

1. State whether each of the following substances is expected to form an IONIC or MOLECULAR solution.

(a) RbBr(s)

(b) $\text{CHCl}_3(\text{l})$

(c) $\text{CSNO}_3(\text{s})$

(d) $\text{CuSO}_4(5)$

(e) $\text{S}_8(\text{s})$

(f) $\text{CrCl}_3(\text{s})$

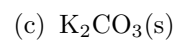
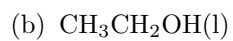
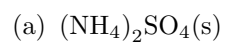
(g) $\text{NaCH}_3\text{COO}(\text{s})$

(h) $\text{ICl}(\text{s})$

(i) $\text{HNO}_3(\text{l})$

(j) $\text{CH}_4(\text{g})$

2. Write equations to show the dissolving of the following substances in water.



3. Write the equation for the equilibrium reaction existing in each of the following saturated aqueous solutions.

(a) K_3PO_4

(b) NH_4Cl

(c) $\text{Al}(\text{NO}_3)_3$

4. Write the crystallization reaction involving $\text{MgBr}_2(\text{s})$.

5. Write the dissolving reaction involving $\text{C}_6\text{H}_{12}\text{O}_6(\text{s})$.

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