

Modern Physics Final Study Guide

PHYS-122

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1 Special Relativity

1.1 Postulates of Special Relativity

Postulates:

- The laws of physics are the same in all inertial frames.
- The speed of light is constant in a vacuum for all observers, regardless of their motion relative to the source.

1.2 Relativity of Simultaneity

Relativity of Simultaneity: Two events that appear simultaneous to one observer may not appear simultaneous to another observer moving relative to the first. This is a direct consequence of the constancy of the speed of light and the Lorentz transformations.

1.3 Lorentz Transformations

Lorentz Transformations: The Lorentz transformations relate spacetime coordinates of an event in two inertial reference frames moving at a constant velocity relative to each other. These transformations preserve the invariant interval (Δs^2), a cornerstone of special relativity.

Transformations:

- For a frame S' moving with velocity v along the x -axis relative to S :

- Time: $t' = \gamma \left(t - \frac{vx}{c^2} \right)$
- Space (x -direction): $x' = \gamma (x - vt)$
- Space (y - and z -directions): $y' = y, \quad z' = z$

- The Lorentz factor is $\gamma = \frac{1}{\sqrt{1-v^2/c^2}}$.

Lorentz Transformation Matrix: The Lorentz transformation can be expressed using matrix multiplication:

$$x' = \Lambda x,$$

where $x = \begin{bmatrix} ct \\ x \\ y \\ z \end{bmatrix}$ and x' are the spacetime coordinates in S and S' , respectively. For a boost along the x -axis, the Lorentz transformation matrix Λ is:

$$\Lambda = \begin{bmatrix} \gamma & -\gamma\beta & 0 & 0 \\ -\gamma\beta & \gamma & 0 & 0 \\ 0 & 0 & 1 & 0 \\ 0 & 0 & 0 & 1 \end{bmatrix},$$

where $\beta = v/c$.

Key Properties:

- The inverse transformation is obtained by negating v , flipping the sign of β :

$$\Lambda^{-1} = \begin{bmatrix} \gamma & \gamma\beta & 0 & 0 \\ \gamma\beta & \gamma & 0 & 0 \\ 0 & 0 & 1 & 0 \\ 0 & 0 & 0 & 1 \end{bmatrix}.$$

- The Lorentz transformations preserve the invariant interval:

$$\Delta s^2 = x \cdot x = x' \cdot x'.$$

Applications:

- Lorentz transformations explain phenomena like time dilation, length contraction, and the relativity of simultaneity.
- They form the basis for defining and transforming four-vectors, such as four-momentum and four-velocity.

1.4 The Relativistic Invariant Interval

Invariant Interval: The relativistic invariant interval is a spacetime quantity that remains unchanged under Lorentz transformations.

Invariant Interval:

$$\Delta s^2 = (c\Delta t)^2 - (\Delta x)^2 - (\Delta y)^2 - (\Delta z)^2$$

Classifications of the Invariant Interval:

- **Time-like:** $\Delta s^2 > 0$, related to proper time:

$$\Delta s^2 = c^2 \Delta \tau^2$$

- **Space-like:** $\Delta s^2 < 0$, related to proper length:

$$-\Delta s^2 = L^2$$

- **Light-like:** $\Delta s^2 = 0$, events are connected by light:

$$c^2 \Delta t^2 = (\Delta x)^2 + (\Delta y)^2 + (\Delta z)^2$$

1.5 Relativistic Velocity Addition

Relativistic Velocity Addition: When combining velocities in different inertial frames, the relativistic formula must be used. Typically, these problems involve 3 objects and the following reference frames and velocities are defined:

- S : The stationary reference frame.
- S' : The moving reference frame.
- u : The velocity of an object relative to S .
- v : The velocity of S' relative to S .
- u' : The velocity of the object relative to S' .

Relativistic Velocity Addition Formula:

$$u' = \frac{u + v}{1 + \frac{uv}{c^2}}$$

Here, u' is the velocity of the object as measured in S' , given the velocity u of the object relative to S and the velocity v of S' relative to S .

1.6 Four-Vectors in Special Relativity

Four-Vectors: Physical quantities in relativity are represented as four-vectors that transform under Lorentz transformations.

General Form of a Four-Vector:

$$a^\mu = (a^0, \vec{a}) = (a^0, a^1, a^2, a^3), \quad a^0 = c\Delta t$$

Dot Product of Four-Vectors: The invariant dot product between two four-vectors a^μ and b^μ is:

$$a \cdot b = a^0 b^0 - \vec{a} \cdot \vec{b}$$

This quantity is Lorentz invariant (unchanged under Lorentz transformations).

