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chapter 4 atomic structure guided practice problems answers ... Chapter 4 Atomic Structure Guided About This Quiz & Worksheet. Although the theory of an atom was discussed by the ancient Greeks, the various pieces that make up atomic structure were not discovered until the 19th and 20th

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name, element symbol, atomic number, and average atomic mass. 2. List three things, other than the name, symbol, atomic number, and average atomic mass, you can discover about an element using the periodic table in Figure 6.9. a. \_\_\_\_ b.

### **SECTION 6.1 ORGANIZING THE ELEMENTS (pages 155-160) (page 155)**

CHEMISTRY 12. Determine the total number of electrons that are used in bonding.  $20 e^- - 14 e^- = 6 e^-$ , or 3 covalent bonds. Step 3 Determine the number of non-bonding electrons.  $14 \text{ valence } e^- - 6 \text{ bonding } e^- = 8 \text{ non-bonding } e^-$ , or 4 lone pairs Hydrogen has no lone pair, and each oxygen atom has 2 lone pairs.

### **Chapter 4 ANSWER KEY - Quia**

CHAPTER 5, Atomic Structure and the Periodic Table(continued) Just How Small Is an Atom? (page 108) 6. Suppose you could grind a sample of the element copper into smaller and smaller particles. The smallest particle that could no longer be divided, yet still has the properties of copper, is \_\_\_\_ . 7.

### 5 ATOMIC STRUCTURE AND THE PERIODIC TABLE - Bergen

Chemistry--Unit 1: Atomic Structure and the Periodic Table Practice Problems I. Atoms 1)

100,000,000 copper atoms would form a line 1 cm long. How long would a line formed by  $1 \times 10^7$  copper atoms be? Express your answer in millimeters. II. Structure of the Atom 2) A sulfur-32 atom contains 16 protons, 16 neutrons, and 16 electrons.

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