The table lists standard reduction potentials, $E^{\ominus}$ values, at 298.15 K (25 °C), and at a pressure of 101.325 kPa (1 atm) (not the standard pressure of 1 bar). The activity of all soluble species is assumed to be 1.000 mol L^{-1}. This is in particular important when *p*H (H^+ or OH^-) take part in the equilibrium. The reliability of the potentials is not the same for all the data. Typically, the values with fewer significant figures have lower reliability. The values of reduction potentials, in particular those of less common reactions, are not definite; they are subject to occasional revisions.

Abbreviations: ac = acetate; bipy = 2,2'-dipyridine, or bipyridine; en = ethylenediamine; phen = 1,10-phenanthroline.

Reference: Rumble, J. *CRC Handbook of Chemistry and Physics*, 98th Edition, CRC Press LLC, 2017.

Reduction half-reaction	$E^{\ominus}$ / V
$Ag^+ + e^- \rightleftharpoons Ag$	+0.7996
$Ag^{2+} + e^- \rightleftharpoons Ag^+$	+1.980
$AgBr + e^- \rightleftharpoons Ag + Br^-$	+0.071 33
$AgCl + e^- \rightleftharpoons Ag + Cl^-$	+0.222 33
$Ag_2CrO_4 + 2\,e^- \rightleftharpoons 2\,Ag + {CrO_4}^{2-}$	+0.4470
$AgF + e^- \rightleftharpoons Ag + F^-$	+0.779
$AgI + e^- \rightleftharpoons Ag + I^-$	−0.152 24
$Al^{3+} + 3\,e^- \rightleftharpoons Al$	−1.676
$H_3AsO_4 + 2\,H^+ + 2\,e^- \rightleftharpoons HAsO_2 + 2\,H_2O$	+0.560
$Au^+ + e^- \rightleftharpoons Au$	+1.692
$Au^{3+} + 2\,e^- \rightleftharpoons Au^+$	+1.401
$Au^{3+} + 3\,e^- \rightleftharpoons Au$	+1.498
$Ba^{2+} + 2\,e^- \rightleftharpoons Ba$	−2.912
$Be^{2+} + 2\,e^- \rightleftharpoons Be$	−1.847
$Bi^{3+} + 3\,e^- \rightleftharpoons Bi$	+0.308
$Bi_2O_3 + 3\,H_2O + 6\,e^- \rightleftharpoons 2\,Bi + 6\,OH^-$	−0.46
$BiO^+ + 2\,H^+ + 3\,e^- \rightleftharpoons Bi + H_2O$	+0.320
$Br_2 + 2\,e^- \rightleftharpoons 2\,Br^-$	+1.0873
$BrO^- + H_2O + 2\,e^- \rightleftharpoons Br^- + 2\,OH^-$	+0.761
$2\,HBrO + 2\,H^+ + 2\,e^- \rightleftharpoons Br_2 + 2\,H_2O$	+1.574
$2\,{BrO_3}^- + 12\,H^+ + 10\,e^- \rightleftharpoons Br_2 + 6\,H_2O$	+1.482
$CO_2 + 2\,H^+ + 2\,e^- \rightleftharpoons HCOOH$	−0.199
$Ca^{2+} + 2\,e^- \rightleftharpoons Ca$	−2.868
$Cd^{2+} + 2\,e^- \rightleftharpoons Cd$	−0.4030
$Cd(OH)_2 + 2\,e^- \rightleftharpoons Cd + 2\,OH^-$	−0.809
$Ce^{3+} + 3\,e^- \rightleftharpoons Ce$	−2.336
$Ce^{4+} + e^- \rightleftharpoons Ce^{3+}$	+1.72
$Cl_2 + 2\,e^- \rightleftharpoons 2\,Cl^-$	+1.358 27
$ClO^- + H_2O + 2\,e^- \rightleftharpoons Cl^- + 2\,OH^-$	+0.81
${ClO_4}^- + 2\,H^+ + 2\,e^- \rightleftharpoons {ClO_3}^- + H_2O$	+1.189
