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Chapter 16 Solutions 167. SECTION 16.1 PROPERTIES OF SOLUTIONS (pages 471–477) This section identifies the factors that affect the solubility of a substance and determine the rate at which a solute dissolves. Solution Formation (pages 471–472) Look at Figure 16.1 on page 471 to help you answer Questions 1 and 2.

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i am looking for the same answer of this question for my science class will somebody please answer it all nitrates are soluble, therefore they can make aqueous solutions. ... The five properties ...

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16.1 Properties of Solutions 16

Solutions Chapter 13 Properties of Solutions Chemistry, The Central Science, 10th edition Theodore L. Brown; H. Eugene LeMay, Jr.; and Bruce E. Bursten. Solutions Solutions • Solutions are homogeneous mixtures of two or more pure substances. • In a solution, the solute is dispersed uniformly

Chapter 13 Properties of Solutions

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Chapter 16: Colligative Properties of Solutions 45 16-4. The mole fraction of (NH 4) 2SO 4(aq) is given by x (NH 4) 2SO 4 = n (NH 4) 2SO 4 n (NH 4) 2SO 4 + n H 2O Because (NH 4) 2SO 4(aq) is a strong electrolyte, it dissociates completely into NH + 4 (aq) and SO2- 4 (aq) ions. Assume a one kilogram solution. The number of moles of ions in one ...

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CHAPTER 16. Colligative Properties of Solutions

Sample answer: About 15.3 g of potassium nitrate would dissolve. 3. The volume of water increases by a factor of 10, so the amount of solute will increase by about the same factor. For the sample data, this will be 153 g of potassium nitrate, or ten times the amount in 10 mL of water at 60°C.

Acids, Bases, and Solutions ANSWER KEY - Lab35

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Properties of Solutions: Help & Review - Practice Test ...

1. pH=pKa + log (base/acid), best with equimolar concentrations 2. C6H8O7 + NaOH = NaC6H7O7 + H2O C6H7O7 + NaOH = NaC6H6O7 + H2O C6H6O7 + NaOH = NaC6H5O7 + H2O 3. a. Equal molar concentrations of C6H8O7 and NaC6H7O7 b. Equal molar concentrations of C6H6O7 and NaC6H5O7 4. Ideal

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