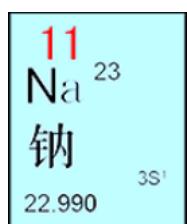
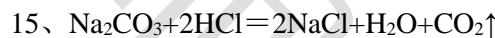
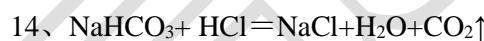
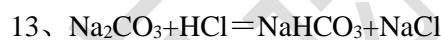
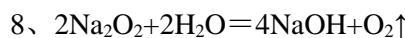
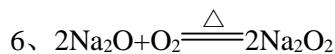
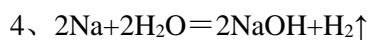
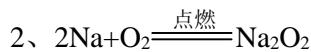
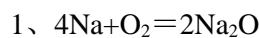




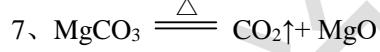
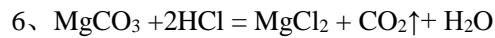
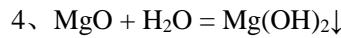
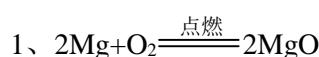
高中无机化学方程式汇总



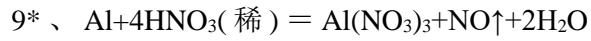
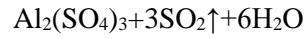
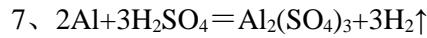
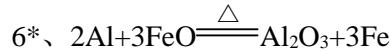
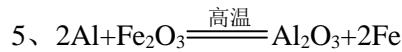
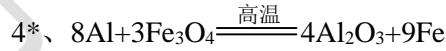
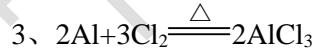
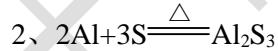
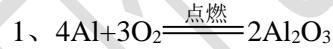
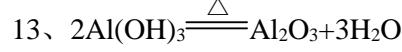
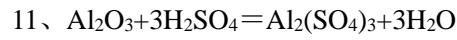
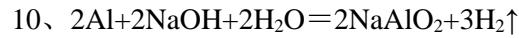
钠及其化合物

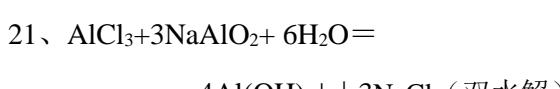
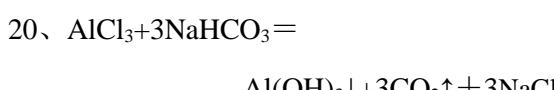
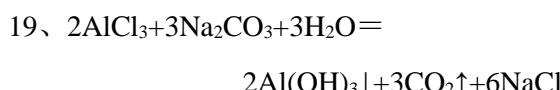
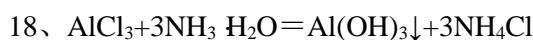
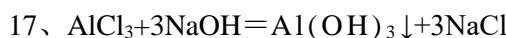


镁及其化合物

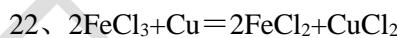
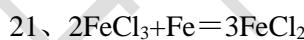
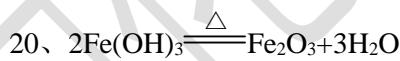
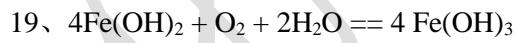
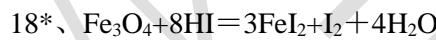
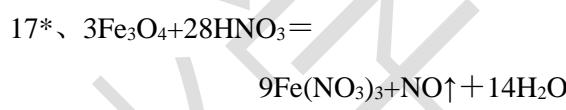
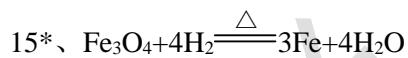
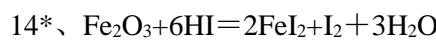
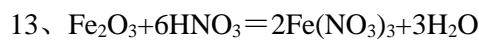
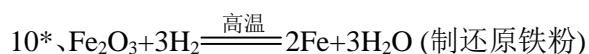
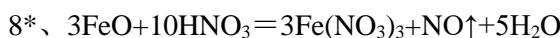
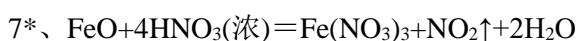
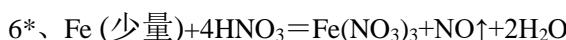
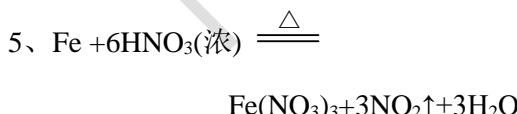
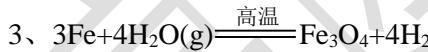
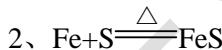
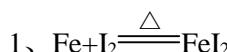


