

# Lab 1: Energy and Respiration

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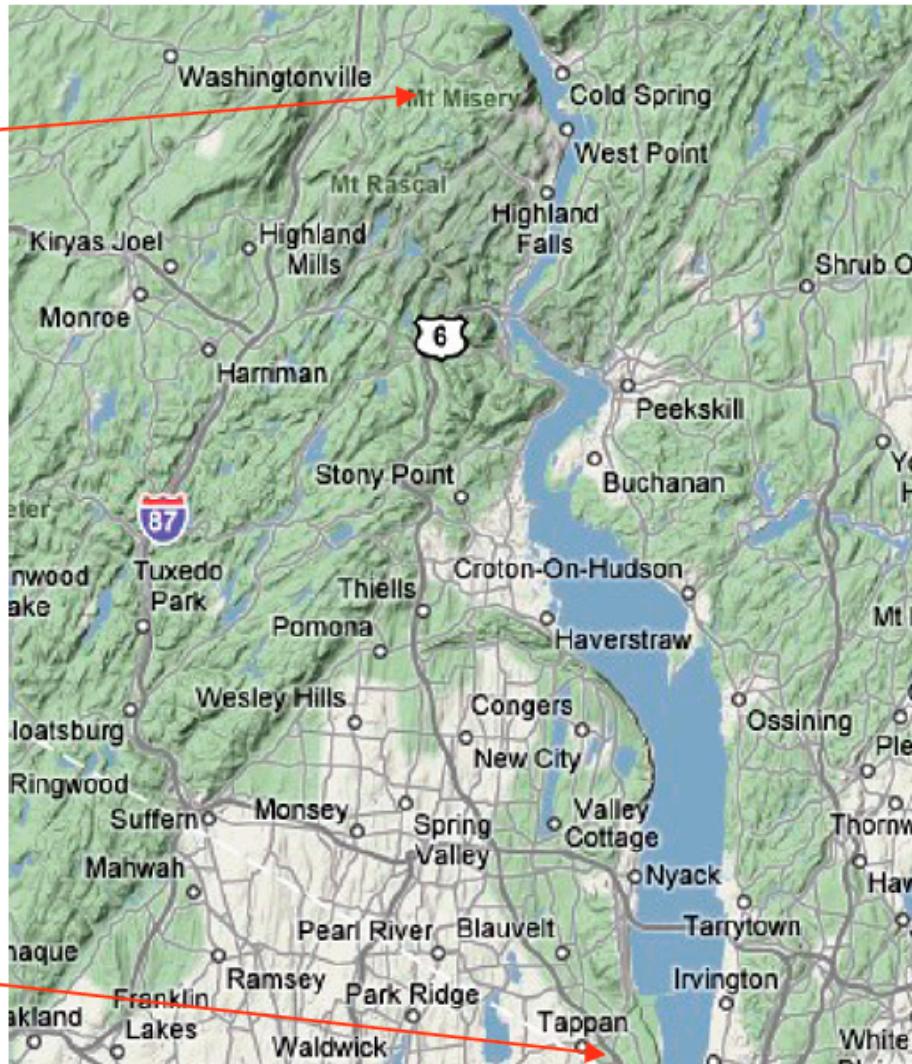
Wednesday 2:30-4:00  
554 Schermerhorn



- ❖ 9 labs (Wednesday 4-7 pm, 555 Schermerhorn)
- ❖ 3 Field trips (Fridays)
- ❖ Reports are due one week after the lab (by midnight the next Wednesday), and you'll get feedbacks one week after
- ❖ ALL lab reports are uploaded through CourseWorks
- ❖ Late reports will be deducted 10% automatically