1.7 Examples of Four-Vectors

Common Four-Vectors:

- **Four-Position:**

$$x^\mu = (ct, x, y, z)$$

- **Four-Velocity:**

$$u^\mu = \frac{dx^\mu}{d\tau} = \gamma(c, \vec{v}), \quad \gamma = \frac{1}{\sqrt{1 - v^2/c^2}}$$

- **Four-Momentum:**

$$p^\mu = mu^\mu = \gamma m(c, \vec{v})$$

For massless particles, like photons:

$$p \cdot p = 0, \quad E = |\vec{p}|c$$

1.8 Relativistic Conservation Laws

Relativistic Conservation Laws: The principles of conservation of energy and momentum extend to special relativity, incorporating the total energy (rest energy + kinetic energy) and relativistic momentum.

Relativistic Total Energy:

- The total energy E includes both the rest energy and the relativistic kinetic energy:

$$E = \gamma m_0 c^2.$$

- The rest energy is given by $E_0 = m_0 c^2$, where m_0 is the rest mass.
- In any physical process, the total energy is conserved.

Relativistic Momentum:

- The relativistic momentum is defined as:

$$\vec{p} = \gamma m_0 \vec{v},$$

where $\gamma = \frac{1}{\sqrt{1-v^2/c^2}}$.

- Both the magnitude and direction of total momentum in an isolated system are conserved.

Four-Momentum Conservation:

- Energy and momentum form components of the four-momentum vector:

$$p^\mu = \left(\frac{E}{c}, \vec{p} \right).$$

- In all inertial frames, the conservation of four-momentum holds:

$$\sum p_{\text{initial}}^\mu = \sum p_{\text{final}}^\mu.$$

Relativistic Collisions:

- Total energy and total momentum are conserved in collisions, including elastic and inelastic cases.
- Unlike classical mechanics, the concept of conserved "mass" is not generally applicable. For example:
 - In pair annihilation, rest mass converts to energy.
 - In particle creation, energy converts into rest mass.

Energy-Momentum Relation:

$$E^2 = (pc)^2 + (m_0c^2)^2$$

This relation connects the total energy E , the momentum p , and the rest mass m_0 of a particle.

Invariant Four-Momentum Magnitude:

$$p^\mu \cdot p_\mu = m_0^2 c^2$$

This invariant magnitude holds in all inertial frames, ensuring the consistency of relativistic conservation laws.

2 Quantum Mechanics

2.1 Quantum Theory of Light

2.1.1 Blackbody Radiation

Blackbody Radiation: A blackbody is an idealized object that absorbs all incident electromagnetic radiation and re-emits it as a function of its temperature. The spectral intensity of radiation emitted by a blackbody is described by Planck's Law.

Key Challenge: Classical physics predicted the "ultraviolet catastrophe," where energy radiated at high frequencies diverged. Planck resolved this by introducing the concept of quantized energy:

$$E = nhf, \quad n = 1, 2, 3, \dots$$

This assumption restricted the energy of oscillators to discrete levels, successfully explaining blackbody radiation.

Planck's Law: Describes the spectral radiance of a blackbody at a given temperature T :

$$I(\lambda, T) = \frac{2\pi hc^2}{\lambda^5} \frac{1}{e^{hc/\lambda k_B T} - 1}, \quad I(f, T) = \frac{8\pi h f^3}{c^3} \frac{1}{e^{hf/k_B T} - 1}.$$

Here:

- $I(\lambda, T)$: Intensity as a function of wavelength λ .
- $I(f, T)$: Intensity as a function of frequency f .
- h : Planck's constant.
- c : Speed of light.
- k_B : Boltzmann constant.

Stefan-Boltzmann Law: The total power radiated per unit area of a blackbody is proportional to the fourth power of its absolute temperature:

$$P = \sigma T^4, \quad \sigma = \frac{2\pi^5 k_B^4}{15c^2 h^3} \approx 5.67 \times 10^{-8} \text{ W/m}^2 \cdot \text{K}^4.$$

Here:

- P : Power per unit area emitted by the blackbody.
- σ : Stefan-Boltzmann constant.

Wien's Displacement Law: The wavelength λ_{max} at which the blackbody radiation is most intense is inversely proportional to the temperature:

$$\lambda_{\text{max}} = \frac{b}{T}, \quad b \approx 2.897 \times 10^{-3} \text{ m} \cdot \text{K}.$$

Here:

- λ_{max} : Wavelength at peak intensity.
- T : Absolute temperature.
- b : Wien's constant.

Key Insights:

- Planck's quantization hypothesis was a pivotal step toward quantum mechanics.
- The Stefan-Boltzmann law integrates Planck's law over all wavelengths to find total radiated power.
- Wien's displacement law helps identify the peak emission wavelength of a blackbody at a given temperature.

2.1.2 Photoelectric Effect

Photoelectric Effect: The photoelectric effect describes the emission of electrons from a material when light of sufficient frequency shines on it. The energy of the emitted electrons depends on the frequency of the light, not its intensity, providing evidence for the particle-like nature of light.

