

PROPULSION

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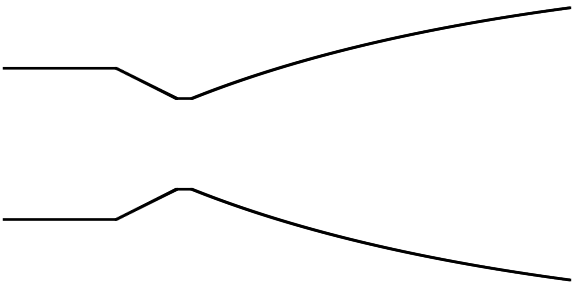
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ABSTRACT. Everything about Propulsion, with a focus on rocket propulsion.
I also look at (rocket) propulsion for engineers from a (theoretical and mathematical) physicists’ point of view. I would like to seek more cross-pollination between physicists and mathematicians and engineers in thermodynamics and fluid mechanics.



Part 1. Notes and Solutions for *Rocket Propulsion Elements* by George Sutton and Oscar Biblarz

1. DEFINITIONS AND FUNDAMENTALS

Ch. 2 Definitions and Fundamentals of Biblarz and Sutton (2001) [2]

1.1. **Definitions.** Section 2.1. “Definitions”, Ch. 2 Definitions and Fundamentals of Biblarz and Sutton (2001) [2]

total impulse $I_s = \int_0^t F_{\text{thrust}} dt$

$t \equiv$ burning time $\equiv t_p$

specific impulse I_s , total impulse per unit weight of propellant

$$I_s = \frac{\int_0^t F_{\text{thrust}} dt}{g_0 \int \dot{m} dt} = \frac{I_t}{m_p g_0}$$

$$I_s = \frac{\dot{m} u_e t_p}{\dot{m} g t_p} = \frac{u_e}{g}$$

If $F_{\text{thrust}} = \dot{m} u_e$, and constant propellant mass flow,

2. NOZZLE THEORY AND THERMODYNAMIC RELATIONS

2.1. **Isentropic Flow Through Nozzles.** Subsection 3.3 “Isentropic Flow Through Nozzles” of Biblarz and Sutton (2001) [2]

The Bernoulli invariant for compressible flow ¹ gives us this:

$$h_1 + \frac{1}{2}v_1^2 = h_2 + \frac{1}{2}v_2^2 \text{ or } v_2 = \sqrt{2(h_1 - h_2) + v_1^2}$$
$$\implies h_1 - h_2 = \frac{C_p}{MN}(\tau_1 - \tau_2) = \frac{\gamma}{\gamma - 1} \frac{\tau_1}{M} \left(1 - \frac{\tau_2}{\tau_1}\right) = \frac{\gamma}{\gamma - 1} \frac{\tau_1}{M} \left(1 - \left(\frac{p_2}{p_1}\right)^{\frac{\gamma-1}{\gamma}}\right) = \frac{\gamma RT_1}{\gamma - 1} \left(1 - \left(\frac{p_2}{p_1}\right)^{\frac{\gamma-1}{\gamma}}\right)$$

so that

$$v_2 = \sqrt{\frac{2\gamma RT_1}{\gamma - 1} \left(1 - \left(\frac{p_2}{p_1}\right)^{\frac{\gamma-1}{\gamma}}\right) + v_1^2}$$

When chamber section is large compared to nozzle throat section, chamber velocity or nozzle approach velocity comparatively small and v_1^2 can be neglected.[2]

Chamber temperature T_1 at nozzle inlet; under isentropic conditions, T_1 differs little from stagnation temperature or (for chemical rocket) combustion temperature.[2]

Ch. 4 Flight Performance of Biblarz and Sutton (2001) [2] should probably be read before Ch. 3 Nozzle Theory and Thermodynamic Relations or before. In AE121 Fall 2015, the material for Ch. 4 was covered in lectures (I think, from the problem sets) before Thermodynamics and Nozzle Theory.

[2]

3. FLIGHT PERFORMANCE

cf. Chapter 4 Flight Performance of Biblarz and Sutton (2001) [2]

3.1. **Gravity-Free, Drag-Free Space Flight.** Let

$m_p \equiv$ (total) propellant mass (initially)

$t_p \equiv$ propellant burning duration time

$$m(0) = 0$$

Let mass of rocket + propellant $M = M(t) = M(0) - m(t)$ s.t. $M(0) = M_0 + m_p$

$$m(t_p) = m_p$$

$$M(t_p) = M_0$$

If $\dot{m} = \frac{m_p}{t_p}$ (assume constant propellant flow rate),

$$M = M(0) - \frac{m_p}{t_p}t = M(0) \left(1 - \frac{m_p}{M(0)} \frac{t}{t_p}\right) = M(0) \left(1 - \left(1 - \frac{M(0) - m_p}{M(0)}\right) \frac{t}{t_p}\right)$$

cf. Eq. 2-7 of Biblarz and Sutton (2001) [2].

Define some quantities: mass ratio $= \frac{m_f}{m_0}$ where

$m_f :=$ final mass (after rocket operation had consumed all usable propellant), which is M_0 above

$m_0 :=$ mass before rocket operation, which is $M(0)$

propellant mass fraction $\frac{m_p}{M(0)}$ cf. Eq. 2-8 of Biblarz and Sutton (2001) [2].

¹“Euler equations (fluid dynamics)”, Wikipedia

For thrust F_{thrust} ,

$$F_{\text{thrust}} = \dot{m}u_e = M \frac{du}{dt} \implies \frac{\dot{m}dt}{M} = \frac{du}{u_e} \text{ or } \frac{\Delta u}{u_e} = \int \frac{\frac{m_p}{t_p} dt}{M(0) - \frac{m_p}{t_p}t} = -\ln(M(0) - \frac{m_p}{t_p}t) =$$
$$= -\ln(M(0) - \frac{m_p}{t_p}t_p) + \ln(M(0)) = \ln\left(\frac{M(0)}{M(0) - \frac{m_p}{t_p}t_p}\right) = \ln\left(\frac{M(0)}{M(0) - m_p}\right)$$

Thus

$$\Delta u = u_e \ln\left(\frac{M(0)}{M(0) - m_p}\right) = u_e \ln\left(\frac{M(0)}{M_0}\right) \text{ or}$$

$$\exp\left(\frac{\Delta u}{u_e}\right) = \frac{M(0)}{M(0) - m_p} = \frac{M(0)}{M_0} \text{ or } \frac{M_0}{M(0)} = \exp\left(\frac{-\Delta u}{u_e}\right)$$

Also remember that $F = \dot{m}u_e = \dot{m}I_{\text{sp}}g_0$, so $u_e \equiv$ effective exhaust velocity can be related directly to the specific impulse I_{sp} .

Also note that for propellant mass fraction $\frac{m_p}{M(0)} < 1$, $\frac{m_p}{M(0)} = \frac{M(0) - M_0}{M(0)} = 1 - \frac{M_0}{M(0)}$

3.2. **Forces Acting on a Vehicle in the Atmosphere.** assume starting and stopping transients very short and neglected,

$F = u_e \dot{m}$

if mass rate of propellant consumption \dot{m} constant, $\dot{m} = \frac{m_p}{t_p}$ so $F = \frac{m_p}{t_p}u_e$

drag D opposite to flight path due to resistance of body to motion in fluid

lift L normal to flight path

$$L = C_L \frac{1}{2} \rho A u^2$$

$$D = C_D \frac{1}{2} \rho A u^2$$

cf. Eqns. (4-10), (4-11) of Biblarz and Sutton (2001) [2]

density of earth’s atmosphere can vary by factor up to 2 (for altitudes of 300 to 1200 km)

depending on solar activity and night-to-day temperature variations. major unknown in drag

Assume neglect variation of gravity with geographical features and oblate shape of earth,

$$F = \frac{GM_e m}{r^2}$$
$$mg_0 = \frac{GM_e}{R_0^2} m \implies g = g_0 \frac{R_0^2}{R^2} = g_0 \left(\frac{R_0}{R_0 + h}\right)^2$$

3.3. **Basic Relations of Motion.** cf. Section 4.3 “Basic Relations of Motion,” Chapter 4 Flight Performance of Biblarz and Sutton (2001) [2]

$$m\dot{\mathbf{u}} = \mathbf{F} - \mathbf{D} - m\mathbf{g}$$

$$m\dot{u} = F \cos(\psi - \theta) - D - mg \sin \theta$$

$$\implies mu \frac{d\theta}{dt} = F \sin(\psi - \theta) + L - mg \cos \theta$$

$\psi =$ direction of thrust angle from horizontal reference

Consider perturbation effects (cf. Sec. 4.6, listed), drag and gravity

$$\dot{u} = \frac{F}{m} \cos(\psi - \theta) - \frac{C_D}{2m} \rho A u^2 - g \sin \theta$$

$$\implies u\dot{\theta} = \frac{F}{m} \sin(\psi - \theta) + \frac{C_L}{2m} \rho A u^2 - g \cos \theta$$

C_D, C_L are functions of velocity or Mach number (!!!)

For actual trajectory analyses, perturbation effects in Sec. 4.6 must be considered
3-body theory considered
when propellant flow and thrust not constant
from optical or radar tracking data, thrust or actual specific impulse during actual vehicle flights determined from accurately observed trajectory data.

make assumption or measurement on propellant flow (which usually varies in a predetermined manner)
If $L = 0$ (wingless rocket projectile), $\psi - \theta = 0$ (flight direction θ same as thrust direction), and
 $M(t) = M(0) - \dot{m}t$; assume constant $\dot{m} = \frac{t}{t_p} m_p$
 $M(t) = M(0) - \frac{tm_p}{t_p} = M(0)(1 - \frac{m_p}{M(0)} \frac{t}{t_p})$, with $\xi \equiv \frac{m_p}{M(0)}$ propellant mass ratio

$$\begin{aligned} \dot{u} &= \frac{F}{M} - \frac{C_D}{2M} \rho A u^2 - g \sin \theta = \frac{u_e \dot{m}}{M} - \frac{C_D \rho A u^2}{2M} - g \sin \theta = \\ &= \frac{u_e \xi \frac{t}{t_p}}{1 - \xi \frac{t}{t_p}} - \frac{C_D \rho A u^2}{2M(0)(1 - \xi \frac{t}{t_p})} - g \sin \theta \end{aligned}$$

and

$$u \dot{\theta} = -g \cos \theta$$

Example 4-1. Consider “a simple-stage rocket for a rescue flare has the following characteristics and its flight path nomenclature is shown in the sketch.”
Neglect drag “since flight velocities are low,” assume no wind, assume local acceleration of gravity to be equal to sea level g_0 , invariant throughout flight.

