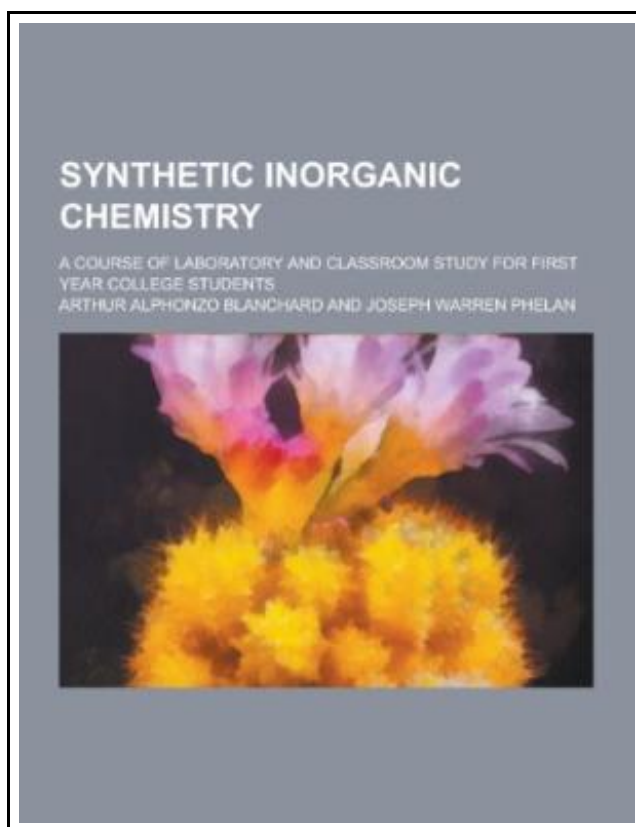


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Theclassics.Us, United States, 2013. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can usually download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1922 edition. Excerpt: . into the flame and note the deposit of sulphur. (c) Generate hydrogen gas in exactly the same way as hydrogen sulphide in (a). Place about gram of sulphur in the end of the hard glass tube nearest the generator. When it is ascertained that the hydrogen coming off is pure, heat the further end of the glass tube to redness and gradually move the flame towards the sulphur until this begins to be volatilized a little. By this arrangement a mixture of hydrogen and sulphur vapor is made to pass through a red-hot tube. Test the escaping gases with lead acetate paper. Parts (o) and (6) show that sulphur is deposited when hydrogen sulphide is strongly heated. In the freely burning flame the ultimate products are sulphur dioxide and water: $\text{H}_2\text{S} + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{SO}_2$. But if the unburned vapors in the interior of the flame are cooled by the porcelain before they can burn we find that sulphur is present there. Although we have not directly proved the presence of hydrogen we are pretty well justified in thinking that it is the other decomposition product of hydrogen sulphide. Part (c) shows that some hydrogen sulphide is formed when a mixture of hydrogen and sulphur vapor is heated to redness and it is thus clear that neither the synthesis nor the decomposition of hydrogen sulphide is complete at this temperature but that hydrogen sulphide reaches an equilibrium with its products according to the...



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