

## *Thermodynamics Statistical Kinetics Solutions*

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where  $R$  is the ideal gas constant,  $n$  the molar number,  $\Phi$  the volume fraction, and the subscripts 1 and 2 denote the solvent and the polymer, respectively. With this equation,  $\Delta S_{\text{com}}$  is always positive because  $\Phi_1 + \Phi_2 = 1$  and  $\partial \Delta S_{\text{com}} / \partial n_2$ , the partial molar entropy of the polymer, is also positive when the amount of the solvent is much larger than the polymer.

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### **Physical chemistry - Wikipedia**

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**Transport phenomena - Wikipedia**

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