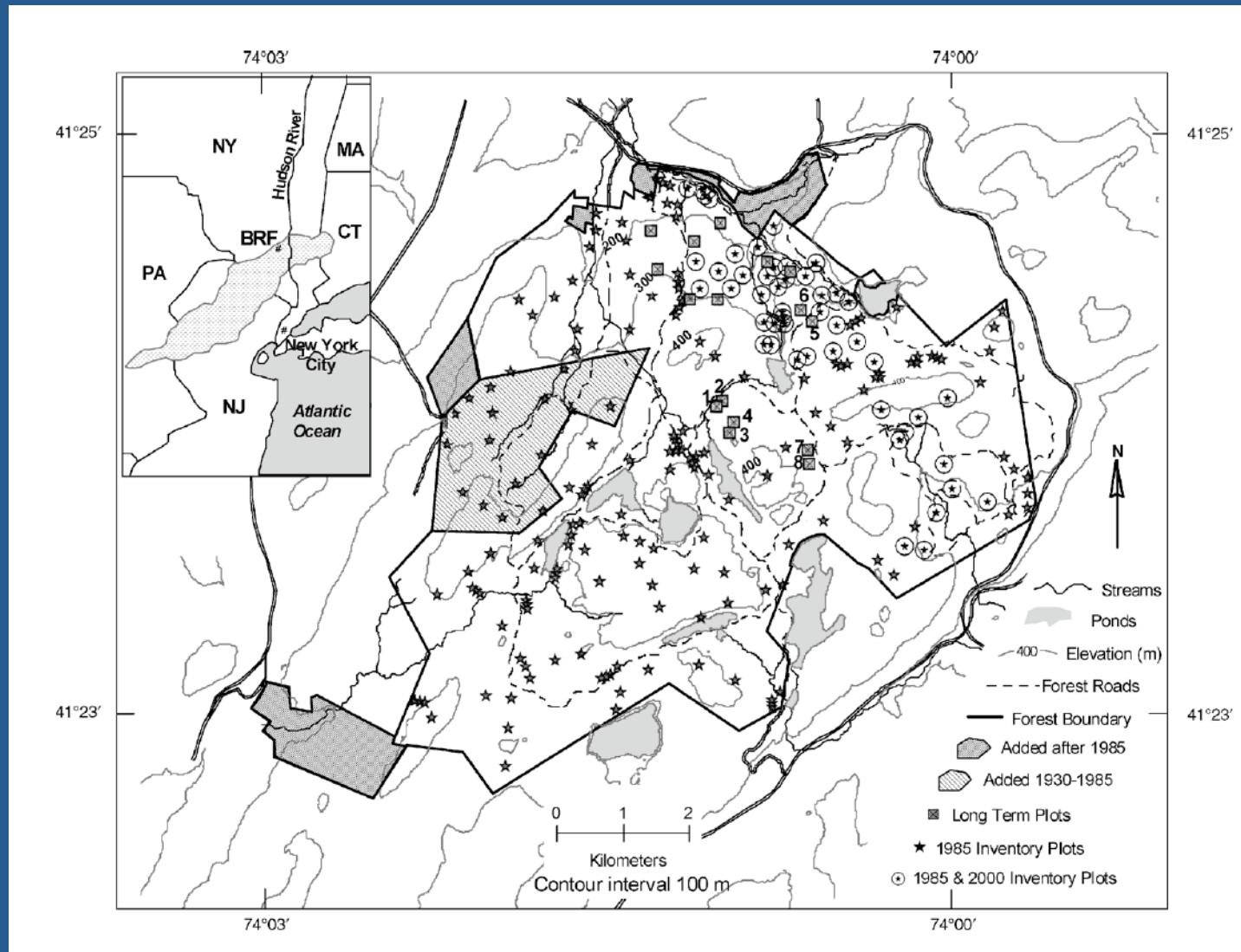


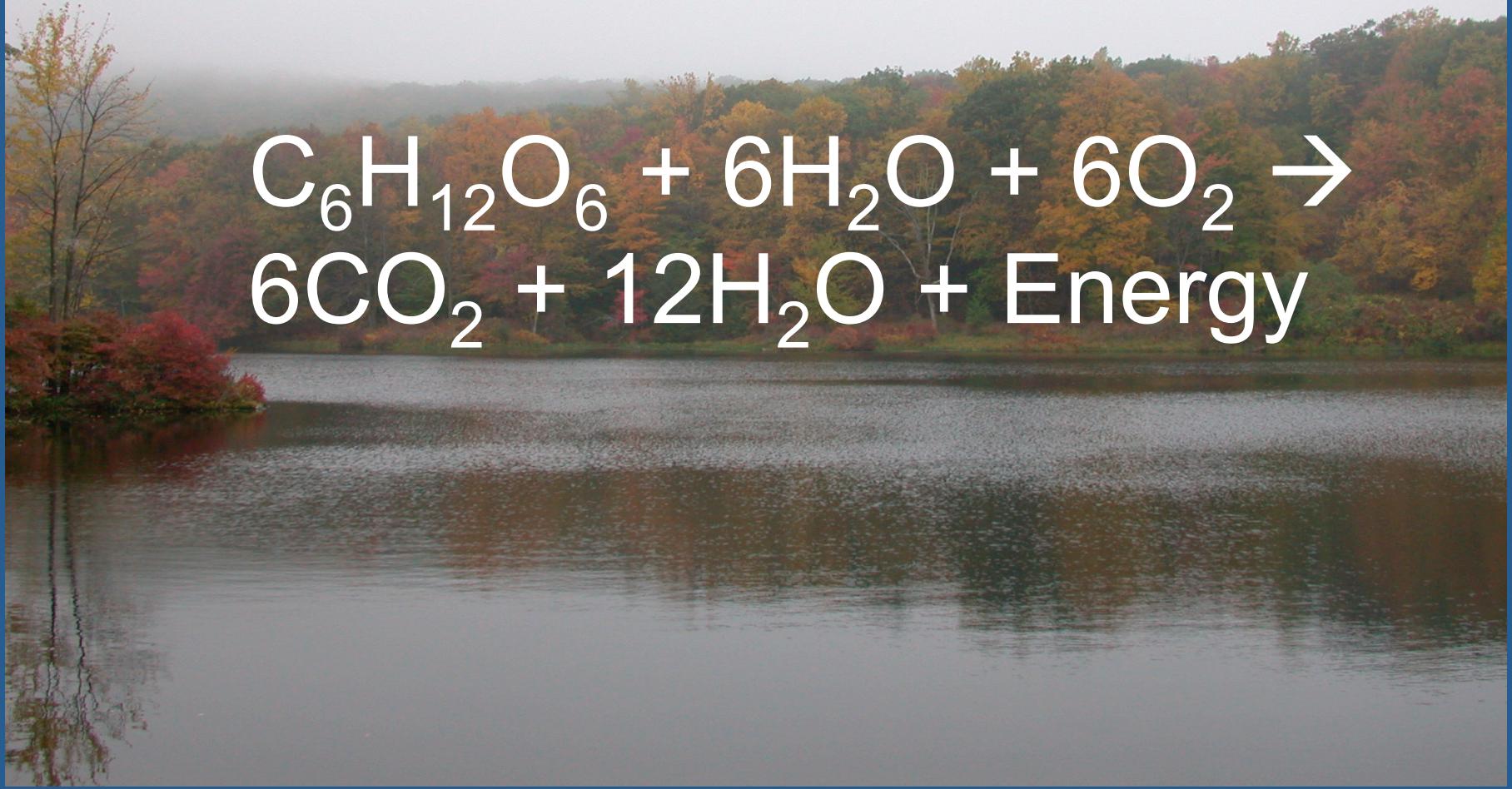
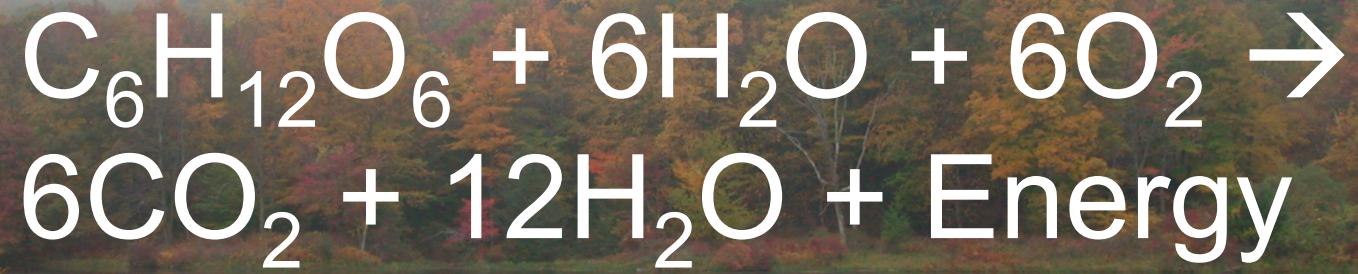
[www.blackrockforest.org](http://www.blackrockforest.org)