Einstein's Photoelectric Equation: The energy of the incident photons is used to overcome the work function (ϕ) of the material and provide the kinetic energy (KE_{max}) to the ejected electrons:

$$hf = \phi + KE_{max},$$

where:

- h : Planck's constant.
- f : Frequency of the incident light.
- ϕ : Work function (minimum energy needed to eject an electron).
- KE_{max} : Maximum kinetic energy of the emitted electron.

Threshold Frequency: Electrons are only emitted if the frequency of light f exceeds the threshold frequency f_c :

$$hf_c = \phi.$$

If $f < f_c$, no electrons are emitted regardless of the light's intensity.

Stopping Potential: To measure KE_{max} , a stopping potential V_s is applied to halt the emitted electrons:

$$KE_{max} = eV_s,$$

where e is the elementary charge.

Key Observations:

- The kinetic energy of emitted electrons depends on the light's frequency, not its intensity.
- There is a threshold frequency below which no electrons are emitted.
- Increasing light intensity increases the number of emitted electrons but not their kinetic energy.
- The stopping potential is independent of light intensity and depends solely on KE_{max} .

Experimental Verification: The photoelectric effect was experimentally validated by observing:

- The emission of electrons from metals when illuminated by ultraviolet light.
- A linear relationship between the frequency of light and KE_{max} .

This experiment strongly supported Einstein's photon theory of light, which extended Planck's quantization hypothesis to all electromagnetic radiation.

Applications:

- Photoelectric cells (e.g., solar panels) convert light into electricity.
- Light sensors and detectors in various scientific instruments.
- Validation of quantum theory and photon-based models of light.

2.1.3 Compton Scattering

Compton Scattering: Compton scattering describes the interaction between high-energy photons (e.g., X-rays) and electrons, leading to a measurable change in the photon's wavelength. This phenomenon provides strong evidence for the particle-like properties of light.

Compton Wavelength Shift: The change in the wavelength of the photon after scattering is given by:

$$\lambda' - \lambda = \frac{h}{m_e c} (1 - \cos \theta),$$

where:

- λ : Initial wavelength of the photon.
- λ' : Wavelength of the scattered photon.
- h : Planck's constant.
- m_e : Mass of the electron.
- c : Speed of light.
- θ : Angle of scattering (angle between the initial and scattered photon).

Energy Relation: The relationship between the photon's initial and final energies is:

$$\frac{1}{E'} - \frac{1}{E} = \frac{1 - \cos \theta}{m_e c^2},$$

where:

- E : Initial photon energy ($E = hf$).
- E' : Energy of the scattered photon.
- $m_e c^2$: Rest energy of the electron.

Key Observations:

- The scattered photon has lower energy and longer wavelength than the incident photon, consistent with energy conservation.
- The shift in wavelength depends only on the scattering angle θ and fundamental constants, not the material or initial photon energy.
- Demonstrates that photons carry momentum, supporting their particle-like nature.

Experimental Verification:

- Arthur Compton's experiments in 1923 confirmed the theoretical prediction of the wavelength shift.
- Results agreed precisely with the derived equation for the Compton wavelength shift, providing strong evidence for quantum theory.

Applications:

- Used in X-ray and gamma-ray spectroscopy to study material properties.
- Plays a key role in medical imaging techniques such as Compton scattering tomography.
- Foundational to understanding photon interactions in particle physics and astrophysics.

2.1.4 Wave-Particle Duality

Wave-Particle Duality: Light and matter exhibit dual behavior, displaying characteristics of both waves and particles depending on the experimental conditions.

Key Features:

- Travels as a wave, interacting with itself, and demonstrates phenomena such as interference and diffraction.
- Interacts as a particle, transferring discrete packets of energy, known as photons or quanta.
- Confirmed by experiments:
 - **Double-Slit Experiment:** Demonstrates interference patterns for light and electrons, showing wave-like behavior.
 - **Photoelectric Effect:** Demonstrates the particle nature of light, as photons eject electrons from a material.
 - **Blackbody Radiation:** Explained by Planck's quantization hypothesis, resolving the ultraviolet catastrophe.
 - **Compton Scattering:** Demonstrates photon momentum through scattering with electrons.

De Broglie Wavelength: $\lambda = \frac{h}{p} = \frac{h}{mv}$

Wave Number: $k = \frac{2\pi}{\lambda}$

2.1.5 Heisenberg Uncertainty Principle

Heisenberg Uncertainty Principle: A fundamental principle of quantum mechanics stating that certain pairs of physical observables cannot be simultaneously measured with arbitrary precision.

Key Relations:

- **Position-Momentum Uncertainty:**

$$\Delta x \Delta p \geq \frac{\hbar}{2},$$

where Δx and Δp are the standard deviations of position and momentum, respectively.

- **Energy-Time Uncertainty:**

$$\Delta E \Delta t \geq \frac{\hbar}{2},$$

where ΔE and Δt are the uncertainties in energy and time, respectively.

