CHM129

Acid-Base Equilibrium: Salts

Determine whether aqueous solutions of the following salts will be acidic, basic or neutral:

(a) $Ca(C_2H_3O_2)_2$ Basic

(b) NH_4NO_3 Acidic

(c) KCl Neutral

(d) $Fe(ClO_4)_3$ Acidic

(e) NaClO Basic

(f) $Ba(NO_3)_2$ Neutral

(g) NH_4NO_2 $Ka(NH_4^+)=5.6x10^{-10}$, $Kb(NO_2^-)=2.2x10^{-11}$

$Ka>Kb$ → Acidic

(h) CH_3NH_3Cl Acidic

(i) NaF Basic