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# Map of Black Rock Forest

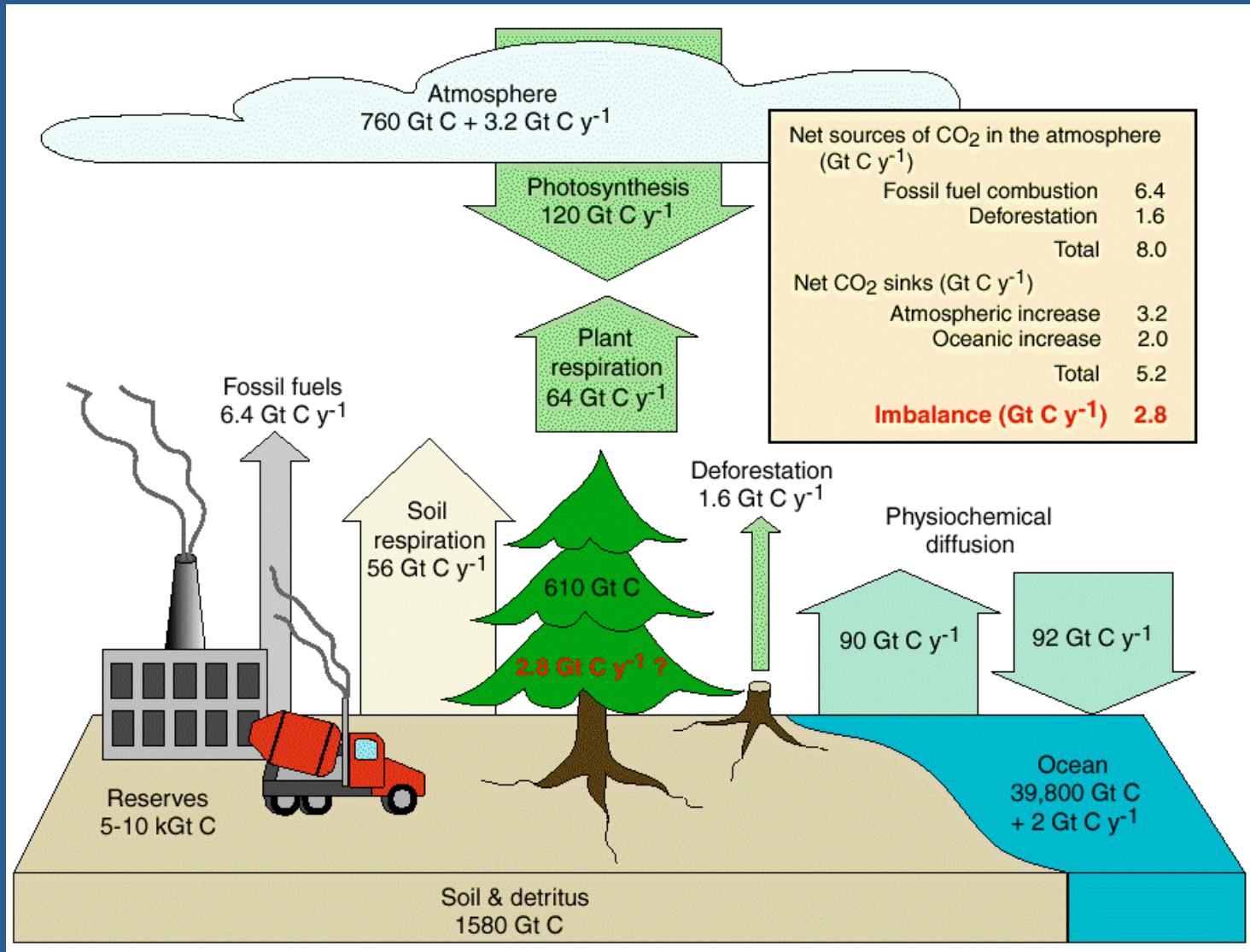




# Lab Goals

- Think quantitatively
- Think about (across) scale
- What we understand at one level informs our understanding at other levels
- Consider environmental influences
  - Variation
  - Global change
- Think independently and creatively

# Carbon Cycle



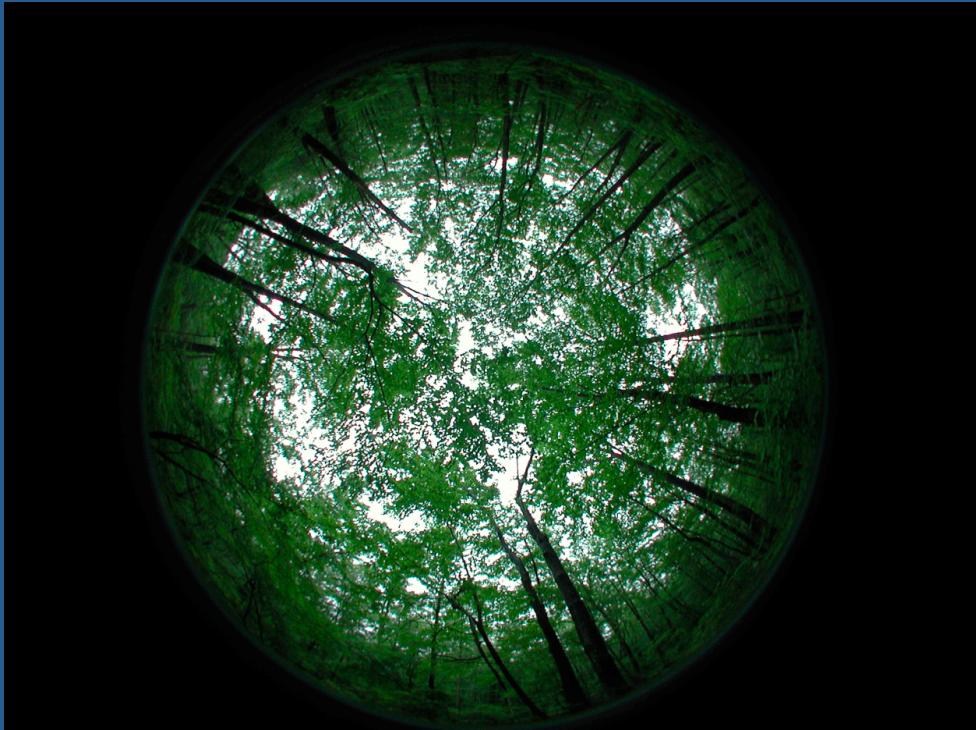
Respiration can be measured as an instantaneous  
How can we scale this in both space and time?