${ClO_4}^- + H_2O + 2\,e^- \rightleftharpoons {ClO_3}^- + 2\,OH^-$	+0.36
$2\,HClO + 2\,H^+ + 2\,e^- \rightleftharpoons Cl_2 + 2\,H_2O$	+1.611
$2\,{ClO_3}^- + 12\,H^+ + 10\,e^- \rightleftharpoons Cl_2 + 6\,H_2O$	1.47
$Co^{2+} + 2\,e^- \rightleftharpoons Co$	−0.28
$Co^{3+} + e^- \rightleftharpoons Co^{2+}$	+1.92
$Cr^{2+} + 2\,e^- \rightleftharpoons Cr$	−0.913
$Cr^{3+} + e^- \rightleftharpoons Cr^{2+}$	−0.407
$Cr^{3+} + 3\,e^- \rightleftharpoons Cr$	−0.744
${Cr_2O_7}^{2-} + 14\,H^+ + 6\,e^- \rightleftharpoons 2\,Cr^{3+} + 7\,H_2O$	+1.36
${CrO_4}^{2-} + 4\,H_2O + 3\,e^- \rightleftharpoons Cr(OH)_3 + 5\,OH^-$	−0.13
$Cs^+ + e^- \rightleftharpoons Cs$	−3.026
$Cu^+ + e^- \rightleftharpoons Cu$	+0.521
$Cu^{2+} + e^- \rightleftharpoons Cu^+$	+0.153
$Cu^{2+} + 2\,e^- \rightleftharpoons Cu$	+0.3419
$Cu(OH)_2 + 2\,e^- \rightleftharpoons Cu + 2\,OH^-$	−0.222
$F_2 + 2\,e^- \rightleftharpoons 2\,F^-$	+2.866
$Fe^{2+} + 2\,e^- \rightleftharpoons Fe$	−0.447
$Fe^{3+} + 3\,e^- \rightleftharpoons Fe$	−0.037
$Fe^{3+} + e^- \rightleftharpoons Fe^{2+}$	+0.771
$[Fe(CN)_6]^{3-} + e^- \rightleftharpoons [Fe(CN)_6]^{4-}$	+0.358
$[Fe(bipy)_3]^{3+} + e^- \rightleftharpoons [Fe(bipy)_3]^{2+}$	+1.03
$Fe(OH)_3 + e^- \rightleftharpoons Fe(OH)_2 + OH^-$	−0.56
$[Fe(phen)_3]^{3+} + e^- \rightleftharpoons [Fe(phen)_3]^{2+}$	+1.147
$Ga^{3+} + e^- \rightleftharpoons Ga$	−0.549
$Ga^+ + e^- \rightleftharpoons Ga$	−0.2
$2\,H^+ + 2\,e^- \rightleftharpoons H_2$	0
$2\,H_2O + 2\,e^- \rightleftharpoons H_2 + 2\,OH^-$	−0.8277
${HO_2}^- + H_2O + 2\,e^- \rightleftharpoons 3\,OH^-$	+0.878
$H_2O_2 + 2\,H^+ + 2\,e^- \rightleftharpoons 2\,H_2O$	+1.776
${Hg_2}^{2+} + 2\,e^- \rightleftharpoons 2\,Hg$	+0.7973
$Hg^{2+} + 2\,e^- \rightleftharpoons Hg$	+0.851
$2\,Hg^{2+} + 2\,e^- \rightleftharpoons {Hg_2}^{2+}$	+0.920
$Hg_2Cl_2 + 2\,e^- \rightleftharpoons 2\,Hg + 2\,Cl^-$	+0.268 08
$I_2 + 2\,e^- \rightleftharpoons 2\,I^-$	+0.5355
${I_3}^- + 2\,e^- \rightleftharpoons 3\,I^-$	+0.536
${IO_3}^- + 12\,H^+ + 10\,e^- \rightleftharpoons I_2 + 6\,H_2O$	+1.195
$In^+ + e^- \rightleftharpoons In$	−0.14
$In^{2+} + e^- \rightleftharpoons In^+$	−0.40
$In^{3+} + e^- \rightleftharpoons In^{2+}$	−0.49
$In^{3+} + 2\,e^- \rightleftharpoons In^+$	−0.443
$In^{3+} + 3\,e^- \rightleftharpoons In$	−0.3382
$K^+ + e^- \rightleftharpoons K$	−2.931
$La^{3+} + 3\,e^- \rightleftharpoons La$	−2.379
$Li^+ + e^- \rightleftharpoons Li$	−3.0401
$Mg^{2+} + 2\,e^- \rightleftharpoons Mg$	−2.372
$Mn^{2+} + 2\,e^- \rightleftharpoons Mn$	−1.185
$Mn^{3+} + e^- \rightleftharpoons Mn^{2+}$	+1.5415
