25. For each of the following combinations of equal volumes of 0.2 M aqueous solutions,

i) identify the possible products by formula.
ii) state which (if any) product has a low solubility.
iii) if there IS a precipitate write the formula equation, total ionic equation, and net ionic equation for the reaction.

(a) $MgS + Sr(OH)_2$

(b) $CuBr_2 + Pb(NO_3)_2$

(c) $FeBr_3 + SrI_2$

(d) $Ba(NO_3)_2 + Li_2SO_4$

(e) $K_3PO_4 + CuCl_2$

(f) $(NH_4)_2SO_3 + Al_2(SO_4)_3$

(g) silver nitrate and sodium phosphate

(h) zinc sulphate and iron (II) chloride

(i) cobalt (II) sulphate and lithium carbonate

(j) iron (III) nitrate and magnesium sulphide

(k) beryllium sulphate and ammonium carbonate

(l) magnesium sulphate and strontium hydroxide