


Elastic softening and defect-mediated diffusion in superionic Li_2O revealed by molecular dynamicsZhifeng Wu¹, Yiwei You¹, Fangfang Zhang¹, Tie-Yu Lü¹, Xinrui Cao¹, Yang Sun¹, Zi-Zhong Zhu¹, and Shunqing Wu^{1*}*Department of Physics, OSED, Key Laboratory of Low Dimensional Condensed Matter Physics
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The superionic transition in lithium oxide (Li_2O) presents significant challenges in establishing the structure-property relationship between its microscopic dynamical behavior to macroscopic physical properties. Additionally, the fundamental mechanism driving this transition remains contentious due to competing theoretical interpretations. Using machine learning potentials that maintain *ab initio*-level precision while enabling extended in both temporal and spatial scales, we systematically investigate lithium-ion diffusion dynamics spanning from crystalline to superionic phases. Our results give a clear physical pattern of Li_2O Arrhenius plot throughout three distinct regions. The elastic softening serves as a signal of accelerating of Li-ions movement. Superionic transition is manifested in the melting of the Li sublattice. Dynamical analysis indicates that the superionic transition is driven by the formation of Frenkel pairs. Octahedral interstitial Li-ions, increasing sharply near the superionic state, act as primary diffusion carriers. The tetrahedral Li-ions exhibit much lower individual mobility and play a secondary role in diffusion under superionic state. This highlights the unique defect-mediated nature of diffusion in superionic state.

DOI: [10.1103/4z6n-zfdr](https://doi.org/10.1103/4z6n-zfdr)**I. INTRODUCTION**

The deuterium(D)-tritium(T) nuclear fusion reaction stands as a pivotal candidate for addressing global energy sustainability, with tritium breeding blanket materials serving as critical components for fuel regeneration in fusion reactors [1,2]. Among lithium-containing ceramics—including Li_2O , Li_2TiO_3 , LiAlO_2 , and Li_2ZrO_3 [1,2]— Li_2O emerges as a prime candidate due to its exceptional combination of high melting point and high ionic conductivity at specific temperature, and low LiOH solubility [3]. Crucially, the tritium release kinetics in Li_2O -dominated systems are governed by lithium vacancy (V_{Li}) diffusion barriers [3], making atomic-scale understanding of Li^+ transport mechanisms imperative for both material design and reactor operation.

Similar to various fluorite or antifluorite structured compounds, such as UO_2 [4,5], Li_2S [6], and CaF_2 [7], Li_2O undergoes a superionic transition at elevated temperature. There exists two types of superionic transition [8]. Type I superionic transition occurs with a sudden change of diffusion coefficient (as seen in CuI [9], $\alpha\text{-KAg}_3\text{Se}_2$ [10]) accompanied by a distinct phase transition, during which some types of ions rearrange while others exhibit fast ionic diffusion. Type II superionic transition occurs without any rearrangement, and is associated with a second-order phase transition, as found in Li_3N [11], Li_3YCl_6 [12], and H_2O [13]. In the superionic state, materials exhibit properties intermediate between liquid and solid, such as liquidlike ionic diffusion behavior in a stable solid framework. These characteristics make superionic materials highly promising for applications in solid-state electrolyte [14–16] and solid oxide fuel cells [17,18].

The distinctive properties of Li_2O have been well investigated through both experiment and computational approaches. Notably, Li_2O has been reported to undergo a superionic transition at temperatures ranging from ~ 1200 K (as reported in inelastic neutron experiments [19,20]) to $1400\text{--}1600$ K (as predicted by molecular dynamics (MD) simulations [21]), and with *ab initio* molecular dynamics (AIMD) simulation suggesting a transition threshold above 1300 K [22]. Nuclear magnetic resonance (NMR) has been employed to probe Li_2O self-diffusion at specific temperature [2]. Concurrently, lattice dynamics study revealed rapid Li-ion diffusion at 1000 K [23]. MD simulations have further characterized the role of defects in Li_2O by quantifying their influence on the diffusion coefficient [24]. More recently, a cluster expansion method combined with kinetic Monte Carlo has extended Li_2O Arrhenius plot to room temperature [25].

