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Membrane/electrolyte interplay on ammonia motion inside a flow-cell for electrochemical nitrogen and nitrate reduction /  
Pirrone, N.; Garcia Ballesteros, S.; Hernández, S.; Bella, F.. - In: ELECTROCHIMICA ACTA. - ISSN 0013-4686. - ELETTRONICO. - 493:(2024), pp. 1-12. [10.1016/j.electacta.2024.144415]

*Availability:*

This version is available at: 11583/2991098 since: 2024-07-22T13:01:14Z

*Publisher:*

Elsevier

*Published*

DOI:10.1016/j.electacta.2024.144415

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# Membrane/electrolyte interplay on ammonia motion inside a flow-cell for electrochemical nitrogen and nitrate reduction

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## ARTICLE INFO

### Keywords:

Ammonia electrosynthesis  
Cation exchange Nafion membrane  
Celgard membrane  
Ammonium crossover  
Flow-cell reactor

## ABSTRACT

Electrochemical ammonia production from molecular nitrogen and nitrate reduction reactions at ambient conditions has gained a lot of attention in recent years, making this topic more and more appealing. The race towards good and quick results in terms of Faradaic efficiency and productivity is not always focused on the possible source of ammonia contamination. In particular, Nafion membrane is the most commonly used in this field as cell separator, discarding the possible known disadvantages coming from ammonium ions absorption and release. The wettable microporous Celgard membrane has been proposed as a substitute for Nafion membranes, despite the separation mechanism, in this case, is only dimension-driven, so it does not assure ammonium ions retention. This paper reveals that the mechanism of ammonium ions absorption and release by Nafion 117 is strongly related to the cations present in the electrolyte and to a lesser extent by its pH value. On the other hand, Celgard membrane does not show any relevant ammonium ions absorption. Moreover, the different trend of ammonium ions motion from catholyte to anolyte solution inside a flow-cell reactor shows that none of the membranes is able to avoid ammonium ions crossover and that there is a correlation between the applied potential and the motion trend. Electrochemical nitrogen and nitrate reduction tests confirm how Nafion membrane can have a big impact on the final result of ammonium ions production, especially when dealing with low production quantities, leading to mistakes in the real quantity of ammonium ions coming from the reduction reaction.

## 1. Introduction

Ammonia ( $\text{NH}_3$ ) is a backbone source of nitrogen fertilisers playing a crucial role in global agriculture and food production, and with an annual production of *ca.* 150 Mt is the second largest manufactured chemical. In fact,  $\text{NH}_3$  is estimated to be responsible for the existence of nearly half of the world population [1] and its synthesis at the industrial scale is surely the biggest scientific discovery of the 20th century [2]. In recent years, due to its peculiar properties (*i.e.*, high energy density equal to  $4.32 \text{ kWh L}^{-1}$  at liquid state, significant hydrogen content up to 17.6 wt%, and easy liquefaction for handling, storage, and transportation),  $\text{NH}_3$  has emerged as a potential fuel on the route towards decarbonisation [3–5]. Several authors even talk about the implementation of a  $\text{NH}_3$ -based economy in the future [6–9].

Up to now,  $\text{NH}_3$  is industrially almost exclusively produced (90%) *via* the Haber-Bosch process, which relies on fossil fuels as the main source of hydrogen ( $\text{H}_2$ ) precursor [10]. This process requires high

temperatures (300–500 °C) to off-set the sluggish kinetics of the reaction and high pressures (200–300 atm) to avoid  $\text{NH}_3$  decomposition at such elevated temperatures, making each industrial plant highly energy demanding ( $485 \text{ kJ mol}^{-1}$ ) [11]. As a consequence, *ca.* 2% of the world energy production is consumed and 400 Mt of carbon dioxide ( $\text{CO}_2$ ) are annually released to the environment. Therefore, approaching an  $\text{NH}_3$ -based economy makes the development of environmentally friendly and energy-efficient alternative production processes mandatory.

In such a scenario, one approach that holds great potential is the electrochemical  $\text{NH}_3$  synthesis under mild conditions based on the nitrogen ( $\text{N}_2$ ) or nitrate ( $\text{NO}_3^-$ ) reduction reactions (E-NRR or E- $\text{NO}_3$ RR) [12,13]. Those processes can offer many advantages apart from the obvious elimination of fossil fuels and  $\text{CO}_2$  emissions: (1) converting sustainable energy in remote areas to transportable chemicals; (2) smaller plant infrastructure favouring the on-site  $\text{NH}_3$  generation and, thus, its decentralisation; (3) E- $\text{NO}_3$ RR provides a pathway for

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efficiently converting one of the most widespread groundwater pollutant (*i.e.*,  $\text{NO}_3^-$ ) into a high-valued-added product (*i.e.*,  $\text{NH}_3$ ). [12, 14,15] Although substantial research work focuses on the study of both E-NRR/E- $\text{NO}_3\text{RR}$ , they remain in preliminary stages, facing challenges mostly related to the low selectivity of the system, which leads to low efficiencies [16,17]. On the other hand, part of the know-how developed in recent years in the field of  $\text{CO}_2$  electroreduction can be precious for the electrochemical conversion of nitrogenous species [18–20].

Among the different electrolytic cells employed to perform E-NRR/E- $\text{NO}_3\text{RR}$ , the use of divided cells (*e.g.* H-type cell or flow-cells) is considered the most suitable approach, since it minimises the possible re-oxidation at the anode of the reduced species from the cathode. [21] In those systems, catholyte and anolyte are usually divided by an ion exchange membrane, being Nafion the most commonly used [22]. Such membranes are made of a poly(tetrafluoroethylene) (PTFE) backbone and sulfonic groups at the side chains, which are actively involved in the proton transport [23]. However, the interaction of Nafion membranes with ammonium ions ( $\text{NH}_4^+$ ) is known and some studies question its use, especially for E-NRR, where the amount of  $\text{NH}_4^+$  produced is still low and a minimum source of pollution can easily lead to false positives [22, 24–27]. Despite the transport properties of cation-exchange membranes are quite important, with particular relevance to the contamination problems, in-depth understanding of the influence that the membrane and its interactions with the electrolyte have in the E-NRR/E- $\text{NO}_3\text{RR}$  have not yet received enough attention.

On the other side, Celgard membranes have always been employed for energy storage device separators [28–31]. Contrary to Nafion membranes, such separators have a microporous structure, commonly made in poly(propylene) and poly(ethylene). The transport mechanism does not involve an interaction between the species in the electrolyte and the structure of the membrane itself, making it free from any possible contaminations [32,33]. For such property, this separator became attractive for the world of E-NRR and E- $\text{NO}_3\text{RR}$ , and it was proposed as a substitute for Nafion membranes in the rigorous protocol for electrochemical  $\text{NH}_3$  synthesis published in 2019 by Andersen et al [25]. Since then, the number of works employing Celgard membranes as separators has increased in both E-NRR and E- $\text{NO}_3\text{RR}$  applications [34–38]. However, information regarding its influence on the system and how  $\text{NH}_4^+$  can migrate between anolyte and catholyte is still missing.

Some studies addressing the membrane- $\text{NH}_4^+$  interaction in E-NRR/E- $\text{NO}_3\text{RR}$  systems can be found in the literature [22,25,27,39–42]. For example, in the context of this work, Leonardi et al. [42] compared the  $\text{NH}_4^+$  uptake by Nafion and Zifron membranes in  $\text{NH}_4^+$  enriched water and  $\text{Na}_2\text{SO}_4$  0.1 M electrolyte (pH 6.5) and transport inside an H-type cell; Gyenge et al. [40] employed Nafion membranes of different thicknesses in combination with three electrolytes at pH values equal to 1 (with  $\text{H}_2\text{SO}_4$  0.05 M), *ca.* 6 (with  $\text{Na}_2\text{SO}_4$  0.1 M), and 13 (with  $\text{NaOH}$  0.1 M) in a membrane electrode assembly-type electrolytic cell, thus testing the  $\text{NH}_4^+$  crossover and volatilization as  $\text{NH}_3$ . Wilder et al. [39] studied the influence of three different kinds of membranes (*i.e.*, anion exchange membrane, cation exchange membrane, and porous membrane) on the  $\text{NH}_4^+$  motion inside two different cell architectures (*i.e.*, H-type and zero-gap gas diffusion electrode (GDE)-type). However, there is still disagreement among the reported results, values for the  $\text{NH}_3$  crossover from the catholyte to the anolyte varied from 2% [27] to 25–49% [22,39,40] depending on the cell architecture, membrane thickness, electrolyte pH, etc. Moreover, in most of the studies,  $\text{Na}_2\text{SO}_4$ , strong acids (*e.g.*  $\text{H}_2\text{SO}_4$ ,  $\text{HCl}$ ) or bases (*e.g.*  $\text{NaOH}$ ) are used as electrolytes and detailed information about how the cation species present in the electrolyte or applied potential can affect the membrane- $\text{NH}_4^+$  interaction is still missing.

This paper aims to give a full understanding of how  $\text{NH}_3$  interacts with both Nafion and Celgard membranes, as well as its motion inside a flow-cell reactor during E-NRR and E- $\text{NO}_3\text{RR}$ . For the first time, the influence of the electrolyte composition (in terms of type of cation and pH) on the membrane- $\text{NH}_4^+$  interaction is assessed. In particular, the

dependence of  $\text{NH}_4^+$  absorption on the membranes is investigated under different conditions to find whether there is a correlation between the other cationic species present in the solution and the pH, or not. Moreover, the  $\text{NH}_4^+$  passage across each membrane inside a flow-cell microreactor (three-chamber configuration) without and with applying a potential has also been considered. Such experiments allow us to have a better understanding of the possible source of  $\text{NH}_4^+$  contamination in different operating conditions and the effect they have on the overall  $\text{NH}_4^+$  production performance assessment.

## 2. Experimental

### 2.1. Materials

Nafion membranes (Nafion 117, N-117) with a nominal thickness of 183  $\mu\text{m}$  were purchased from Quintech. Celgard 3401 separator with nominal thickness of 25  $\mu\text{m}$  was purchased from Celgard.

Anhydrous lithium, sodium, and potassium sulphates ( $\text{Li}_2\text{SO}_4$ ,  $\text{K}_2\text{SO}_4$ , and  $\text{Na}_2\text{SO}_4$ ,  $\geq 99.0\%$ ) were purchased from Sigma-Aldrich and used to prepare electrolytes with a concentration of 0.1 M Sulfuric acid ( $\text{H}_2\text{SO}_4$ , 95.0%–98.0%), used to modify the pH of the solutions, as trap for  $\text{NH}_3$ , and to activate N-117, potassium hydroxide ( $\geq 85.0\%$ ), used to modify the pH of the solutions, and hydrogen peroxide ( $\text{H}_2\text{O}_2$ , 30 wt%), used to clean N-117 membrane, were purchased from Sigma-Aldrich. Ammonium chloride ( $\text{NH}_4\text{Cl}$ , 99.99%), used as  $\text{NH}_4^+$  source, was also bought from Sigma-Aldrich. Sodium citrate dihydrate ( $\text{HOC}(\text{COONa})(\text{CH}_2\text{COONa})_2 \cdot 2\text{H}_2\text{O}$ ,  $\geq 99.0\%$ ), sodium salicylate ( $\text{C}_7\text{H}_5\text{NaO}_3$ , 99.5%), sodium nitroferricyanide(III) dihydrate ( $\text{Na}_2[\text{Fe}(\text{CN})_5\text{NO}] \cdot 2\text{H}_2\text{O}$ , 99.0%), sodium hypochlorite ( $\text{NaClO}$ , 5% active chlorine), and sodium hydroxide ( $\text{NaOH}$ ,  $\geq 97.0\%$ ), used to measure the  $\text{NH}_4^+$  concentration in the electrolyte via the salicylate method (described in Section 2.2), were purchased from Sigma-Aldrich. Potassium nitrate ( $\text{KNO}_3$ ,  $\geq 99.0\%$ ) was purchased from Sigma-Aldrich and used as nitrate source in E- $\text{NO}_3\text{RR}$  experiments. Commercial molybdenum disulfide ( $\text{MoS}_2$ ,  $<2\ \mu\text{m}$ , 99.0%) powder, used as catalyst in E-NRR and E- $\text{NO}_3\text{RR}$ , was also purchased from Sigma-Aldrich. All the solutions were prepared with ultrapure water (Milli-Q).