Problems. Problem 1. Recall that

$$\frac{M_0}{M(0)} = \exp \left(\frac{-\Delta u}{u_e} \right)$$

and so (in Python)

```
import sympy
from sympy import *
>>> exp(-1600/2000.)
0.449328964117222
```

Problem 2.

$$\frac{m_p}{M(0)} = \frac{M(0) - M_0}{M(0)} = 1 - \frac{M_0}{M(0)} = 1 - \frac{1}{5} = 4/5 = 0.8$$

Problem 3. dragless projectile, so $D = 0$.

$$\dot{u} = -g_0 + \frac{u_e}{t_p(M(0)/m_p - t/t_p)} \implies \Delta u = -g_0 t - u_e \ln \left(1 - \frac{m_0 t}{M(0)t_p} \right)$$

Plugging in $u_e = 2209 \text{ m/sec}$, $m_p/M(0) = 0.57$, $t_p = 5.0 \text{ sec}$, $u_0 = h_0 = 0$,

```
Propulsion.py
>>> integrate( flightpathdirection.subs(Drag,0).subs(psi,theta).subs(theta,pi/2).subs(F_thrust, m_p/t_p*u_e).
subs(M,M_constantflow).subs(m_p,M0*0.57).subs(t_p,5.).subs(u_e,2209.).subs(g_0,9.8).factor(M0).rhs, (t,0,5.0) )
1815.32988528061

Problem0403 = flightpathdirection.subs(Drag,0).subs(psi,theta).subs(theta,pi/2).subs(F_thrust, m_p/t_p*u_e).
subs(M,M_constantflow).subs(m_p,M0*0.57).subs(t_p,5.).subs(u_e,2209.).subs(g_0,9.8).factor(M0).rhs
```

$u_p = 1815 \text{ m/sec}$

```
>>> integrate( integrate(Problem0403,(t,0,t) ),(t,0,5.0) )
3890.37850288891
```

$h_p = 3.89 \times 10^3$

Problem 4. How to estimate A of the projectile?

Problem 5.
Now

$$\begin{aligned} M(t)a &= F_{\text{thrust}} \\ I_{sp} &= \frac{F_{\text{thrust}} t_p}{g_0 \frac{m_p}{t_p} t_p} = \frac{F_{\text{thrust}} t_p}{g_0 m_p} \end{aligned}$$

and for $M(t) = M(0) - \frac{m_p}{t_p} t$,

$$a = \frac{I_{sp} \left(\frac{g_0 m_p}{t_p} \right)}{M(0) - \frac{m_p}{t_p} t} \leq a(t = t_p) \implies a(t = t_p) = \frac{I_{sp} \left(\frac{g_0 m_p}{t_p} \right)}{M_0} = \frac{I_{sp} g_0}{t_p} \left(\frac{1}{\frac{M(0)}{m_p} - 1} \right)$$

Plugging $a = 50 \text{ m/sec}^2$ and solving for t_p ,

```
>>> 260.*9.8/50.*(1/( 1/0.88 - 1 ) )
373.70666666666637
```

(a)

373.71 sec

maximum allowable burn time, assuming steady propellant mass flow

(b)

$$a = \frac{I_{sp} g_0 \xi / t_p}{1 - \xi t / t_p} \implies \Delta u = -I_{sp} g_0 \ln(1 - \xi t / t_p)$$

so

$$\Delta u = 5402.4 \text{ m/sec}$$

for maximum velocity relative to the launch vehicle

Problem 6. Satellite in circular orbit

$$\begin{aligned} F &= \frac{GM_e m}{(R_0 + h)^2} = \frac{mv^2}{(R_0 + h)} \implies \sqrt{\frac{GM_0}{R_0 + h}} = v \\ \frac{2\pi(R_0 + h)}{T} &= v \implies T = \frac{2\pi(R_0 + h)^{3/2}}{\sqrt{GM_e}} \\ \frac{1}{2}mv^2 + \frac{-GM_e m}{R_0 + h} - \left(\frac{-GM_e m}{R_0} \right) &= m \left[\frac{1}{2} \frac{GM_e}{R_0 + h} - \frac{GM_e}{R_0 + h} + \frac{GM_e}{R_0} \right] = m \left[GM_e \left(\frac{-1}{2(R_0 + h)} + \frac{1}{R_0} \right) \right] \end{aligned}$$

Then run **Propulsion.py** which now imports (**import**) in Physique, a small package with the NIST (National Institute of Standards and Technology) Fundamental Constants **FundConst**, NIST SI conversions **conv**, and NASA Planetary Fact Sheet **plnfacts** as Python **pandas** DataFrames.

```
M_earth = plnfacts.loc[plnfacts['Planet']=='EARTH',"Mass_(1024kg)"].values[0]*10**(24) # in kg
R_earth = plnfacts.loc[plnfacts['Planet']=='EARTH',"Diameter_(km)"].values[0]/Decimal(2)
```

```
Gconst = FundConst[ FundConst["Quantity"].str.contains("gravitation") ].loc[243,"Value"]
v0406 = sqrt( Gconst*M_earth/((R_earth + Decimal(500))*10**3) )
# velocity of satellite v of Chapter 4, Problem 6 of Biblarz and Sutton
7611.17633707692
```

```
T0406 = (2.*N(pi)*float(((R_earth + Decimal(500))*10**3 )**(3./2))/float(sqrt( Gconst*M_earth))
# 5678 secs. or 1.58 hours
```

```
Eperm0406 = Gconst*M_earth*(-1/(2*((R_earth+Decimal(500))*10**3)) + 1/(R_earth*10**3))
# Energy per mass
'%6E' % Eperm0406 # 33.51 MJ/kg
```

$$v = \boxed{7611 \text{ m/sec} \quad T = 5678 \text{ s or 1.58 hours} \quad 33.51 \text{ MJ/kg}}$$

Part 2. 1-dim. propulsion (revisited)

Looking at pp. 40-41 of [1], Chapter 2, “Thermodynamics and Fluid Flows”,

$$(\rho u^2 A)_2 - (\rho u^2 A)_1 = \int_{A_2 \amalg A_1 \amalg \text{side}} T^{ij} dS_j \otimes e_1$$

$\int_{A_2 \amalg A_1 \amalg \text{side}} T^{ij} dS_j \otimes e_1$ comprise of the contribution from the fluid being a perfect fluid, $T = -pg$, so consider that *only* first:

$$\int_{A_2 \amalg A_1} -pg^{1j} dS_j \otimes e_1 = -pA_2 + pA_1$$

Suppose shear stress (which happens at solid boundary; by no slip condition, there’s a boundary layer), occurs at wall.

$$\int_{\text{side}} T^{1j} dS_j \otimes e_1 = \int_{\text{side}} (T^{12}(2dx^3 \wedge dx^1) + T^{13}(2dx^1 \wedge dx^2))e_1$$

Assume average (constant) shear over “infinitesimal” or differential cross section: $T^{12} = T^{13} = \frac{\tau}{2}$

$$\implies \tau \int_{\text{side}} dx^3 \wedge dx^1 + dx^1 \wedge dx^2 = \tau(\text{circumference}) \int dx^1$$

Thus, this explains the last equation on pp. 40, and first 2 equations, including Eq. (2.62) on pp.41 of “Thermodynamics and Fluid Flows”, Chapter 2, of Gas Turbine and Rocket Propulsion of Oates (1997) [1]

Oates (1997) [1]

$$\begin{aligned} \tau d\sigma &= dU + pdV \\ \implies \tau d\sigma &= \left(\frac{\partial U}{\partial \tau} \right)_V d\tau + \left(\frac{\partial U}{\partial V} \right)_\tau dV + pdV \equiv C_V d\tau + \left(\frac{\partial U}{\partial V} \right)_\tau dV + pdV \end{aligned}$$

with $U = U(\tau, V)$.

For (calorically) perfect gas, $U = U(\tau)$ and so

$$\begin{aligned} d\sigma &= C_V \frac{d\tau}{\tau} + \frac{N}{V} dV = C_V \frac{d\tau}{\tau} + \frac{-Nd\rho}{\rho} \\ \implies \sigma_2 - \sigma_1 &= C_V \ln \left(\frac{\tau_2}{\tau_1} \right) + N \ln \left(\frac{V_2}{V_1} \right) = \ln \left[\left(\frac{\tau_2}{\tau_1} \right)^{C_V} \left(\frac{V_2}{V_1} \right)^N \right] \\ \implies \exp(\sigma_2 - \sigma_1) &= \left[\left(\frac{\tau_2}{\tau_1} \right)^{\frac{C_V}{N}+1} \frac{p_1}{p_2} \right]^N \quad \text{or} \quad \frac{p_2}{p_1} = \left(\frac{\tau_2}{\tau_1} \right)^{\frac{\gamma}{\gamma-1}} \exp \left[\frac{\sigma_2 - \sigma_1}{-N} \right] \end{aligned}$$

Starting from Sec. 2.16 *The Channel Flow Equations* of Ch. 2 of Oates (1997) [1], pp. 39-40, consider how heating Q' inside an (control volume) element of fluid in a duct changes the (specific) enthalpy h' :

$$\begin{aligned} Q' &\in \Omega^1(\Sigma) \otimes \Omega^n(N) \\ h' &\in \Omega^1(\Sigma) \otimes \Omega^{n-1}(N) \end{aligned}$$

where it is now clear that the integration over the (control volume) element of the fluid on spatial manifold N is distinguished from the integration (usually along a curve representing a thermodynamic process) over the manifold Σ of *thermodynamic states*.

By energy conservation,

$$Q' = \mathbf{d}h'$$

where \mathbf{d} is the exterior derivative *on* N , not on Σ . h' , the “specific enthalpy”, is a *Bernoulli invariant* of the flow ².

h' is also called *stagnation enthalpy*.

²“Euler equations (fluid dynamics)” *Wikipedia*, [https://en.wikipedia.org/wiki/Euler_equations_\(fluid_dynamics\)](https://en.wikipedia.org/wiki/Euler_equations_(fluid_dynamics))

For instance, integrate over the control volume, and then over a 1-dimensional flow (cylindrical volume element)

$$\int_{V_0} Q' = \int_{V_0} \mathbf{d}h' = (h')_2 A_2 - (h')_1 A_1 = (h_0 + \frac{u^2}{2})_2 \rho_2 A_2 - (h_0 + \frac{u^2}{2})_1 \rho_1 A_1$$

But one also have to account for the thermodynamic process during the flow across the element control volume, on Σ :

$$\tau d\sigma = dU + pdV = dH - Vdp = dH - \frac{dp}{\rho/NM}$$

So for the Bernoulli invariant h' , h' has a piece $h \equiv H/NM \in C^\infty(\Sigma)$, that lives on Σ .

In conclusion,

$$(1) \quad Q' = \mathbf{d} \left(\frac{u^2}{2} \right) + \mathbf{d}h = \mathbf{d} \left(\frac{u^2}{2} \right) + \frac{\tau}{NM} d\sigma + \frac{dp}{\rho} \equiv \mathbf{d} \left(\frac{u^2}{2} \right) + \tau ds + \frac{dp}{\rho}$$

which is Eq. (2.61) of Oates (1997) [1].