铝及其化合物

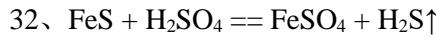
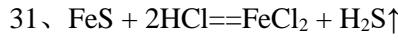
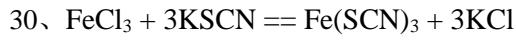
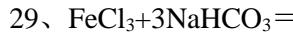
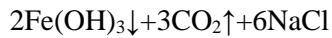
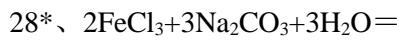
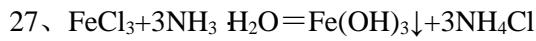
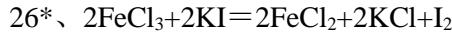
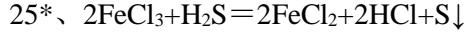
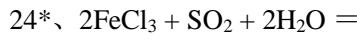
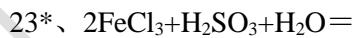
(Al、Fe 在冷的浓 H_2SO_4 、浓 HNO_3 中钝化)

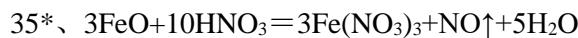
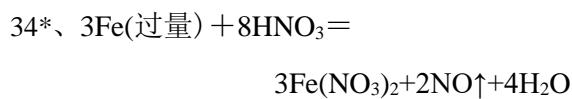
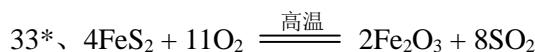


铁及其化合物



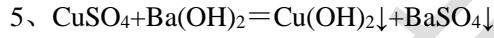
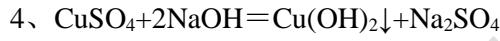
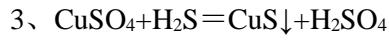
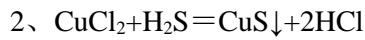
(用于雕刻铜线路版)





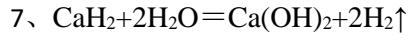
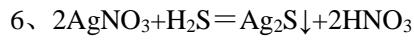
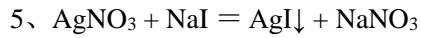
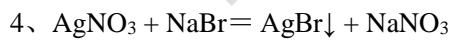
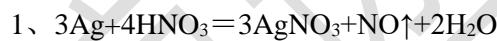
29 Cu 铜 $\frac{63}{65}$ 63.546 $3d^{10}4s^1$

铜及其化合物



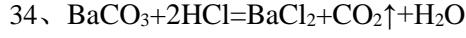
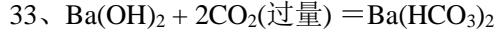
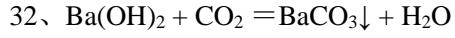
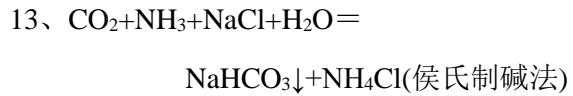
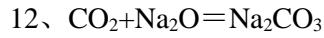
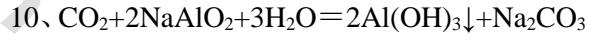
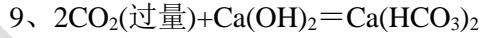
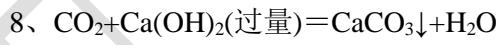
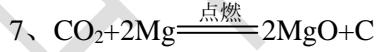
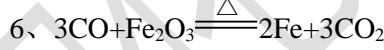
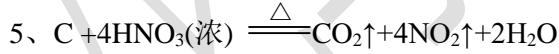
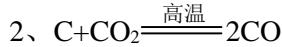
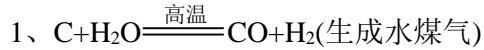
20 Ca 钙 $\frac{40}{42}$ $\frac{44}{46}$ $\frac{43}{48}$ 40.078 $4s^2$	47 Ag 银 $\frac{107}{109}$ 107.87 $4d^{10}5s^1$
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钙银及其化合物



