NE585 NUCLEAR FUEL CYCLES Nuclear engineering basics 1

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Learning objectives

Demonstrating fundamental concepts in nuclear physics

Analyzing decay data to identify radionuclides

Summarizing different decay mechanisms

Book – Chapter 2

Learning nodes

Relativity

Atomic structure

Binding energy

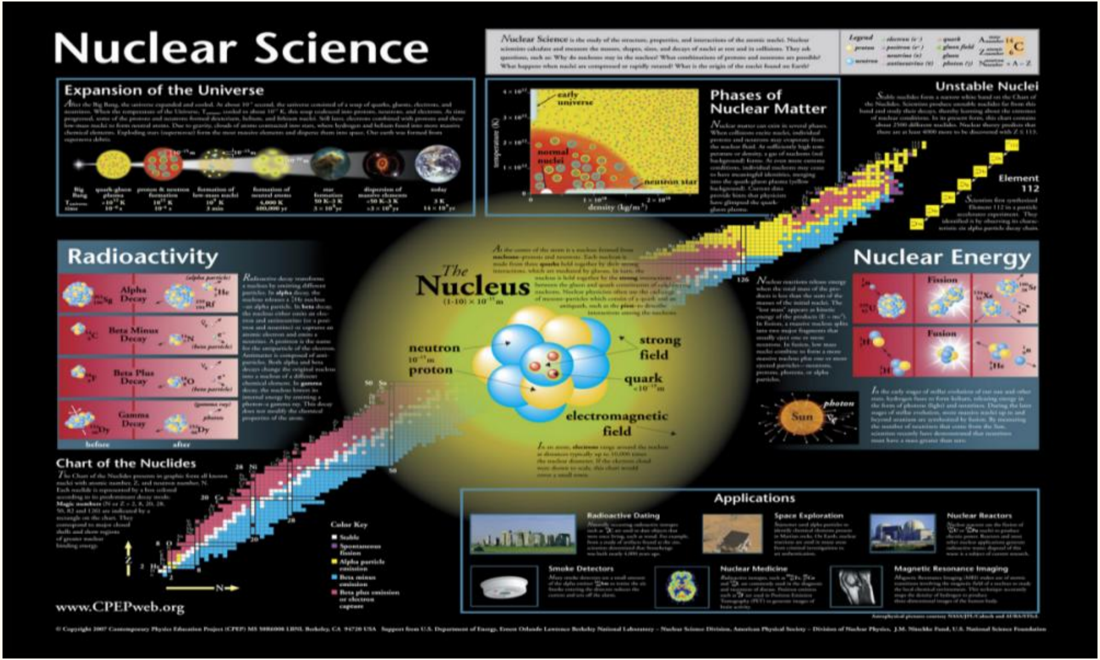
Q-value

Radioactivity

Chart of the nuclides

Decay chains

Half life


Nuclear Science
Expansion of the Universe
Phases of Nuclear Matter
Unstable Nuclei
Element 112
Radioactivity
Alpha Decay
Beta Minus Decay
Beta Plus Decay
Gamma Decay
before
after
The Nucleus
neutron
proton
strong field
quark
electromagnetic field
Nuclear Energy
Fission
Fusion
Sun
photon
Chart of the Nuclides
Color Key
Stable
Spontaneous fission
Alpha particle emission
Beta minus emission
Beta plus emission or electron capture
Applications
Radioactive Dating
Space Exploration
Nuclear Reactors
Smoke Detectors
Nuclear Medicine
Magnetic Resonance Imaging
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Relativity