Consider mass conservation for fluid flow across a control volume element:

$$\dot{m} = \frac{d}{dt} \int_{V_0} \rho \text{vol}^n = \int_{V_0} \frac{\partial \rho}{\partial t} \text{vol}^n + \int_{V_0} di_u \rho \text{vol}^n = 0 + \int_{\partial V_0} i_u \rho \text{vol}^n = (\rho u A)_2 - (\rho u A)_1$$

Then for $\rho u A = \text{const.}$, $\frac{dp}{\rho} + \frac{du}{u} + \frac{dA}{A} = 0$.

Assuming an ideal gas, $p = \frac{N\tau}{V} = \frac{\rho\tau}{M}$

$$\frac{dA}{A} + \frac{du}{u} + \frac{dp}{\rho} = \frac{dA}{A} + \frac{du}{u} + \frac{d(pM/\tau)}{\rho}$$

Then for $\frac{d(pM/\tau)}{\rho}$ only, and using Eq. 1, which is from energy conservation where Q' heating (or heat dissipation) is also included,

$$\begin{aligned} \frac{d(pM/\tau)}{\rho} &= \frac{M}{\tau} \frac{dp}{\rho} + \frac{pM}{-\tau^2} \frac{d\tau}{\rho} = \\ &= \left(Q' - \mathbf{d} \left(\frac{u^2}{2} \right) - \frac{\tau d\sigma}{NM} \right) \frac{M}{\tau} + \frac{pM}{-\tau^2} \frac{d\tau}{\rho} \end{aligned}$$

The last term, $\frac{pM}{-\tau^2} \frac{d\tau}{\rho}$, is resolved by considering again the Bernoulli invariant h , and the fact that it remains constant along the flow:

$$\begin{aligned} \frac{pM}{-\tau^2} \frac{d\tau}{\rho} &= \frac{pMV}{-\tau^2 MN} d\tau = -\frac{d\tau}{\tau} \quad (\text{ideal gas}) \\ h &= h_1 + \frac{u_1^2}{2} = h_2 + \frac{u_2^2}{2} \implies dh = C_p/(MN) d\tau = u du \\ \implies \frac{d\tau}{\tau} &= \frac{u du}{(C_p/MN)\tau} = \frac{u du}{\left(\frac{\gamma}{M(\gamma-1)} \right) \tau} = M \frac{\gamma-1}{\gamma} \frac{u du}{\tau} \end{aligned}$$

Thus

$$Q' \frac{M}{\tau} - \frac{d\sigma}{N} - \frac{M}{\gamma} \frac{u du}{\tau} = Q' \frac{M}{\tau} - \frac{d\sigma}{N} - \mathfrak{M}^2 \frac{du}{u}$$

with Mach number $\mathfrak{M} := \frac{u}{a}$ and speed of sound $a^2 = \frac{\gamma\tau}{M}$ which comes from the adiabatic process of compressing and the expansion of gas longitudinally $\left(\frac{\partial p}{\partial \rho} \right)_{\text{adiabatic}} = a^2$.

And so for fluid (ideal gas) flow through a (control volume, differential) element of a (area) duct, including heating (or heat dissipation),

$$(2) \quad \frac{dA}{A} + (1 - \mathfrak{M}^2) \frac{du}{u} = \frac{d\sigma}{N} - Q' \frac{M}{\tau}$$

3.3.1. *Adiabatic Flow of Ideal Gas.* Let $Q' = 0$ for adiabatic flow (by definition of “adiabatic”). Then from Eq. 2

$$\frac{dA}{A} + (1 - \mathfrak{M}^2) \frac{du}{u} = \frac{d\sigma}{N}$$

Now $d\sigma \geq 0$, always (2nd. law of Thermodynamics). $N > 0$.

$dA < 0$ for *converging* nozzle. $dA > 0$ for *diverging* nozzle.

For isentropic flow $d\sigma = 0$,

for $dA < 0$, for $\mathfrak{M} < 1$, $du > 0$ (flow is accelerating)

for $\mathfrak{M} > 1$, $du < 0$ (flow is decelerating)

At the throat, $dA = 0$, $\mathfrak{M} = 1$.

However, for $d\sigma > 0$, at the throat, $dA = 0$, for acceleration $du > 0$, $\mathfrak{M} < 1$ and for deceleration $du < 0$, $\mathfrak{M} > 1$. So from subsonic flow in the initial part (usually the converging part), then the “effective throat” of the nozzle is shifted slightly downstream.

cf. Ch. 2 Problems, Oates (1997) [1]

Problem 2.1.

Starting from the thermodynamic identity,

$$d\sigma = \frac{dU}{\tau} + \frac{p}{\tau} dV = \frac{dU}{\tau} + \frac{NdV}{V}$$

$U = U(\tau)$ for an (perfect) ideal gas, and $C_V = \left(\frac{\partial U}{\partial \tau}\right)_V = \frac{dU}{d\tau}$ so

$$d\sigma = C_V \frac{d\tau}{\tau} + \frac{N}{V} dV$$

For an isentropic process, $\int d\sigma = 0$, so

$$\int d\sigma = 0 = \int C_V \frac{d\tau}{\tau} + N \ln \frac{V_2}{V_1}$$

For $C_V = A + B\tau + C\tau^2 + D\tau^3$,

$$\begin{aligned} N \ln \frac{V_1}{V_2} &= N \ln \frac{\rho_2}{\rho_1} = A \ln \left(\frac{\tau_2}{\tau_1} \right) + B(\tau_2 - \tau_1) + \frac{C(\tau_2 - \tau_1)^2}{2} + D \frac{(\tau_2 - \tau_1)^3}{3} \\ \implies \frac{\rho_2}{\rho_1} &= \left(\left(\frac{\tau_2}{\tau_1} \right) \exp \left[\frac{B}{A}(\tau_2 - \tau_1) + \frac{C}{2A}(\tau_2 - \tau_1)^2 + \frac{D}{3A}(\tau_2 - \tau_1)^3 \right] \right)^{A/N} \end{aligned}$$

Over a “wide” temperature range, C_p is usually given as a polynomial, from empirically measured parameters.

For C_p , consider

$$dH = dU + pdV + Vdp = \tau d\sigma + Vdp = \left(\frac{\partial H}{\partial \tau} \right)_p d\tau + \left(\frac{\partial H}{\partial p} \right)_\tau dp$$

$$Q = dU + pdV = dH - Vdp$$

$$C_p = \left(\frac{\partial H}{\partial \tau} \right)_p$$

$$Q = \tau d\sigma = dH - Vdp \text{ or } d\sigma = \frac{dH}{\tau} - \frac{V}{\tau} dp = \frac{dH}{\tau} - \frac{Ndp}{p}$$

$$\implies \int d\sigma = \int \frac{C_p d\tau}{\tau} - N \ln \frac{p_2}{p_1}$$

$$\implies \frac{p_2}{p_1} = \left(\left(\frac{\tau_2}{\tau_1} \right) \exp \left[\frac{B}{A}(\tau_2 - \tau_1) + \frac{C}{2A}(\tau_2 - \tau_1)^2 + \frac{D}{3A}(\tau_2 - \tau_1)^3 \right] \right)^{A/N}$$

cf. Ch. 3 Chemical Rockets, Oates (1997) [?]

rocket volume R_0

∂R_0 includes

internal surfaces wetted by fluid (chamber, pipes, pumps, etc.)

Consider rocket on thrust stand. Consider force on thrust stand.

\sum_0 outer surface area.

\sum_c inner chamber surface area.

Now consider the total mass of the gas propellant, in the chamber and nozzle, represented by a differential n -form on (spatial) manifold N , representing spatial points, of dimension $\dim N = n$, denoted by m :

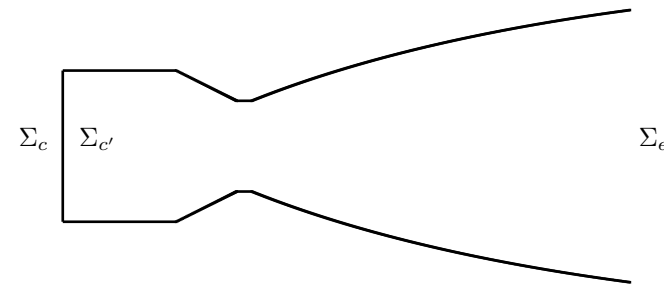
$$m = \int_{B(t)} \rho \text{vol}^n$$

where $B(t)$ is the control volume, as a function of time t , and $\text{vol}^n \in \Omega^n(N)$ is the volume form, $\Omega^n(N)$ is the set of all n -forms on N .

Then, by differential geometry,

$$\dot{m} = \int_{B(t)} \left[\frac{\partial \rho}{\partial t} \text{vol}^n + \mathcal{L}_u \rho \text{vol}^n \right] = \int_{B(t)} \left[\frac{\partial \rho}{\partial t} \text{vol}^n + di_u \rho \text{vol}^n \right] = \int_{B(t)} \frac{\partial \rho}{\partial t} \text{vol}^n + \int_{\partial B} i_u \rho \text{vol}^n = \int_{B(t)} \frac{\partial \rho}{\partial t} \text{vol}^n + \int_{\partial B} \rho u^i dS_i$$

where the last equality was obtained by Stoke’s law.



Consider the total momentum, the sum of all the momentums of the fluid particles, Π , and its time derivative, $\dot{\Pi}$,

$$\Pi = \int_{B(t)} \rho u^i \text{vol}^n \otimes e_i$$

$$\dot{\Pi} = \int_{B(t)} \frac{\partial(\rho u^i)}{\partial t} \text{vol}^n \otimes e_i + \int_{B(t)} d(\rho u^i i_u \text{vol}^n) \otimes e_i = \int_{B(t)} \frac{\partial(\rho u^i)}{\partial t} \text{vol}^n \otimes e_i + \int_{\partial B(t)} \rho u^i i_u \text{vol}^n \otimes e_i$$

For steady state $\frac{\partial(\rho u^i)}{\partial t} = 0$, $\dot{\Pi} = \int_{\partial B(t)} \rho u^i i_u \text{vol}^n \otimes e_i$.

Now

$$\partial B(t) = A_e \coprod \Sigma_{c'}$$

$\Sigma_{c'} \equiv$ internal surface of chamber (orientation, or outward normal, chosen to point into rocket stages, i.e. outward from engine).

