# Oxidation Number Rules

1. For an atoms in its elemental form: O.N. = 0
2. For a monoatomic ion: O.N. = ion charge
3. The sum of the O.N. values for the atoms in a neutral compound equals zero. In a polyatomic ion, it equals the ion's charge.
4. Specific groups:
   a. **1A**: O.N. = +1
   b. **2A**: O.N. = +2
   c. **Hydrogen**: O.N. = +1 (with nonmetals)
      O.N. = -1 (with metals)
   d. **Fluorine**: O.N. = -1
   e. **Oxygen**: O.N. = -2 (except with F)
      In peroxide oxygen's O.N. = -1
   f. **7A**: O.N. = -1 (with metals, nonmetals (except O), and halogen lower in the group)