

Thermochemistry Workbook Answers

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Thermochemistry Answers - Worksheet Number One. We will ignore any heats losses to the walls of the container and losses to the air. These is a typical position to take since, in a real experiment, both would have to be accounted for, making for much more complexity.

Thermochemistry Answers - Worksheet Number One

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184 Guided Reading and Study Workbook CHAPTER 17, Thermochemistry (continued) 6. In thermochemical calculations, is the direction of heat flow given from the point of view of the system, or of the surroundings? 7. What universal law states that energy can neither be created nor destroyed

SECTION 17.1 THE FLOW OF ENERGY HEAT AND WORK (pages 505-510)

How can you describe the specific heat of olive oil if it takes approximately 420 J of heat to raise the temperature of 7 g of olive oil by 30°C? a. greater than the specific heat of water c. equal to the specific heat of water b. less than the specific heat of water d. Not enough information is given. ____
32. The specific heat of silver is 0 ...

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Prentice Hall Chemistry Chapter 17: Thermochemistry ...

Rules of Thermochemistry 1. The magnitude of ΔH is directly proportional to the amount of reactant or product 2. ΔH for the reaction is equal in magnitude but opposite in sign for ΔH for the reverse of the reaction 3. The value of ΔH is the same whether the reaction occurs in one step or as a series of steps

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