Suppose $u = 0$ on $\Sigma_{c'}$. Then

$$\dot{\Pi} = \int_{A_e} \rho u^i i_u \text{vol}^n \otimes e_i$$

Consider the external forces on a fluid. The external forces on a fluid, acting on the surface of the fluid, is “wrapped up” in the stress tensor $T = T^{ij} e_i \otimes e_j \in TN \otimes TN$. On a surface,

$$\int T^{ij} dS_j \otimes e_i$$

Free space case: $g = D = 0$

$$\begin{aligned}\frac{dv}{dt} &= \frac{-\dot{M}u_{\text{exh}}}{M} + \frac{(p_{\text{exh}} - p_a)A_e}{M} \\ \Delta v_{\text{tot}} &= -u_{\text{exh}} \ln \left(\frac{M(0) - m_p}{M(0)} \right) + \frac{(p_{\text{exh}} - p_a)A_e}{\frac{-m_p}{t_f}} \ln \left(M(0) - \frac{m_p}{t_f} t \right) \Bigg|_0^{t_f} = \\ &= -u_{\text{exh}} \ln \left(\frac{M(0) - m_p}{M(0)} \right) + \frac{(p_{\text{exh}} - p_a)A_e}{\frac{-m_p}{t_f}} \ln \left(\frac{M(0) - m_p}{M(0)} \right) = \\ &= \left[u_{\text{exh}} + \frac{(p_{\text{exh}} - p_a)A_e}{\dot{m}} \right] \ln \left(\frac{M(0)}{M(0) - m_p} \right)\end{aligned}$$

For an atmosphere-free planet, $D = 0$,

$$\begin{aligned}\dot{v} &= \frac{-\dot{M}u_{\text{exh}}}{M} + \frac{(p_{\text{exh}} - p_a)A_e}{M} - g \cos \theta \\ \implies \Delta v &= \left(u_{\text{exh}} + \frac{(p_{\text{exh}} - p_a)A_e}{\dot{m}} \right) \ln \left(\frac{M(0)}{M(0) - m_p} \right) - g \cos \theta t_f\end{aligned}$$

4. MULTIPLE STAGING

cf. pp. 67, **Multiple-Stage Rockets**, Ch. 3, Oates (1997) [1]

5. LAGRANGIAN

5.1. **Review.** Consider a solid cylinder with surface area $A_{\text{cyl}} = A_{\text{sides}} \coprod A_{\text{top}} \coprod A_{\text{bottom}}$ s.t.

$$\begin{aligned}dA_{\text{sides}} &= ad\varphi dz & \hat{n} &= \frac{x\mathbf{e}_x + y\mathbf{e}_y}{\|x\mathbf{e}_x + y\mathbf{e}_y\|} = \mathbf{e}_r = \cos \varphi \mathbf{e}_x + \sin \varphi \mathbf{e}_y \\ dA_{\text{top}} &= dA_{\text{bottom}} = 2\pi r dr & \hat{n} &= \pm \mathbf{e}_z \quad (\text{for top and bottom, respectively})\end{aligned}$$

for a fixed, *body-frame* coordinates $\mathbf{e}_x, \mathbf{e}_y, \mathbf{e}_z$.

Let

$$A_{\text{projected}} = \int_A (\cos \beta) dA \chi_{\cos \beta \geq 0}$$

where A is the original area

β is the angle between normal to surface A and normal to arbitrary plane onto which we project and

$\chi_{\cos \beta \geq 0}$ is an indicator function, a fancy way/terminology for saying s.t. $\chi_{\cos \beta \geq 0} = \begin{cases} 1 & \text{if } \cos \beta \geq 0 \\ 0 & \text{otherwise} \end{cases}$.

Consider $\frac{\mathbf{u}}{\|\mathbf{u}\|} \equiv \hat{u} = \hat{u}_x \mathbf{e}_x + \hat{u}_y \mathbf{e}_y + \hat{u}_z \mathbf{e}_z$.

Consider now these sanity checks: consider $\hat{u} = \mathbf{e}_z$.

$$\begin{aligned}\int_{A_{\text{top}}} \hat{u} \cdot \hat{n} 2\pi r dr &= 2\pi \int_0^a r dr = \pi a^2 \\ \int_{A_{\text{bottom}}} \hat{u} \cdot \hat{n} 2\pi r dr \chi_{\hat{u} \cdot \hat{n} \geq 0} &= 0 \text{ since } \hat{u} \cdot \hat{n} = -1 \\ \int_{A_{\text{side}}} \hat{u} \cdot \hat{n} 2\pi r dr &= 0 \text{ since } \hat{u} \cdot \hat{n} = 0 \\ \implies A_{\text{projected}} &= \pi a^2\end{aligned}$$

Consider $\hat{u} = \mathbf{e}_x$ or \mathbf{e}_y or \mathbf{e}_r . $\hat{u} \cdot \hat{n} = 0$ on A_{top} and on A_{bottom} .

³<http://ice.as.arizona.edu/~dpsaltis/Phys422/chapter6.pdf>

$$\int_{A_{\text{side}}} \hat{u} \cdot \hat{n} ad\varphi dz \chi_{\hat{u} \cdot \hat{n} \geq 0} = ah \int_{-\pi}^{\pi} d\varphi \cos \varphi \chi_{\cos \varphi \geq 0} = ah \int_{-\pi/2}^{\pi/2} d\varphi \cos \varphi = 2ah$$

Indeed, this result is rotationally symmetric about \mathbf{e}_z , for if $\hat{u} = \mathbf{e}_y$.

$$\int_{A_{\text{side}}} \sin \varphi d\varphi \chi_{\sin \varphi \geq 0} = \int_0^{\pi} d\varphi \sin \varphi = -\cos \varphi \Big|_0^{\pi} = 2$$

and so $A_{\text{projected}} = 2ah$.

Consider $\hat{u} = \hat{u}_\rho \mathbf{e}_\rho + \hat{u}_z \mathbf{e}_z$. Then $A_{\text{projected}} = \hat{u}_z \pi a^2 + \hat{u}_\rho 2ah$.

Returning back to the drag

$$\mathbf{F}_d = \frac{-1}{2} \rho C_a A |\mathbf{u}| \mathbf{u}$$

The components of the drag force in 2-dim. Cartesian coordinates are

$$(\mathbf{F}_d)_1 = \frac{-1}{2} \rho C_a A (\dot{x}_1^2 + \dot{x}_2^2)^{1/2} \dot{x}_1$$

$$(\mathbf{F}_d)_2 = \frac{-1}{2} \rho C_a A (\dot{x}_1^2 + \dot{x}_2^2)^{1/2} \dot{x}_2$$

In spherical coordinates,

$$Q_r = F_1 \frac{\partial x^1}{\partial r} + F_2 \frac{\partial x^2}{\partial r} = \frac{-1}{2} \rho C_a A (\dot{r}^2 + (r\dot{\varphi})^2)^{1/2} \dot{r}$$

$$Q_\varphi = F_1 \frac{\partial x^1}{\partial \varphi} + F_2 \frac{\partial x^2}{\partial \varphi} = \frac{-1}{2} \rho C_a A (\dot{r}^2 + (r\dot{\varphi})^2)^{1/2} r^2 \dot{\varphi}$$

since $(\dot{r}c_\varphi - rs_\varphi\dot{\varphi})c_\varphi + (\dot{r}s_\varphi + rc_\varphi\dot{\varphi})s_\varphi = \dot{r}$ (cf. Chapter 6 “Lagrangian Dynamics” ³)

$$(\dot{r}c_\varphi - rs_\varphi\dot{\varphi})(-s_\varphi) + (\dot{r}s_\varphi + rc_\varphi\dot{\varphi})c_\varphi = r\dot{\varphi}$$

Clearly, the so-called “Rayleigh’s dissipation function” (Eq. (1.67) on pp. 23 of Goldstein, et. al. [7]) is

$$\mathcal{F} = \frac{1}{3} \rho C_a A (\dot{x}_1^2 + \dot{x}_2^2)^{3/2}$$

One can also write down a generalized force for the thrust

$$F_{\text{thrust}} = T = \dot{m}_p \mathbf{u}_e + (p_e - p_a) A_e \mathbf{e}_z$$

where \mathbf{u}_e is in the direction opposite of the z axis of symmetry of the *body-frame*, as with \mathbf{e}_z .

One should also consider \mathbf{e}_z to be a vector in the *body-frame* that’s arbitrary to represent cold-gas thrusters:

$$F_{\text{thrust}} \equiv T = (\dot{m}_p u_e + (p_e - p_a) A_e) \mathbf{e}_z$$

Part 3. AE121

Let’s translate $\frac{\gamma}{\gamma-1}$ between physicists and engineers:

(6)

$$\boxed{\frac{\gamma}{\gamma-1} = \frac{\frac{C_p}{C_V}}{\frac{N}{C_V}} = \frac{C_p}{N} = \frac{c_p M N}{N} = \frac{c_p k_B}{R}}$$

5.1.1. *Speed of sound.* From pp. 179, Problem 10 “Isnetropic relations of ideal gas” of Chapter 6: Ideal Gas of Kittel and Kroemer [4], recall isentropic bulk moduli B_σ

$$B_\sigma := -v \left(\frac{\partial p}{\partial V} \right)_\sigma = \gamma \frac{p_i V_i^\gamma}{V^\gamma} = \gamma p$$

(8)

with $p = \frac{p_i V_i^\gamma}{V^\gamma}$.
Very little heat transfer in sound wave. For velocity (magnitude) i.e. speed of sound, a

$$a = \left(\frac{B_\sigma}{\rho} \right)^{1/2} = \left(\frac{\gamma p}{\rho} \right)^{1/2}$$

Now

$$p = \frac{N\tau}{V}$$
$$\frac{p}{\rho} = \frac{N\tau}{V} \frac{1}{\left(\frac{MN}{V}\right)} = \frac{\tau}{M}$$

Now $p = \rho RT$ (outside of theoretical physics, people use the so-called universal gas constant R).

$$p = \rho RT = \frac{N\tau}{V} = \frac{MN}{V} R \frac{\tau}{k_B}$$

and so

(7)

$$R := \frac{k_B}{M}$$

so, as Polk says, R is different for different gases.

And so

$$\frac{\gamma p}{\rho} = \frac{\gamma \tau}{M}$$

6. ISENTROPIC FLOW EQNS. WITH AREA CHANGE

$$\frac{p}{p_i} = \left(\frac{\tau}{\tau_i} \right)^{\frac{\gamma}{\gamma-1}}$$
$$\frac{\tau}{\tau_i} = \left(\frac{V_i}{V} \right)^{\gamma-1} = \left(\frac{\rho}{\rho_i} \right)^{\gamma-1}$$
$$\frac{p}{p_i} = \left(\frac{V_i}{V} \right)^\gamma = \left(\frac{\rho}{\rho_i} \right)^\gamma$$

cf. “Isentropic relations of ideal gas” of Kittel and Kroemer [4]

cf. 20151105 AE121 Polk

$$h_0 = h + \frac{v^2}{2}$$
$$C_p T_0 = C_p T + \frac{v^2}{2}$$
$$\frac{T_0}{T} = 1 + \frac{v^2}{2C_p T}$$

Define $\mathfrak{M} = \frac{v}{a}$, Mach #

sound speed $a = (\gamma RT)^{1/2}$

$$\frac{T_0}{T} = 1 + \frac{\gamma-1}{2} \mathfrak{M}^2$$

$$\frac{p_0}{p} = \left(1 + \frac{\gamma-1}{2} \mathfrak{M}^2 \right)^{\frac{\gamma}{\gamma-1}}$$

$$\frac{\rho_0}{\rho} = \left(1 + \frac{\gamma-1}{2} \mathfrak{M}^2 \right)^{\frac{1}{\gamma-1}}$$

(EY : 20151118 says

$$h := \frac{H}{MN}$$
$$c_p := \frac{C_p}{MN} \implies h_0 = h + \frac{u^2}{2}$$
$$H_0 = H + \frac{MNu^2}{2} \quad c_p \tau_0 = c_p \tau + \frac{u^2}{2}$$
$$H = C_p \tau$$

and now

$$\mathfrak{M} = \frac{u}{a} = \frac{u}{(\gamma \tau / M)^{1/2}} \quad C_p = C_V + N$$
$$a = \left(\frac{\gamma \tau}{M} \right)^{1/2} = \left(\frac{\gamma k_B T}{M} \right)^{1/2} = (\gamma RT)^{1/2} \implies \gamma - 1 = \frac{N}{C_V}$$

and so

$$\frac{\tau_0}{\tau} = 1 + \frac{u^2}{2c_p \tau} = 1 + \frac{1}{2c_p \tau} \mathfrak{M}^2 \frac{\gamma \tau}{M} = 1 + \frac{\gamma \mathfrak{M}^2}{2C_p / N} = 1 + \frac{\mathfrak{M}^2}{2C_V / N} = 1 + \frac{\gamma-1}{2} \mathfrak{M}^2$$

)

Now we need to get \mathfrak{M} in terms of area so we can apply these to a nozzle.