We apply energy from nuclear reactions and radiation interactions with matter

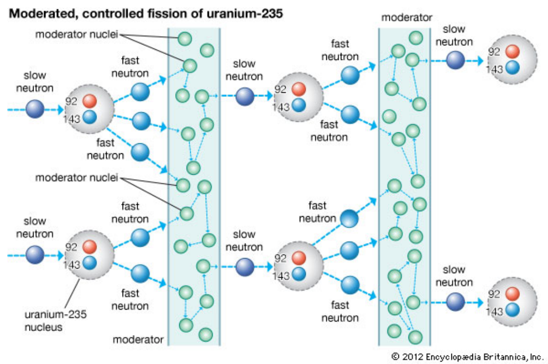
$$E_{REST} = m_0 c^2 \tag{1}$$

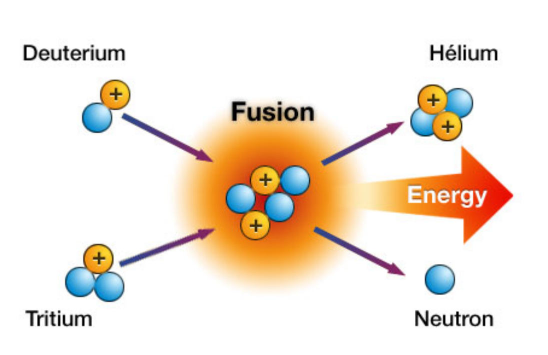
$$E = mc^2 \tag{2}$$

Rest mass energy for an electron – 0.511 MeV

Save that for later

Energy released or energy injected





Determine the energy of a fusion neutron that is not at rest

$$m = \frac{m_0}{\sqrt{1 - (\frac{v^2}{c})^2}} \rightarrow E = mc^2 \quad (3)$$

With $v = 5.2 \times 10^7 \; m/s$, that is about $0.17c$; not negligible

So we have to compute the not rest mass

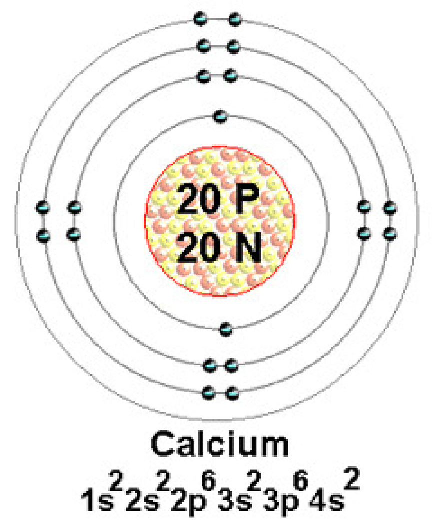
$m_0(n) = 1.674927211 \times 10^{-27} \; kg$

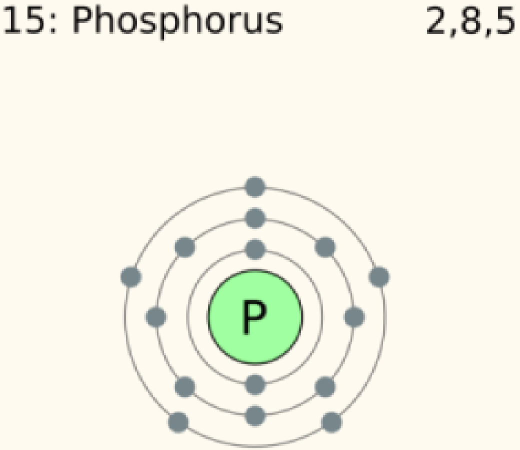
Fusion neutron at that speed has $E = 14.5 \; MeV$

For relativistic calculations use MeV

Atomic structure

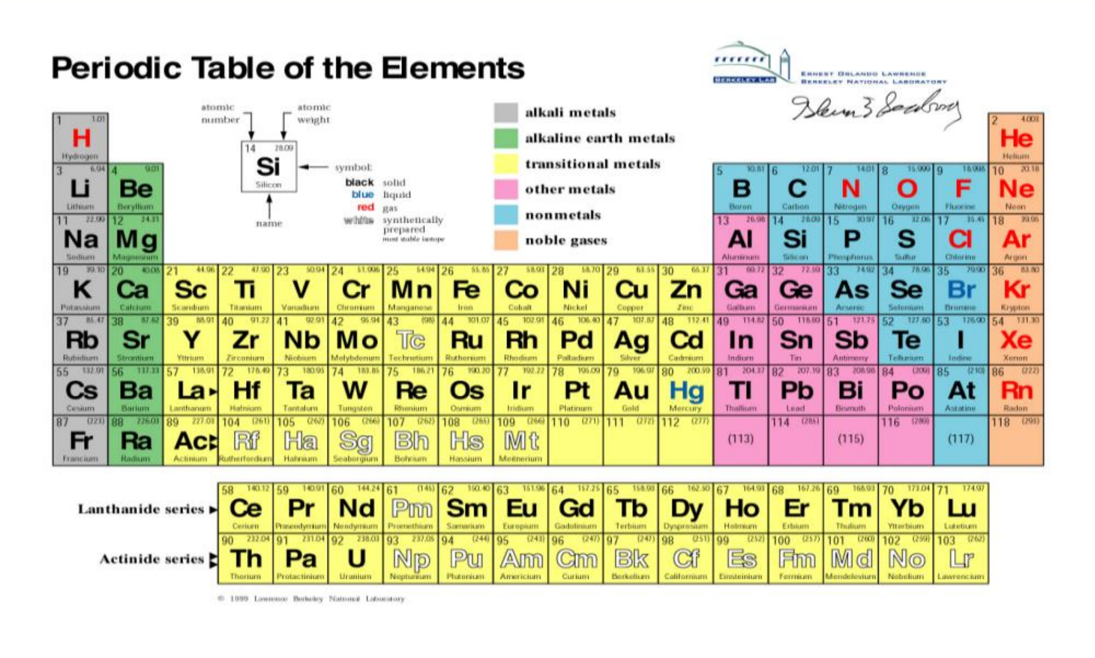
The shell model describes how subatomic particles are arranged





Notation

$$^{A}_{Z}Q_{A-Z} = ^{235}_{92}U_{143}$$



Periodic Table of the Elements
Ernest Orlando Lawrence Berkeley National Laboratory
atomic number
atomic weight
symbol:
black solid
blue liquid
red gas
white synthetically prepared
name
alkali metals
alkaline earth metals
transitional metals
other metals
nonmetals
noble gases
H He
Li Be B C N O F Ne
Na Mg Al Si P S Cl Ar
K Ca Sc Ti V Cr Mn Fe Co Ni Cu Zn Ga Ge As Se Br Kr
Rb Sr Y Zr Nb Mo Tc Ru Rh Pd Ag Cd In Sn Sb Te I Xe
Cs Ba La Hf Ta W Re Os Ir Pt Au Hg Tl Pb Bi Po At Rn
Fr Ra Ac Rf Ha Sg Bh Hs Mt (113) (115) (117)
Lanthanide series Ce Pr Nd Pm Sm Eu Gd Tb Dy Ho Er Tm Yb Lu
Actinide series Th Pa U Np Pu Am Cm Bk Cf Es Fm Md No Lr
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Binding energy

There are four fundamental forces of interaction

Gravity – Obvious

Strong nuclear – Holds the nucleus together

Electromagnetic – Basically electrons & magnetism

Weak nuclear – Changes the flavor of quarks

To blow out a nucleus, a lot of energy is needed

Would it be constant or variable per nucleus?