6 C 碳 $\frac{12}{13}$ $2s^22p^2$ 12.011
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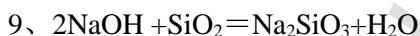
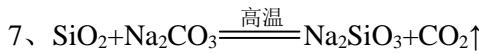
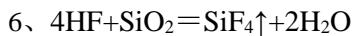
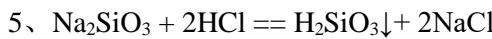
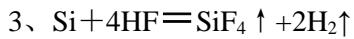
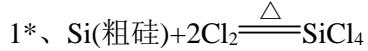
碳及其化合物





14	Si	28 29 30	$3s^23p^2$
	硅		28.086

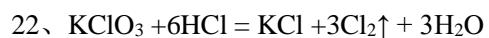
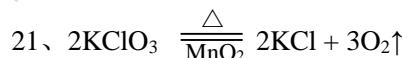
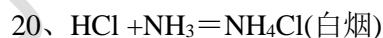
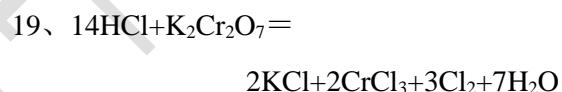
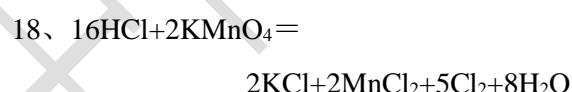
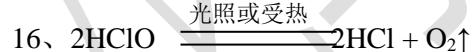
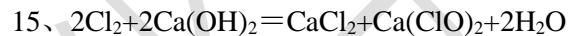
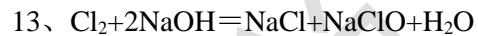
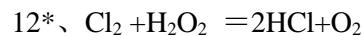
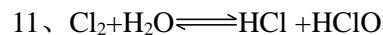
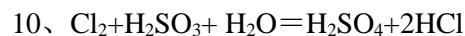
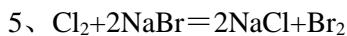
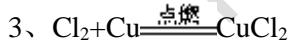
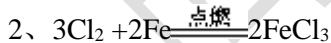
硅及其化合物



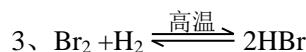
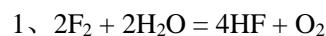
(常温下强碱缓慢腐蚀玻璃)

17	Cl	35 37	$3s^23p^5$
	氯		35.453

氯及其化合物



9	F	19	氟溴碘及其化合物	
	氟		$2s^22p^5$	18.998
	Br	79 81	$4s^24p^5$	79.904
	I	127 129	$5s^25p^5$	126.90





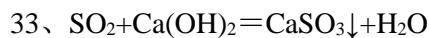
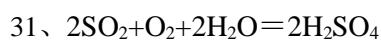
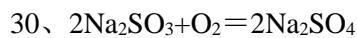
- 5、 $\text{Br}_2 + 2\text{FeBr}_2 = 2\text{FeBr}_3$
- 6、 $\text{Br}_2 + 2\text{NaI} = 2\text{NaBr} + \text{I}_2$
- 7、 $\text{Br}_2 + 2\text{HI} = 2\text{HBr} + \text{I}_2$
- 8、 $\text{Br}_2 + \text{H}_2\text{O} \rightleftharpoons \text{HBr} + \text{HBrO}$
- 9、 $\text{Br}_2 + 2\text{NaOH} = \text{NaBr} + \text{NaBrO} + \text{H}_2\text{O}$
- 10、 $\text{Br}_2 + \text{SO}_2 + 2\text{H}_2\text{O} = \text{H}_2\text{SO}_4 + 2\text{HBr}$
- 11、 $\text{Br}_2 + \text{H}_2\text{S} = \text{S} \downarrow + 2\text{HBr}$
- 12、 $\text{HBr} + \text{NaOH} = \text{NaBr} + \text{H}_2\text{O}$
- 13、 $2\text{HBr} + \text{Cl}_2 = 2\text{HCl} + \text{Br}_2$
- 14、 $2\text{HI} + \text{Cl}_2 = 2\text{HCl} + \text{I}_2$
- 15、 $\text{HI} + \text{NaOH} = \text{NaI} + \text{H}_2\text{O}$
- 16、 $\text{I}_2 + 2\text{NaOH} = \text{NaI} + \text{NaIO} + \text{H}_2\text{O}$