Variance-Squared Average Relationship: The uncertainty in a measurable quantity A is defined as:

$$\Delta A = \sqrt{\langle A^2 \rangle - \langle A \rangle^2},$$

where $\langle A \rangle$ is the expectation value, and $\langle A^2 \rangle$ is the expectation value of A^2 .

Implications:

- Establishes the probabilistic nature of quantum mechanics, replacing deterministic classical physics.
- Demonstrates the impossibility of assigning definite trajectories to particles as in classical mechanics.

2.1.6 Group and Phase Velocities

Group and Phase Velocities: These describe two distinct velocities associated with wave propagation. While group velocity represents the speed of the overall wave packet (and the particle), phase velocity corresponds to the speed of individual wave crests.

Group Velocity (v_g): The velocity of the wave packet, which often represents the velocity of energy or information transfer. It is defined as:

$$v_g = \frac{d\omega}{dk},$$

where:

- ω : Angular frequency of the wave.
- k : Wave number ($k = 2\pi/\lambda$).

Key Features:

- For non-dispersive media (constant v), $v_g = v_p$.
- In quantum mechanics, v_g is often associated with the particle's velocity.
- Relevant for wave packets, where multiple frequencies combine to form a localized pulse.

Phase Velocity (v_p): The velocity of individual wave crests, which is distinct from the velocity of the overall wave packet. It is defined as:

$$v_p = \frac{\omega}{k},$$

where:

- ω : Angular frequency.
- k : Wave number.

Key Features:

- v_p can exceed the speed of light (c) in certain media without violating causality.
- Represents the velocity of harmonic waves in a wave train.
- In dispersive media, $v_p \neq v_g$.

Relation Between Group and Phase Velocities: In dispersive media, the group velocity depends on the phase velocity:

$$v_g = v_p - \lambda \frac{dv_p}{d\lambda},$$

where:

- λ : Wavelength of the wave.
- $\frac{dv_p}{d\lambda}$: Rate of change of phase velocity with respect to wavelength.

Key Observations:

- v_g is the velocity of energy or particle propagation, making it more relevant for physical phenomena.
- v_p can provide insights into the nature of wave propagation in a medium.
- The relationship between v_g and v_p becomes significant in understanding wave dispersion.

Applications:

- Group velocity is crucial in quantum mechanics for describing particle motion through the de Broglie wavelength.
- Phase velocity plays a role in characterizing electromagnetic waves in media, including fiber optics and plasma physics.
- Both are essential for understanding dispersion in waves, which occurs in optics, acoustics, and quantum systems.

2.2 Schrödinger Equation

2.2.1 Overview

Schrödinger Equation: Governs the quantum mechanical behavior of particles. It describes how the wave function of a particle evolves under the influence of potential energy.

General Forms:

- **1D Time-Independent Schrödinger Equation:**

$$-\frac{\hbar^2}{2m} \frac{d^2\psi(x)}{dx^2} + V(x)\psi(x) = E\psi(x).$$

- **3D Schrödinger Equation (spherical coordinates):**

$$-\frac{\hbar^2}{2m} \nabla^2 \psi(r, \theta, \phi) + V(r)\psi(r, \theta, \phi) = E\psi(r, \theta, \phi).$$

The solution is separable:

$$\psi(r, \theta, \phi) = R(r)Y(\theta, \phi),$$

where $Y(\theta, \phi)$ are spherical harmonics.

2.2.2 Normalization of the Wave Function

Normalization Condition: The probability of finding a particle in the entire space must be equal to 1. The wave function $\psi(x)$ (or $\psi(r, \theta, \phi)$ in 3D) is normalized if:

$$\int_{-\infty}^{\infty} |\psi(x)|^2 dx = 1 \quad (1D) \quad \text{or} \quad \int |\psi(r, \theta, \phi)|^2 d\tau = 1 \quad (3D).$$

Here:

- $|\psi(x)|^2$ is the probability density in 1D.
- $|\psi(r, \theta, \phi)|^2$ is the probability density in 3D.
- $d\tau = r^2 \sin \theta dr d\theta d\phi$ is the volume element in spherical coordinates.

Key Insight: If the wave function is not normalized, it can be normalized by multiplying it by a constant A , such that:

$$\psi_{\text{norm}}(x) = A\psi(x), \quad A = \frac{1}{\sqrt{\int |\psi(x)|^2 dx}}.$$

2.2.3 Different Potentials

Free Particle:

- For $V(x) = 0$, solutions are plane waves:

$$\psi(x) = Ae^{ikx} + Be^{-ikx},$$

where $k = \frac{p}{\hbar}$, $\omega = \frac{E}{\hbar}$, and $p = \hbar k = \sqrt{2mE}$.

- e^{ikx} is right moving plane wave and e^{-ikx} is left moving plane wave.
- For full time dependent, $\Psi(x, t) = \psi(x)e^{-i\omega t}$
- Wave function is not normalizable!