Use continuity:

$$\rho_1 v_1 A_1 = \rho_2 A_2 v_2$$

Thus

$$\frac{A_2}{A_1} = \frac{\rho_1 v_1}{\rho_2 v_2} = \frac{\mathfrak{M}_1}{\mathfrak{M}_2} \left(\frac{T_1 \rho_1^2}{T_2 \rho_2^2} \right)^{1/2}$$

EY : 20151120 One can also relate a point in the flow, 1, to another point “downstream” to the flow, 2:

$$\frac{T_1}{T_2} = \frac{T_1/T_0}{T_2/T_0}$$

Substitute for T_1/T_2 and p_1/p_e , e or exh for exhaust,

(9)

$$\frac{A_2}{A_1} = \frac{\mathfrak{M}_1}{\mathfrak{M}_2} \left[\left(\frac{1 + \frac{\gamma-1}{2} \mathfrak{M}_2^2}{1 + \frac{\gamma-1}{2} \mathfrak{M}_1^2} \right)^{\frac{\gamma+1}{\gamma-1}} \right]^{1/2}$$

To get thrust, we need an expression for mass flow rate

$$\dot{m} = \rho v A$$
$$\frac{\dot{m}}{A} = \rho v \quad (\text{Recall}) \ a = (\gamma RT)^{1/2}$$

$$v = \mathfrak{M} a = \mathfrak{M} (\gamma RT)^{1/2} = \mathfrak{M} (\gamma RT_0)^{1/2} \left(\frac{T}{T_0} \right)^{1/2} = \mathfrak{M} \left[\frac{\gamma RT_0}{1 + \frac{\gamma-1}{2} \mathfrak{M}^2} \right]^{1/2}$$

Now

$$\rho = \rho_0 \left(\frac{\rho}{\rho_0} \right) = \rho_0 \left[\frac{1}{1 + \frac{\gamma-1}{2} \mathfrak{M}^2} \right]^{\frac{1}{\gamma-1}}$$

$$\rho v = \frac{\dot{m}}{A} = \frac{\rho_0 \mathfrak{M} (\gamma R T_0)^{1/2}}{\left(1 + \frac{\gamma-1}{2} \mathfrak{M}^2\right)^{1/2} \left(1 + \frac{\gamma-1}{2} \mathfrak{M}^2\right)^{\frac{1}{\gamma-1}}}$$

Using the ideal gas law

$$\frac{\dot{m}}{A} = \frac{p_0 \gamma^{1/2}}{(R T_0)^{1/2}} \mathfrak{M} \frac{1}{\left(1 + \frac{\gamma-1}{2} \mathfrak{M}^2\right)^{\frac{\gamma+1}{2(\gamma-1)}}}$$

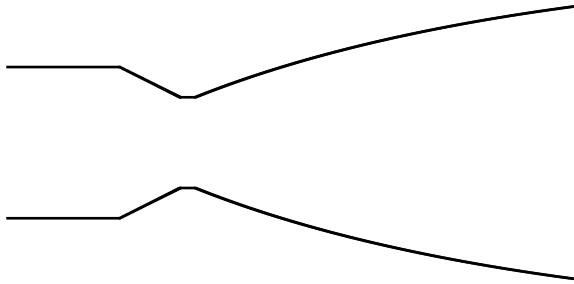
At the throat, $\mathfrak{M} = 1$,

$$(10) \quad \frac{\dot{m}}{A^*} = \frac{p_0 \gamma^{1/2}}{(R T_0)^{1/2}} \left(\frac{2}{\gamma + 1} \right)^{\frac{\gamma+1}{2(\gamma-1)}}$$

Thus, from Eq. 10 above (giving the mass flow from throat area and stagnation p_0 , T_0), and plugging this into thrust (force) equation,

$$(11) \quad T = \dot{m} v_e + (p_e - p_a) A_e =$$

$$= \frac{A^* p_0 \gamma^{1/2}}{(R T_0)^{1/2}} \left(\frac{2}{\gamma + 1} \right)^{\frac{\gamma+1}{2(\gamma-1)}} \sqrt{\frac{2 \gamma R T_0}{\gamma - 1} \left(1 - \left(\frac{p_e}{p_0} \right)^{\frac{\gamma-1}{\gamma}} \right)} + (p_e - p_a) A_e$$



7. PSs

7.1. PS2. Problem 1: Ares I launch vehicle. Consider a 2-stage rocket. It's also interesting to explore the properties of polybutadiene acrylonitrile (PBAN) (first stage solid rocket motor fuel), liquid oxygen/liquid hydrogen (LOX/LH2) (second stage liquid rocket engine fuel).

ϵ , according to wikipedia ⁴, is the “ratio between the empty mass of the stage, and the combined empty mass and propellant mass”, which is, in the notation of Biblarz and Sutton (2001) [2], $\zeta \equiv \frac{m_p}{M(0)}$.

Neglect drag and earth's gravity.

Assume constant mass flow.

Let

$$M(t) = M_2(t) + M_{\text{payl}} + M_{01} - m_1(t)$$

with $M = M(t)$ being the total mass of the entire system during the first stage, with $m_1 = m_1(t) \in \mathbb{R}$ s.t.

$$m_1(0) = 0$$

⁴https://en.wikipedia.org/wiki/Multistage_rocket

PROPULSION

$m_1(t_{p1}) = m_{p1}$ (total mass of propellant of stage 1)

$t_{p1} \equiv$ burn time of first stage

$M_{01} \equiv$ total mass of empty stage 1 + propellant mass for stage 1

$M_{01} - m_{p1} \equiv$ mass of empty stage 1.

Also,

$$\epsilon_1 = \frac{M_{01} - m_{p1}}{M_{01}} = 1 - \frac{m_{p1}}{M_{01}}$$

and

$$I_{sp} = \frac{\int F_{\text{thrust}} dt}{W} = \frac{F_{\text{thrust}} t_p}{g_0 \dot{m} t_p} = \frac{u_e}{g_0}$$

Now for the second stage,

$$M_2(t) = \begin{cases} M_{02} & \text{if } 0 \leq t \leq t_{p1} \\ M_{02} - m_2(t) & \end{cases}$$

for $m_2(t_{p1}) = 0$

$$m_2(t_{p2} + t_{p1}) = m_{p2}$$

From physics, equating kinematics, dynamics $M\dot{u}$ to the external force, F_{thrust} ,

$$M\dot{u} = F_{\text{thrust}} = \dot{m}_1(I_{sp1}g_0) = \frac{m_{p1}}{t_{p1}} I_{sp1}g_0 \text{ so}$$

$$\dot{u} = \frac{\frac{m_{p1}}{t_{p1}} I_{sp1}g_0}{M_2(t) + M_{\text{payl}} + M_{01} - \frac{m_{p1}}{t_{p1}} t} \implies \Delta u_1 = -I_{sp1}g_0 \ln \left(\frac{M_2(t) + M_{\text{payl}} + M_{01} - m_{p1}}{M_2(t) + M_{\text{payl}} + M_{01}} \right)$$

Also,

$$\Delta u_2 = -I_{sp2}g_0 \ln \left(\frac{M_{\text{payl}} + M_{02} - m_{p2}}{M_{\text{payl}} + M_{02}} \right)$$

```
(a) gstd = FundConst[ FundConst["Quantity"] . str.contains("gravity") ].loc[303,:].Value
# get standard acceleration of gravity
M_0 = Symbol('M_0', positive=True)
Deltat = -I_sp*g_0*ln( (M_0 -m_p)/M_0)
# part (a)
Deltat.subs(I_sp,268.8).subs(g_0,gstd).subs(M_0,805309.).subs(m_p, (1-0.1396)*586344)
# 2595.74521034101 m/s
```

$$\Delta u_1 = 2595.7 \text{ m/s}$$

```
(b) Deltat.subs(I_sp,452.1).subs(g_0,gstd).subs(M_0,183952+35013.).subs(m_p, (1-0.1110)*183952)
# 6090.68716730318 m/s
```

$$\Delta u_2 = 6090.7 \text{ m/s and so}$$

$$u_{2f} = 8686.4 \text{ m/s}$$

(c) Now

$$\frac{F_{\text{thrust}}}{W} = \frac{\dot{m}_1 u_{e1}}{g_0 M(0)} = 1.5$$

is the thrust to weight ratio at the instant of takeoff.

Now $u_{e1} = I_{sp1}g_0$ and so

$$\dot{m}_1 = \frac{1.5M(0)}{I_{sp1}}$$

and so

```
>>> 1.5*805309./268.8
# 4493.911830357143
```

so

$$\dot{m}_1 = 4493.9 \text{ kg/s}$$

Over 4 tons of propellant reactants is dumped out per second!

Problem 2. Continuous staging

- (a) $M_{\text{payl}} \equiv$ payload mass; $M_{\text{payl}} \in \mathbb{R}^+$.
Assume structure mass discarded at 0 velocity.
structure mass continuously jettisoned during the burn.
 $M = M(t) \in C^\infty(\mathbb{R})$ represents mass of structure undiscarded at time t , i.e. system of structure + propellant, at “control volume” at time t . So consider
 $M = M_{\text{payl}} + m_s(t) + m_p(t)$ s.t.
 $m_s = m_s(t) \in C^\infty(\mathbb{R})$, mass of structure not yet thrown out, and propellant mass $m_p = m_p(t) \in C^\infty(\mathbb{R})$
If we assume constant propellant burn (out), constant propellant flow rate, then

$$\dot{m}_p(t) = \frac{-m_p}{t_p}$$

with $m_p \in \mathbb{R}^+$ total mass of propellant
 $t_p \in \mathbb{R}^+$ burn time of propellant fuel.

