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[2] (b) When sulfur is heated under pressure with chlorine, the major product is  $\text{SCI}_2$  ( $\text{Cl-S-Cl}$ ).  $\text{S}_8(\text{g}) + 8\text{Cl}_2(\text{g}) \rightarrow 8\text{SCI}_2(\text{g})$  Use data from the Data Booklet to calculate the enthalpy change,  $\Delta H$ , for this reaction. The eight sulfur atoms in the  $\text{S}_8$  molecule are all joined in a single ring by single bonds.

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