### 2.2. Ammonia quantification

The concentration of  $\text{NH}_4^+$  was determined spectrophotometrically by using the salicylate method optimized by Giner-Sanz et al [43]. In a typical way, 2 mL of samples were added with 240  $\mu\text{L}$  of salicylate catalyst solution (2.75 M  $\text{C}_7\text{H}_5\text{NaO}_3$  and 0.95 mM  $\text{Na}_2[\text{Fe}(\text{CN})_5\text{NO}] \cdot 2\text{H}_2\text{O}$ ), followed by the addition of 400  $\mu\text{L}$  of alkaline  $\text{NaClO}$  solution (previously prepared by diluting 1:10 the  $\text{NaClO}$  5% in a aqueous solution containing  $\text{HOC}(\text{COONa})(\text{CH}_2\text{COONa})_2 \cdot 2\text{H}_2\text{O}$  340 mM and  $\text{NaOH}$  465 mM). After being kept under dark for 45 min, the absorbence at 650 nm of the mixed samples was measured using a HITACHI U-500 UV spectrophotometer. The concentration was determined according to the calibration lines obtained before (see Figure S1).

### 2.3. Ammonia absorption/releasing tests

The absorption tests were performed by immersing the membranes ( $5 \times 5\ \text{cm}$  size) in a beaker containing 70 mL of Milli-Q water or the studied electrolytes (*i.e.*,  $\text{K}_2\text{SO}_4$ ,  $\text{Na}_2\text{SO}_4$ , and  $\text{Li}_2\text{SO}_4$ ) and  $\text{NH}_4^+$  2 mg  $\text{L}^{-1}$  at three different pH values (*i.e.*, 3.5, 7.0, and 11.0) for 24 h. The salt concentration for all the electrolytes was 0.1 M and the pH was adjusted to the desired value by dropwise addition of diluted  $\text{H}_2\text{SO}_4$  or KOH. Regarding the desorption tests, the previously absorbed membranes were immersed for another 24 h in 70 mL of fresh solutions (Milli-Q water,  $\text{K}_2\text{SO}_4$ ,  $\text{Na}_2\text{SO}_4$ , and  $\text{Li}_2\text{SO}_4$ ). In both cases, samples were periodically taken and analysed to monitor the  $\text{NH}_4^+$  concentration evolution.

Before tests, N-117 membranes were activated following a

commonly used method consisting of 3 boiling steps at *ca.* 80 °C in 3 wt % H<sub>2</sub>O<sub>2</sub> for 1 h, then in Milli-Q water, and finally in H<sub>2</sub>SO<sub>4</sub> 0.5 M, rinsing between and after the boiling step with Milli-Q water.

#### 2.4. Electrochemical setup and measurements

A commercial micro-flow-cell reactor (ElectroCell A/S) equipped with a GDE and assembled in a three-chambers configuration (*i.e.*, gas and liquid on the cathodic side, liquid in the anodic one, see Figure S2) was used for NH<sub>4</sub><sup>+</sup> motion experiments and E-NRR/E-NO<sub>3</sub>RR tests. When needed, the electrochemical tests were performed with a VSP-300 potentiostat (Biologic). The electrode area was 10.2 cm<sup>2</sup> and the distance between the electrode and the membrane was 6.2 mm. The membranes were placed as the separator of the two liquid chambers. The respective electrolytes (*i.e.*, K<sub>2</sub>SO<sub>4</sub>, Na<sub>2</sub>SO<sub>4</sub>, or Li<sub>2</sub>SO<sub>4</sub>) were continuously recirculated from reservoir bottles through a peristaltic pump at a flow rate of 20 mL min<sup>-1</sup>. The catholyte and anolyte total volume was 50 mL and reservoir bottles were sealed to avoid possible NH<sub>3</sub> losses, especially when working under alkaline conditions. For NH<sub>4</sub><sup>+</sup> motion experiments, 2 mg L<sup>-1</sup> of NH<sub>4</sub><sup>+</sup> were added to the catholyte reservoir bottle. All the experiments lasted for 3 h and samples were periodically taken from both catholyte and anolyte to monitor the NH<sub>4</sub><sup>+</sup> concentration; after the tests, membranes were immersed in fresh electrolyte to desorb the possible absorbed NH<sub>4</sub><sup>+</sup>, and the liquid was analysed after 24 h. When no potential was applied to the cell, a bare carbon paper (Toray, 0.19 mm thickness - 20 wt% PTFE treated) was placed as working electrode (WE). Instead, the tests conducted applying a potential difference were carried out employing a GDE consisting of commercial MoS<sub>2</sub> catalyst deposited on the carbon paper through air-brushing. Commercial MoS<sub>2</sub> was chosen as active material; indeed, even if it is not the best catalyst for E-NRR/E-NO<sub>3</sub>RR, due to its intrinsic catalytic activity for water reduction, it is still expected to be active towards E-NRR/E-NO<sub>3</sub>RR, considering that molybdenum and sulphur play key roles in N<sub>2</sub> fixation [44]. The counter electrode (CE) at the anode was a mixed metal oxide electrode with iridium, while a leak-free Ag/AgCl electrode, placed close to the cathode (at 5 mm), was used as reference electrode. Ar or N<sub>2</sub> gas (purity 99.9999%), further purified by passing through a commercial filter (Agilent) following an already published procedure [45], were supplied at 5 mL min<sup>-1</sup> in the gas chamber at the back side of the cathode, to allow a simpler release of the gaseous products (H<sub>2</sub> and NH<sub>3</sub>).

All electrochemical NH<sub>4</sub><sup>+</sup> synthesis (E-NRR and E-NO<sub>3</sub>RR) studies were performed according to the following protocol: (i) open circuit voltammetry (OCV) until reaching stability; (ii) cyclic voltammetry (CV) at a scan rate of 10 mV s<sup>-1</sup> and linear sweep voltammetry (LSV) at 5 mV s<sup>-1</sup> within 0 and -2 V vs Ag/AgCl; (iii) chronoamperometry (CA) or chronopotentiometry (CP) at the selected potential or current density; (iv) OCV until reaching stability again. For E-NRR, steps (i) and (ii) were carried out first with Ar and then again under N<sub>2</sub>. The full protocol was repeated using only Ar atmosphere as a control test. Samples were taken at the beginning, after the LSV (ii), and after the last OCV (iv) from both cathodic and anodic sides, and membranes were immersed in fresh electrolyte to desorb the possible absorbed NH<sub>4</sub><sup>+</sup>. In the case of E-NRR, the sample called "background" refers to a sample taken after (i) and (ii) steps in Ar (protocol scheme in Figure S3). Moreover, in those experiments, a trap consisting of 20 mL of H<sub>2</sub>SO<sub>4</sub> 0.02 M was placed in the gas exit to recover possible volatilised NH<sub>3</sub>. The total height of the 20 mL acid trap was 5 cm and it was open to air, thus no back-pressure was applied to the cell. E-NO<sub>3</sub>RR was performed in K<sub>2</sub>SO<sub>4</sub> 0.1 M + KNO<sub>3</sub> 0.01 M as catholyte and K<sub>2</sub>SO<sub>4</sub> 0.1 M as anolyte, while E-NRR was performed in K<sub>2</sub>SO<sub>4</sub> 0.1 M both as catholyte and anolyte. All electrolytes were saturated with Ar before entering the cell.

Faradaic efficiency (FE) and productivity (P) were used as figures of merit to assess the system performance. FE is defined as follows:

$$FE = \frac{n_{NH_3} \cdot F \cdot n_{e^-}}{Q}$$

where  $n_{NH_3}$  is the quantity of NH<sub>3</sub> produced (mol),  $F$  is the Faraday's constant,  $n_{e^-}$  is the number of electrons needed to produce a mole of NH<sub>3</sub> (*i.e.*, 3 mol of e<sup>-</sup> per mol NH<sub>3</sub> for E-NRR, 8 mol of e<sup>-</sup> per mol NH<sub>3</sub> for E-NO<sub>3</sub>RR), and  $Q$  is the total charge applied (C). The NH<sub>4</sub><sup>+</sup> considered for FE calculations is that produced during the CA or CP, disregarding that produced during the CV and LSV. Instead,  $P$  is defined as:

$$P = \frac{n_{NH_3}}{t}$$

where  $t$  is the reaction time, expressed in h.

#### 2.5. Conductivity measurements

Conductivity of N-117 membranes as a function of the electrolyte used was measured through electrochemical impedance spectroscopy (EIS) technique. The membranes were cut into 1 cm diameter disks using a manual puncher. Then, they were cleaned in H<sub>2</sub>SO<sub>4</sub> 1 M solution in an ultrasonic bath for 1 h and rinsed with Milli-Q water. Before electrochemical testing, membranes were ion-exchanged in the desired solution in an orbital mixer for 7 days changing the solution to a fresh one every 3 days. In particular, it was chosen to compare the conductivity of the membranes exchanged in Milli-Q, Li<sub>2</sub>SO<sub>4</sub>, K<sub>2</sub>SO<sub>4</sub>, and Na<sub>2</sub>SO<sub>4</sub> at 0.1 M concentration with and without NH<sub>4</sub><sup>+</sup> in solution. The measurements were carried out in a symmetrical T-cell (depicted in Figure S4), equipped with a glassy carbon disk as both cathode and anode, supported by a stainless-steel piston. The membrane was dried with a paper to remove the excess electrolyte and thickness was measured with a thickness gauge. An oscillating voltage of 10 mV at OCV was applied between the WE and the CE in a frequency range between 100 kHz and 10 mHz. The value of the impedance at high frequency is considered as the resistance of the membrane and conductivity is calculated using the following equation:

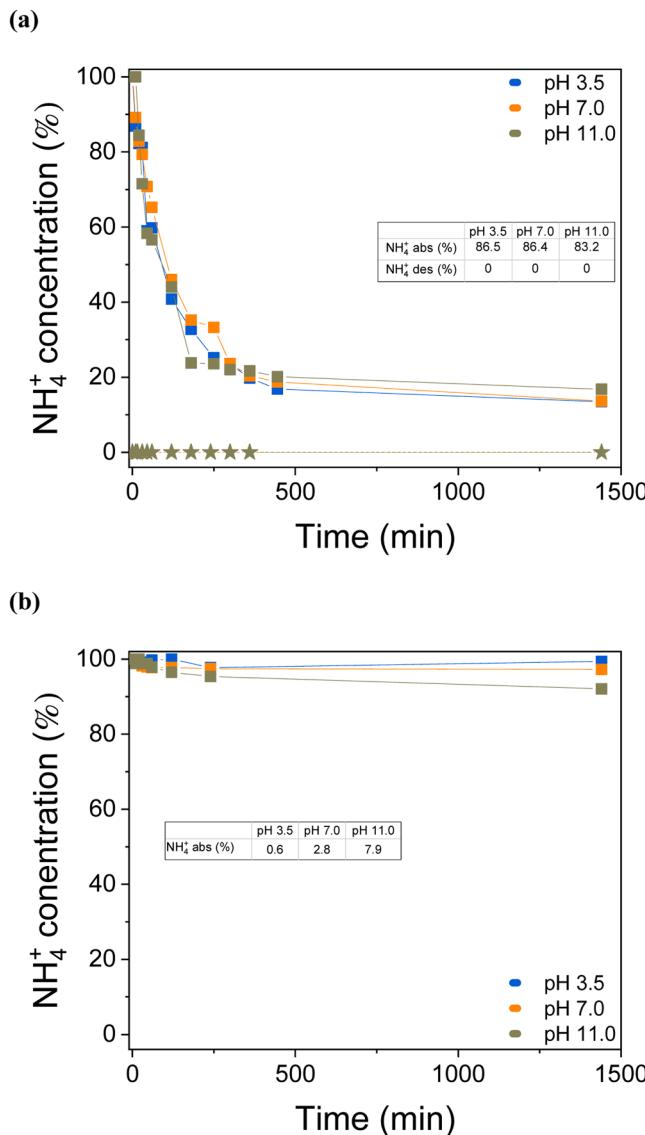
$$\kappa = \frac{l}{R_m \cdot A}$$

where  $\kappa$  is the conductivity (mS cm<sup>-1</sup>),  $l$  is membrane thickness (cm),  $A$  is the membrane area (given 1 cm diameter, it was 0.78 cm<sup>2</sup>), and  $R_m$  (Ω) is the membrane resistance, extrapolated from the Nyquist's plot using Zview software. When carrying out these experiments, the operating conditions were maintained the same to have consistency and be able to compare the conductivity in different electrolytes. This work is not focused on proton conductivity calculations and does not pretend to have a precise quantification of that value. The aim is to investigate if the presence of NH<sub>4</sub><sup>+</sup> can affect proton conductivity when compared to the values here obtained when no NH<sub>4</sub><sup>+</sup> is added.