Particle in a Box/Infinite Square Well Potential:

- For $V(x) = 0$ inside the box and $V(x) = \infty$ outside:

- General Solution 1D:

$$\psi_n(x) = A \sin\left(\frac{n\pi x}{L}\right)$$

- General Solution 3D:

$$\psi_{n_1, n_2, n_3} = A \sin\left(\frac{n_1\pi x}{L}\right) \sin\left(\frac{n_2\pi y}{L}\right) \sin\left(\frac{n_3\pi z}{L}\right)$$

- Particular Solution 1D:

$$\psi_n(x) = \sqrt{\frac{2}{L}} \sin\left(\frac{n\pi x}{L}\right), \quad E_n = \frac{n^2\pi^2\hbar^2}{2mL^2} = \frac{n^2\hbar^2}{8mL^2}$$

- Particular Solution 3D:

$$\psi_{n_1, n_2, n_3} = \left(\frac{2}{L}\right)^{3/2} \sin\left(\frac{n_1\pi x}{L}\right) \sin\left(\frac{n_2\pi y}{L}\right) \sin\left(\frac{n_3\pi z}{L}\right)$$

$$E_n = \frac{\hbar^2\pi^2}{2mL^2}(n_1^2 + n_2^2 + n_3^2)$$

Step Potential:

$$V(x) = \begin{cases} 0, & x < 0 \\ U_0, & x \geq 0 \end{cases}$$

The particle's behavior depends on whether its energy E is greater than or less than U_0 .

Case 1: $E > U_0$

Wave Function Solutions:

- $x < 0$: $\psi(x) = Ae^{ik_0x} + Be^{-ik_0x}$, $k_0 = \sqrt{\frac{2mE}{\hbar^2}}$
- $x \geq 0$: $\psi(x) = Ce^{ik_1x}$, $k_1 = \sqrt{\frac{2m(E-U_0)}{\hbar^2}}$

Reflection and Transmission Coefficients:

$$R = \left(\frac{k_0 - k_1}{k_0 + k_1} \right)^2, \quad T = \frac{4k_0k_1}{(k_0 + k_1)^2}, \quad R + T = 1$$

Case 2: $E < U_0$

Wave Function Solutions:

- $x < 0$: $\psi(x) = Ae^{ik_0x} + Be^{-ik_0x}$, $k_0 = \sqrt{\frac{2mE}{\hbar^2}}$
- $x \geq 0$: $\psi(x) = De^{-\kappa x}$, $\kappa = \sqrt{\frac{2m(U_0-E)}{\hbar^2}}$

Tunneling Insight: The particle cannot propagate in the $x \geq 0$ region but has a finite probability of being found in the barrier due to tunneling.

Key Insights:

- For $E > U_0$, the particle has probabilities of transmission and reflection at the step.
- For $E < U_0$, the particle exhibits tunneling with an exponentially decaying wave function in the barrier.

2.3 Hydrogen Atom

2.3.1 Wave Functions

Wave Functions:

$$\psi(r, \theta, \phi) = R_{nl}(r)Y_l^m(\theta, \phi),$$

where $R_{nl}(r)$ is the radial wave function, and $Y_l^m(\theta, \phi)$ are spherical harmonics.

Radial Probability Distribution: The probability density $P(r)$ is given by:

$$P(r) = r^2 |R_{nl}(r)|^2.$$

2.3.2 Quantum Numbers

Quantum Numbers:

- **Principal quantum number (n):** Determines energy level ($n = 1, 2, 3, \dots$).

- Tells us the **energy** of the Hydrogen atom

$$E(\psi_{n,\ell,m}) = -\frac{13.6 \text{ eV}}{n^2}$$

- **Orbital angular momentum quantum number (ℓ):** Determines the shape of the orbital ($\ell = 0, 1, \dots, n - 1$).

- Tells us the **total angular momentum** of the Hydrogen atom

$$L(\psi_{n,\ell,m}) = \hbar\sqrt{\ell(\ell+1)}$$

- **Magnetic quantum number (m):** Determines the orientation ($m = -\ell, -\ell + 1, \dots, \ell$).

- Tells us the **z-component of angular momentum** of the Hydrogen atom

$$L_z(\psi_{n,\ell,m}) = \hbar m$$

- **Spin quantum number (m_s):** Intrinsic angular momentum ($m_s = \pm \frac{1}{2}$).

- Has to be rotated **720 degrees** for it to return to original state
- Fermions have half integer spin (electrons, muons, protons, neutrons, etc.)
- Bosons have integer spin (photons, W and Z bosons, Higgs boson, etc.)
- Stern-Gerlach Experiment

2.3.3 Multi-Electron Atoms and the Pauli Exclusion Principle

Multi-Electron Atoms: In atoms with more than one electron, the arrangement of electrons is determined by their energies and quantum states:

- Electrons occupy orbitals starting from the lowest energy levels, but no more than two electrons can share the same orbital.
- Inner electrons shield outer electrons from the full nuclear charge, affecting orbital energies.
- Example: In Potassium ($Z = 19$), the electron configuration is $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$. The last electron occupies the $4s$ orbital instead of $3d$ due to shielding.