Why would you want to blow out the nucleus?

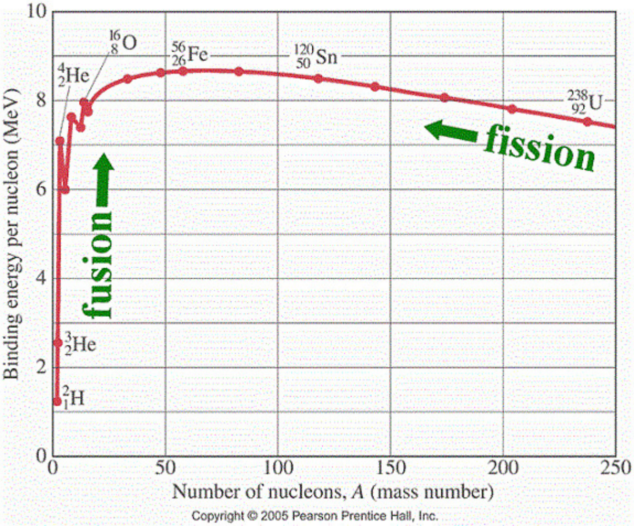
$$\Delta m = (A - Z)m_n - Zm_p - M \tag{4}$$

$$\Delta m({}^{239}_{94}Pu) = 1759\ MeV = 7.40\ \frac{MeV}{nucleon} \tag{5}$$

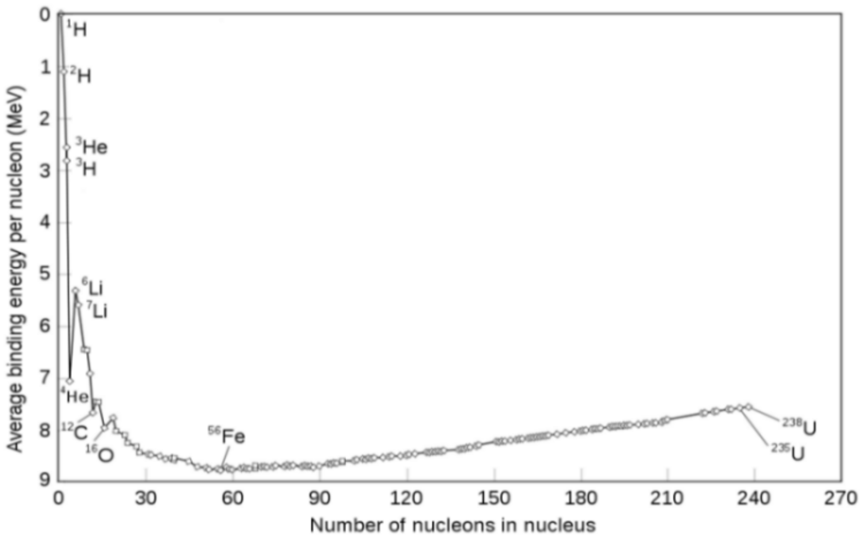
This energy corresponds to a velocity of 0.88c

$931.5\ \frac{MeV}{amu}$

Binding energy is used to compare atomic stability, reaction energy, probability of fission/fusion



Light nuclei get more stable by fusion; heavy nuclei get more stable by fission



Q-value

The Q-value for a reaction indicates energy requirements

$$ {}^{6}_{3}Li\ (n, \alpha)\ {}^{3}_{1}H \tag{6} $$

$$ {}^{6}_{3}Li + {}^{1}_{0}n \rightarrow {}^{4}_{2}\alpha + {}^{3}_{1}H \tag{7} $$

(1) Conservation of nucleons

(2) Conservation of charge

(3) Conservation of momentum

(4) Conservation of energy

Compute the Q value for this reaction

$$^6_3Li +^1_0 n \rightarrow^4_2 \alpha +^3_1 H \tag{8}$$

$$Q = [(M_{Li} + M_n) - (M_\alpha + M_H)]c^2 \tag{9}$$

$$Q = 4.78\ MeV \tag{10}$$

$Q > 0 \rightarrow ?$

$Q < 0 \rightarrow ?$

What is the energy of fission from thorium and uranium fuel cycles?