### 3. Results and discussion

#### 3.1. Ammonium ion absorption/release by Nafion/celgard membranes

The capability of Nafion and Celgard membranes to absorb NH<sub>4</sub><sup>+</sup> ions was first tested in Milli-Q water without any supporting electrolyte (apart from the H<sub>2</sub>SO<sub>4</sub> or KOH added to adjust the pH, the total concentrations were never higher than 0.16 mM or 1 mM, respectively) by immersing them in a solution containing 2 mg L<sup>-1</sup> of NH<sub>4</sub><sup>+</sup>. Since pH 11 exceeds the NH<sub>4</sub><sup>+</sup> pK<sub>a</sub> (9.25) and NH<sub>4</sub><sup>+</sup> deprotonates into NH<sub>3</sub>, which can easily pass into the gas phase, a 24 h control test on NH<sub>4</sub><sup>+</sup> enriched (2 mg L<sup>-1</sup>) Milli-Q water at pH 11.0 was done to evaluate the NH<sub>3</sub> lost into the gas phase, and a decrease in the NH<sub>4</sub><sup>+</sup> concentration of only 5% was found (see Figure S5). As it can be observed in Fig. 1(a), NH<sub>4</sub><sup>+</sup> concentration decreases up to 86.5% of the initial value for the N-117 membrane, while almost no variations were observed for the Celgard one



**Fig. 1.**  $\text{NH}_4^+$  absorption in and release from N-117 (a) and Celgard (b) membranes in Milli-Q water at pH 3.5, 7, and 11.  $\text{NH}_4^+$  concentration refers to the ratio between  $\text{NH}_4^+$  measured at the immersion time  $t$  and the initial one ( $2 \text{ mg L}^{-1}$ ). The solid lines and -■- symbols represent the absorption, while dashed lines and -★- symbol the release. Note that, for Celgard membrane, release test was not done due to the low amount of total  $\text{NH}_3$  absorbed on the membrane.

(Fig. 1(b)). Such behaviour is expected considering that N-117 is a cation exchange membrane and Celgard is an inert and size-selective one. These results are thus in line with those already reported in the literature [40,42]. Regarding pH, negligible differences in the  $\text{NH}_4^+$  absorption percentage are observed along the test for the three studied values, being the trend after 24 h as follows: absorption at pH 3.5  $\approx$  7.0  $>$  11.0 (Fig. 1(a)). After the absorption, the releasing tests were performed by immersing the previously absorbed membranes in fresh Milli-Q water. As it can be appreciated in Fig. 1(a), no releasing was obtained, since in Milli-Q water no ion exchange can take place. Next, membranes previously absorbed in  $\text{NH}_4^+$  enriched ( $2 \text{ mg L}^{-1}$ ) Milli-Q water were immersed in fresh electrolytes (*i.e.*,  $\text{K}_2\text{SO}_4$ ,  $\text{Na}_2\text{SO}_4$ , and  $\text{Li}_2\text{SO}_4$  0.1 M) at pH 3.5, 7.0, and 11.0. As shown in Table 1, the amount of  $\text{NH}_4^+$  released increases with the atomic radius of the cation in the electrolyte ( $\text{Li}^+ < \text{Na}^+ < \text{K}^+$ ) and the pH value; a release of 89.6% is observed for  $\text{K}_2\text{SO}_4$  at pH 11, while only a 50.9% of  $\text{NH}_4^+$  release is obtained for  $\text{Li}_2\text{SO}_4$  at pH 3.5. The desorption trend is presented in

**Table 1**

$\text{NH}_4^+$  absorption-desorption (%) values from N-117 membrane previously absorbed in Milli-Q water containing  $2 \text{ mg L}^{-1}$  of  $\text{NH}_4^+$  at pH 3.5, 7, and 11 and desorbed in  $\text{Li}_2\text{SO}_4$ ,  $\text{Na}_2\text{SO}_4$ , and  $\text{K}_2\text{SO}_4$  0.1 M at pH 3.5, 7, and 11.

	pH 3.5		pH 7.0		pH 11.0	
	%Ass (in $\text{H}_2\text{O}$ )	%Des	%Ass (in $\text{H}_2\text{O}$ )	%Des	%Ass (in $\text{H}_2\text{O}$ )	%Des
Li	88.0	50.9	88.3	55.7	87.7	61.5
Na	88.8	66.5	90.0	78.5	90.0	79.6
K	90.1	82.8	91.3	83.2	92.0	89.6

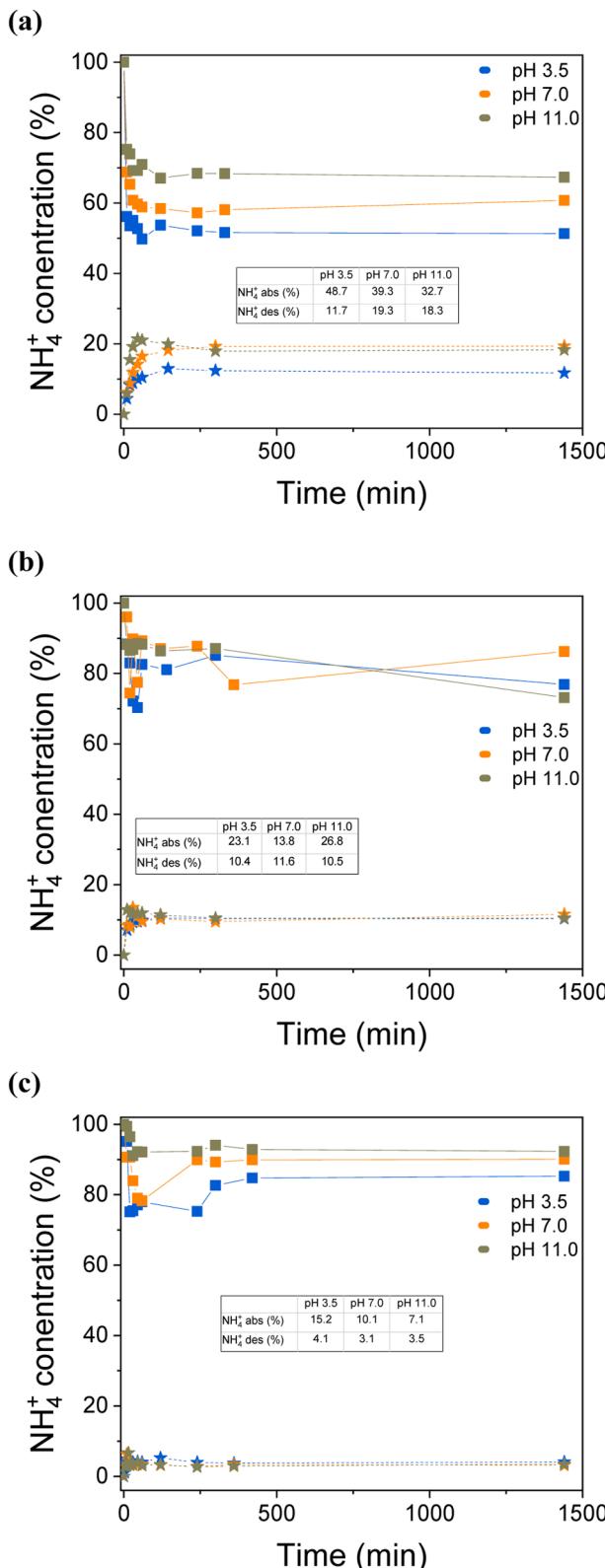
Figure S6. As for the Celgard membrane, no releasing tests were performed since no  $\text{NH}_4^+$  absorption was observed.

To get a deeper understanding of the role that the different cations (*i.e.*,  $\text{Li}^+$ ,  $\text{Na}^+$ , and  $\text{K}^+$ ) have in the  $\text{NH}_4^+$  exchange equilibrium, absorption/release tests were later performed employing all the three electrolytes previously used, again at acidic, neutral, and basic pH (*i.e.*, 3.5, 7.0, and 11.0). Regarding the  $\text{NH}_4^+$  absorption, as illustrated in Fig. 2, it decreases with the increase of the ionic radius, namely  $\text{Li}^+ > \text{Na}^+ > \text{K}^+$ , and the trend is the same independently of the solution pH. Furthermore,  $\text{NH}_4^+$  absorption seems to be correlated also to the cation hydration number, which increases in the order  $\text{Li}^+ > \text{Na}^+ > \text{K}^+$ . It is possible to assume that, when the cation is more solvated by water molecules, it is less reactive with N-117 membrane, facilitating  $\text{NH}_4^+$  absorption. On the contrary, when the cation is less solvated by water molecules, it more easily reacts with the membrane, hindering  $\text{NH}_4^+$  absorption. Moreover, an oscillation of the  $\text{NH}_4^+$  concentration in the first hour of absorption, followed by a stabilization of the value, is observed, meaning that at the beginning the cation in the electrolyte (*i.e.*,  $\text{Li}^+$ ,  $\text{Na}^+$  or  $\text{K}^+$ ) and  $\text{NH}_4^+$  are in continuous competition for the N-117 membrane sulfonic groups until an equilibrium is reached at *ca.* 180 min. However, for the  $\text{NH}_4^+$  release (as previously observed), the opposite trend is obtained, that is, it decreases with increasing the ionic radius. All these results reveal the preference of the membrane for the cations and the increase of this affinity with the ionic radius, as it has been already stated in different works [46–49]. Regarding the pH, a minimal effect with no defined trend on the  $\text{NH}_4^+$  absorption and release from N-117 membrane was observed. Regarding adsorption, it generally increases with the decreases of the pH, being the difference higher for  $\text{Li}_2\text{SO}_4$ , while for the desorption no trend was observed. Moreover, a release of 100% of the absorbed  $\text{NH}_4^+$  was obtained only when  $\text{K}_2\text{SO}_4$  0.1 M was employed, confirming the difficulty of removing the  $\text{NH}_4^+$  from the N-117 membrane.

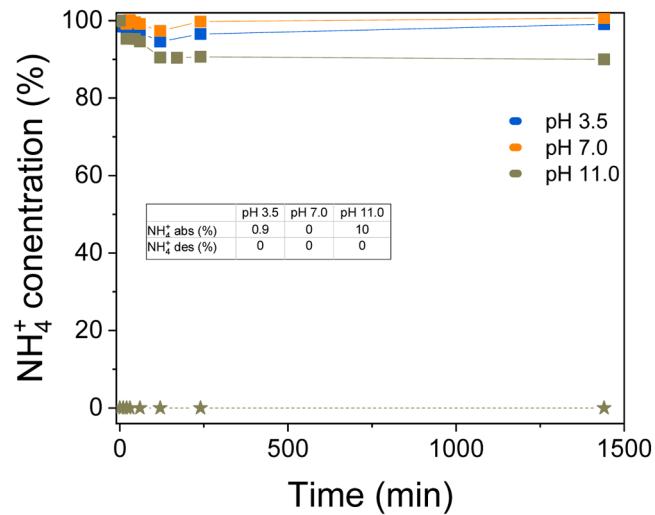
Finally, the same experiments were carried out for Celgard membrane, but only for the  $\text{Li}_2\text{SO}_4$  solution, since it is the one that shows the highest  $\text{NH}_4^+$  uptake. As obtained for Milli-Q water experiments, no  $\text{NH}_4^+$  absorption was detected after the 24 h immersion, except at pH 11.0 where a decrease of 10% (*ca.* three times lower than with N-117) was obtained (Fig. 3), again confirming the inert character of the membrane.

