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42.50^@"C" The answer to this problem depends on whether or not you should approximate the density of water to be equal to "1.0 g mL"^(-1). Since no information about density was provide, I assume that this is what you must do. However, it's important to note that water's density varies with temperature, and that the value "1.0 g mL"^(-1) is only an approximation.

# If you combine 360.0 mL of water at 25.00°C and 120.0 mL ...

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