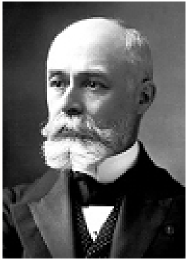
$$ {}^1_0n + {}^{235}_{92}U \rightarrow {}^{92}_{36}Kr + {}^{141}_{56}Ba + {}^1_0n + {}^1_0n + {}^1_0n \tag{11} $$

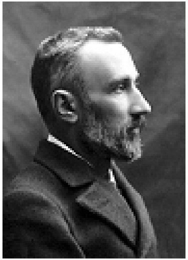
$$ {}^1_0n + {}^{233}_{92}U \rightarrow {}^{92}_{36}Kr + {}^{141}_{56}Ba + {}^1_0n \tag{12} $$

For comparison, burning a carbon atom releases 4 eV

Radioactivity

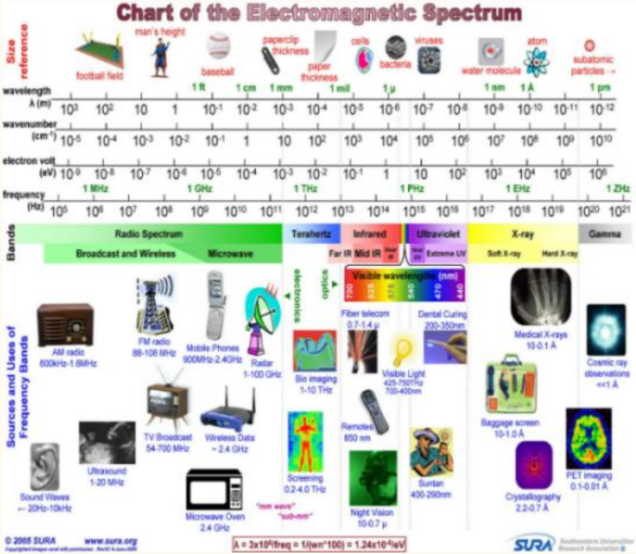
Scientists in late 1800s made discoveries which would change the course of science, medicine, and history in the 20th Century



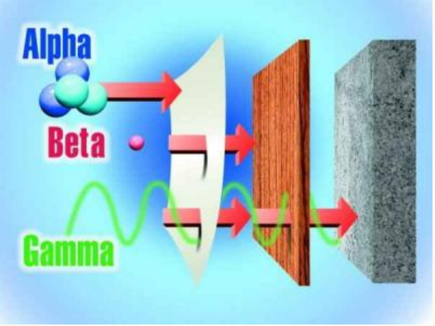


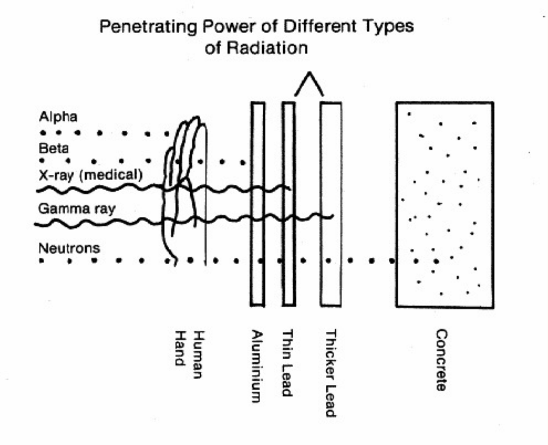
The Nobel Prize in Physics 1903 was divided, one half awarded to Antoine Henri Becquerel *"in recognition of the extraordinary services he has rendered by his discovery of spontaneous radioactivity"*, the other half jointly to Pierre Curie and Marie Curie, née Sklodowska *"in recognition of the extraordinary services they have rendered by their joint researches on the radiation phenomena discovered by Professor Henri Becquerel"*.

Radiation is energy as particles or electromagnetic waves, like sunshine



Radiation is energy as particles or electromagnetic waves





Radiation is energy as particles or electromagnetic waves

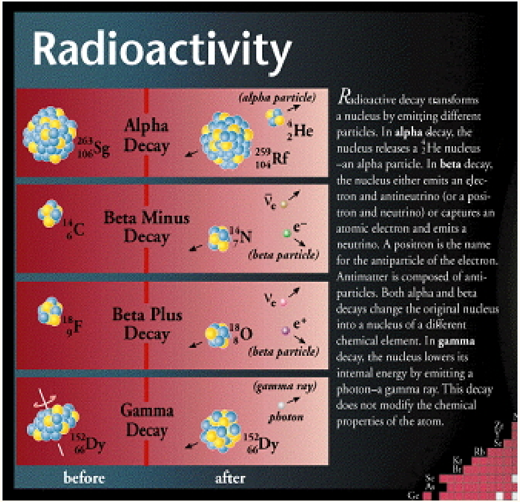
Charged particles are –

- protons (+)
- alpha (++)
- beta (+/-)
- heavy ions (varying)
- U,Pu (3+,4+,5+,6+)

Neutrons have no charge, but are also ionizing

Ionizing rays are x-rays, g-rays, cosmic rays

Ernest Rutherford discovered the three main kinds of radioactive decay (Chemistry Nobel 08)



Alpha particle decay is the ejection of a helium atom from a heavy nucleus

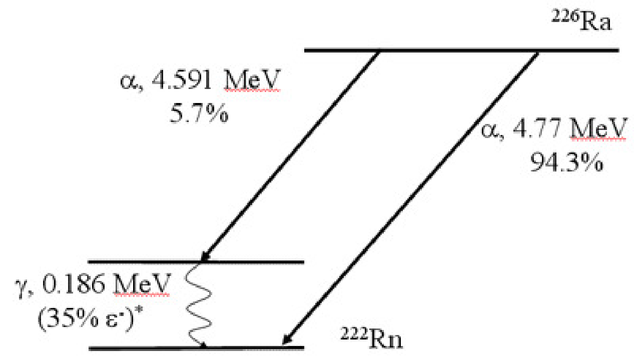
$$^{238}_{92}U \rightarrow ^{234}_{90} Th + ^{4}_{2} \alpha \tag{13}$$

$$Q = 4.268\ MeV \tag{14}$$

Sometimes nucleus decays to excited state and then emits a gamma ray (for any kind of radiation)