$m_p(t) = m_p - \frac{m_p}{t_p}t = m_p(1 - \frac{t}{t_p})$
Assume constant dead mass ratio $\delta = m_s(t)/m_p(t)$ (EY: my intuition is we’re throwing out as much propellant out in fixed proportion to structure being discarded out; the name “dead” refers to what’s still left that’s being propelled forward by thrust, I think (???))
 $m_s(t) = \delta m_p(t) = \delta(m_p)(1 - t/t_p)$
Assume I_{sp} constant

$$I_{sp} \equiv \frac{\int F_{\text{thrust}} dt}{W_{\text{propellant}}} = \frac{(-\dot{m}_p)u_e t_p}{m_p g_0} = \frac{u_e}{g_0}$$

Consider the instantaneous rest frame of spacecraft + propellant fuel system $\sum_\alpha \Delta p_{i\alpha} = 0$
Consider before and after, after an instant. So for $\sum_\alpha p_{f\alpha}$,
Now $M = M(t) = M_{\text{payl}} + m_s(t) + m_p(t)$
 $M(t + \delta t) = M(t) + \dot{M}\delta t$
Then the momentum of the part that’s going to be propelled by the thrust at time $t + \delta t$ is

$$M(t + \delta t)u(t + \delta t) = (M(t) + \dot{M}\delta t)(u(t) + \dot{u}\delta t) = (M(t) + \dot{M}\delta t)(0 + \dot{u}\delta t) = M\dot{u}\delta t + O((\delta t)^2)$$

The momentum of the propellant expelled out + structure that’s discarded is

$$(-\dot{m}_p dt)(-u_e) + (-\dot{m}_s dt) \cdot 0 = -\dot{m}_p dt(-u_e)$$

$$\implies M\dot{u}\delta t + \dot{m}_p u_e \delta t = 0 \implies \dot{u} = -\frac{\dot{m}_p u_e}{M_{\text{payl}} + m_s(t) + m_p(t)} = \frac{\frac{m_p}{t_p} u_e}{M_{\text{payl}} + (\delta + 1)m_p(t)} = \frac{\frac{m_p u_e}{t_p}}{M_{\text{payl}} + (1 + \delta)m_p(1 - \frac{t}{t_p})}$$

$$\implies \Delta u = \frac{-u_e}{1 + \delta} \left[\ln(M_{\text{payl}} + (1 + \delta)m_p(1 - \frac{t}{t_p})) - \ln(M_{\text{payl}} + (1 + \delta)m_p) \right] = \frac{I_{sp}g_0}{1 + \delta} \ln \left[\frac{M_{\text{payl}} + (1 + \delta)m_p}{M_{\text{payl}} + (1 + \delta)m_p(1 - \frac{t}{t_p})} \right]$$

For $t = t_p$, $\Delta u = \frac{I_{sp}g_0}{1 + \delta} \ln \left[\frac{M_{\text{payl}} + (1 + \delta)m_p}{M_{\text{payl}}} \right]$

⁵ “Payload fraction”, *Wikipedia* https://en.wikipedia.org/wiki/Payload_fraction

- (b) Payload fraction is defined as the weight of payload over the takeoff weight (cf. wikipedia ⁵. Then for part (a), the Δu , Δu_a , is

$$(\Delta u)_a = \frac{I_{sp}g_0}{1 + \delta} \ln \left[\frac{1}{\lambda} \right] = \frac{-I_{sp}g_0}{1 + \delta} \ln \lambda$$

- (i) $M = M(t) = M_{\text{payl}} + M_s + m_p(t)$ s.t. $m_p(0) = M_p$ if assume constant mass flow out. $m_p(t) = M_p(1 - \frac{t}{t_p})$
 $m_p(t_p) = 0$

$$M\dot{u} = -\dot{m}_p u_{\text{exh}} \implies \dot{u} = \frac{\frac{M_p}{t_p} u_e}{M_{\text{payl}} + M_s + M_p(1 - \frac{t}{t_p})}$$

So

$$u(t) = -I_{sp}g_0 [\ln(M_{\text{payl}} + M_s + M_p(1 - \frac{t}{t_p})) - \ln(M_{\text{payl}} + M_s + M_p)]$$

$$\Delta u = I_{sp}g_0 \ln \left[\frac{M_{\text{payl}} + M_s + M_p}{M_{\text{payl}} + M_s} \right]$$

Then it’s just algebra to put the ratio of the masses above in terms of λ, δ (there’s 3 unknowns, M_{payl} , M_s , M_p masses, and we’re given λ, δ and a ratio we want, $\frac{M_{\text{payl}} + M_s + M_p}{M_{\text{payl}} + M_s}$). Instead of doing the algebra entirely by hand, let’s use Python and the sympy library:

```
import sympy
from sympy import *
>>> M_payl = Symbol('M_payl', positive=True)
>>> M_s = Symbol('M_s', positive=True)
>>> ratio_bi = (M_payl + M_s)/(M_payl + M_s + M_p)
>>> payloadfrac = Symbol('payloadfrac', positive=True)
>>> delta = Symbol('delta', positive=True)
>>> ratio_bi_new = ratio_bi.subs(M_payl, (M_s + M_p)/(1/payloadfrac - 1)).subs(M_s, M_p*delta)
>>> ratio_bi_new.expand().factor(M_p).simplify().factor()
(delta + payloadfrac)/(delta + 1)
```

and so

$$(\Delta u)_i = I_{sp}g_0 \ln \left[\frac{1 + \delta}{\lambda + \delta} \right]$$

(ii)

$$(\Delta u)_{ii} = I_{sp}g_0 \ln \left[\frac{M_{\text{payl}} + M_p}{M_{\text{payl}}} \right] = I_{sp}g_0 \ln \left(\frac{1}{\lambda} \right)$$

(iii) For 1 (total number of) stage(s),

$$\Delta u = I_{sp}g_0 \ln \left[\frac{1 + \delta}{\lambda + \delta} \right]$$

as in part (b)(i).
Let N be the number of stages.

7.2. Lagrangian point of view for gravity-free, drag-free rocket. Let $M = M(t) = M(0) - m(t)$ s.t. $m(0) = 0$ where

$$m(t_p) = M_p$$

M_p is the total mass of the propellant to be used, and $M(0)$ represents the (initial) mass of the propellant + spacecraft or payload or the interesting part we want to launch out. t_p is the burn time of the propellant.

Then kinetic energy of M at instantaneous time t is $\frac{1}{2}Mu^2$. Also note that $\int_0^{t_p} \dot{m} dt = M_p$.

Now take the exterior derivative of m : $dm = \dot{m}dt$. Consider the infinitesimal piece of mass dm at the (instantaneous) time t ; its kinetic energy will be

$$\frac{1}{2}dm(u - u_e)^2$$

Note that its velocity is $(u - u_e)$ because mass dm is being expunged out of the rocket at constant exhaust velocity u_e *relative* to the rocket (i.e. change, transform, or “boost” into the instantaneous inertial reference frame of the rocket, so that the rocket has 0 velocity in this frame; now return back to the “lab” frame-the propellant expunged has gained velocity u as well).

Now

$$\frac{1}{2}dm(u - u_e)^2 \xrightarrow{\int_0^t dt} \frac{1}{2} \int_0^t dm(u - u_e)^2 = \frac{1}{2} \int_0^t \dot{m}d\tau(u - u_e)^2$$

The full Lagrangian at time t is

$$\mathcal{L} = \frac{1}{2}Mu^2 + \frac{1}{2} \int_0^t \dot{m}d\tau(u - u_e)^2$$

Notice that \mathcal{L} in this specific case does not depend on position. So $\frac{\partial \mathcal{L}}{\partial x^i} = 0$.

Now

$$\frac{\partial \frac{1}{2}Mu^2}{\partial u^i} = Mu_i \xrightarrow{\frac{d}{dt}} \frac{d}{dt} \frac{\partial \frac{1}{2}Mu^2}{\partial u^i} = \dot{M}u_i + M\dot{u}_i = -\dot{m}u_i + M\dot{u}_i$$

and

$$\begin{aligned} \frac{\partial}{\partial u^i} \frac{1}{2} \int_0^t \dot{m}dt(u - u_e)^2 &= \frac{1}{2} \int_0^t \dot{m}d\tau 2(u_i - u_e) = \int_0^t \dot{m}d\tau(u_i - u_e) \\ \xrightarrow{\frac{d}{dt}} \frac{d}{dt} \frac{\partial}{\partial u^i} \frac{1}{2} \int_0^t \dot{m}dt(u - u_e)^2 &= \frac{d}{dt} \int_0^t \dot{m}d\tau(u_i - u_e) = \dot{m}(t)(u_i - u_e) \end{aligned}$$

for clearly, if $\int_0^t \dot{m}d\tau(u_i - u_e) = F(t) - F(0)$, then applying the total time derivative, will result in $F'(t)$ and so we can read off the antiderivative, namely $\dot{m}(t)(u_i - u_e)$ at time t .

The Euler-Lagrange equation tells us that $\frac{\partial \mathcal{L}}{\partial x^i} - \frac{d}{dt} \frac{\partial \mathcal{L}}{\partial \dot{x}^i} = 0$ and so

$$-\dot{m}u_i + M\dot{u}_i + \dot{m}u_i - \dot{m}u_e = 0 \text{ or } M\dot{u}_i = \dot{m}u_e$$

$$\implies \boxed{\dot{u}_i = \frac{\dot{m}u_e}{M}}$$

Problem 4. Ballistic trajectories with atmospheric drag

(a) Now

$$M_0\dot{u} = -M_0g + F_d$$

$$\dot{u} = -g + \frac{F_d}{M_0}$$

In components,

$$\begin{aligned} \dot{u}_x &= \frac{F_d}{M_0} \left(\frac{-u_x}{|u|} \right) \\ \dot{u}_y &= -g + \frac{F_d}{M_0} \left(\frac{-u_y}{|u|} \right) \end{aligned}$$

And so, for

$$F_d = \frac{1}{2}\rho C_D u^2 A = \frac{1}{2}\rho C_D (u_x^2 + u_y^2) A$$

and

$$\begin{aligned} \rho &= \rho_0 e^{-y/\lambda} \\ \ddot{x} &= \frac{1}{2M_0} \rho_0 e^{-y/\lambda} C_D (\dot{x}^2 + \dot{y}^2) A \left(\frac{-\dot{x}}{(\dot{x}^2 + \dot{y}^2)^{1/2}} \right) \\ \ddot{y} &= -g + \frac{1}{2M_0} \rho_0 e^{-y/\lambda} C_D (\dot{x}^2 + \dot{y}^2) A \left(\frac{-\dot{y}}{(\dot{x}^2 + \dot{y}^2)^{1/2}} \right) \end{aligned}$$