Despite extensive research on Li_2O , notable gaps remain. The structure-property relationship between its microscopic dynamic behavior and macroscopic physical properties, such as ionic conductivity and elastic modulus, remains unclear. Additionally, the microscopic mechanism underlying the superionic transition is poorly understood and continues to be a topic of debate, hindering a deeper understanding of this phenomenon. In the meanwhile, the existing methodologies exhibit significant limitations. Density functional theory (DFT) method only supports small models, leading to a finite-size effect [26–30]. Conventional MD simulation, often relying on empirical potentials, are difficult to accurately capture the dynamical behavior across wide temperature ranges. To address these challenges, machine learning potentials (MLPs) have emerged as a transformative solution, offering DFT-level accuracy while enabling access to larger systems and longer simulation timescales [31–37].

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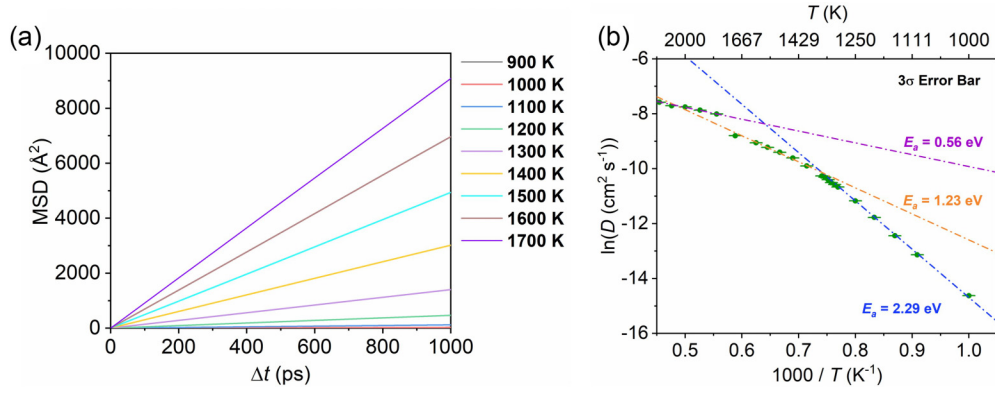


FIG. 1. Li_2O diffusion behavior. (a) The calculated time-averaged mean square displacement of Li-ions in Li_2O . (b) Li-ions diffusion coefficients. The color lines exhibit exponential fits at different regions. Error bars are within the symbol size.

In this work, we combine deep potential [36,38] with MD simulations to comprehensively investigate the dynamical characteristics of Li_2O , with a focus on its superionic transition. This approach allows for the precise quantification of atomic-level behavior across a wide temperature range. Most notably, our work highlights the distinct dynamical regimes across a wide temperature range, the understanding of the structure-property relationship between dynamical characteristics and macroscopic physical properties, and uncovers the key mechanisms driving fast-ion diffusion in the superionic state. These insights provide critical guidelines for designing Li_2O -based materials for energy storage and conversion applications.

II. METHODS

Our machine learning dataset was constructed by first-principles calculations. The first-principles calculations were conducted by Vienna *ab initio* simulation package (VASP) [39,40]. The calculations employed the projected augmented wave (PAW) method [41] with the Perdew-Burke-Ernzerhof (PBE) [42] generalized gradient approximation (GGA) for the exchange-correlation functional. A plane-wave cutoff energy of 520 eV was adopted, and Brillouin-zone sampling was conducted using a Γ -centered Monkhorst-Pack k -point grid with a spacing of 0.3 \AA^{-1} for structure optimization.

The deep potential model was trained using the DeePMD-kit package [38] and deployed for molecular dynamics simulations via used deep potential molecular dynamics (DPMD) framework [36]. The model incorporated a cutoff radius of 6.0 \AA with a smoothing width of 0.5 \AA to capture local atomic environments. The neural network architecture consisted of an embedding net with three hidden layers [10,20,40] and a fitting net of [120, 120, 120]. Training was performed for 6 000 000 steps using the Adam stochastic gradient descent method [43], leading to an exponential decrease in the learning rate starting from 0.001. The energy prefactor was scaled from 0.02 to 2, while the virial and force prefactors were annealed from 1000 to 1 during training. This model was initialized from our pretrained deep potential model [44,45], with training strategies and dataset selection detailed in Note 1 in the Supplemental Material [46].