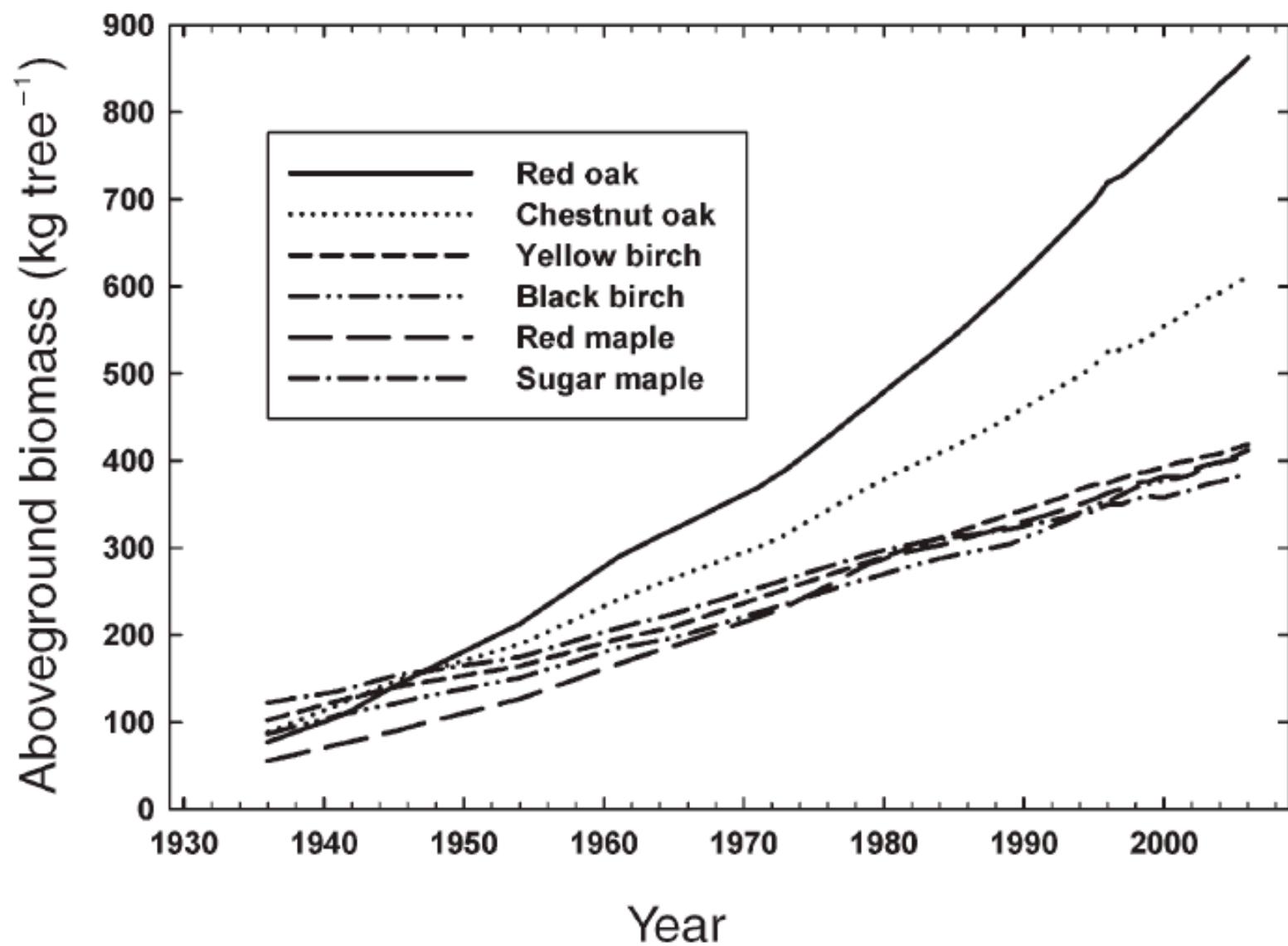


How can we estimate respiration  
from a forest?