### 3.2. Ammonia motion inside a flow-cell

The  $\text{NH}_4^+$  cross-over from the catholyte to the anolyte employing N-117 and Celgard membranes as separators and the influence of the electrolyte composition were first studied without applying any potential. Since  $\text{NH}_4^+$  absorption in batch conditions is more affected by the type of electrolyte than the pH, in the following tests the pH was always neutral and only the effect of the salt cation was studied. Comparing both membranes (Fig. 4),  $\text{NH}_4^+$  losses at the cathodic side were similar for N-117 and Celgard, and for different electrolyte solutions, being the decrease only slightly bigger when employing N-117 membrane as separator and  $\text{Na}_2\text{SO}_4$  and  $\text{Li}_2\text{SO}_4$  0.1 M as electrolyte. Despite the catholyte  $\text{NH}_4^+$  losses are similar for both membranes, the increase at the anodic side is considerably bigger when N-117 (Figure 4(a)) is used as separator. Moreover, membranes were desorbed in fresh electrolytes



**Fig. 2.** NH<sub>4</sub><sup>+</sup> absorption and release from N-117 membrane in Li<sub>2</sub>SO<sub>4</sub> 0.1 M (a), Na<sub>2</sub>SO<sub>4</sub> 0.1 M (b), and K<sub>2</sub>SO<sub>4</sub> 0.1 M (c) electrolytes at pH 3.5, 7, and 11. NH<sub>4</sub><sup>+</sup> concentration refers to the ratio between NH<sub>4</sub><sup>+</sup> measured at the immersion time *t* and the initial one (2 mg L<sup>-1</sup>). The solid lines and -■- symbols represent the absorption, while dashed lines and -★- symbol the release.



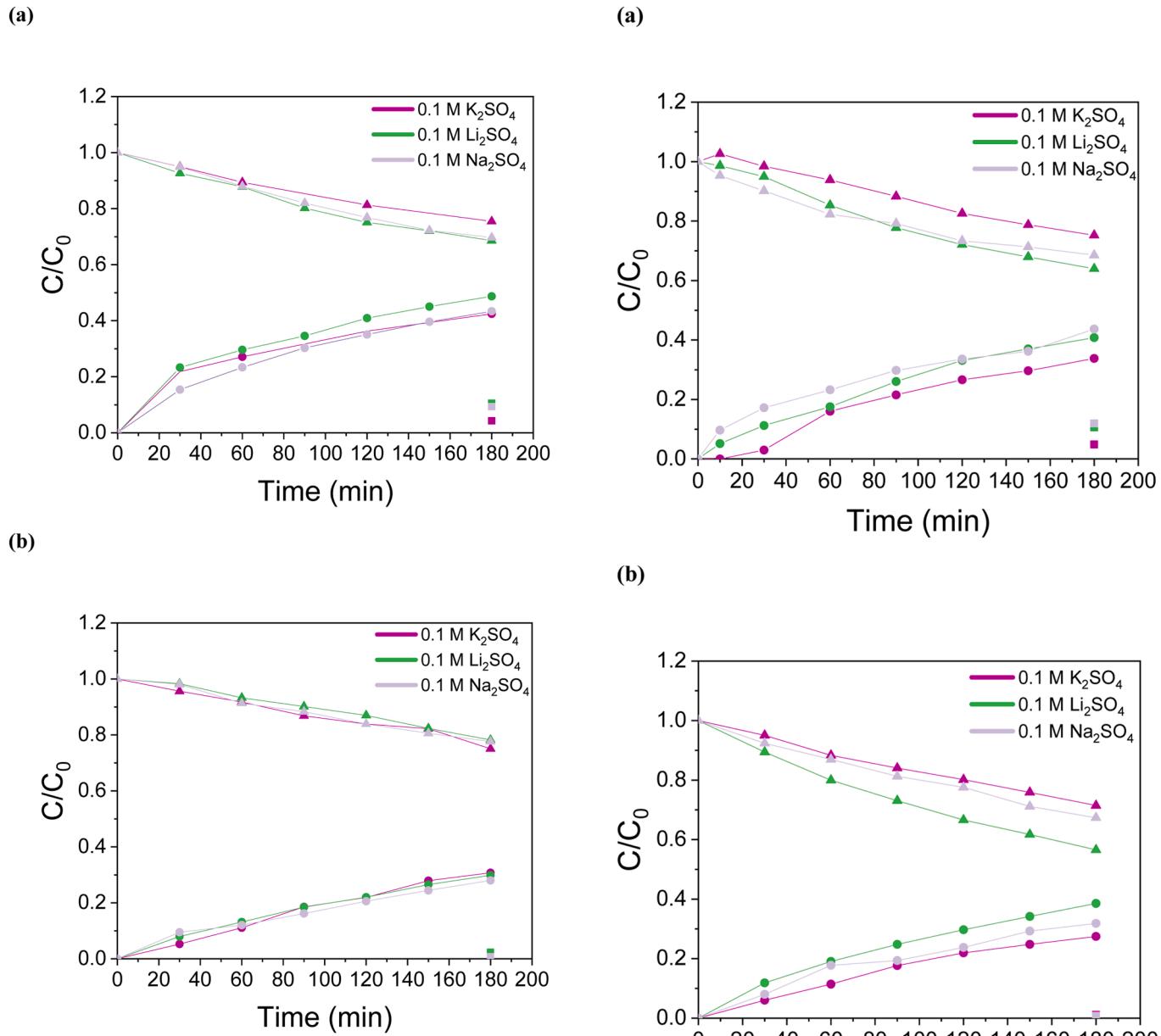
**Fig. 3.** NH<sub>4</sub><sup>+</sup> absorption in and release from Celgard membrane in Li<sub>2</sub>SO<sub>4</sub> 0.1 M electrolyte at pH 3.5, 7, and 11. NH<sub>4</sub><sup>+</sup> concentration refers to the ratio between NH<sub>4</sub><sup>+</sup> measured at the immersion time *t* and the initial one (2 mg L<sup>-1</sup>). The solid lines and -■- symbols represent the absorption, while dashed lines and -★- symbol the release.

and only N-117 released a considerable amount of NH<sub>4</sub><sup>+</sup>. In general, NH<sub>4</sub><sup>+</sup> motion in the case of Celgard membrane - when no potential is applied - is independent of the cation nature, suggesting that NH<sub>4</sub><sup>+</sup> has small enough dimensions to cross Celgard pores and that the concentration gradient is the only driving force.

Finally, if we sum the amount of NH<sub>4</sub><sup>+</sup> found in the different compartments together with the one desorbed from the membrane, the value obtained for N-117 is bigger than the initially added amount, confirming the introduction of NH<sub>4</sub><sup>+</sup> in the system by the N-117, even if they are new and cleaned right before the experiment in H<sub>2</sub>SO<sub>4</sub> 1 M [22,41]. Additionally, it appears that the NH<sub>4</sub><sup>+</sup> contaminations present in the membrane are released in the anolyte. Such behaviour could be explained by the concentration gradient driving force: NH<sub>4</sub><sup>+</sup> is released in the anolyte within the first 30 min, where the initial concentration of NH<sub>4</sub><sup>+</sup> is zero.

Next, we repeat the same experiments, but applying a low potential commonly adopted in literature for E-NRR/E-NO<sub>3</sub>RR, *i.e.*, -1.4 V vs Ag/AgCl (J-t curves for these experiments are provided in Figure S7). As illustrated in Fig. 5(b), in the case of Celgard membrane, the NH<sub>4</sub><sup>+</sup> cross-over is higher compared to the case in which no potential is applied (Figure 4(b)), especially when employing Li<sub>2</sub>SO<sub>4</sub> 0.1 M as electrolyte. Moreover, it seems that, when a slight electric field is applied, a dependence of NH<sub>4</sub><sup>+</sup> motion from the cation shows up, making NH<sub>4</sub><sup>+</sup> motion dependent on electronegativity values of the cation in the electrolyte (Li<sup>+</sup> > Na<sup>+</sup> > K<sup>+</sup>). Indeed, Li<sup>+</sup> tendency to be attracted towards the negative side (cathode) of the cell is higher, being membrane pores less occupied by Li<sup>+</sup>. This aspect, together with the acidification of the anolyte, favours the diffusion of NH<sub>4</sub><sup>+</sup> to the anode. Indeed, Celgard membrane is a porous separator, and no chemical bond arises between the polymer matrix and the cations in the electrolyte. As both NH<sub>4</sub><sup>+</sup> and Li<sup>+</sup> have dimensions small enough to cross that separator, the only driving forces governing the cross-over phenomena are concentration gradient, pH, and electronegativity. Regarding N-117 membrane (Fig. 5(a)), results are mostly similar to those obtained when no potential is applied, probably because the potential is so low that it does not influence the NH<sub>4</sub><sup>+</sup> motion.

Finally, experiments by increasing the potential to -2 V vs Ag/AgCl (Fig. 6) were performed with the K<sub>2</sub>SO<sub>4</sub> 0.1 M electrolyte; for N-117 membrane, no NH<sub>4</sub><sup>+</sup> crossover was detected. Such results confirm that the NH<sub>4</sub><sup>+</sup> crossover not only depends on the interaction with the membrane itself and the competition with the other cations present in the electrolytes, but also on the applied electrical field; regarding N-117

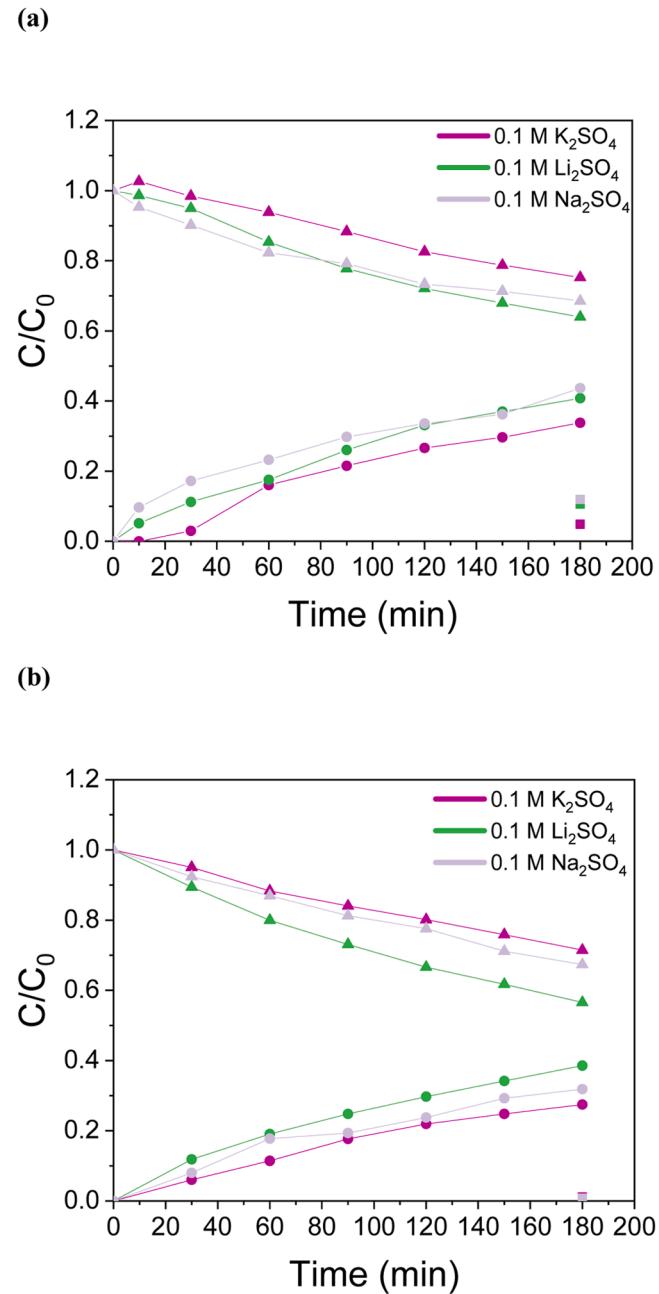


**Fig. 4.**  $NH_4^+$  motion inside a flow-cell employing different electrolytes without applying any potential, using (a) N-117 and (b) Celgard membranes as separator. The electrolytes tested are  $Li_2SO_4$ ,  $Na_2SO_4$ , and  $K_2SO_4$  with a 0.1 M concentration and a pH value of 7. The initial  $NH_4^+$  concentration is  $2 \text{ mg L}^{-1}$ . In the plot, (-▲-) is used to identify the cathodic side, while (-●-) is used to identify the anodic side on the cell. (-■-) refers to the  $NH_4^+$  quantity found absorbed on the membrane. That value is the ratio between the quantity found and the initial quantity present in the cathodic side.

membrane, at increased potential, all positive charges are dragged to the cathodic side charged negatively, overcoming the diffusion across the membrane. On the contrary, in the case of Celgard, the  $NH_4^+$  cross-over increases as the potential increases, until an equilibrium between catholyte and anolyte is reached, as already experienced by Andersen et al [25]. Accordingly, it is not recommended to look at the application of a high potential as a general method to prevent  $NH_4^+$  crossover.