Other calculation methods can be seen in Note 2 in the Supplemental Material [46]. The reliability of the deep potential function can be seen in Note 3 in the Supplemental Material [46].

III. RESULTS AND DISCUSSION

A. Ionic diffusion

The time-averaged mean square displacement (MSD) of Li ions is presented in Fig. 1(a). At 900 K (see Fig. S7 [46]), the MSD reaches only 4.0 \AA^2 after 1 ns of simulation, which is smaller than both the squared nearest-neighbor $\text{Li}^+ - \text{Li}^+$ distance ($\sim 5.4 \text{ \AA}^2$) and the Li^+ -to-nearest-octahedral-interstitial distance ($\sim 3.9 \text{ \AA}^2$), where subdiffusive behavior is observed. This confinement indicates that Li^+ vibrations dominate below $\sim 900 \text{ K}$, with few diffusion events (termed “MSD limitation” [47]). Consequently, calculating diffusion coefficients from the MSD slope is invalid in this regime. Notably, the MSD exhibits linearity at all temperatures except 900 K.

The temperature dependence of the diffusion coefficient is shown in the Arrhenius plot [Fig. 1(b)]. Linear fitting identifies three distinct regions in Li_2O : the normal state (activation energy $E_a = 2.29 \text{ eV}$), the fast ion region—superionic state ($E_a = 1.23 \text{ eV}$, onset at $\sim 1300 \text{ K}$),—and the liquid phase region ($E_a = 0.56 \text{ eV}$, transition at $\sim 1700 \text{ K}$). This behavior aligns with superionic materials like UO_2 [5,48], Li_3N [11], and Li_3YCl_6 [12]. The decrease in E_a at 1300 K signals the superionic transition, while the discontinuity at $\sim 1700 \text{ K}$ corresponds to melting of Li_2O . In liquid region, the reduced activation energy facilitates a faster Li-ions diffusion. Besides, it’s worth noting that the calculated ionic conductivity exceeds 1 S/cm above 1300 K (Fig. S8 [46]), further confirming the superionic transition in Li_2O , consistent with prior reports [8].

The Arrhenius behavior of lithium oxide has been widely studied. While Gupta [19] reported a detailed analysis of Li-ion diffusion, the assumption of constant activation energy across all temperatures remains debated. Oishi *et al.* [49] divided the Arrhenius plot into two regions: a low-temperature regime ($E_a = 23.4 \times 10^3 \text{ kcal/mol}$) and a high-temperature region ($E_a = 58.2 \times 10^3 \text{ kcal/mol}$) as shown in Fig. S9 [46]. However, this abrupt increase in activation energy contradicts

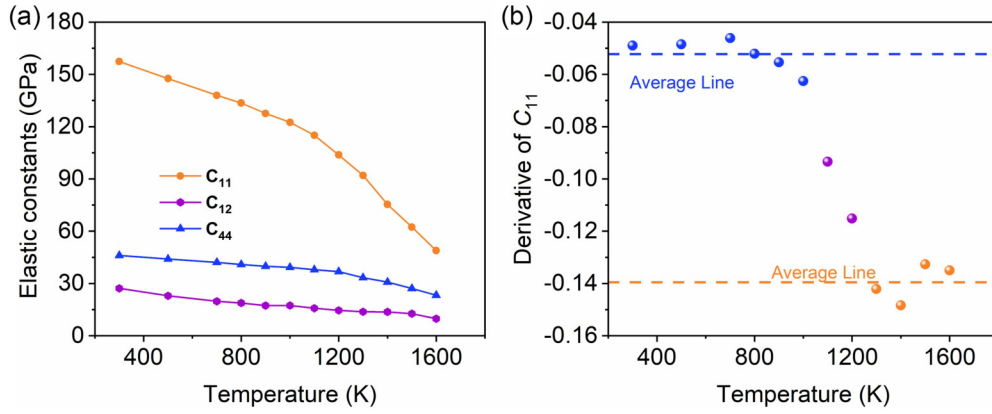


FIG. 2. Li_2O Elastic properties. (a) Elastic constants of Li_2O . (b) The temperature derivative of C_{11} in (a). The horizontal dashed lines represent the average derivative in the range of 300–1000 K and 1300–1600 K, respectively.