Pauli Exclusion Principle: No two electrons in an atom can occupy the same quantum state. This means:

- No two electrons can have identical quantum numbers (n, l, m, m_s) .
- This principle determines the filling of electron orbitals and gives rise to electron configurations.

Electron Configuration: The distribution of electrons in orbitals follows these rules:

- Electrons are added one at a time to the lowest energy subshell available.
- The periodic table reflects the outer electron configurations, which dictate chemical properties.

3 Statistical Mechanics

3.1 Microscopic vs. Macroscopic States

Microscopic States: The state of a system described by the position and momentum of every particle. For N particles, this requires $6N$ values (3 spatial coordinates and 3 momentum components per particle).

Macroscopic States: When N is large (e.g., Avogadro's number), it is impractical to track each particle. Instead, systems are described using averaged quantities like energy, volume, and temperature.

Example: For 10 coins:

- Total **microstates**: $2^{10} = 1024$.
- Total **macrostates**: 11 (e.g., all heads, 1 tail and 9 heads, etc.).

3.2 Multiplicity and Degeneracy

Multiplicity (g): The number of microstates corresponding to a macrostate. Calculated as:

$$g = \binom{N}{k} = \frac{N!}{k!(N-k)!},$$

where k is the number of particles in a specific state.

Example: For 10 coins with 3 tails and 7 heads:

$$g = \binom{10}{3} = \frac{10!}{3!7!} = 120.$$

Degeneracy: The number of quantum states associated with a given energy level.

Microcanonical Ensemble: A set of all microstates with the same energy, where each is equally probable.

3.3 Statistical Distributions

Probability Distributions: The likelihood of finding a particle in a state with energy E_i depends on temperature (T) and energy. The three primary distributions are:

- **Maxwell-Boltzmann (Distinguishable Particles):** Applicable when $T \gg T_c$ (thermal energy kT is much larger than quantum energy level spacing) and particle density is low.

$$f(E) = f_B(E) = Ae^{-E/kT}$$

- **Bose-Einstein (Indistinguishable Bosons):** Below the critical temperature (T_c), a Bose-Einstein condensate forms as many bosons occupy the ground state.

$$f(E) = f_{BE}(E) = \frac{1}{e^{\alpha} e^{E/k_B T} - 1}.$$

- **Fermi-Dirac (Indistinguishable Fermions):** Governed by the Pauli exclusion principle, where no two fermions occupy the same quantum state.

$$f(E) = f_{FD}(E) = \frac{1}{e^{\alpha} e^{E/k_B T} + 1}.$$

Occupation Number:

- The **occupation number**, $n(E)dE$, represents the average number of particles in a quantum state with energy between E and $E + dE$.
- It is calculated as:

$$n(E)dE = f(E)g(E)dE,$$

where:

- $f(E)$ is the probability distribution (Maxwell-Boltzmann, Bose-Einstein, or Fermi-Dirac).
- $g(E)dE$ is the degeneracy, or the number of states available in the energy range E to $E + dE$.

3.4 Normalization Conditions

Normalization Conditions:

- The total number of particles in the system:

$$N = \int f(E)g(E)dE,$$

where $f(E)$ is the distribution function, and $g(E)$ is the density of states.

- The total energy of the system:

$$E_{\text{tot}} = \int f(E)g(E)E dE.$$

3.5 Maxwell-Boltzmann Distribution

Maxwell-Boltzmann Distribution: Describes the distribution of particles in a system of distinguishable particles.

Degeneracy of States: For an ideal monatomic gas:

$$g(p)dp = 4\pi p^2 dp \rightarrow g(E)dE = 4\pi m\sqrt{2mE}dE$$

where:

- m : Particle mass.
- E : Energy of the particle.

Key Properties:

- Valid for classical gases where quantum effects are negligible.
- Particles are distinguishable.
- No restriction on the number of particles in a given state.

Applications: Used to describe the behavior of ideal gases in classical thermodynamics and kinetic theory.

3.6 Bose-Einstein Distribution

Bose-Einstein Distribution: Describes indistinguishable bosons, such as photons, which can occupy the same quantum state.

Degeneracy of States for Bosons:

$$g(n)dn = \frac{1}{8}4\pi n^2 dn \rightarrow g(E)dE = \frac{\pi}{4E_0^{3/2}}\sqrt{E}dE$$

where $E_0 = \frac{\hbar^2 \pi^2}{2mL^2}$

Key Properties:

- Bosons can occupy the same quantum state.
- Leads to phenomena like Bose-Einstein condensation at low temperatures.

Applications: Used to describe systems like photons (blackbody radiation) and liquid helium (superfluidity).

3.7 Bose-Einstein Condensation

Bose-Einstein Condensation: Occurs when a system of bosons is cooled below a critical temperature (T_c), causing many particles to occupy the ground state.