Discrete energy spectrum

Some % decay at certain energy

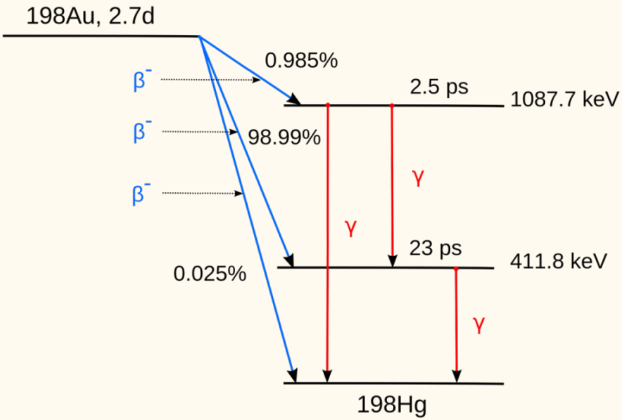

226Ra
α, 4.591 MeV
5.7%
α, 4.77 MeV
94.3%
γ, 0.186 MeV
(35% ε⁻)*
222Rn

β^- particle decay occurs for unstable nuclei with excessive neutrons

$$ {}^{19}_{8}O \rightarrow {}^{19}_{9}F + \beta^- + \overline{\nu} \tag{15}$$

The neutron is converted to a proton and a beta negative (electron) and anti-neutrino (momentum) are ejected

Sometimes nucleus decays to excited state and then emits a gamma ray (for any kind of radiation)

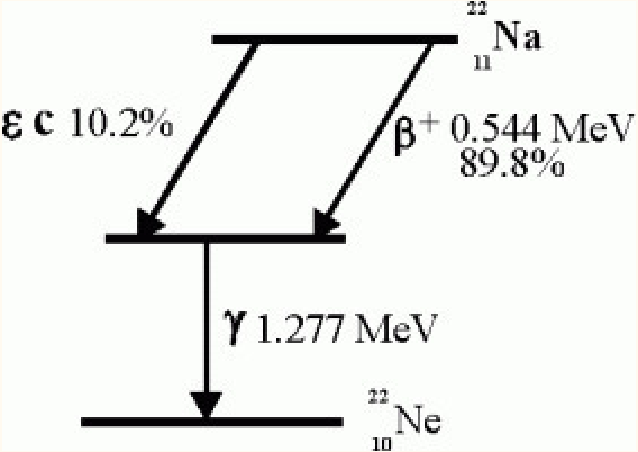

198Au, 2.7d
0.985%
β-
2.5 ps
1087.7 keV
β-
98.99%
γ
β-
γ
23 ps
411.8 keV
0.025%
γ
198Hg

β^+ particle decay occurs for unstable nuclei with deficient neutrons

$$ {}^{11}_{6}C \rightarrow {}^{11}_{5}B + \beta^+ + \nu \tag{16} $$

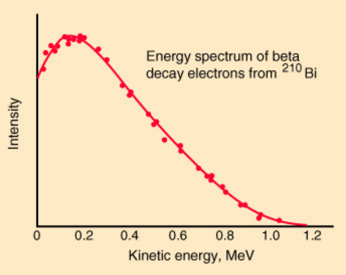
The proton is converted to a neutron and a beta positive (positron) and neutrino (momentum) are ejected

Sometimes nucleus decays to excited state and then emits a gamma ray (for any kind of radiation)



22
11
Na
ε c 10.2%
β+ 0.544 MeV
89.8%
γ 1.277 MeV
22
10
Ne

In both forms of beta decay, the energies are exhibited by a continuous energy spectrum



$$\overline{E}(\beta^-) = 0.3 E_{MAX} \tag{17}$$

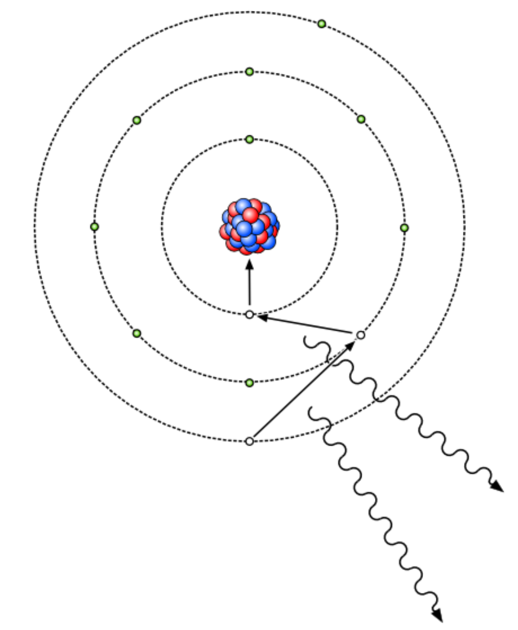
$$\overline{E}(\beta^+) = 0.4 E_{MAX} \tag{18}$$

