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Set up two algebraic equations, one expressing the mass of carbon dioxide produced in terms of each reagent and the other expressing the mass of sample burned in terms of each reagent. What is the mass of methyl alcohol ( $\text{CH}_3\text{OH}$ ) in the sample? An organic liquid is a mixture of methyl alcohol ( $\text{CH}_3\text{OH}$ ) and ethyl alcohol ( $\text{C}_2\text{H}_5\text{OH}$ ). A 0.220-g sample of the liquid is burned in an excess of  $\text{O}_2(\text{g})$  and ...

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