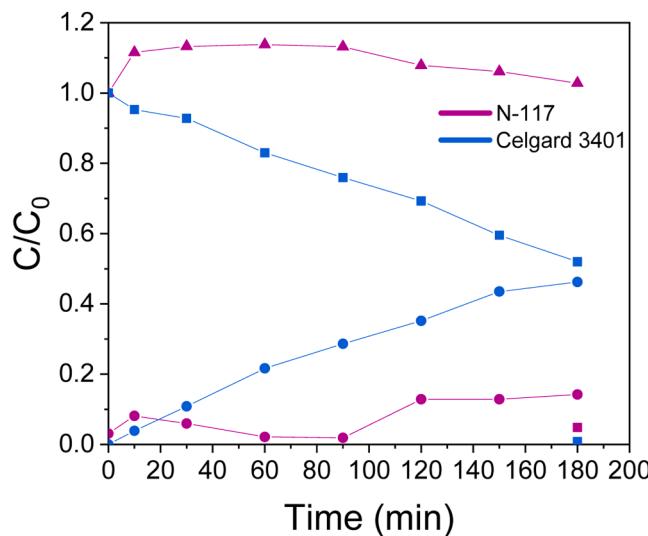
### 3.3. E- $NO_3RR$ experiment

The ability of  $MoS_2$  to convert  $NO_3^-$  into  $NH_4^+$  was studied using both N-117 and Celgard membranes. At  $-1.4 \text{ V vs Ag/AgCl}$  applied potential for 3 h, the FE and  $P$  values obtained using the N-117 membrane were



**Fig. 5.**  $NH_4^+$  motion inside a flow-cell employing different electrolytes during an electrolysis test at  $-1.4 \text{ V vs Ag/AgCl}$  for 3 h, employing (a) N-117 and (b) Celgard membranes as separator. The electrolytes tested are  $Li_2SO_4$ ,  $Na_2SO_4$ , and  $K_2SO_4$  with a 0.1 M concentration. The initial  $NH_4^+$  concentration is  $2 \text{ mg L}^{-1}$ . In the plot, (-▲-) is used to identify the cathodic side, while (-●-) is used to identify the anodic side on the cell. (-■-) refers to the  $NH_4^+$  quantity found absorbed on the membrane. That value is the ratio between the quantity found and the initial quantity present in the cathodic side.

10.6% and  $0.79 \mu\text{mol h}^{-1}$ , respectively, while those obtained using Celgard were 17.6% and  $0.85 \mu\text{mol h}^{-1}$ , respectively. The discrepancy between those values can be attributed to two main factors: (i) the  $NH_4^+$  produced in the case of N-117 is in part absorbed by the membrane, and this amount cannot be considered in the calculations; (ii) in the test using the Celgard membrane, the activity was lower (see LSV and CA plots in Figure S8) and, at the same applied potential, the current density measured was  $-0.1 \text{ mA cm}^{-2}$  on average, while it was  $-0.16 \text{ mA cm}^{-2}$  for N-117. Such a difference in the total charge can have a big impact on the final calculation of the FE for the same amount of  $NH_4^+$  produced.



**Fig. 6.**  $\text{NH}_4^+$  motion inside a flow-cell employing  $\text{K}_2\text{SO}_4$  0.1 M as electrolyte during an electrolysis test at  $-2$  V vs  $\text{Ag}/\text{AgCl}$  for 3 h, employing N-117 (purple) and Celgard (blue) membranes as separator. The initial  $\text{NH}_4^+$  concentration is  $2 \text{ mg L}^{-1}$ . In the plot, (-▲-) is used to identify the cathodic side, while (-●-) is used to identify the anodic side on the cell. (-■-) refers to the  $\text{NH}_4^+$  quantity found absorbed on the membrane. That value is the ratio between the quantity found and the initial quantity present in the cathodic side.

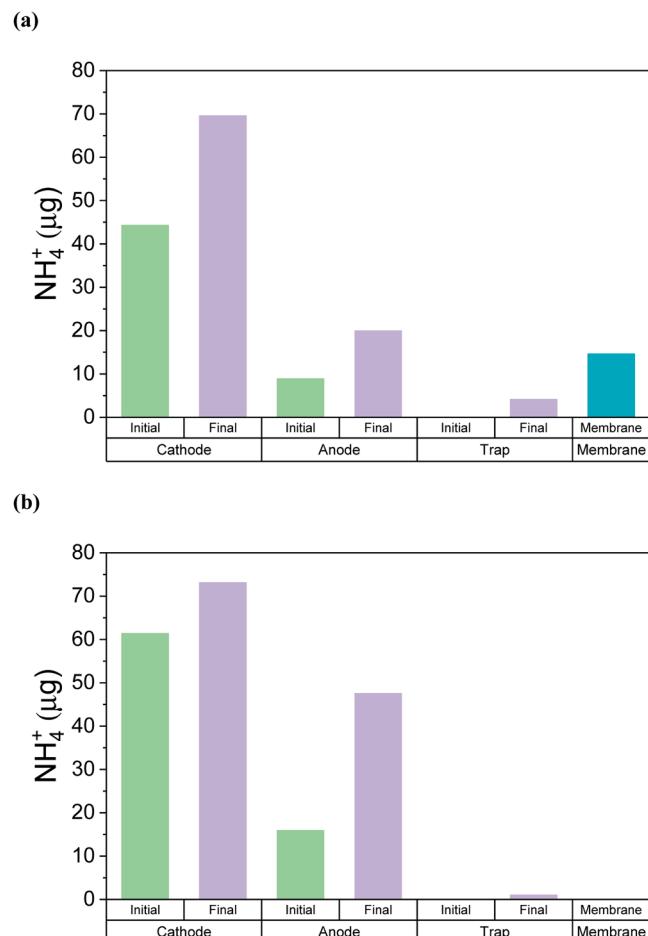
The uncompensated resistance between the WE and reference electrode in the case of N-117 and Celgard were  $4.98 \Omega$  and  $4.13 \Omega$ , respectively. That difference is negligible and should not affect the resulting current density measured during the CA. In Fig. 7(b) it is possible to see that, in the case of Celgard membrane,  $\text{NH}_4^+$  is distributed more uniformly in the anolyte and catholyte, while no  $\text{NH}_4^+$  was present on the membrane after the test. On the other side, in the case of N-117,  $\text{NH}_4^+$  is much more present in the catholyte side, together with a considerable amount on the membrane itself (Fig. 7(a)).

#### 3.4. E-NRR experiment

The performances of Nafion and Celgard membranes were tested for E-NRR. Unfortunately,  $\text{MoS}_2$  commercial catalyst did not show any relevant activity towards  $\text{NH}_4^+$  production. Indeed, using both N-117 and Celgard, only a FE equal to 0.02% towards  $\text{NH}_4^+$  production was reached at  $-5 \text{ mA cm}^{-2}$  for 2 h, and the control test under argon atmosphere showed the same result, meaning that all the  $\text{NH}_4^+$  present in the system is coming from ambient contamination.

From Fig. 8(a) we can see that, in the case of N-117,  $\text{NH}_4^+$  quantity increases at the cathodic side, but having a look at the anodic one, it is evident that  $\text{NH}_4^+$  coming from initial impurities is gradually crossing the membrane, reaching the catholyte chamber. Such behaviour is in agreement with the test on  $\text{NH}_4^+$  motion presented in Section 3.2. In this case,  $\text{NH}_4^+$  is moving from the anolyte to the catholyte due to the big electric field applied during the test. In addition, having a look at the control test under argon, shown in Fig. 8(b), the difference between the total final and initial  $\text{NH}_4^+$  quantity is the same as the one obtained when using  $\text{N}_2$ , which highlights that no  $\text{N}_2$  is effectively reduced to  $\text{NH}_4^+$  and all the quantities found come from system impurities.

On the other side, from Fig. 9 it is clear that, using Celgard membrane, the obtained quantities of  $\text{NH}_4^+$  are significantly lower than those detected with N-117. In this case, quantities are coherent both under  $\text{N}_2$  (Figure 9(a)) and Ar (Fig. 9(b)) environments and remain stable during all the tests. This further confirms that the majority of the contamination is coming from the membrane itself. Even if never used and thoroughly cleaned with  $\text{H}_2\text{SO}_4$  1 M, the N-117 membrane releases some  $\text{NH}_4^+$ . It is the authors' opinion that N-117 can absorb  $\text{NH}_3$  from the ambient in the

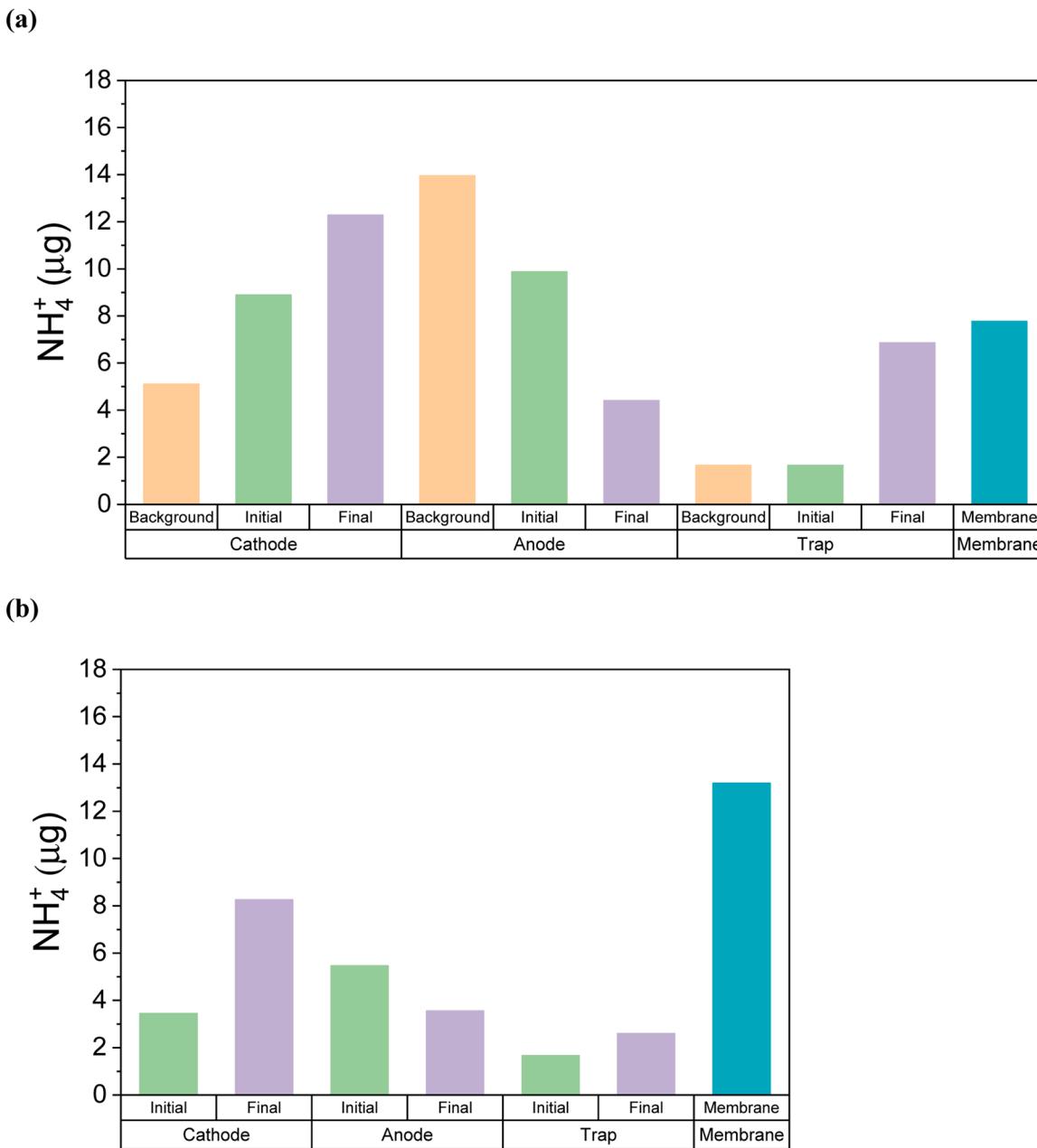


**Fig. 7.**  $\text{NH}_4^+$  produced through E- $\text{NO}_3$ RR in a flow-cell reactor, employing  $\text{MoS}_2$  as a catalyst in combination with  $\text{K}_2\text{SO}_4$  0.1 M +  $\text{KNO}_3$  0.01 M as electrolyte and (a) N-117 and (b) Celgard membranes as separator. Initial value refers to the quantity of  $\text{NH}_4^+$  found in the sample taken after CV and LSV measurements, right before the CA at  $-1.4$  V vs  $\text{Ag}/\text{AgCl}$  for 3 h (see Figure S3 for protocol details).

time interval needed to assemble the cell. In this framework, Celgard seems to be a good alternative as it is not a source of contamination. However, such a membrane is size-selective, allowing the passage of all the species that have a small enough dimension. Such conditions do not allow the holding of different electrolyte compositions on the catholyte and anolyte sides.

