14 The Behavior Of Gases Chapter Quiz

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14 The Behavior Of Gases
Ch 14 The Behavior of Gases. it creates an opening between the inside of the can and the air outside causing the gas to flow through the opening to the lower pressure region outside. this movement of the gas forces the paint outside of the can.

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Chapter 14 The Behavior of Gases149 12. When the volume is reduced by one half, what happens to the pressure? 13. Is the following sentence true or false? Raising the temperature of a contained gas causes its pressure to decrease. _____ 14. Circle the letter next to each sentence that correctly describes how gases behave when the temperature increases. a.

SECTION 14.1 PROPERTIES OF GASES(pages 413–417)
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The Behavior Of Gases Chapter 14. The large relative distances between gas particles means that there is considerable empty space between them.The assumption that gas particles are far apart explains gas compressibility.Compressibility is a measure of how much the volume of matter decreasesunder pressure.

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Unformatted text preview: Chapter 14 “The Behavior of Gases” Section 14.1 The Properties of Gases OBJECTIVES: Explain why gases are easier to compress than solids or liquids are. Describe the three factors that affect gas pressure.

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After you claim an answer you’ll have 24 hours to send in a draft. An editor will review the submission and either publish your submission or provide feedback. Next Answer Chapter 14 - The Behavior of Gases - 14.2 The Gas Laws - 14.2 Lesson Check - Page 463: 22 Previous Answer Chapter 14 - The ...

Chapter 14 - The Behavior of Gases - 14.2 The Gas Laws ...
Chapter 14 - The Behavior of Gases - 14.3 Ideal Gases - 14.3 Lesson Check - Page 468: 34 Answer An ideal gas is a gas that follows the gas laws at all conditions of pressure and temperature.

Chapter 14 - The Behavior of Gases - 14.3 Ideal Gases - 14 ...
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Chapter 14
exerted by a mixture of gases is equal to the sum of the partial pressures of the component gases. c. d. the contribution each gas in a mixture makes to the total pressure e. A gas tends to move to an area of lower concentration until the concentration is uniform throughout. f. P 1 T 1 V 1 P 2 T 2 V 2 g. P V n R T h.

05 CTR ch14 7/12/04 8:13 AM Page 361 THE BEHAVIOR OF GASES 14
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The behavior of gases is explained by what scientists call the Kinetic Molecular Theory. According to this theory, all matter is made of constantly moving atoms or molecules. Because of their mass and velocity, they possess kinetic energy, (K.E. = 1/2mv).

The Theories and Behavior of Gas | Owlcation
THE PROPERTIES OF GASES 14.1 Section Review Objectives why gases are easier to compress than solids or liquids are Describe the three factors that affect gas pressure Vocabulary compressibility Part A Completion Use this completion exercise to check your understanding ofthe concepts and terms that are introduced in this section.

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Chapter 14. The Behavior of Gases The Properties of Gases OBJECTIVES: Explain why gases are easier to compress than solids or liquids are. Section 14.1 The Properties of Gases OBJECTIVES: Describe the three factors that affect gas pressure. Compressibility Gases can expand to fill its container, unlike solids or liquids The reverse is also true: They are easily compressed, or squeezed into a ...

Chapter 14 the Behavior of Gases | Gases | Temperature
Chapter 14 The Behavior of Gases147 SECTION 14.1 PROPERTIES OF GASES(pages 413–417) This section uses kinetic theory to explain the properties of gases. This section also explains how gas pressure is affected by the amount of gas, its volume, and its temperature.

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Chemistry Ch 14A - Name Date Class THE BEHAVIOR OF GASES ...
Chapter 14 - Behavior of Gases Chapter 15 - Water and Aqueous Systems Chapter 16 - Solutions Chapter 17 -Thermochemistry Chapter 18 - Reaction Rates and Equilibrium Chapter 19 - Acids, Bases and Salts Chapter 20 - Oxidation-Reduction Reactions Chapter 25 - Nuclear Chemistry

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Chapter 14 Multiple Choice ... Under what conditions of temperature and pressure is the behavior of real gases most like that of ideal gases? a. low temperature and low pressure c. high temperature and low pressure b. ... Chapter 14 example test () ...

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Chapter 13: The Behavior of Gases b. Ptotal = P1 + P2 + P3 + ...Pn, where n = the total number of gases in the mixture c. Suppose that three different gases in three separate but identically-sized containers are all combined into one container.

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• Relate the total pressure of a mixture of gases to the partial pressures of the component gases • Explain how the molar mass of a gas affects the rate at which

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