Electron capture also occurs for unstable nuclei with deficient neutrons

$$^{22}_{11}Na \rightarrow \epsilon^{-} + ^{22}_{10}Ne + \nu \tag{19}$$

The K shell electron (not free) is absorbed by the nucleus to convert the proton to a neutron

Sometimes nucleus decays to excited state and then emits a gamma ray (for any kind of radiation)



Electron capture and β^+ are competing processes

$$^{22}_{11}Na \rightarrow \epsilon^- + ^{22}_{10}Ne + \nu \tag{20}$$

$$^{22}_{11}Na \rightarrow ^{22}_{10}Ne + \beta^+ + \nu \tag{21}$$

$$Q = 2.842\ MeV \tag{22}$$

$Q > 1.022\ MeV$ positron favored

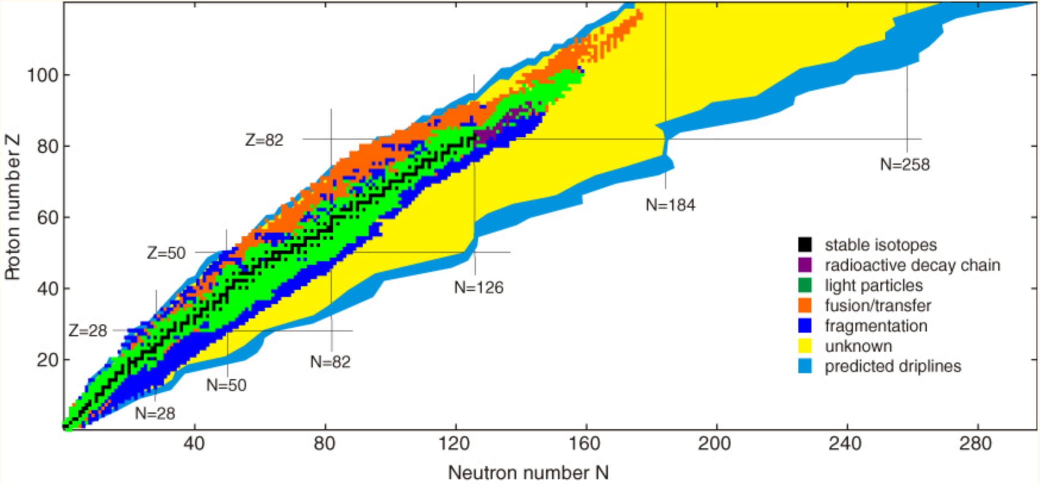
$Q < 1.022\ MeV$ electron capture favored

Why?

γ-ray emission is not a decay mode, but superfast de-excitation of the nucleus

Interactions of gamma rays with matter is a very important branch of nuclear physics

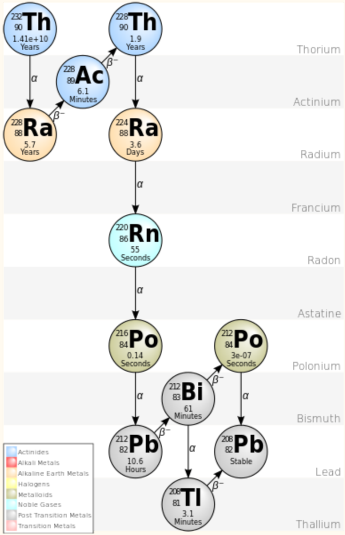
Chart of the nuclides



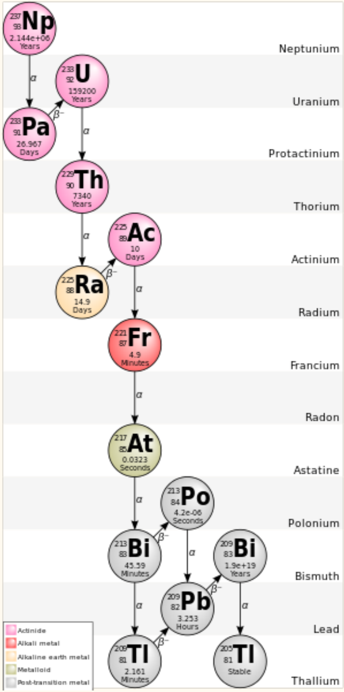
Proton number Z
Neutron number N
100
80
60
40
20
40
80
120
160
200
240
280
Z=82
Z=50
Z=28
N=28
N=50
N=82
N=126
N=184
N=258
stable isotopes
radioactive decay chain
light particles
fusion/transfer
fragmentation
unknown
predicted driplines

Ra 224 3.66 d α 5.6854 5.4486... γ 241..., C14 σ 12.0	**Ra 225** 14.8 d β^- 0.3, 0.4 γ 40 e^-	**Ra 226** 1600 a α 4.7843, 4.601... γ 186... C14 σ 12.8 σ_f < 5E-5
Fr 223 21.8 m β^- 1.1... α 5.34 γ 50, 80, 235...	**Fr 224** 3.3 m β^- 2.6, 2.8... γ 216, 132, 837 1341...	**Fr 225** 4.0 m β^- 1.6... γ 182, 32, 225 200...
Rn 222 3.825 d α 5.48948... γ (510) σ 0.74	**Rn 223** 23.2 m β^- γ 593, 417, 636 655...	**Rn 224** 1.78 h β^- γ 261, 266...