the typical decrease expected during superionic transitions. Lörger *et al.* [50] reported an Arrhenius profile distinct from ours, likely due to their use of cation doping—an extrinsic vacancy source—contrasted with our focus on thermally induced intrinsic defects. Our results align with [47] but exhibit higher precision, revealing that the Arrhenius slope magnitude is smaller in the superionic state and larger in the normal state, directly correlated with the activation energies of the respective states. Importantly, we confirm that lithium oxide followed Arrhenius kinetics across three distinct temperature intervals, with each interval characterized by unique activation energies. The activation energy extracted from the Arrhenius plot represents an effective average over all possible Li-ions migration events, rather than the barrier of a single path. Detailed migration barriers along specific crystallographic directions have been reported in previous studies [51,52].

B. Elastic constants and dynamics behavior

Elastic constants are key indicators of superionic transition. In Li_2O , high temperature C_{11} softening has been identified as a signature of superionic behavior [5,13,20,22,53,54]. Figure 2(a) shows our calculated elastic constants. While C_{12} and C_{44} decrease smoothly and linearly across all temperatures, C_{11} exhibits two distinct regions. Below 1000 K, C_{11} decreased linearly; between 1100–1200 K, pronounced softening occurs; above 1300 K, it resumes decreases linearly but with a shallower slope. The temperature derivative of C_{11} , as shown in Fig. 2(b), highlights these regions via dashed lines denoting average slopes for 300–1000 K and 1300–1600 K.

The O sublattice remains rigid up to 1700 K, as evidenced by the vibration-dominated O MSD and negligible diffusion at 1700 K (Fig. S6 [46]), isolating Li-ion dynamics as the origin of elastic softening. To correlate with Li mobility, Figs. 3(a)–3(c) show atomic trajectories: at 900 K, Li ions are mostly localized; by 1100 K, enhanced diffusion coincides with C_{11} softening; and at 1300 K, Li delocalization and sublattice melting confirm a type II superionic transition, with Li-ions diffusing through O octahedral sites. By 1800 K, the distribution of Li and O is shown in Fig. 3(d), where the entire system completes melting into the liquid phase.

The van Hove self-correlation function can be used to observe vibration and diffusion behavior. It is defined as follows:

$$G_s(\mathbf{r}, \Delta t) = \frac{1}{N} \left\langle \sum_i^N \delta(\mathbf{r} + \mathbf{r}_i(0) - \mathbf{r}_i(\Delta t)) \right\rangle. \quad (1)$$

N is the calculated atom number. Δt is a period of time, which should be selective containing a complete diffusion