#### 3.5. Effect of cations and $\text{NH}_4^+$ on the conductivity of the Nafion membrane

As previously reported in literature, cations and  $\text{NH}_4^+$  can interact with the external sulfonic group of the Nafion chains and contaminate Nafion membranes, hindering the correct proton transport [27]. This phenomenon can affect the overall system not only in terms of  $\text{NH}_4^+$  misleading quantification, but also in the control of the bulk pH at the cathode and anode side. Indeed, during the  $\text{H}_2$  evolution reaction,  $\text{H}^+$  are consumed at the cathode side and local pH rises, while  $\text{H}^+$  are produced at the anode side during  $\text{O}_2$  evolution reaction. In an ideal situation, ionic equilibrium should be maintained thanks to proton transport across the cation exchange membrane. However, other species present in the electrolyte can negatively affect proton transport, leading to an increase in the catholyte and a decrease in the anolyte bulk pH. E-NRR and E- $\text{NO}_3$ RR have a very strong dependence on the pH [50,51] and such low control on that variable could lead to worse overall



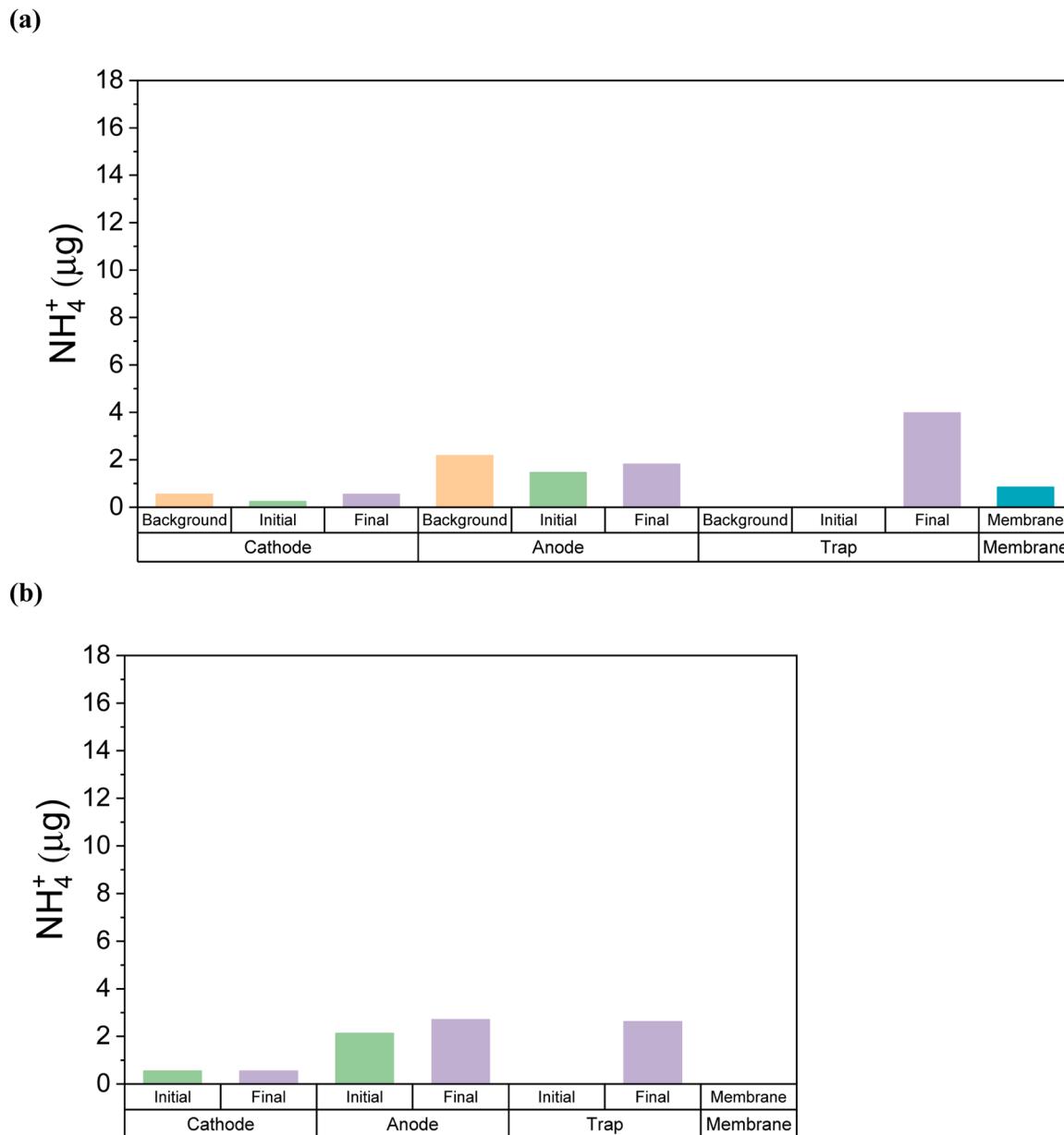
**Fig. 8.**  $\text{NH}_4^+$  produced through E-NRR in a flow-cell reactor, employing  $\text{MoS}_2$  as a catalyst in combination with  $\text{K}_2\text{SO}_4$  0.1 M as electrolyte and N-117 membrane as separator. (a) Test under  $\text{N}_2$ , background refers to a sample taken after the OCV, LSV, and CV in Ar. Initial value refers to the quantity of  $\text{NH}_4^+$  found in the sample taken after CV an LSV measurements under  $\text{N}_2$ , right before the CP technique at  $-5 \text{ mA cm}^{-2}$  for 2 h (b) Control test under Ar. Initial value refers the sample taken after CV an LSV measurements in Ar, right before the CP at  $-5 \text{ mA cm}^{-2}$  for 2 h (see Figure S3 for protocol details).

performances as well as catalyst stability issues [52,53]. Thus, the purpose of this set of experiments is to assess how proton conductivity of Nafion membrane is affected by the presence of the cations and  $\text{NH}_4^+$  in the electrolyte.

The comparison between the conductivity of Nafion membrane exchanged in  $\text{Li}^+$ ,  $\text{K}^+$ ,  $\text{Na}^+$ , and sole Milli-Q water is presented in Fig. 10. The conductivity decreases up to 83% for the cation-exchanged membrane compared to the proton-exchanged one. The trend shown is in agreement with previous literature data [54] and can be explained through a parallelism with the cation mobility mechanism in water electrolytes: while protons are transported through the hopping mechanism, other cations are transported through the vehicle mechanism, which is slower than the previous one [55].

The variability of conductivity between the different exchanged

cations can be explained through the cation correlation with membrane water content, as well as with water molecule mobility inside the membrane [56]. Indeed, as shown in Fig. 10, conductivity decreases in the order  $\text{K}^+ < \text{Na}^+ \leq \text{Li}^+$ , likewise water uptake trend found in literature, obtained in similar conditions [46]. Moreover, such results are in agreement with Saito et al.'s work. They suggested that water mobility inside the Nafion membrane is affected by the interaction between water molecules and cations present in the solution.  $\text{Li}^+$  and  $\text{Na}^+$  possess higher interaction with water molecules in the membranes compared to  $\text{H}^+$ , slowing down water molecules transport through the membrane itself [57]. Conductivity values differ from the ones present in the literature, which are usually in the order of magnitude of  $10 \text{ mS cm}^{-1}$  [58–60], due to the different experimental setups, but the general trend is maintained. Indeed, the experimental conductivity determination is



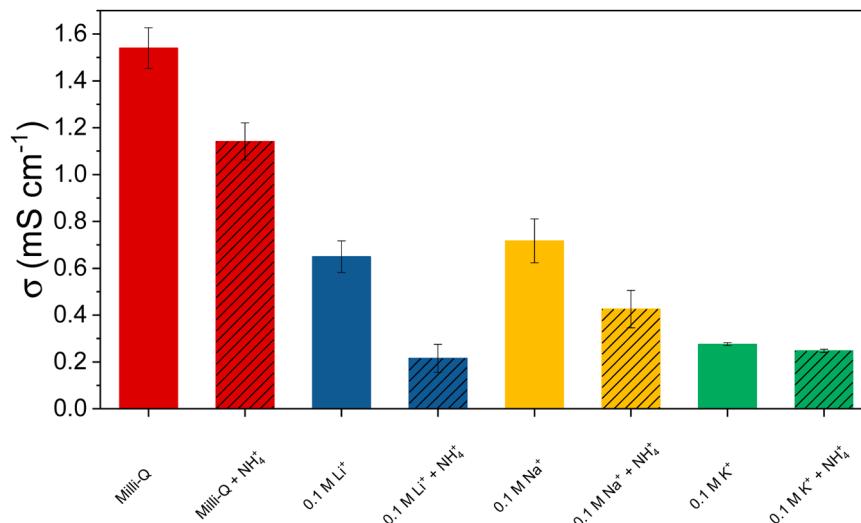
**Fig. 9.**  $NH_4^+$  produced through E-NRR in a flow-cell reactor, employing  $MoS_2$  as a catalyst in combination with  $K_2SO_4$  0.1 M as electrolyte and Celgard membrane as separator. (a) Test under  $N_2$ . Background refers to a sample taken after the OCV, LSV, and CV in Ar. Initial value refers to the quantity of  $NH_4^+$  found in the sample taken after CV an LSV measurements under  $N_2$ , right before the CP technique at  $-5\text{ mA cm}^{-2}$  for 2 h (b) Control test under Ar. Initial value refers to the sample taken after CV an LSV measurements in Ar, right before the CP at  $-5\text{ mA cm}^{-2}$  for 2 h (see Figure S3 for protocol details).

carried out in low relative humidity conditions, in which conductivity values can be less than  $10\text{ mS cm}^{-1}$ , as reported by other studies [61,62]. Moreover, a recent statistical study, that collects all the literature on the topic of proton conductivity of Nafion membranes, clarifies how parameters such as relative humidity and temperature can have a big impact on the final conductivity value [63].

When  $NH_4^+$  is added to the exchanged solution, conductivity is further affected, with a percentage of decrease in the order  $Li^+ < Na^+ < K^+$ . In the case of  $K^+$ , the conductivity decreases in the presence of  $NH_4^+$  is negligible, thus confirming that  $NH_4^+$  poorly interacts with membrane in  $K^+$  containing electrolyte. Such behaviour supports the data on absorption presented till now. Indeed,  $NH_4^+$  absorbed on the Nafion membrane follows the opposite order (*i.e.*,  $Li^+ > Na^+ > K^+$ ), meaning that  $NH_4^+$  has a negative effect on the conductivity with a magnitude strictly dependent on the quantity absorbed.

#### 4. Conclusions

Till now, Nafion 117 and Celgard membranes have been the most commonly used separators for E-NRR and E- $NO_3$ RR experiments, even if they both present some disadvantages. In particular, N-117 is known to absorb  $NH_4^+$ , becoming a source of contamination. In this paper, we analysed how different cations (*i.e.*,  $Li^+$ ,  $Na^+$ , and  $K^+$ ) and pH value in the electrolyte influence the  $NH_4^+$  ions absorbed by N-117 membrane. It seems that 86% of  $NH_4^+$  is absorbed when no cations are present in the solution, while when a cation is added, the absorption values lower to 40, 14, and 10% for  $Li_2SO_4$ ,  $Na_2SO_4$ , and  $K_2SO_4$ , respectively, while the pH has no significant effect. Such a trend reflects the increase in the ionic radius as well as the decrease in the cation hydration number, while the trend in  $NH_4^+$  release is the opposite. Contrary to what is stated in other literature works, none of the cations tested can completely exchange the  $NH_4^+$  trapped in N-117. Instead,  $NH_4^+$  is not absorbed by Celgard



**Fig. 10.** Conductivity measurements on N-117 membranes exchanged in different electrolytes. Conductivity values are extrapolated from impedance value at high frequency, obtained through EIS. Comparison of all electrolytes in the absence of  $\text{NH}_4^+$  (solid fill) and with  $2 \text{ mg L}^{-1}$  of  $\text{NH}_4^+$  (pattern fill). The data referred to as Milli-Q are N-117 membranes activated in  $\text{H}_2\text{SO}_4$  1 M and then exchanged in Milli-Q, which explains the high conductivity.

membrane in any condition.