Key Relations:

- Critical temperature:

$$T_c = \left(\frac{\hbar^2}{2\pi mk} \right) \left(\frac{N}{V\zeta(3/2)} \right)^{2/3}.$$

where $\zeta(3/2) = 2.612$

- Fraction in the ground state:

$$\frac{N_0}{N} = 1 - \left(\frac{T}{T_c} \right)^{3/2}.$$

3.8 Fermi-Dirac Distribution

Fermi-Dirac Distribution: Describes indistinguishable fermions (particles with half-integer spin) that obey the Pauli Exclusion Principle.

Degeneracy of States for Fermions:

$$g(E)dE = \frac{\pi}{4E_0^{3/2}} \sqrt{E} dE = 8\sqrt{2}\pi \left(\frac{m^{3/2}V}{h^3} \right) \sqrt{E} dE$$

where:

- V : Volume of the system.
- m : Mass of the particle.
- $E_0 = \frac{h^2}{8mL^2}$

Key Properties:

- Fermions cannot occupy the same quantum state (Pauli Exclusion Principle).
- At absolute zero, all states up to the Fermi energy (E_F) are occupied, while higher energy states are empty.
 - The Fermi energy forms the boundary between occupied and unoccupied states
 - at $T = 0$, all states with energies below E_F are occupied, and those above are not

$$E_F = \frac{h^2}{2m} \left(\frac{3N}{8\pi V} \right)^{2/3}$$

Applications: Used to study electrons in metals, semiconductors, and the behavior of neutron stars.

3.9 Degeneracy Pressure

Degeneracy Pressure: Prevents the collapse of fermionic systems due to quantum mechanical effects, critical in white dwarfs and neutron stars.

Key Relation:

$$P \propto \left(\frac{\hbar^2}{m} \right) \left(\frac{N}{V} \right)^{5/3}.$$

4 Quick Reference

4.1 Summary of Equations

Special Relativity:

- Time Dilation:

$$\Delta t = \gamma \Delta t_0, \quad \gamma = \frac{1}{\sqrt{1 - v^2/c^2}}$$

- Length Contraction:

$$L = L_0 \sqrt{1 - v^2/c^2}$$

- Relativistic Velocity Addition:

$$u' = \frac{u + v}{1 + uv/c^2}$$

- Invariant Interval:

$$\Delta s^2 = (c\Delta t)^2 - (\Delta x)^2 - (\Delta y)^2 - (\Delta z)^2$$

- Energy-Momentum Relation:

$$E^2 = (pc)^2 + (m_0 c^2)^2$$

- Four-Momentum Conservation:

$$\sum p_{\text{initial}}^\mu = \sum p_{\text{final}}^\mu$$

Quantum Theory of Light:

- Planck's Law:

$$I(\lambda, T) = \frac{2\pi hc^2}{\lambda^5} \frac{1}{e^{hc/\lambda kT} - 1}$$

- Einstein's Photoelectric Equation:

$$hf = \phi + KE_{max}$$

- Threshold Frequency:

$$hf_c = \phi$$

- Stopping Potential:

$$KE_{max} = eV_s$$

- Compton Wavelength Shift:

$$\lambda' - \lambda = \frac{h}{m_e c} (1 - \cos \theta)$$

- Energy Relation in Compton Scattering:

$$\frac{1}{E'} - \frac{1}{E} = \frac{1 - \cos \theta}{m_e c^2}$$

- De Broglie Wavelength:

$$\lambda = \frac{h}{p} = \frac{h}{mv}$$

- Wave Number:

$$k = \frac{2\pi}{\lambda}$$

Heisenberg Uncertainty Principle:

- Position-Momentum Uncertainty:

$$\Delta x \Delta p \geq \frac{\hbar}{2}$$

- Energy-Time Uncertainty:

$$\Delta E \Delta t \geq \frac{\hbar}{2}$$

- Variance-Squared Average Relationship:

$$\Delta A = \sqrt{\langle A^2 \rangle - \langle A \rangle^2}$$

Schrödinger Equation:

- 1D Time-Independent Schrödinger Equation:

$$-\frac{\hbar^2}{2m} \frac{d^2\psi(x)}{dx^2} + V(x)\psi(x) = E\psi(x)$$

- 3D Schrödinger Equation (Spherical Coordinates):

$$-\frac{\hbar^2}{2m} \nabla^2 \psi(r, \theta, \phi) + V(r)\psi(r, \theta, \phi) = E\psi(r, \theta, \phi)$$

- Wave function Normalization:

$$\psi_{\text{norm}}(x) = A\psi(x), \quad A = \frac{1}{\sqrt{\int |\psi(x)|^2 dx}}.$$

