses the common astronomical convention $\epsilon_X = 12 + \log_{10}(n_X/n_H)$, where n_X is the number density of nuclei of element X. The procedure for determining these abundances is described in section 3.2 of [Rimmer et al. \(2021\)](#) which estimates Venus’ surface composition using *Vega 2* measurements of oxide mass fractions ([Surkov et al. 1986](#)).

Element	ϵ	Element	ϵ
H	12.0000	Mg	18.5693
C	16.0985	Ca	18.2402
N	14.9591	Al	18.6107
O	19.5591	Na	17.9238
F	10.3444	K	16.4410
S	17.8827	Ti	16.5126
Cl	9.73721	He	11.1931
Fe	18.1441	Ne	10.8129
Mn	16.4092	Ar	11.9591
Si	18.9942		

anorthite and diopside to desulfidate the atmosphere with a crust ‘skin depth’ of many kilometres. In such a scenario, it would be unlikely that the atmospheric regimes we have identified would be able to occur in practice; rock kilometers deep in the crust would likely be shielded from atmospheric interaction on the time-scales of hundreds of millions of years. However, on sub-metre scales, it is reasonable that fractures and permeability of the crust would allow it to be in effective contact with the atmosphere to this depth, facilitating the reactions needed to reach equilibrium.

The number of sulfur atoms in a given atmospheric column is given by N_S , where

$$p_0 = \int \rho g dz = \mu g \int n dz = \frac{\mu g}{x_S} \int n_S dz = \frac{\mu g N_S}{x_S} \quad (\text{A1})$$

where x_S is the combined mixing ratio of all sulfur-bearing gases. As the Ca:S stoichiometry of reactions 3 and 4 is 1:1, each atmospheric sulfur atom could in principle combine with one calcium atom in the formation of CaSO₄. The skin depth is therefore:

$$L_{Ca} = \frac{N_S}{n_{Ca}} = \frac{x_S p_0}{\mu g n_{Ca}}$$

where n_{Ca} is the number density of calcium atoms in the solids. We find that over the pressure–temperature conditions investigated, $L_{Ca} < 15$ cm. Over this range we would expect fractures in the crust to enable equilibrium to be reached with the lower atmosphere.

APPENDIX B: ELEMENTAL COMPOSITION OF VENUS

The elemental composition of Venus that we have used throughout is estimated using section 3.2 of [Rimmer et al. \(2021\)](#). This procedure uses oxide mass fractions as measured by the *Vega 2* lander, and assumes an SO₂ VMR of 150 ppm at the surface.

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