Then, tests on  $\text{NH}_4^+$  motion inside an electrochemical flow-cell reactor revealed that, for N-117 membrane,  $\text{NH}_4^+$  motion to the anode side is more prominent in  $\text{Li}_2\text{SO}_4$ , when low potential is applied; conversely, when a potential is applied, the effect of the electric field prevails on the interaction between  $\text{NH}_4^+$  and N-117 membrane. On the other side, using Celgard membrane,  $\text{NH}_4^+$  motion when no potential is applied is independent from the cation nature, while the application of a potential induces an electric field that moves the equilibrium as a function of the cation electronegativity ( $\text{Li}^+ > \text{Na}^+ > \text{K}^+$ ).

These experiments also demonstrated that N-117 is a big source of contamination, as the final amount of  $\text{NH}_4^+$  found in the system was always higher than the initial one. Such contamination was not noticed with Celgard. In addition, E-NRR and E- $\text{NO}_3\text{RR}$  final tests confirmed that N-117 can lead to false positive and fallacious calculations on the final FE and productivity. Such tests seem to lead to the conclusion that Celgard membranes are the best option for E-NRR and E- $\text{NO}_3\text{RR}$ , but such separators allow the crossing of multiple species from the catholyte and the anolyte, making it impossible to maintain different environments in the two compartments of the cell.

It is the authors' opinion that much more effort must be focused on the synthesis of a cationic-exchange membrane with specific characteristics for E-NRR and E- $\text{NO}_3\text{RR}$  applications.

#### Declaration of competing interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

#### Acknowledgements

This project has received funding from the European Research Council (ERC) under the European Union's Horizon 2020 research and innovation program (grant agreement No. 948769, project title: SuN2-rise). This study was also carried out within the «HYDREAM» project – funded by European Union - Next Generation EU – within the PRIN 2022 program (D.D. 104 - 02/02/2022 Ministero dell'Università e della Ricerca). Finally, the project was also supported by the European Union's Horizon 2020 research and innovation programme under the Marie Skłodowska-Curie grant agreement No 101107906. This manuscript reflects only the authors' views and opinions, and the Ministry

cannot be considered responsible for them.

#### Supplementary materials

Supplementary material associated with this article can be found, in the online version, at [doi:10.1016/j.electacta.2024.144415](https://doi.org/10.1016/j.electacta.2024.144415).

#### References

- [1] D.C. Massuela, S. Munz, J. Hartung, P.M. Nkebiwe, S. Graeff-Hönninger, Cannabis hunger games: nutrient stress induction in flowering stage – impact of organic and mineral fertilizer levels on biomass, cannabidiol (CBD) yield and nutrient use efficiency, *Front. Plant Sci.* 14 (2023), <https://doi.org/10.3389/fpls.2023.123232>.
- [2] V. Smil, Detonator of the population explosion, *Nature* 400 (415) (1999) 415, <https://doi.org/10.1038/22672>.
- [3] C. Smith, A.K. Hill, L. Torrente-Murciano, Current and future role of Haber–Bosch ammonia in a carbon-free energy landscape, *Energy Environ. Sci.* 13 (2020) 331–344, <https://doi.org/10.1039/C9EE02873K>.
- [4] L. Zhai, S. Liu, Z. Xiang, Ammonia as a carbon-free hydrogen carrier for fuel cells: a perspective, *Indus. Chem. Mater.* 1 (2023) 332–342, <https://doi.org/10.1039/D3IM00036B>.
- [5] S.Joseph Sekhar, M.S. Samuel, G. Glivin, T. Le, T. Mathimani, Production and utilization of green ammonia for decarbonizing the energy sector with a discrete focus on Sustainable Development Goals and environmental impact and technical hurdles, *Fuel* 360 (2024) 130626, <https://doi.org/10.1016/j.fuel.2023.130626>.
- [6] A. Valera-Medina, F. Amer-Hatem, A.K. Azad, I.C. Dedoussi, M. de Joannon, R. X. Fernandes, P. Glarborg, H. Hashemi, X. He, S. Mashruk, J. McGowan, C. Mounaim-Rousselle, A. Ortiz-Prado, A. Ortiz-Valera, I. Rossetti, B. Shu, M. Yehia, H. Xiao, M. Costa, Review on ammonia as a potential fuel: from synthesis to economics, *Energy Fuels* 35 (2021) 6964–7029, <https://doi.org/10.1021/acs.energyfuels.0c03685>.
- [7] N. Morlanés, S.P. Katikaneni, S.N. Paglieri, A. Harale, B. Solami, S.M. Sarathy, J. Gascon, A technological roadmap to the ammonia energy economy: current state and missing technologies, *Chem. Eng. J.* 408 (2021) 127310, <https://doi.org/10.1016/j.cej.2020.127310>.
- [8] E. Spatolisano, L.A. Pellegrini, A.R. de Angelis, S. Cattaneo, E. Roccaro, Ammonia as a carbon-free energy carrier: NH<sub>3</sub> cracking to H<sub>2</sub>, *Ind. Eng. Chem. Res.* 62 (2023) 10813–10827, <https://doi.org/10.1021/acs.iecr.3c01419>.
- [9] D.R. MacFarlane, P.V. Cherepanov, J. Choi, B.H.R. Suryanto, R.Y. Hodgetts, J. M. Bakker, F.M. Ferrero Vallana, A.N. Simonov, A roadmap to the ammonia economy, *Joule* 4 (2020) 1186–1205, <https://doi.org/10.1016/J.JOULE.2020.04.004>.
- [10] B.H.R. Suryanto, H.L. Du, D. Wang, J. Chen, A.N. Simonov, D.R. MacFarlane, Challenges and prospects in the catalysis of electroreduction of nitrogen to ammonia, *Nat. Catal.* 2 (2019) 290–296, <https://doi.org/10.1038/s41929-019-0252-4>.
- [11] M.A. Shipman, M.D. Symes, Recent progress towards the electrosynthesis of ammonia from sustainable resources, *Catal. Today* 286 (2017) 57–68, <https://doi.org/10.1016/j.cattod.2016.05.008>.
- [12] Z.Y.Y. Wu, M. Karanam, X. Yong, Q. Huang, D.A. Cullen, P. Zhu, C. Xia, Q. Xiao, M. Shakouri, F.Y.Y. Chen, J.Y. (Timothy) Kim, Y. Xia, K. Heck, Y. Hu, M.S. Wong, Q. Li, I. Gates, S. Siahrostami, H. Wang, Electrochemical ammonia synthesis via

