

21. Determine the solubility of the COMPOUNDS IN WATER". following ionic compounds from the table
"SOLUBILITY OF COMMON

(a) AgCl

(b) NaOH

(c) FeS

(d) FeSO_4

(e) $\text{Al}_2(\text{CO}_3)_3$

(f) $\text{Fe}(\text{NO}_3)_3$

(g) Na_3PO_4

(h) CuI

(i) CuCl_2

(j) PbBr_2

22. Determine whether 0.2 M solutions of the following mixtures form a precipitate when mixed, and give the formula of any precipitate formed.

(a) AgNO_3 and NH_4Br

(b) SrBr_2 and NaNO_3

(c) KOH and AlCl_3

(d) NaI and $\text{Pb}(\text{NO}_3)_2$

(e) BaS and Na_2SO_4

(f) CaS and NH_4Cl

23. Look at the two "Facts" on the first page of this unit. Suggest answers to the questions asked.

24. You wish to make some precipitates. You will make up 0.2 M solutions of certain soluble salts, mix the solutions and filter off the resulting precipitates. Give the complete chemical formulae for the soluble salts you would select to make up the necessary solutions for the following precipitates.

(a) PbCl_2

(b) AgBr

(c) Cr_2S_3

(d) SrSO_4