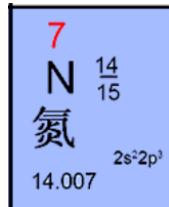
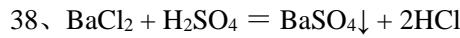
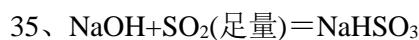
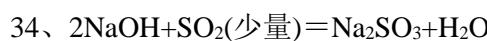
硫及其化合物

1. $2\text{H}_2\text{S} + 3\text{O}_2 \text{(足量)} \xrightarrow{\text{点燃}} 2\text{SO}_2 + 2\text{H}_2\text{O}$
2. $2\text{H}_2\text{S} + \text{O}_2 \text{(少量)} \xrightarrow{\text{点燃}} 2\text{S} + 2\text{H}_2\text{O}$
3. $2\text{H}_2\text{S} + \text{SO}_2 = 3\text{S} + 2\text{H}_2\text{O}$
4. $\text{H}_2\text{S} + \text{H}_2\text{SO}_4 \text{(浓)} = \text{S} \downarrow + \text{SO}_2 + 2\text{H}_2\text{O}$
5. $\text{H}_2\text{S} + 3\text{H}_2\text{SO}_4 \text{(浓)} \xrightarrow{\Delta} \text{S} + 4\text{H}_2\text{O}$
6. $3\text{Na}_2\text{SO}_3 + 2\text{HNO}_3 \text{(稀)} = 3\text{Na}_2\text{SO}_4 + 2\text{NO} \uparrow + \text{H}_2\text{O}$
7. $\text{H}_2\text{SO}_3 + \text{X}_2 + \text{H}_2\text{O} = \text{H}_2\text{SO}_4 + 2\text{HX}$
(X 表示 Cl_2 、 Br_2 、 I_2)
8. $\text{SO}_2 + \text{X}_2 + 2\text{H}_2\text{O} = \text{H}_2\text{SO}_4 + 2\text{HX}$
(X 表示 Cl_2 、 Br_2 、 I_2)
9. $2\text{H}_2\text{SO}_3 + \text{O}_2 = 2\text{H}_2\text{SO}_4$

