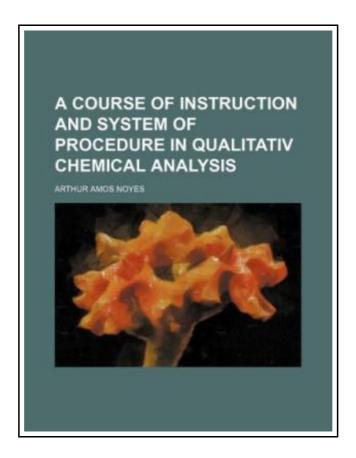
# A Course of Instruction and System of Procedure in Qualitativ Chemical Analysis



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## A COURSE OF INSTRUCTION AND SYSTEM OF PROCEDURE IN QUALITATIV CHEMICAL ANALYSIS



Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English. Brand New Book \*\*\*\*\* Print on Demand \*\*\*\*\*. This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1912 Excerpt: .(potassium cadmicyanide), since its complex anion has a much smaller ionization tendency than the cation Cd(NH3)4++; thus in a normal KCN solution the ratio of the concentration of the complex anion to that of the simple Cd++ ion is about 1017. Yet this complex anion is sufficiently dissociated into Cd++ to cause the solution to become supersaturated with CdS when an alkaline sulfide is added. 4. A very small black precipitate (which may be due to HgS or PbS) may sometimes be produced in the final test for cadmium with H2S; but, provided the analysis has been properly conducted, not in sufficient quantity to prevent the yellow color of 1 or 2 mg. of CdS from being seen. In case a black precipitate is produced, and (hus prevents a positive conclusion as to the presence or absence of a little cadmium, the precipitate may be treated, in order to eliminate the black sulfide, as follows: Boil the precipitate gently for about 5 minutes in a covered casserole with about 15 cc. of a mixture of one volume of 6-normal H2SO4 with four volumes of water, filter, cool the filtrate, add to it three times its volume of water, and pass H2S into it for 5-10 minutes. A yellow precipitate of CdS should then be obtained, if cadmium is present. REPRECIFITATION OF THE TIN-GROUP. Procedure 41.--Reprecipitation of the Tin-Group.--Dilute in a small flask the first portion of the ammonium sulfide solution (P. 22) with about 20 cc....

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