# Leaf Area Index

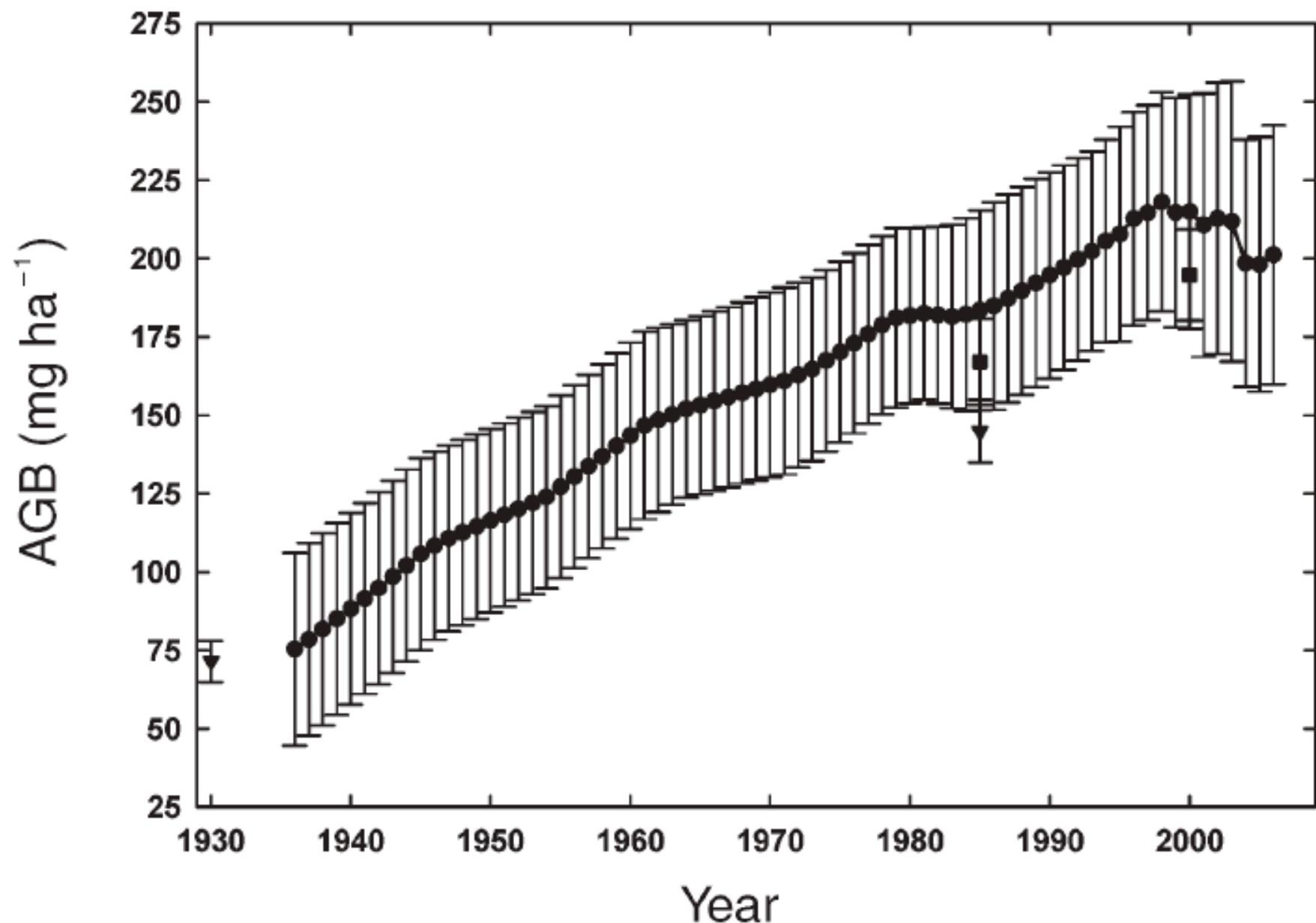


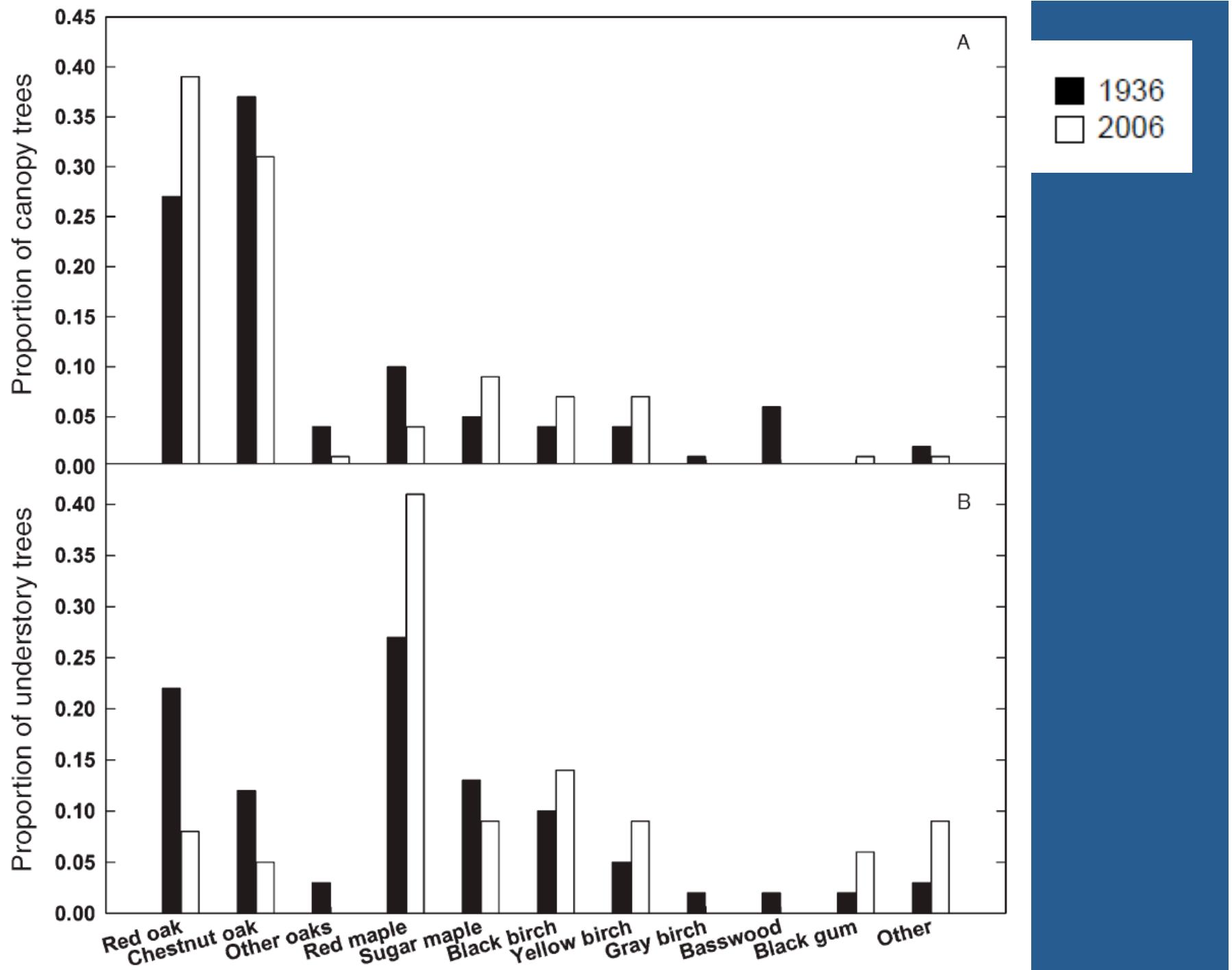
- Ratio of the total leaf surface area to the ground area its growing on
  - $\text{m}^2$  Leaf,  $\text{m}^{-2}$  ground
  - Estimated optically
  - Ranges from 1-15



## Additional Findings From Schuster *et al* 2008

- The forest canopy is dominated by red oak and chestnut oak, but the understory tree community has changed substantially from mixed oak–maple to red maple–black birch
- Over 76 years, red oak canopy trees stored carbon at about twice the rate of similar-sized canopy trees of other species
- Mean AGB on the long-term plots increased from 75 Mg ha<sup>-1</sup> in 1936 to 218 Mg ha<sup>-1</sup> in 1998
- There has been a significant loss of live tree biomass as a result of canopy tree mortality since 1999