- Particle in a Box Wave Function Normalized Solution 1D:

$$\psi_n(x) = \sqrt{\frac{2}{L}} \sin\left(\frac{n\pi x}{L}\right)$$

- Particle in a Box Wave Function Normalized Solution 3D:

$$\psi_{n_1, n_2, n_3}(x, y, z) = \left(\frac{2}{L}\right)^{3/2} \sin\left(\frac{n_1\pi x}{L}\right) \sin\left(\frac{n_2\pi y}{L}\right) \sin\left(\frac{n_3\pi z}{L}\right)$$

- Particle in a Box Energy Levels 1D:

$$E_n = \frac{n^2\pi^2\hbar^2}{2mL^2}$$

- Particle in a Box Energy Levels 3D:

$$E_{n_1, n_2, n_3} = \frac{\pi^2\hbar^2}{2mL^2} (n_1^2 + n_2^2 + n_3^2)$$

- Step Potential Reflection and Transmission Coefficients:

$$R = \left(\frac{k_0 - k_1}{k_0 + k_1}\right)^2, \quad T = \frac{4k_0k_1}{(k_0 + k_1)^2}$$

Statistical Mechanics:

- Maxwell-Boltzmann Distribution:

$$f_{MB}(E) = Ae^{-E/kT}$$

- Bose-Einstein Distribution:

$$f_{BE}(E) = \frac{1}{e^{\alpha} e^{E/kT} - 1}$$

- Fermi-Dirac Distribution:

$$f_{FD}(E) = \frac{1}{e^{\alpha} e^{E/kT} + 1}$$

- Occupation Number:

$$n(E)dE = f(E)g(E)dE$$

- Critical Temperature for Bose-Einstein Condensation:

$$T_c = \left(\frac{h^2}{2\pi mk} \right) \left(\frac{N}{V\zeta(3/2)} \right)^{2/3}$$

- Fraction in Ground State for Bose-Einstein Condensation:

$$\frac{N_0}{N} = 1 - \left(\frac{T}{T_c} \right)^{3/2}$$

- Fermi Energy:

$$E_F = \frac{h^2}{2m} \left(\frac{3N}{8\pi V} \right)^{2/3}$$

Constants:

- Stefan-Boltzmann Law:

$$P = \sigma T^4, \quad \sigma = 5.67 \times 10^{-8} \text{ W/m}^2 \cdot \text{K}^4$$

- Wien's Displacement Law:

$$\lambda_{\max} = \frac{b}{T}, \quad b \approx 2.897 \times 10^{-3} \text{ m} \cdot \text{K}$$

4.2 Mathematical Tools

- Trigonometric Identities:

$$\cos^2 x = \frac{1 + \cos(2x)}{2}$$

$$\sin^2 x = \frac{1 - \cos(2x)}{2}$$

$$\sin(2x) = 2 \sin x \cos x$$

- Series:

- Taylor Series Expansion: $\sum_{n=0}^{\infty} C_n (x-a)^n$ where $C_n = \frac{f^{(n)}(a)}{n!}$

- Binomial Series Expansion: $(1 \pm x)^r = 1 \pm rx \pm \dots$

- Integration Tips:

- For even functions: $\int_{-a}^a f(x)dx = 2 \int_0^a f(x)dx$

- For odd functions: $\int_{-a}^a f(x)dx = 0$

- Gaussian integrals often appear in wave packets

- Complex Numbers Review

- Complex conjugate: $z^* = a - bi$ for $z = a + bi$

- Modulus: $|z| = \sqrt{a^2 + b^2}$

- Euler's formula: $e^{i\theta} = \cos \theta + i \sin \theta$

- Argument: $\arg(z) = \tan^{-1} \left(\frac{y}{x} \right)$

- Second Order Differential Equations

- If ODE has this form: $\frac{d^2 f}{dx^2} = k^2 f$ where k is a constant

The solution is:

$$f(x) = Ae^{kx} + Be^{-kx}$$

4.3 Important Constants

- Stefan-Boltzmann constant: $\sigma = 5.67 \times 10^{-8} \text{ W/m}^2 \cdot \text{K}^4$
- Planck's constant: $h = 6.626 \times 10^{-34} \text{ J} \cdot \text{s} = 4.136 \times 10^{-15} \text{ eV} \cdot \text{s}$
- Reduced Planck's constant: $\hbar = 1.055 \times 10^{-34} \text{ J} \cdot \text{s}$
- Electron mass: $m_e = 9.11 \times 10^{-31} \text{ kg}$
- Speed of light: $c = 3.0 \times 10^8 \text{ m/s}$
- Bohr radius: $a_0 = 0.0529 \times 10^{-9} \text{ m}$
- Compton wavelength of the electron: $\frac{h}{m_e c} = 0.002426 \text{ nm}$

4.4 Unit Conversions

- $1 \text{ eV} = 1.602 \times 10^{-19} \text{ J}$
- $hc = 1240 \text{ eV} \cdot \text{nm}$
- $\hbar c = \frac{1240}{2\pi} \text{ eV} \cdot \text{nm}$
- Temperature: Convert between Kelvin and energy using $k_B = 1.38 \times 10^{-23} \text{ J/K}$.