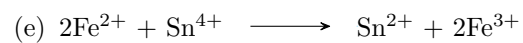
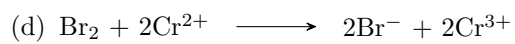
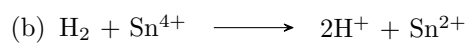
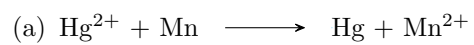


1. In the following reactions, indicate the:
i) species oxidized ii) species reduced iii) oxidizing agent iv) reducing agent.



2. When cesium metal is exposed to chlorine gas, a bright flash occurs as the elements react. The product, cesium chloride, is a white solid composed of cesium ions and chloride ions.
- (a) Write the balanced overall reaction which occurs between chlorine and cesium.
- (b) Write the half-reactions which occur and identify which half-reaction is the reduction and which is the oxidation.
- (c) Identify the reducing and oxidizing agents.