Note that we can rewrite this as the following system of equations:

$$\begin{aligned} \dot{u}_x &= \frac{1}{2M_0} \rho_0 C_D A e^{-y/\lambda} (u_x^2 + u_y^2)^{1/2} (-u_x) \\ \dot{u}_y &= -g_0 + \frac{1}{2M_0} \rho_0 C_D A e^{-y/\lambda} (u_x^2 + u_y^2)^{1/2} (-u_y) \\ \dot{x} &= u_x \\ \dot{y} &= u_y \end{aligned}$$

In `Propulsion.py`, (just run it with `python -i Propulsion.py` in its directory)

```
import scipy
from scipy import exp, array
from scipy.integrate import ode

import matplotlib.pyplot as plt

M_cannonball = (7.8*(10**2)**3/(10**3))*4./3.*N(pi)*(15./2./100.）**3
(1.225)*(0.1)/(2.*M_cannonball)*(N(pi)*(15./2./100.）**2) # 7.85256410256411e-5

# to use scipy.integrate.ode, the ODE system must be declared as a "callable", a Python function
def deriv(t,u): # return derivatives of the array u
    """
    cf. http://bulldog2.redlands.edu/facultyfolder/deweerd/tutorials/Tutorial-ODEs.pdf
    """
    uxdot = (7.853*10**(-5))*exp(-u[3]/(10000.))*(u[0]**2 + u[1]**2)**(0.5)*(-u[0])
    uydots = -9.8 + (7.853*10**(-5))*exp(-u[3]/(10000.))*(u[0]**2 + u[1]**2)**(0.5)*(-u[1])
    return array([ uxdot, uydots, u[0], u[1] ])

# initial conditions
u0 = [300.*cos(50./180.*N(pi)), 300.*sin(50./180.*N(pi)), 0, 0]

# declare the ODE to be integrated
Prob0203 = ode(deriv).set_integrator('dopri5') # Problem 3 from Problem Set 2 for AE121 Fall 2015
# cf. http://stackoverflow.com/questions/26738676/does-scipy-integrate-ode-set-solout-work
Prob0203.set_initial_value(u0)

t1 = 41.575
dt = 0.005
# print out the solution to the ODE for various times as a sanity check
while Prob0203.successful() and Prob0203.t < t1:
    Prob0203.integrate(Prob0203.t+dt)
    print("%g_" % Prob0203.t)
    print Prob0203.y

# store the solutions in a Python list for plotting
Prob0203.set_initial_value(u0)
Prob0203.solution = []
while Prob0203.successful() and Prob0203.t < t1:
    Prob0203.solution.append([ Prob0203.t+dt, ] + list( Prob0203.integrate(Prob0203.t+dt) ) )
# take the transpose of a list of lists
Prob0203.solution = map(list, zip(*Prob0203.solution))

plt.figure(1)
```

```
plt.plot( Prob0203_solution[3] , Prob0203_solution[4])
plt.xlabel( 'x_(m) ' )
plt.ylabel( 'y_(m) ' )
plt.title( 'Cannonball_trajectory_with_Drag:_Variable_density' )
```

Horizontal range is 6.11 *km*

```
(b) def deriv_b(t,u): # return derivatives of the array u
    """
    """
    """cf. http://bulldog2.redlands.edu/facultyfolder/deweerd/tutorials/Tutorial-ODEs.pdf
    """
    uxdot = (7.853*10**(-5)) *(u[0]**2 + u[1]**2)**(0.5)*(-u[0])
    uydots = -9.8 + (7.853*10**(-5)) *(u[0]**2 + u[1]**2)**(0.5)*(-u[1])
    return array([ uxdot , uydots , u[0] , u[1] ])

Prob0203b = ode(deriv_b).set_integrator( 'dopri5' )
Prob0203b.set_initial_value(u0)
Prob0203b.integrate(41.23)

t1b = 41.225
Prob0203b.set_initial_value(u0)
Prob0203b.solution = []
while Prob0203b.successful() and Prob0203b.t < t1b:
    Prob0203b.solution.append( [Prob0203b.t+dt,] + list( Prob0203b.integrate(Prob0203b.t+dt) ) )
Prob0203b.solution = map(list , zip(*Prob0203b.solution))

plt.figure(2)
plt.plot( Prob0203b_solution[3] , Prob0203b_solution[4])
plt.xlabel( 'x_(m) ' )
plt.ylabel( 'y_(m) ' )
plt.title( 'Cannonball_trajectory_with_Drag:_Constant_density' )
```

Horizontal range is 5.89 *km*. This makes sense because the cannonball finds it easier to fly “through the air” at higher altitudes, higher up the atmosphere, because the “air is thinner” in the “upper atmosphere.”

(c) For no drag, this can be solved analytically:

$$\begin{aligned} \dot{u}_x = 0 &\implies u_x = u_0 \cos \theta & x(t) = u_0 \cos \theta t \\ \dot{u}_y = -g &\implies u_y = -gt + u_0 \sin \theta & y(t) = -\frac{1}{2}gt^2 + u_0 \sin \theta t = t(u_0 \sin \theta - \frac{gt}{2}) \end{aligned}$$

So the horizontal range is $x(t_f) = u_0 \cos \theta \left(\frac{2u_0 \sin \theta}{g} \right) = \frac{u_0^2}{g} \sin(2\theta) = 9044.m$ for

```
>>> 300.**2/9.8*sin(2.*50./180.*N(pi) )
9044.15283378558
```

```
#parabola trajectory data
Prob0203c_x = [i*10 for i in range(905)]
Prob0203c_y = [ tan(50./180.*N(pi))*x - (9.8/2.)*x**2/(300.*cos(50./180.*N(pi)))*x**2 for x in Prob0203c_x]
```

```
# plot all 3 trajectories together
plt.figure(3)
plt.plot( Prob0203_solution[3] , Prob0203_solution[4] , label="Drag:_Variable_density")
plt.plot( Prob0203b_solution[3] , Prob0203b_solution[4] , label="Drag:_Constant_density")
plt.plot( Prob0203c_x , Prob0203c_y , label="No_Drag")
plt.xlabel( 'x_(m) ' )
plt.ylabel( 'y_(m) ' )
plt.title( 'Trajectories_of_cannonball_with_Drag_of_variable_density ,_Drag_of_constant_density ,_and_no_drag' )
plt.legend()
```

If there was no drag, then the cannonball will fly out farther, and higher. It’s important to consider air resistance, as the horizontal range difference between drag and no drag is almost 3000 m (!!!). It’s important to consider variation of atmospheric drag with altitude for horizontal range for precision landing (about a 300 m difference).

7.3. PS4. Problem 1: Kinetic theory connection to thermodynamics properties.

- (a)
- (b) Now the Maxwellian velocity distribution, $P(v)$, where $P(v)dv$ is the probability that the particle has speed in $(v, v + dv)$, is given by

$$P(v) = 4\pi \left(\frac{M}{2\pi\tau} \right)^{3/2} v^2 \exp \left(\frac{-Mv^2}{2\tau} \right)$$

If N particles had the same kinetic energy, then the entire system of N particles would have a total internal energy of $N\frac{1}{2}Mv^2$, with M being the mass of a single particle.

Thus, the total internal energy U is calculated by weighting by $P(v)$ the above total kinetic energy, which in this case of only 3 translational degrees of freedom, coincides with the total internal energy:

$$\begin{aligned} U &= N \int_0^\infty dv \left(\frac{1}{2}Mv^2 \right) P(v) = \int_0^\infty dv \frac{MN}{2} \cdot 4\pi \left(\frac{M}{2\pi\tau} \right)^{3/2} v^4 \exp \left(\frac{-Mv^2}{2\tau} \right) = 2\pi MN \left(\frac{M}{2\pi\tau} \right)^{3/2} \int_0^\infty dv v^4 \exp \left(\frac{-Mv^2}{2\tau} \right) = \\ &= 2\pi MN \left(\frac{M}{2\pi\tau} \right)^{3/2} \frac{3}{4 \left(\frac{M}{2\tau} \right)^2} \frac{\sqrt{\pi}}{2 \left(\frac{M}{2\tau} \right)^{1/2}} = 2\pi MN \frac{3}{4} \frac{1}{\left(\frac{M}{2\tau} \right)} \frac{\sqrt{\pi}}{2\pi^{3/2}} = \frac{3\tau N}{2} \end{aligned}$$

Also, $\frac{U}{N} = \frac{3\tau}{2}$.

EY : 20151117 I want to reiterate that there must be a more systematic and sane and rational way of looking up Physical Constants and other physical data than by manually looking it up a book or manually looking it up a website. People can make a mistake copying and pasting! Thus, I wrote the Python package **Physique** that you copy into your working directory and can import in (this is all in the script **Propulsion.py**:

```
import Physique
from Physique import FCconv , KCconv , FundConst , conv , plnfacts , T_C , T_K , T_F
```

```
k_Boltz = FundConst[ FundConst["Quantity"].str.contains("Boltzmann") ].loc[49,: ]
>>> k_Boltz.Value
Decimal('1.38064852E-23')
>>> k_Boltz.Unit
'J.K^-1'
```

So for $T = 300K$ and $T = 1000K$,

```
>>> k_Boltz.Value *300*Decimal(1.5)
Decimal('6.21291834000E-21')
>>> k_Boltz.Value *1000*Decimal(1.5)
Decimal('2.07097278000E-20')
```

or

```
>>> k_Boltz.Value *300*Decimal(1.5)/ JovereV.Multiplyby
Decimal('0.03877797733958233079116726804')
>>> k_Boltz.Value *1000*Decimal(1.5)/ JovereV.Multiplyby
Decimal('0.1292599244652744359705575601')
```

and so $U/N = 6.213 \times 10^{-21}J$ or $0.0388eV$ for $T = 300K$ and $2.071 \times 10^{-20}J$ or $0.129eV$ for $T = 1000K$

(c) Now

$$C_V = \left(\frac{\partial U}{\partial \tau} \right)_V = \frac{3N}{2}$$

From the definition of C_V . Also, for enthalpy H , $H = U + pV = U + N\tau$, for ideal gas law still holds,

$$C_P = \left(\frac{\partial H}{\partial \tau} \right)_V = \left(\frac{\partial U}{\partial \tau} \right)_V + N = C_V + N$$

and so

$$C_P = \frac{5N}{2}$$

These are all, above, physicists’ quantities. For engineers, *specific heat capacities* are useful with real-world material. *Specific* heat capacities are obtained from physicists’ heat capacities by dividing by what physicists would’ve deemed as M , the (straight-up) mass.

$$c_V := \frac{C_V}{M} = \frac{3N}{2M}$$
$$c_P := \frac{C_P}{M} = \frac{5N}{2M}$$

```
N_Avog = FundConst[FundConst[‘‘Quantity’’].str.contains(‘‘Avogadro’’) ]
>>> c_V = float( Decimal(1.5)*(N_Avog.Value)*(k_Boltz.Value))/M.0
>>> c_P = float( Decimal(2.5)*(N_Avog.Value)*(k_Boltz.Value))/M.0
>>> c_V.subs(M.0, 39.948/1000.)
312.198102337360
>>> c_V.subs(M.0, 131.293/1000.)
94.9912774647001
>>> c_P.subs(M.0, 39.948/1000.)
520.330170562267
>>> c_P.subs(M.0, 131.293/1000.)
158.318795774500
```

So c_V and c_P for argon is $312.2J/(kgK)$ and $520.3J/(kgK)$, respectively, and c_V and c_P for xeon is $94.99J/(kgK)$ and $158.3J/(kgK)$, respectively.