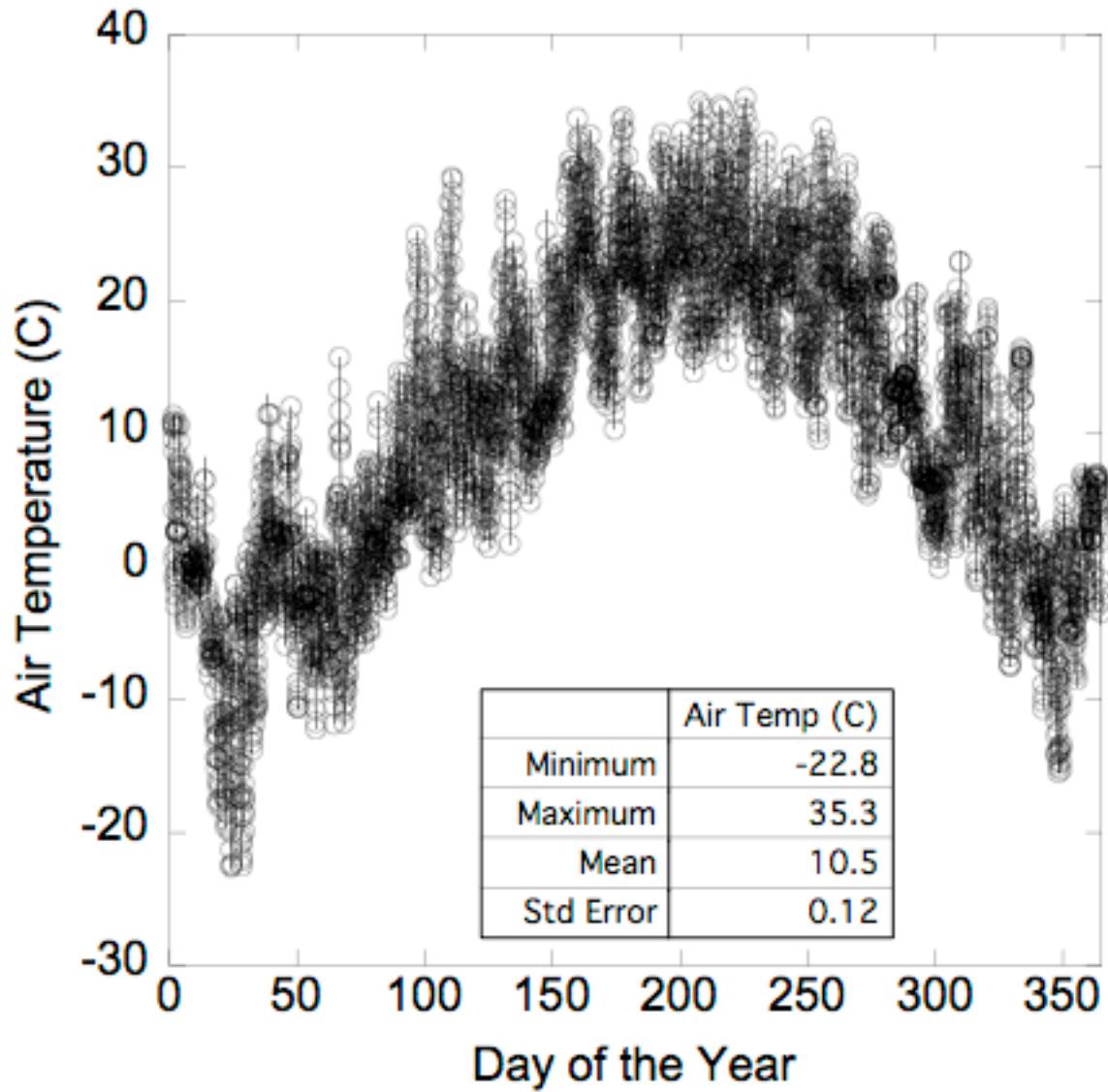


# Arrhenius function

$$R = R_0 e^{\frac{E_0}{R_g} \left( \frac{1}{T_0} - \frac{1}{T_a} \right)}$$

- $R$  = Respiration
- $R_g$  = *Ideal gas constant*
- $R_0$  = Base respiration, measured at the temperature  $T_0$
- $E_0$  = Energy of activation for respiration, measured at the temperature  $T_0$
- $T_0$  = Base temperature
- $T_a$  = Ambient temperature

## BRF Climate 2005



# Enzyme Kinetics

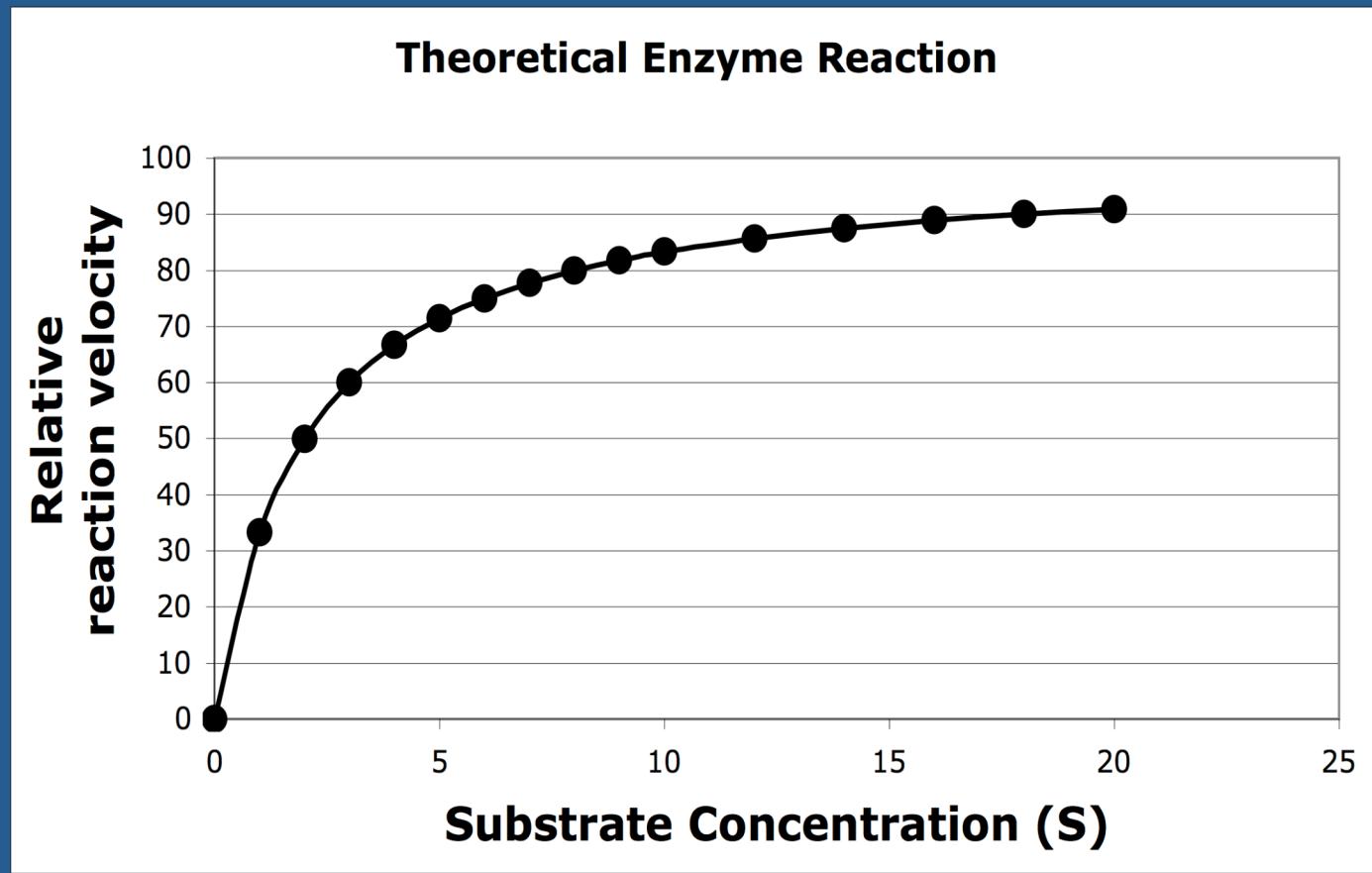
## Michaelis – Menten equation:

$$v = \frac{V_{\max} S}{K_m + S}$$

- Where:
- $v$  = velocity of the reaction
- $V_{\max}$  = Maximal velocity of a reaction,  
occurs at high substrate concentrations  
when the enzyme is saturated (empirical)

# Operational definition of $K_M$

- When  $[S] = K_M$ ,  $v = 1/2V_{max}$ , so that:
- $K_M$  is the substrate concentration at which the reaction velocity is half-maximal.