- nitrate reduction on Fe single atom catalyst, *Nat. Commun.* 12 (2021) 1–10, <https://doi.org/10.1038/s41467-021-23115-x>.
- [13] T. Mou, J. Long, T. Frauenheim, J. Xiao, Advances in electrochemical ammonia synthesis beyond the use of nitrogen gas as a source, *Chempluschem.* 86 (2021) 1211–1224, <https://doi.org/10.1002/cplu.202100356>.
- [14] J.G. Chen, R.M. Crooks, L.C. Seefeldt, K.L. Bren, R.M. Bullock, M.Y. Daresbourg, P.L. Holland, B. Hoffman, M.J. Janik, A.K. Jones, M.G. Kanatzidis, P. King, K. M. Lancaster, S.V. Lymar, P. Pfleiderer, W.F. Schneider, R.R. Schrock, Beyond fossil fuel–driven nitrogen transformations, *Science* 1979 (2018) 360, <https://doi.org/10.1126/science.aar6611>.
- [15] E. Abascal, L. Gómez-Coma, I. Ortiz, A. Ortiz, Global diagnosis of nitrate pollution in groundwater and review of removal technologies, *Sci. Total Environ.* 810 (2022) 152233, <https://doi.org/10.1016/j.scitotenv.2021.152233>.
- [16] C.J. Werth, C. Yan, J.P. Troutman, Factors impeding replacement of ion exchange with (Electro)catalytic treatment for nitrate removal from drinking water, *ACS ES. T. Eng.* 1 (2021) 6–20, <https://doi.org/10.1021/acs.estengg.0c00076>.
- [17] G. Qing, R. Ghazafar, S.T. Jackowski, F. Habibzadeh, M.M. Ashtiani, C.P. Chen, M. R. Smith, T.W. Hamann, Recent advances and challenges of electrocatalytic N 2 reduction to ammonia, *Chem. Rev.* 120 (2020) 5437–5516, <https://doi.org/10.1021/acs.chemrev.9b00659>.
- [18] Z. Chen, T. Wang, B. Liu, D. Cheng, C. Hu, G. Zhang, W. Zhu, H. Wang, Z.J. Zhao, J. Gong, Grain-boundary-rich copper for efficient solar-driven electrochemical CO 2 reduction to ethylene and ethanol, *J. Am. Chem. Soc.* 142 (2020) 6878–6883, <https://doi.org/10.1021/jacs.0c00971>.
- [19] Y. Mi, Y. Qiu, Y. Liu, X. Peng, M. Hu, S. Zhao, H. Cao, L. Zhuo, H. Li, J. Ren, X. Liu, J. Luo, Cobalt–iron oxide nanosheets for high-efficiency solar-driven CO 2 –H 2 O coupling electrocatalytic reactions, *Adv. Funct. Mater.* 30 (2020), <https://doi.org/10.1002/adfm.202003438>.
- [20] X.V. Medvedeva, J.J. Medvedev, S.W. Tatarchuk, R.M. Choueiri, A. Klinkova, Sustainable at both ends: electrochemical CO 2 utilization paired with electrochemical treatment of nitrogenous waste, *Green Chem.* 22 (2020) 4456–4462, <https://doi.org/10.1039/DGCO1754J>.
- [21] S. Garcia-Segura, M. Lanzarini-Lopes, K. Hristovski, P. Westerhoff, Electrocatalytic reduction of nitrate: fundamentals to full-scale water treatment applications, *Appl. Catal. B* 236 (2018) 546–568, <https://doi.org/10.1016/J.APCATB.2018.05.041>.
- [22] Y. Ren, C. Yu, X. Tan, X. Han, H. Huang, H. Huang, J. Qiu, Is it appropriate to use the Nafion membrane in electrocatalytic N 2 reduction? *Small. Methods* 3 (2019) <https://doi.org/10.1002/smtd.201900474>.
- [23] P. Choi, N.H. Jalani, R. Datta, Thermodynamics and proton transport in Nafion, *J. Electrochem. Soc.* 152 (E123) (2005), <https://doi.org/10.1149/1.1859814>.
- [24] C. Tang, S.Z. Qiao, How to explore ambient electrocatalytic nitrogen reduction reliably and insightfully, *Chem. Soc. Rev.* 48 (2019) 3166–3180, <https://doi.org/10.1039/C9CS00280D>.
- [25] S.Z. Andersen, V. Colic, S. Yang, J.A. Schwalbe, A.C. Nielerd, J.M. McEnaney, K. Enemark-Rasmussen, J.G. Baker, A.R. Singh, B.A. Rohr, M.J. Statt, S.J. Blair, S. Mezzavilla, J. Kibsgaard, P.C.K. Vesborg, M. Cargnello, S.F. Bent, T.F. Jaramillo, I.E.L. Stephens, J.K. Nørskov, I. Chorkendorff, A rigorous electrochemical ammonia synthesis protocol with quantitative isotope measurements, *Nature* 570 (2019) 504–508, <https://doi.org/10.1038/s41586-019-1260-x>.
- [26] S.M.M. Hizam, A.M. Al-Dhahebi, M.S. Mohamed Saheed, Recent advances in graphene-based nanocomposites for ammonia detection, *Polymers. (Basel)* 14 (2022) 5125, <https://doi.org/10.3390/polym14235125>.
- [27] X. Cai, H. Iriawan, F. Yang, L. Luo, S. Shen, Y. Shao-Horn, J. Zhang, Interaction of ammonia with nafion and electrolyte in electrocatalytic nitrogen reduction study, *J. Phys. Chem. Lett.* 12 (2021) 6861–6866, <https://doi.org/10.1021/acs.jpclett.1c01714>.
- [28] P. Arora, Z. (John) Zhang, Battery Separators, *Chem. Rev.* 104 (2004) 4419–4462, <https://doi.org/10.1021/cr020738u>.
- [29] W. Xiao, Q. Yang, S. Zhu, Comparing ion transport in ionic liquids and polymerized ionic liquids, *Sci. Rep.* 10 (2020) 7825, <https://doi.org/10.1038/s41598-020-64689-8>.
- [30] P.M. Albert, B. Commarieau, J.C. Daigle, J.P. Claverie, K. Zaghib, A 3D network based on poly( $\epsilon$ -caprolactone) macromonomers as polymer electrolyte for solid state lithium metal batteries, *J. Electrochem. Soc.* 167 (2020) 080527, <https://doi.org/10.1149/1945-7111/ab8de0>.
- [31] C. Zhou, S. Bag, B. Lv, V. Thangadurai, Understanding the role of solvents on the morphological structure and li-ion conductivity of poly(vinylidene fluoride)-based polymer electrolytes, *J. Electrochem. Soc.* 167 (2020) 070552, <https://doi.org/10.1149/1945-7111/ab7c3a>.
- [32] B. Izelaar, D. Ripepi, S. Asperti, A.I. Dugulan, R.W.A. Hendrikx, A.J. Böttger, F. M. Mulder, R. Kortlever, Revisiting the electrochemical nitrogen reduction on molybdenum and iron carbides: promising catalysts or false positives? *ACS. Catal.* 13 (2023) 1649–1661, <https://doi.org/10.1021/acscatal.2c04491>.
- [33] J. Norskov, J. Chen, R. Miranda, T. Fitzsimmons, R. Stack, Sustainable ammonia synthesis – exploring the scientific challenges associated with discovering alternative, sustainable processes for ammonia production, 2016. [10.2172/1283146](https://doi.org/10.2172/1283146).
- [34] E. Murphy, Y. Liu, I. Matanovic, M. Rüschter, Y. Huang, A. Ly, S. Guo, W. Zang, X. Yan, A. Martini, J. Timoshenko, B.R. Cuanya, I.V. Zenyuk, X. Pan, E.D. Spoerke, P. Atanassov, Elucidating electrochemical nitrate and nitrite reduction over atomically-dispersed transition metal sites, *Nat. Commun.* 14 (2023) 4554, <https://doi.org/10.1038/s41467-023-40174-4>.
- [35] D. Liu, L. Qiao, Y. Chen, P. Zhou, J. Feng, C.C. Leong, K.W. Ng, S. Peng, S. Wang, W.F. Ip, H. Pan, Electrocatalytic reduction of nitrate to ammonia on low-cost manganese-incorporated Co3O4 nanotubes, *Appl. Catal. B* 324 (2023) 122293, <https://doi.org/10.1016/j.apcatb.2022.122293>.
- [36] M. Wang, S. Liu, H. Ji, T. Yang, T. Qian, C. Yan, Salting-out effect promoting highly efficient ambient ammonia synthesis, *Nat. Commun.* 12 (2021) 3198, <https://doi.org/10.1038/s41467-021-23360-0>.
- [37] G. Zhou, T. Li, R. Huang, P. Wang, B. Hu, H. Li, L. Liu, Y. Sun, Recharged Catalyst with Memristive Nitrogen Reduction Activity through Learning Networks of Spiking Neurons, *J. Am. Chem. Soc.* 143 (2021) 5378–5385, <https://doi.org/10.1021/jacs.0c12458>.
- [38] Y. Liu, X. Zhang, Z. Chen, X. Zhang, P. Tsikarais, P.K. Shen, Electrocatalytic reduction of nitrogen on FeAg/Si for ammonia synthesis: a simple strategy for continuous regulation of faradaic efficiency by controlling H<sup>+</sup> ions transfer rate, *Appl. Catal. B* 283 (2021) 119606, <https://doi.org/10.1016/j.apcatb.2020.119606>.
- [39] L.M. Wilder, K. Wyatt, C.A. Skangos, W.E. Klein, M.R. Parimuha, J.L. Katsirubas, J. L. Young, E.M. Miller, Membranes matter: preventing ammonia crossover during electrochemical ammonia synthesis, *ACS. Appl. Energy Mater.* 7 (2024) 536–545, <https://doi.org/10.1021/acs.amec.3c02461>.
- [40] W. Bi, E. Gyenge, D.P. Wilkinson, Crossover, volatilization, and adsorption of ammonium ions in a proton-exchange membrane electrolyzer in relation to electrochemical ammonia production, *Chem. Eng. J.* 478 (2023) 147359, <https://doi.org/10.1016/j.cej.2023.147359>.
- [41] F. Hanifpour, A. Sveinbjörnsson, C.P. Canales, E. Skúlason, H.D. Flosadóttir, Preparation of nafion membranes for reproducible ammonia quantification in nitrogen reduction reaction experiments, *Angew. Chem.* 132 (2020) 23138–23142, <https://doi.org/10.1002/ange.202007998>.
- [42] M. Leonardi, G. Tranchida, R. Corso, R.G. Milazzo, S.A. Lombardo, S.M.S. Privitera, Role of the Membrane Transport Mechanism in Electrochemical Nitrogen Reduction Experiments, *Membranes. (Basel)* 12 (2022) 969, <https://doi.org/10.3390/membranes12100969>.
- [43] J.J. Giner-Sanz, G. Leverick, V. Pérez-Herranz, Y. Shao-Horn, Optimization of the salicylate method for ammonia quantification from nitrogen electroreduction, *J. Electroanal. Chemistry* 896 (2021) 115250, <https://doi.org/10.1016/J.JELECHEM.2021.115250>.
- [44] L. Zhang, X. Ji, X. Ren, Y. Ma, X. Shi, Z. Tian, A.M. Asiri, L. Chen, B. Tang, X. Sun, Electrochemical ammonia synthesis via nitrogen reduction reaction on a MoS 2 catalyst: theoretical and experimental studies, *Adv. Mater.* 30 (2018) 1800191, <https://doi.org/10.1002/adma.201800191>.
- [45] J. Ingle, U.D. Patel, Electrochemical reduction of nitrate in the presence of silver-coated polyvinyl alcohol beads as a spatially suspended catalyst, *J. Water. Process. Eng.* 49 (2022) 103082, <https://doi.org/10.1016/j.jwpe.2022.103082>.
- [46] M.A.A. Izquierdo-Gil, V.M.M. Barragán, J.P.G.P.G. Villaluenga, M.P.P. Godino, Water uptake and salt transport through Nafion cation-exchange membranes with different thicknesses, *Chem. Eng. Sci.* 72 (2012) 1–9, <https://doi.org/10.1016/j.ces.2011.12.040>.
- [47] M. Schalenbach, L. Keller, B. Janotta, A. Bauer, H. Tempel, H. Kungl, M. Bonnet, R. A. Eichel, The Effect of Ion Exchange Poisoning on the Ion Transport and Conduction in Polymer Electrolyte Membranes (PEMs) for Water Electrolysis, *J. Electrochem. Soc.* 169 (2022) 094510, <https://doi.org/10.1149/1945-7111/ac9087>.
- [48] A. Paspureddi, Z. Zhang, V. Ganesan, M.M. Sharma, L.E. Katz, Mechanism of monovalent and divalent ion mobility in Nafion membrane: an atomistic simulation study, *J. Chem. Phys.* 158 (2023), <https://doi.org/10.1063/5.0145205>.
- [49] M.J. Kelly, G. Faflik, J.O. Besenhard, H. Kronberger, G.E. Nauer, Contaminant absorption and conductivity in polymer electrolyte membranes, *J. Power. Sources.* 145 (2005) 249–252, <https://doi.org/10.1016/j.jpowsour.2005.01.064>.
- [50] D. Johnson, A. Djire, Effect of pH on the Electrochemical Behavior and Nitrogen Reduction Reaction Activity of Ti 2 N Nitride MXene, *Adv. Mater. Interfaces.* 10 (2023), <https://doi.org/10.1002/admi.202202147>.
- [51] X. Zhao, G. Hu, G. Chen, H. Zhang, S. Zhang, H. Wang, Comprehensive Understanding of the Thriving Ambient Electrochemical Nitrogen Reduction Reaction, *Adv. Mater.* 33 (2021), <https://doi.org/10.1002/adma.202007650>.
- [52] T.F. Beltrame, F.M. Zoppas, J.Z. Ferreira, F.A. Marchesini, A.M. Bernardes, Nitrate reduction by electrochemical processes using copper electrode: evaluating operational parameters aiming low nitrite formation, *Water Sci. Technol.* 84 (2021) 200–215, <https://doi.org/10.2166/wst.2021.215>.
- [53] N.C. Kani, A. Prajapati, B.A. Collins, J.D. Goodpaster, M.R. Singh, Competing Effects of pH, Cation Identity, H 2 O Saturation, and N 2 Concentration on the Activity and Selectivity of Electrochemical Reduction of N 2 to NH 3 on Electrodeposited Cu at Ambient Conditions, *ACS. Catal.* 10 (2020) 14592–14603, <https://doi.org/10.1021/acscatal.0c04864>.
- [54] J. Peng, M. Tian, N.M. Cantillo, T. Zawodzinski, The ion and water transport properties of K<sup>+</sup> and Na<sup>+</sup> form perfluorosulfonic acid polymer, *Electrochim. Acta* 282 (2018) 544–554, <https://doi.org/10.1016/j.electacta.2018.06.035>.
- [55] S. Cukierman, Et tu, Grothuss! and other unfinished stories, *Biochim. Biophys. Acta (BBA) Bioenergetics* 1757 (2006) 876–885, <https://doi.org/10.1016/j.bbabi.2005.12.001>.
- [56] T. Okada, N. Arimura, H. Satou, M. Yuasa, T. Kikuchi, Membrane transport characteristics of binary cation systems with Li<sup>+</sup> and alkali metal cations in perfluorosulfonated ionomer, *Electrochim. Acta* 50 (2005) 3569–3575, <https://doi.org/10.1016/j.electacta.2005.01.002>.
- [57] M. Saito, N. Arimura, K. Hayamizu, T. Okada, Mechanisms of ion and water transport in perfluorosulfonated ionomer membranes for fuel cells, *J. Phys. Chem. B* 108 (2004) 16064–16070, <https://doi.org/10.1021/jp0482565>.
- [58] N.P. Berezina, N.A. Kononenko, O.A. Dyomina, N.P. Gnusin, Characterization of ion-exchange membrane materials: properties vs structure, *Adv. Colloid. Interface Sci.* 139 (2008) 3–28, <https://doi.org/10.1016/j.cis.2008.01.002>.

- [59] N.P. Gnusin, N.P. Berezina, N.A. Kononenko, O.A. Dyomina, Transport structural parameters to characterize ion exchange membranes, *J. Memb. Sci.* 243 (2004) 301–310, <https://doi.org/10.1016/J.MEMSCI.2004.06.033>.
- [60] I.A. Stenina, P. Sistat, A.I. Rebrov, G. Pourcelly, A.B. Yaroslavtsev, Ion mobility in Nafion-117 membranes, *Desalination*. 170 (2004) 49–57, <https://doi.org/10.1016/J.DESAL.2004.02.092>.
- [61] Y. Sone, P. Ekdunge, D. Simonsson, Proton Conductivity of Nafion 117 as Measured by a Four-Electrode AC Impedance Method, *J. Electrochem. Soc.* 143 (1996) 1254–1259, <https://doi.org/10.1149/1.1836625/XML>.
- [62] A.V. Anantaraman, C.L. Gardner, Studies on ion-exchange membranes. Part 1. Effect of humidity on the conductivity of Nafion®, *J. Electroanal. Chem.* 414 (1996) 115–120, [https://doi.org/10.1016/0022-0728\(96\)04690-6](https://doi.org/10.1016/0022-0728(96)04690-6).
- [63] L. Liu, W. Chen, Y. Li, An overview of the proton conductivity of Nafion membranes through a statistical analysis, *J. Memb. Sci.* 504 (2016) 1–9, <https://doi.org/10.1016/J.MEMSCI.2015.12.065>.