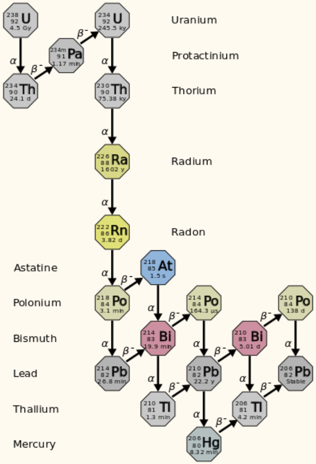
Decay chains

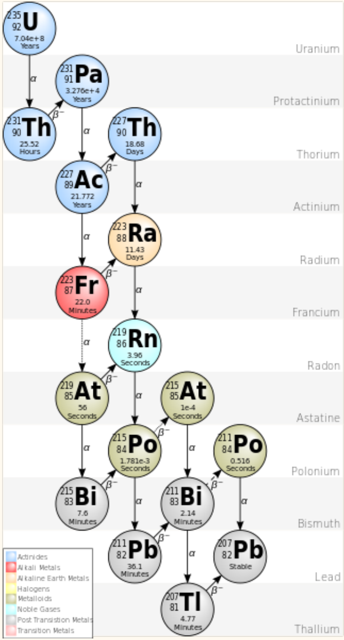


232 90 Th
1.41e+10 Years
228 90 Th
1.9 Years
228 89 Ac
6.1 Minutes
228 88 Ra
5.7 Years
224 88 Ra
3.6 Days
220 86 Rn
55 Seconds
216 84 Po
0.14 Seconds
212 84 Po
3e-07 Seconds
212 83 Bi
61 Minutes
212 82 Pb
10.6 Hours
208 82 Pb
Stable
208 81 Tl
3.1 Minutes
α
β−
Thorium
Actinium
Radium
Francium
Radon
Astatine
Polonium
Bismuth
Lead
Thallium
Actinides
Alkali Metals
Alkaline Earth Metals
Halogens
Metalloids
Noble Gases
Post Transition Metals
Transition Metals



237 93 Np 2.144e+06 Years
233 92 U 159200 Years
233 91 Pa 26.967 Days
229 90 Th 7340 Years
225 89 Ac 10 Days
225 88 Ra 14.9 Days
221 87 Fr 4.9 Minutes
217 85 At 0.0323 Seconds
213 84 Po 4.2e-06 Seconds
213 83 Bi 45.59 Minutes
209 83 Bi 1.9e+19 Years
209 82 Pb 3.253 Hours
209 81 Tl 2.161 Minutes
205 81 Tl Stable
α
β−
Neptunium
Uranium
Protactinium
Thorium
Actinium
Radium
Francium
Radon
Astatine
Polonium
Bismuth
Lead
Thallium
Actinide
Alkali metal
Alkaline earth metal
Metalloid
Post-transition metal


Uranium
Protactinium
Thorium
Radium
Radon
Astatine
Polonium
Bismuth
Lead
Thallium
Mercury
238 92 U 4.5 Gy
234 92 U 245.5 ky
234m 91 Pa 1.17 min
234 90 Th 24.1 d
230 90 Th 75.38 ky
226 88 Ra 1602 y
222 86 Rn 3.82 d
218 85 At 1.5 s
218 84 Po 3.1 min
214 84 Po 164.3 µs
210 84 Po 138 d
214 83 Bi 19.9 min
210 83 Bi 5.01 d
214 82 Pb 26.8 min
210 82 Pb 22.2 y
206 82 Pb Stable
210 81 Tl 1.3 min
206 81 Tl 4.2 min
206 80 Hg 8.32 min
α
β-



235 92 U 7.04e+8 Years
Uranium
231 91 Pa 3.276e+4 Years
Protactinium
231 90 Th 25.52 Hours
227 90 Th 18.68 Days
Thorium
227 89 Ac 21.772 Years
Actinium
223 88 Ra 11.43 Days
Radium
223 87 Fr 22.0 Minutes
Francium
219 86 Rn 3.96 Seconds
Radon
219 85 At 56 Seconds
215 85 At 1e-4 Seconds
Astatine
215 84 Po 1.781e-3 Seconds
211 84 Po 0.516 Seconds
Polonium
215 83 Bi 7.6 Minutes
211 83 Bi 2.14 Minutes
Bismuth
211 82 Pb 36.1 Minutes
207 82 Pb Stable
Lead
207 81 Tl 4.77 Minutes
Thallium
α
β⁻
Actinides
Alkali Metals
Alkaline Earth Metals
Halogens
Metalloids
Noble Gases
Post Transition Metals
Transition Metals

Half life

A radioactive isotope decays with a unique characteristic time

Decay is stochastic, characterized by Poisson distribution

So the decay of any one radionuclide cannot be predicted

The probability per time that a nucleus will decay is a constant

Decay is described by a Poisson process

Occurrences are randomly distributed in time

What is the 'occurrence' here?

Homogeneous, long term decay rate is constant

$$\mu \equiv \lambda \Delta t \tag{23}$$

$$P(n) = e^{-\mu} \cdot \frac{\mu^n}{n!} \tag{24}$$

What is the expected value and variance of the distribution?