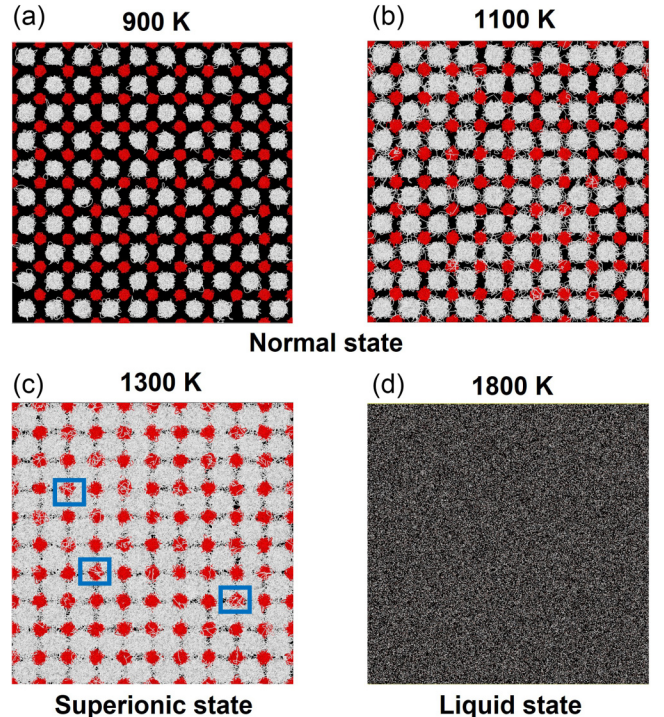


FIG. 3. Li_2O diffusion behaviors. (a)–(c) Trajectories of Li (white) and O (red) during 20 ps at 900 K, 1100 K, and 1300 K, respectively. The blue boxes in (c) highlight the typical phenomenon where Li occupied the octahedral interstitial positions during the diffusion process. (d) Distribution of Li (white) and O (red) at 1800 K. To better reveal the characteristics of the melted state, multiple frames were superimposed and the atoms were plotted with small sizes.

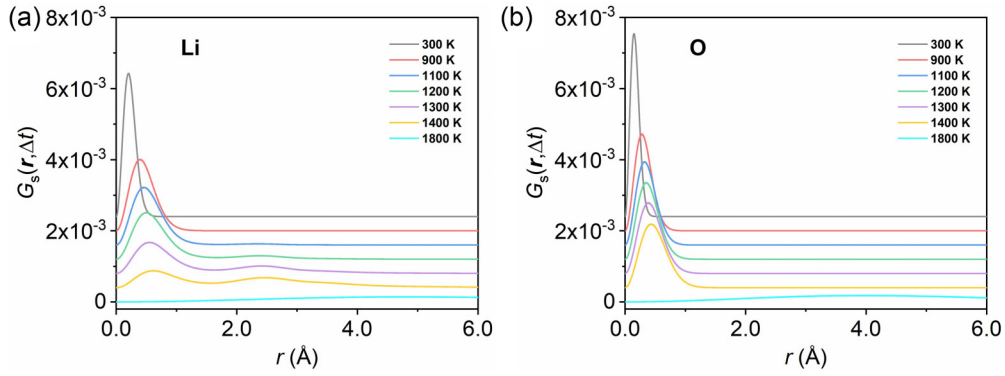


FIG. 4. The van Hove self-correlation function of Li (a) and O (b). Δt was set to be 2 ps. All curves were vertically shifted for clarity and ease of comparison.

or vibration event. $r_i(0)$ is atom's initial position. $r_i(\Delta t)$ is atom's position at $t = \Delta t$, which is shown in Fig. 4. Between 300–1200 K, both Li and O ions exhibit single vibration peaks, confirming ion localization and minimal diffusion. It is consistent with the normal state in Figs. 3(a) and 3(b). At 1300 K, O ions retain a single vibration peak, while Li ions display an additional broad peak at ~ 2.3 Å. This signals a transition in Li ion: Li sublattice melting and the onset of superionic behavior corresponding to Fig. 3(c). At 1800 K, complete melting [Fig. 3(d)] erases these features. Together, these results confirm the superionic transition occurs near 1300 K, with the elastic softening in Fig. 2 directly linked to enhanced Li-ion mobility.

Our results reproduce the C_{11} softening behavior reported by He [22] and Gupta [20], who observed similar trends at specific temperatures. However, the inflection in C_{11} here arises from the accelerated Li ion diffusion rather than a superionic transition. As evidenced by the constant activation energy in Fig. 1(b) and the absence of a board peak in Fig. 4(a), the softening regime (1100–1200 K) remains within the normal state. The superionic transition occurs at a higher temperature, distinct from the intermediate softening state.