# Michaelis – Menten equation

- $K_M$  describes how curved the line is

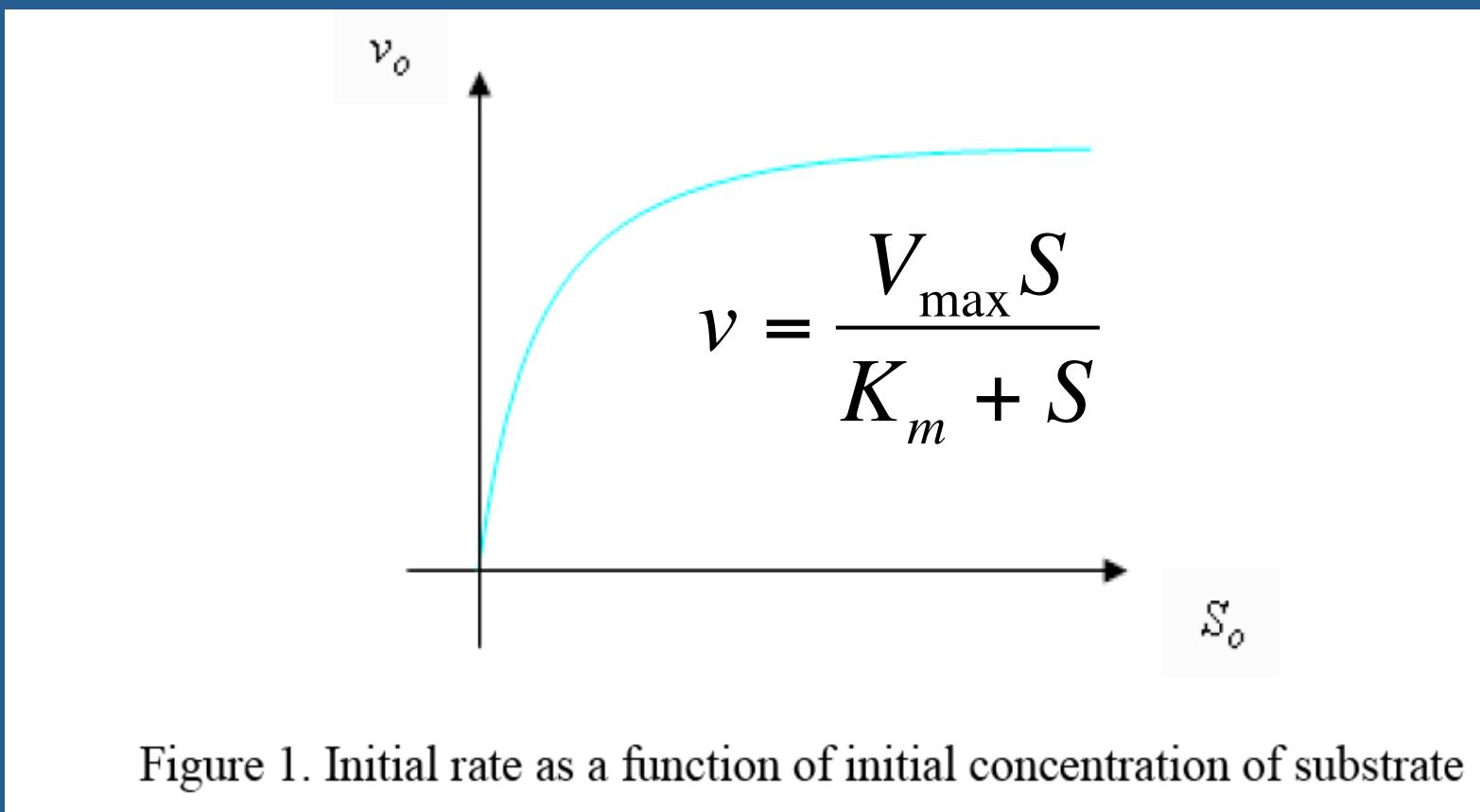
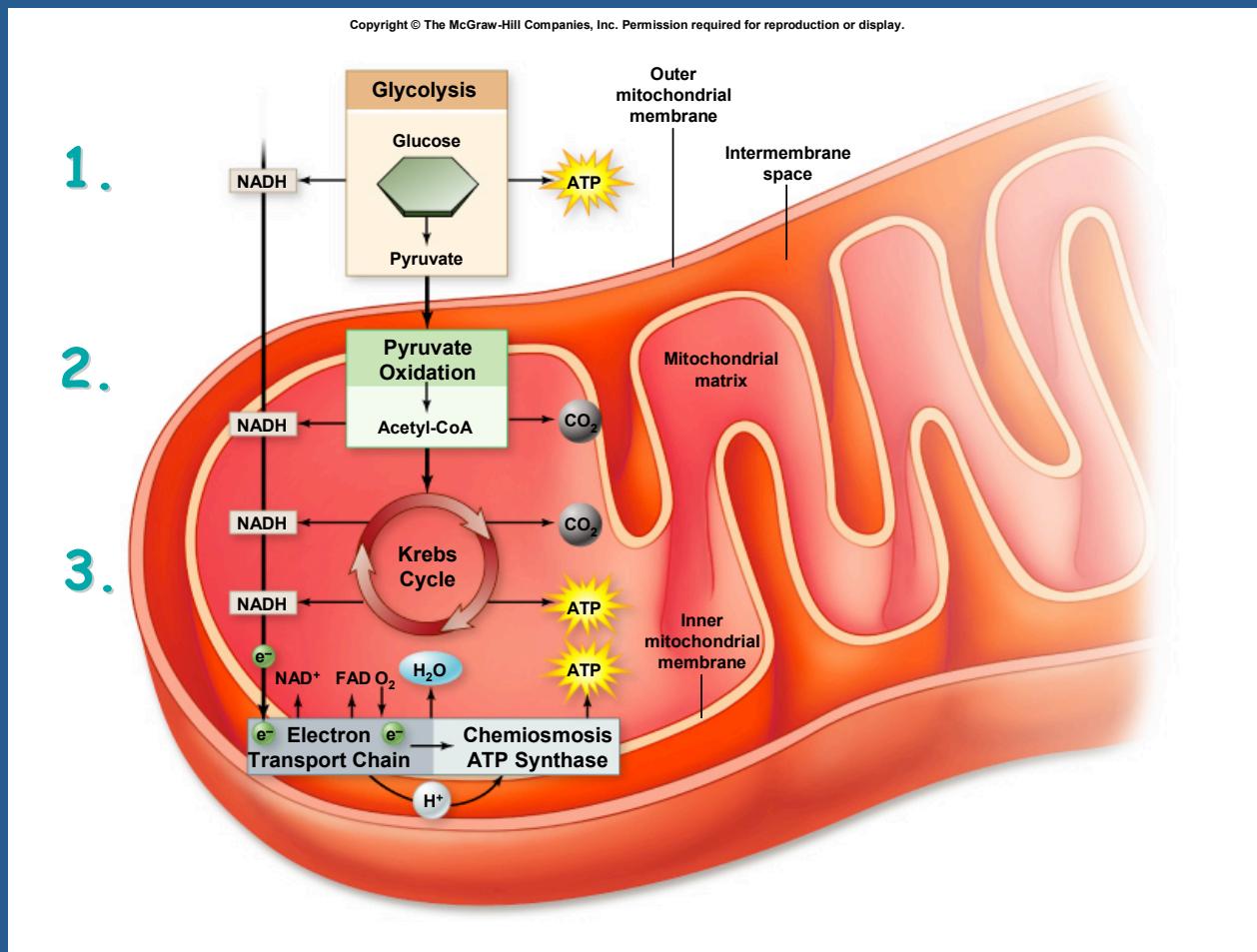