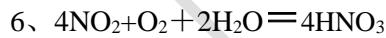
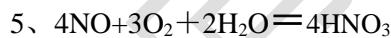
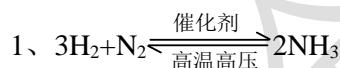
- 10、 $2\text{Na}_2\text{SO}_3 + \text{O}_2 = 2\text{Na}_2\text{SO}_4$
- 11、 $\text{H}_2\text{SO}_3 + \text{H}_2\text{O}_2 = \text{H}_2\text{SO}_4 + \text{H}_2\text{O}$
- 12、 $5\text{H}_2\text{SO}_3 + 2\text{KMnO}_4 = 2\text{MnSO}_4 + \text{K}_2\text{SO}_4 + 2\text{H}_2\text{SO}_4 + 3\text{H}_2\text{O}$
- 13、 $2\text{H}_2\text{SO}_4 \text{(浓)} + \text{S} \xrightarrow{\Delta} 3\text{SO}_2 \uparrow + 2\text{H}_2\text{O}$
- 14、 $2\text{H}_2\text{SO}_4 \text{(浓)} + \text{Cu} \xrightarrow{\Delta} \text{CuSO}_4 + \text{SO}_2 \uparrow + 2\text{H}_2\text{O}$
- 15、 $\text{H}_2\text{SO}_4 \text{(浓)} + 2\text{HBr} \xrightarrow{\Delta} \text{SO}_2 + \text{Br}_2 + 2\text{H}_2\text{O}$
- 16、 $\text{H}_2\text{SO}_4 \text{(浓)} + 2\text{HI} \xrightarrow{\Delta} \text{SO}_2 + \text{I}_2 + 2\text{H}_2\text{O}$
- 17、 $\text{H}_2\text{SO}_4 \text{(浓)} + \text{NaCl} = \text{NaHSO}_4 + \text{HCl} \uparrow$
- 18、 $\text{H}_2\text{SO}_4 \text{(浓)} + 2\text{NaCl} \xrightarrow{\Delta} \text{Na}_2\text{SO}_4 + 2\text{HCl}$
- 19、 $\text{H}_2\text{SO}_4 \text{(浓)} + \text{NaNO}_3 \xrightarrow{\Delta} \text{NaHSO}_4 + \text{HNO}_3 \uparrow \text{(微热)}$
- 20、 $2\text{KAl}(\text{SO}_4)_2 + 6\text{NaOH} = \text{K}_2\text{SO}_4 + 3\text{Na}_2\text{SO}_4 + 2\text{Al}(\text{OH})_3 \downarrow$
- 21、 $2\text{KAl}(\text{SO}_4)_2 + 3\text{Ba}(\text{OH})_2 \text{(少量)} = \text{K}_2\text{SO}_4 + 3\text{BaSO}_4 \downarrow + 2\text{Al}(\text{OH})_3 \downarrow$
(反应后溶液显中性)
- 22、 $\text{KAl}(\text{SO}_4)_2 + 2\text{Ba}(\text{OH})_2 \text{(过量)} = \text{KAlO}_2 + 2\text{BaSO}_4 \downarrow + 2\text{H}_2\text{O}$ (沉淀最大)
- 23、 $\text{Ba}(\text{OH})_2 + \text{H}_2\text{SO}_4 = \text{BaSO}_4 \downarrow + 2\text{H}_2\text{O}$
- 24、 $\text{BaCl}_2 + \text{H}_2\text{SO}_4 = \text{BaSO}_4 \downarrow + 2\text{HCl}$
- 25、 $\text{BaCl}_2 + \text{Na}_2\text{SO}_4 = \text{BaSO}_4 \downarrow + 2\text{NaCl}$
- 26、 $3\text{Na}_2\text{SO}_3 + 2\text{HNO}_3 \text{(稀)} = 3\text{Na}_2\text{SO}_4 + 2\text{NO} \uparrow + \text{H}_2\text{O}$
- 27、 $\text{Na}_2\text{S}_2\text{O}_3 + \text{H}_2\text{SO}_4 = \text{Na}_2\text{SO}_4 + \text{S} + \text{SO}_2 \uparrow + \text{H}_2\text{O}$
- 28、 $\text{Na}_2\text{SO}_3 + \text{SO}_2 + \text{H}_2\text{O} = 2\text{NaHSO}_3$
- 29、 $2\text{SO}_2 + \text{O}_2 \xrightarrow[\text{高温}]{\text{催化剂}} 2\text{SO}_3$



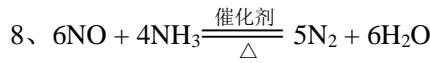
(不能用澄清石灰水鉴别 SO_2 和 CO_2 . 可用品红鉴别)



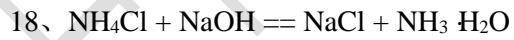
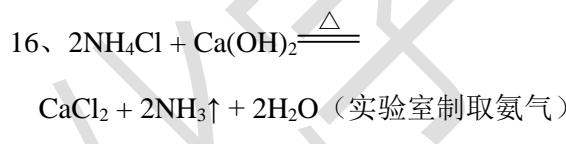
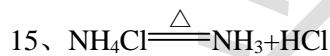
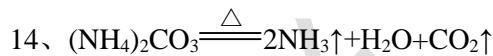
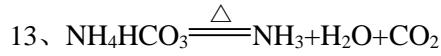
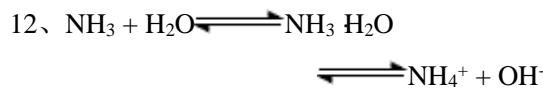
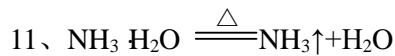
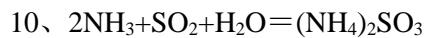
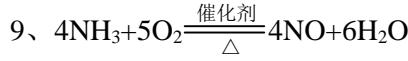
氮及其化合物



(用于制硝酸工业中吸收尾气中的 NO 和 NO_2)



(用氨清除 NO)



高送電