C. Transition mechanism

The thermally induced intrinsic defect Frenkel pair is widely considered the primary mechanism for superionic transition in fluorite and antfluorite materials at elevated temperatures [21,24,25,55–57]. A Frenkel pair can be generated as follows:

$$A_A^\times = A_i^\bullet + V_A'$$

The defect reaction is expressed using Kröger-Vink notation, taking a monovalent ion as an example. Here, A represents an atom, i denotes an interstitial site, and V represents a vacancy. The combination of an interstitial site and a vacancy forms a Frenkel pair. Schematic of a Frenkel pair in Li_2O shows a Li-ion move from a tetrahedral site to an octahedral interstitial, creating a vacancy at its original position [Fig. 5(b)].

According to the definition of Frenkel defects, the population of octahedral interstitial Li ions (Li_{Oct}), serves as a direct measure of Frenkel pair concentration and a key indicator of defect-mediated ion transport. As shown in Fig. 6, Li_{Oct}

is nearly absent between 300–900 K, consistent with limited Li-ion diffusion. Above 1100 K, its population rises sharply, signaling increased Frenkel pair formation and the onset of structural disorder.

To further evaluate its role, we analyzed the residence time of Li ions at different sites (Table I, Fig. S10 [46]). Figure 5 shows the possible diffusion paths, Figs. 5(a) and 5(b) show diffusion via Li_{Oct} and tetrahedral Li ions (Li_{Tet}), while Figs. 5(c) and 5(d) show diffusion only via Li_{Tet} . Even at low temperatures, over 85% of Li_{Oct} diffuse within 1 ps, and increase to nearly 100% above 1200 K. Below 1300 K, Li_{Tet} present less than 30% diffusion events within 1 ps. Above 1300 K, its contribution to superionic state becomes considerable ($\sim 37\%$). It's worth noting that across all temperatures, Li_{Oct} displayed markedly higher mobility than Li_{Tet} . These findings underscore a dual mechanism rooted in Frenkel defect dynamics. Li_{Oct} , which shows high dynamical activity, confirms a primary role in fast ion transport and its contribution to the superionic transition. Li_{Tet} , which shows less mobility than Li_{Oct} , also contributes to the superionic transition above 1300 K. Our results differ from Mulliner *et al.* [55], who reported increasing Li_{Oct} residence times with temperature. Instead, we find Li_{Oct} lifetimes remain short and become even shorter at higher temperatures, indicating they act as highly mobile carriers rather than trapped defects. Compared with the results reported by Mulliner *et al.*, the present study adopts a different computational approach. An empirical potential was employed by Mulliner, whose accuracy in capturing defect and diffusion behavior may be limited. For instance, their simulations indicated an increase in the residence time of intersti-

TABLE I. Percentage of Li ions with residence times shorter than 1 ps at octahedral and tetrahedral sites.

| Temperature (K) | Li_{Oct} | Li_{Tet} |
|-----------------|--------------------------|--------------------------|
| 900 | 89.4 | 7.5 |
| 1000 | 94.9 | 11.2 |
| 1100 | 97.6 | 12.6 |
| 1200 | 99.3 | 21.1 |
| 1300 | 99.9 | 37.1 |
| 1350 | 99.9 | 45.5 |
| 1400 | 100 | 56.6 |
| 1500 | 100 | 73.4 |