Figure 1. Initial rate as a function of initial concentration of substrate

# Factors Affecting Enzymatic Reactions

- Temperature
- pH
- Concentration of Enzyme
- Concentration of Substrate(s)
- Presence and Concentration of Inhibitor(s)
- Concentration cofactor(s)

# Respiration & Energy

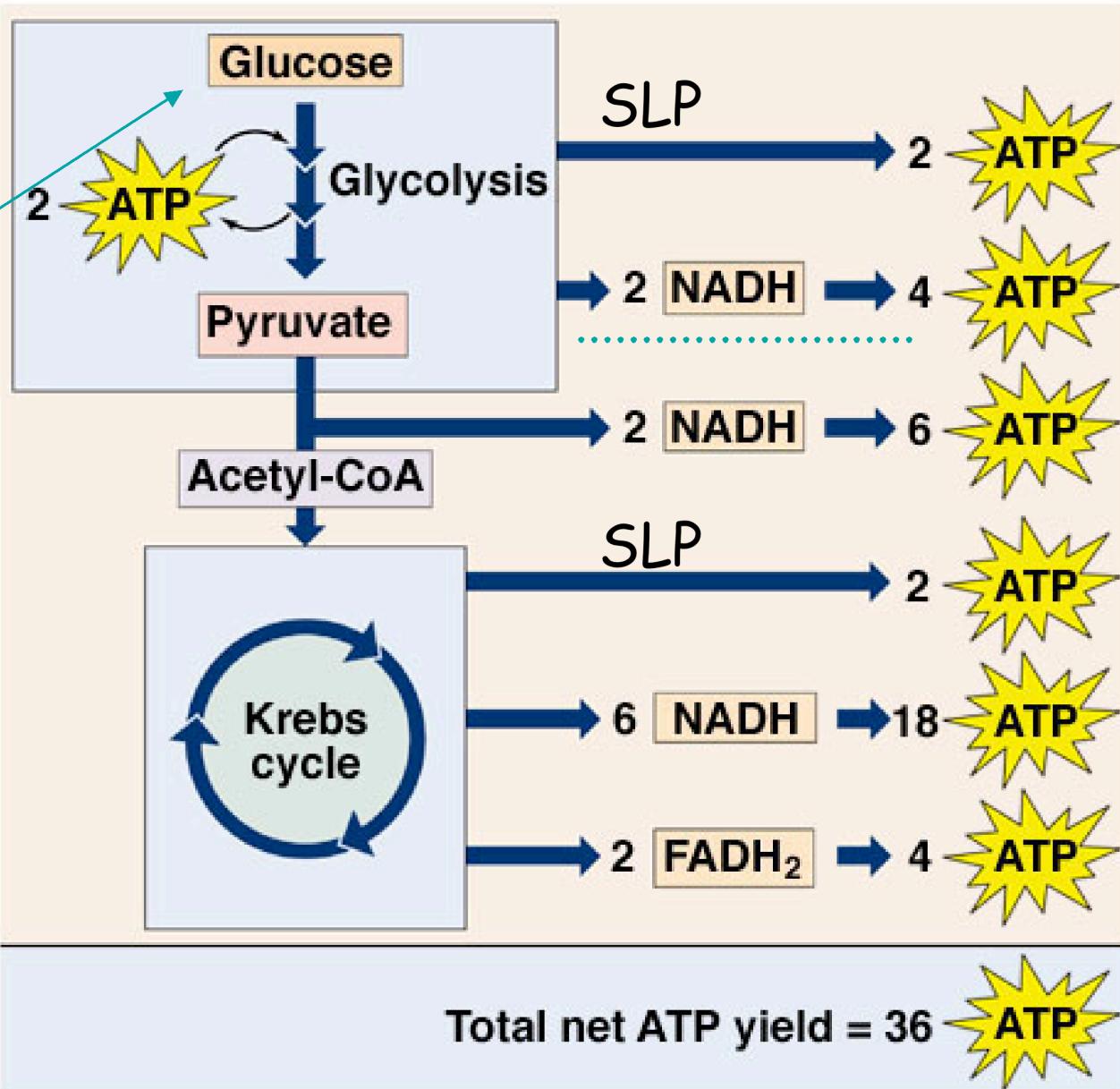


# ATP Theoretical Yield



NADH - 3 H<sup>+</sup>  
FADH<sub>2</sub> - 2 H<sup>+</sup>

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# How much energy is in a chocolate bar?



# Today's Lab

Exercise 1: Energy efficiency

Calculate energy in a candy bar

Exercise 2: Enzymes

Using the Michaelis-Menten equation you will evaluate how changing  $K_m$  and  $V_{max}$  influence the rate of a reaction ( you will produce 3 graphs)

Exercise 3: Forest Respiration and Climate Change

Using a modified Arrhenius function you will examine respiration as a function of temperature and tree species

Please complete a short survey  
for the Tree Respiration module  
you used in Part III

- <http://www.surveymonkey.com/s/HVDBTN2>