A radioactive isotope decays with a unique characteristic time

$n(t)$ atoms at t have decayed in dt

$$\lambda n(t)dt \tag{25}$$

The number of atoms that decay on average in the interval $t, (t + dt)$ can be expressed as –

$$-dn(t) = \lambda n(t)dt \tag{26}$$

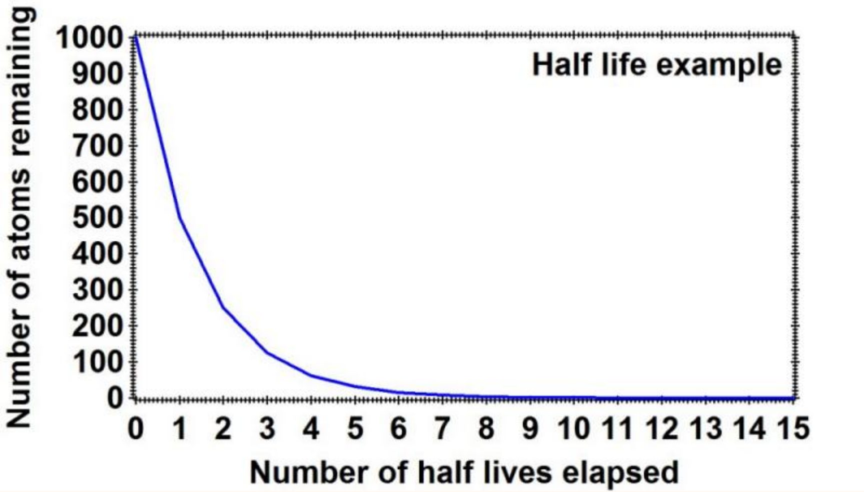
You should be able to solve this for $n(0) = n_0$

Radionuclides are characterized by half life

$$t_{\frac{1}{2}} = \frac{ln2}{\lambda} \tag{27}$$

Half life can be derived from the decay law

A radionuclide is ‘negligible’ after 10 half-lives have elapsed



Why would this be important?

How was half-life discovered?

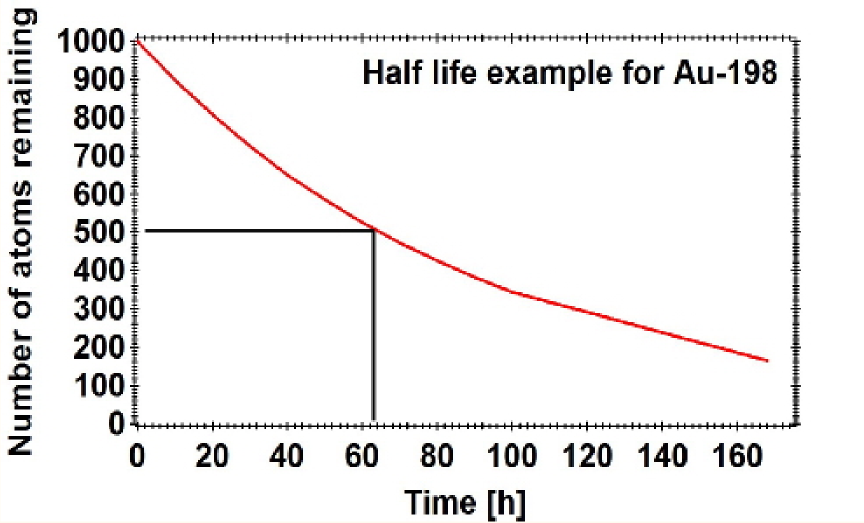
Rutherford discovered alpha and beta particle decay

In the course of, he noticed that there was a characteristic time

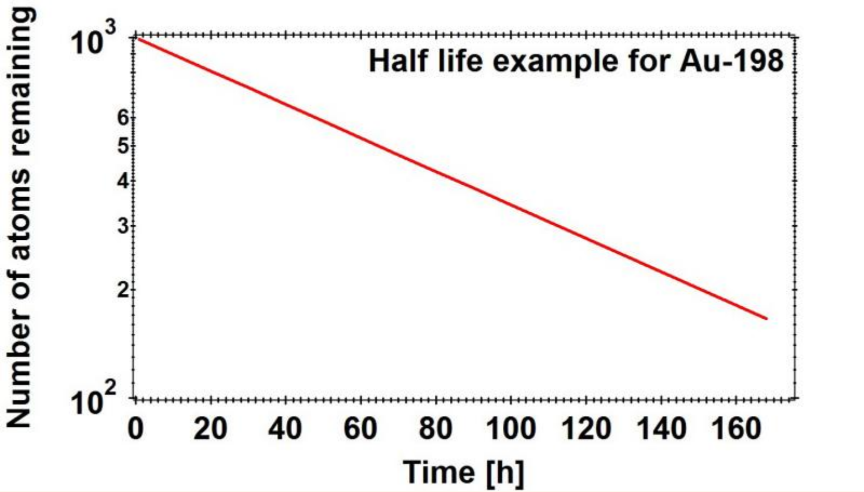
Basically observed decay time was the same for each atom

Experiments with thorium led to the discovery of radon

Then observing the radon gas led to the decay/half-life (65 s)


Half life example for Au-198
Number of atoms remaining
1000
900
800
700
600
500
400
300
200
100
0
0
20
40
60
80
100
120
140
160
Time [h]

Engineers like straight lines



Slope of the line is the decay constant


Pets on leash
only