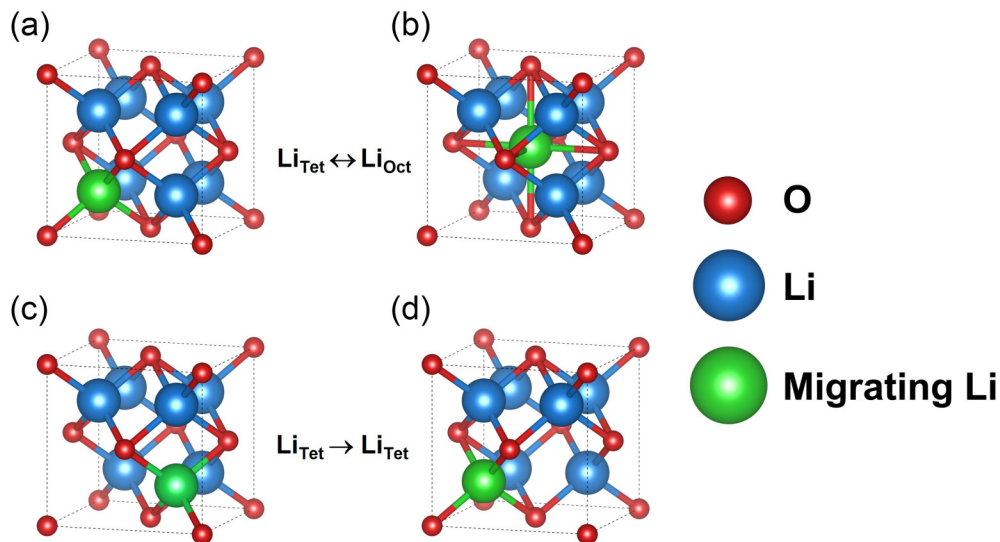


FIG. 5. Schematic illustration of Li₂O and migration paths. (a) perfect crystal. (b) Crystal containing a thermally generated Frenkel pair. Red balls and blue balls denote O ions and Li ions, respectively. In (a) and (b), Li ions migrate via Oct-Tet transitions, whereas in (c) and (d), Li ions migrate between Tet sites.

tial Li ions with rising temperature, which appears counterintuitive for thermally activated diffusion processes. In contrast, the deep potential model used in this work, trained directly on DFT data, is suitable to capture the energy landscape associated with defect formation and Li-ion migration in Li₂O. This enables a more accurate characterization of the temperature dependent diffusion mechanisms and offers a more reliable microscopic interpretation of the observed behavior. In particular, the above insights suggest that strategies to increase Li_{Oct} population, such as doping, could enhance Li-ion mobility and lower the superionic transition temperature.

IV. CONCLUSIONS

Employing a deep potential molecular dynamics (DPMD) framework trained with *ab initio* precision, we investigate the dynamical behavior of lithium oxide. Our DPMD simulations identify three well-defined regimes: a normal crystalline phase (<1300 K); a superionic phase (1300–1700 K) exhibiting gradual type II transition characteristics; and a liquid phase (>1700 K). An obvious softening of elas-

tic constants C_{11} in 1100–1200 K acts as an acceleration of Li-ion diffusion, which is still classified as normal state. This is also proved by dynamics trajectories and van Hove correlation functions at specific temperatures, which also confirm that it's the melting of Li sublattice that leads to the superionic transition. The superionic transition is governed by the formation of Frenkel pairs, where Li ions move from their original tetrahedral sites to interstitial octahedral sites. A unique defect-mediated diffusion is observed, characterized by fast and effective Li_{Oct} diffusion, and much slower Li_{Tet} diffusion. This work provides fundamental insights into the structure-property relationship of Li₂O, linking microscopic ion dynamics to macroscopic properties such as ionic coefficient and elastic constants, and elucidates the microscopic mechanism underlying the superionic transition.

Particularly, these findings can be extended to other materials with fluorite and antifluorite structures, offering a broader framework for investigating ionic diffusion and superionic transitions, and facilitating their practical applications.

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The authors declare no competing interests.

DATA AVAILABILITY

The data that support the findings of this article are not publicly available upon publication because it is not technically feasible and/or the cost of preparing, depositing, and hosting the data would be prohibitive within the terms of this research project. The data are available from the authors upon reasonable request.

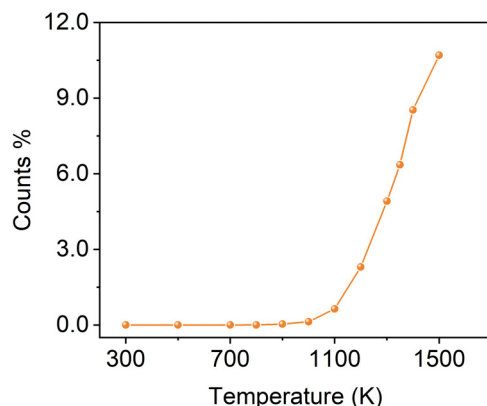


FIG. 6. The number of Li at octahedral interstitial sites.

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