$MnO_2 + 4\,H^+ + 2\,e^- \rightleftharpoons Mn^{2+} + 2\,H_2O$	+1.224
${MnO_4}^- + e^- \rightleftharpoons {MnO_4}^{2-}$	+0.558
${MnO_4}^- + 4\,H^+ + 3\,e^- \rightleftharpoons MnO_2 + 2\,H_2O$	+1.679
${MnO_4}^- + 8\,H^+ + 5\,e^- \rightleftharpoons Mn^{2+} + 4\,H_2O$	+1.507
${MnO_4}^- + 2\,H_2O + 3\,e^- \rightleftharpoons MnO_2 + 4\,OH^-$	+0.595
$HNO_2 + H^+ + e^- \rightleftharpoons NO + H_2O$	+0.983
$2\,{NO_3}^- + 4\,H^+ + 2\,e^- \rightleftharpoons N_2O_4 + 2\,H_2O$	+0.803
${NO_3}^- + 4\,H^+ + 3\,e^- \rightleftharpoons NO + 2\,H_2O$	+0.957
${NO_3}^- + H_2O + 2\,e^- \rightleftharpoons {NO_2}^- + 2\,OH^-$	+0.01
$Na^+ + e^- \rightleftharpoons Na$	−2.71
$Ni^{2+} + 2\,e^- \rightleftharpoons Ni$	−0.257
$O_2 + 4\,H^+ + 4\,e^- \rightleftharpoons 2\,H_2O$	+1.229
$O_2 + H_2O + 2\,e^- \rightleftharpoons {HO_2}^- + OH^-$	−0.076
$O_2 + 2\,H_2O + 4\,e^- \rightleftharpoons 4\,OH^-$	+0.401
$O_2 + 2\,H^+ + 2\,e^- \rightleftharpoons H_2O_2$	+0.695
$O_3 + 2\,H^+ + 2\,e^- \rightleftharpoons O_2 + H_2O$	+2.076
$O_3 + H_2O + 2\,e^- \rightleftharpoons O_2 + 2\,OH^-$	+1.24
$Pb^{2+} + 2\,e^- \rightleftharpoons Pb$	−0.1262
$PbO_2 + {SO_4}^{2-} + 4\,H^+ + 2\,e^- \rightleftharpoons PbSO_4 + 2\,H_2O$	+1.6913
$PbSO_4 + 2\,e^- \rightleftharpoons Pb + {SO_4}^{2-}$	−0.3588
$Pt^{2+} + 2\,e^- \rightleftharpoons Pt$	+1.18
$[PtCl_4]^{2-} + 2\,e^- \rightleftharpoons Pt + 4\,Cl^-$	+0.755
$Pu^{4+} + e^- \rightleftharpoons Pu^{3+}$	+1.006
$Ra^{2+} + 2\,e^- \rightleftharpoons Ra$	−2.8
$Rb^+ + e^- \rightleftharpoons Rb$	−2.98
$S + 2\,e^- \rightleftharpoons S^{2-}$	−0.476 27
$S + 2\,H^+ + 2\,e^- \rightleftharpoons H_2S$	+0.142
${H_2SO_3} + 4\,H^+ + 4\,e^- \rightleftharpoons S + 3\,H_2O$	+0.45
${SO_4}^{2-} + 4\,H^+ + 2\,e^- \rightleftharpoons H_2SO_3 + H_2O$	+0.172
${S_2O_8}^{2-} + 2\,e^- \rightleftharpoons 2\,{SO_4}^{2-}$	+2.010
$Se + 2\,e^- \rightleftharpoons Se^{2-}$	−0.670
$Sn^{2+} + 2\,e^- \rightleftharpoons Sn$	−0.1375
$Sn^{4+} + 2\,e^- \rightleftharpoons Sn^{2+}$	+0.151
$Sr^{2+} + 2\,e^- \rightleftharpoons Sr$	−2.899
$Te + 2\,e^- \rightleftharpoons Te^{2-}$	−1.143
$Ti^{2+} + 2\,e^- \rightleftharpoons Ti$	−1.628
$Ti^{3+} + e^- \rightleftharpoons Ti^{2+}$	−0.369
$Tl^+ + e^- \rightleftharpoons Tl$	−0.336
$U^{3+} + 3\,e^- \rightleftharpoons U$	−1.66
$U^{4+} + e^- \rightleftharpoons U^{3+}$	−0.52
$V^{2+} + 2\,e^- \rightleftharpoons V$	−1.175
$V^{3+} + e^- \rightleftharpoons V^{2+}$	−0.255
${VO_2}^+ + 2\,H^+ + e^- \rightleftharpoons VO^{2+} + H_2O$	+0.991
$H_4XeO_6 + 2\,H^+ + 2\,e^- \rightleftharpoons XeO_3 + 3\,H_2O$	+2.42
$Zn^{2+} + 2\,e^- \rightleftharpoons Zn$	−0.7618