The so-called *molar heat capacity* is the amount of heat needed to increase the temperature of 1 mole of substance. Now, physicists’ C_V and C_P is the heat capacities for the amount of heat needed to raise the temperature of a system of N number of particles. Then certainly, dividing C_V, C_P by N will result in the amount of heat needed to raise the temperature of a *single* particle. Use Avogadro’s number to convert between number of particles and moles.

$$c_V = \frac{C_V}{N} = \frac{3}{2}$$
$$c_P = \frac{C_P}{N} = \frac{5}{2}$$

```
>>> Decimal(1.5)*(N_Avog.Value)*(k_Boltz.Value)
42      12.471689792172872460
Name: Value, dtype: object
>>> Decimal(2.5)*(N_Avog.Value)*(k_Boltz.Value)
42      20.786149653621454100
```

and so $c_V = 12472.J/(\text{kmol}K)$ and $c_P = 20786.J/(\text{kmol}K)$

Problem 2: Mean thermal velocity.

Wikipedia article “Thermal velocity” has 3 mean thermal velocities ⁶

- $P(v) = 4\pi \left(\frac{M}{2\pi\tau}\right)^{3/2} v^2 \exp\left(\frac{-Mv^2}{2\tau}\right)$

Now

$$\int_0^\infty dv v^4 \exp(-\alpha v^2) = \int_0^\infty v^3 \left(\frac{\exp(-\alpha v^2)}{-2\alpha}\right)' = 0 - \int 3v^2 \frac{\exp(-\alpha v^2)}{-2\alpha} = \int_0^\infty \frac{3}{2\alpha} v^2 \exp(-\alpha v^2) =$$
$$= \frac{3}{2\alpha} \int v \left(\frac{e^{-\alpha v^2}}{-2\alpha}\right)' = \frac{3}{2\alpha} \left[0 - \int \frac{e^{-\alpha v^2}}{-2\alpha}\right] = \frac{3}{4\alpha^2} \frac{\sqrt{\pi}}{2\sqrt{\alpha}}$$

So

$$\langle v^2 \rangle = \int_0^\infty v^2 P(v) = 4\pi \left(\frac{M}{2\pi\tau}\right)^{3/2} \frac{3}{4 \left(\frac{M}{2\tau}\right)^2} \frac{\sqrt{\pi}}{\sqrt{\frac{M}{2\tau}}} \frac{1}{2} = \frac{3}{\frac{M}{2}} \frac{1}{2} = \frac{3\tau}{M}$$

⁶“Thermal Velocity”, *wikipedia*, https://en.wikipedia.org/wiki/Thermal_velocity

so

$$v_{rms} = \left(\frac{3\tau}{M}\right)^{1/2}$$

Now

$$\langle v^2 \rangle = 3\langle v_x^2 \rangle = \frac{3\tau}{M} \text{ so } \langle v_x^2 \rangle^{1/2} = \left(\frac{\tau}{M}\right)^{1/2}$$

•

$$\frac{\partial P(v)}{\partial v} = \frac{2P(v)}{v} + P(v) \left(\frac{-M}{\tau}\right) v = P(v) \left[\frac{2}{v} - \frac{M}{\tau} v\right] = 0 \text{ if } v_{mp} = 0 \text{ or } v_{mp} = \sqrt{\frac{2\tau}{M}}$$

where mp stands for most probable.

$$v_{mp} = \sqrt{\frac{2\tau}{M}} < \sqrt{\frac{3\tau}{M}} = v_{rms}$$

•

$$\bar{c} = \int_0^\infty dv v P(v) =: \langle v \rangle = \int_0^\infty dv 4\pi \left(\frac{M}{2\pi\tau}\right)^{3/2} v^3 \exp\left(\frac{-Mv^2}{2\tau}\right) = 4\pi \left(\frac{M}{2\pi\tau}\right)^{3/2} \int_0^\infty dv v^2 \left(\frac{e^{-\frac{Mv^2}{2\tau}}}{\frac{-M}{\tau}}\right)' =$$
$$= 4\pi \left(\frac{M}{2\pi\tau}\right)^{3/2} \left[0 - \int_0^\infty dv 2v \frac{e^{-\frac{Mv^2}{2\tau}}}{\frac{-M}{\tau}}\right] = 4\pi \left(\frac{M}{2\pi\tau}\right)^{3/2} \left[\frac{2\tau}{M} \left(\frac{e^{-\frac{Mv^2}{2\tau}}}{\frac{-M}{\tau}}\right)\right]_0^\infty = 4\pi \left(\frac{M}{2\pi\tau}\right)^{3/2} \frac{2\tau^2}{M^2} =$$
$$= \frac{1}{(2\pi)^{3/2}} 8\pi \sqrt{\frac{\tau}{M}} = \sqrt{\frac{8\tau}{\pi M}}$$

Note

$$\frac{v_{rms}}{\bar{c}} = \frac{\left(\frac{3\tau}{M}\right)^{1/2}}{\left(\frac{8\tau}{\pi M}\right)^{1/2}} = \left(\frac{3\pi}{8}\right)^{1/2}$$

Problem 3: Ideal vs. real rocket analysis.

7.4. PS5. Problem 1: Nozzle flow in liquid rocket engines.

Viking series liquiad rocket engines used on first 2 stages of Ariane 4 launch vehicle.
Rocket engines are storable propellant motors that use
nitrogen tetroxide and UDMH25 (unsymmetrical dimethyl hydrazine with 25 percent hydrazine hydrate) as propellants.
suitable mean value for molecular weight of the combustion product mixture 23 g/mol ratio of specific heats approximately 1.2

Combustion chamber temperatures for both 3350 K approximately.
Viking 5C motor: chamber pressure 5800 kPa
propellant mass flow rate 275.2 kg/s
nozzle expansion ratio 10.5
Viking 4B engine: chamber pressure 5850 kPa
mass flow 278.0 kg/s
expansion ratio 30.8

(a) Use Eq. 10,

$$\frac{\dot{m}}{A^*} = \frac{p_0 \gamma^{1/2}}{(RT_0)^{1/2}} \left(\frac{2}{\gamma + 1}\right)^{\frac{\gamma+1}{2(\gamma-1)}}$$

Opening up `NozzleTheory.py`, using Python’s `sympy` library,


```
Viking5Cnozzle = massflowrateExp.subs(gamma, 1.2).subs(massflow, 275.2).subs(p_0,5800*1000).subs(T_0,3350).\
subs(R, k_Boltz.Value/ (Decimal(23/1000.)/N_Avog.Value ))
Viking4Bnozzle = massflowrateExp.subs(gamma, 1.2).subs(massflow, 278.0).subs(p_0,5850*1000).subs(T_0,3350).\
subs(R, k_Boltz.Value/ (Decimal(23/1000.)/N_Avog.Value ))

solve(Viking5Cnozzle, Astar)[0] # 0.0805128291046479
solve(Viking4Bnozzle, Astar)[0] # 0.0806368550290290
```

So the area of throat for Viking 5C and Viking 4B is $0.08051\,m^2$ and $0.08126\,m^2$, respectively. Then, the expansion ratio gives the expanded area for Viking 5C and Viking 4B of $0.8454\,m^2$ and $2.484\,m^2$, respectively.

(b) We need a relationship relating (cross-sectional) area ratio to Mach number, Eq. 9

$$\frac{A_2}{A_1} = \frac{\mathfrak{M}_1}{\mathfrak{M}_2} \left[\left(\frac{1 + \frac{\gamma-1}{2}\mathfrak{M}_2^2}{1 + \frac{\gamma-1}{2}\mathfrak{M}_1^2} \right)^{\frac{\gamma+1}{\gamma-1}} \right]^{1/2}$$

At the throat, $\mathfrak{M} = 1$, and denote the area by A^* . Thus

$$\frac{A_e}{A^*} = \frac{1}{\mathfrak{M}} \left[\left(\frac{1 + \frac{\gamma-1}{2}\mathfrak{M}^2}{\frac{\gamma+1}{2}} \right)^{\frac{\gamma+1}{\gamma-1}} \right]^{1/2}$$

and so $\frac{A_e}{A^*}$ is the expansion ratio.

This is implemented in `NozzleTheory.py` as a sympy object Eq:

```
# Area Ratio to Mach numbers
Mach_1 = Symbol('Mach_1', positive=True)
Mach_2 = Symbol('Mach_2', positive=True)
A_1 = Symbol('A_1', positive=True)
A_2 = Symbol('A_2', positive=True)
AreastoMachs = Eq( A_2/A_1, Mach_1/Mach_2*( ( (Rat(1) + (gamma-Rat(1)))/(Rat(2) )*Mach_2**2 )/\
(Rat(1) + (gamma-Rat(1)))/(Rat(2) )*Mach_1**2 ) )**(( gamma +1)/(gamma-1) )**0.5 )
```

and one substitutes in the desired, given numbers (parameters):

```
Viking5CMachEq = AreastoMachs.subs(Mach_1,Rat(1) ).subs(gamma,1.2).subs(A_2,A_1*10.5)
# 10.5 == (0.0909090909090909*Mach_2**2 + 0.909090909090909)**5.5/Mach_2
```

Notice that now we have a so-called “root-finding” problem, with non-integer exponents. One should use a numerical solver so that finding this root is done “efficiently”. Also, to do this in Python’s `scipy`, we have to create a Python function object, and so I used sympy’s `lambdify` to turn a sympy expression into an actual Python function:

```
Viking5CMach = lambdify(Mach_2, Viking5CMachEq.rhs - Viking5CMachEq.lhs )
# Remember to move all terms to 1 side, and so the other side equals 0
```

```
scipy.optimize.newton( Viking5CMach, 3) # Newton Raphson method
# 3.3123573073570207
```

```
scipy.optimize.bisect( Viking5CMach, 3,4) # Bisection method
# 3.312357307356251
```

AIMS Senegal had a nice, easy introduction to scipy’s root finding methods ⁷.

Next, the static temperature, static pressure, static density at the nozzle exit can be easily calculated from the isentropic relations:

$$\frac{\tau}{\tau_0} = \left(1 + \frac{\gamma-1}{2}\mathfrak{M}^2 \right)^{-1} \quad \frac{p}{p_0} = \left(\frac{\tau}{\tau_0} \right)^{\frac{\gamma}{\gamma-1}} \quad \rho_0 = \frac{p_0}{RT_0} \quad \frac{\rho}{\rho_0} = \left(\frac{\tau}{\tau_0} \right)^{\frac{1}{\gamma-1}}$$

⁷Roots finding, Numerical integrations and differential equations, AIMS Senegal, AIMS Senegal

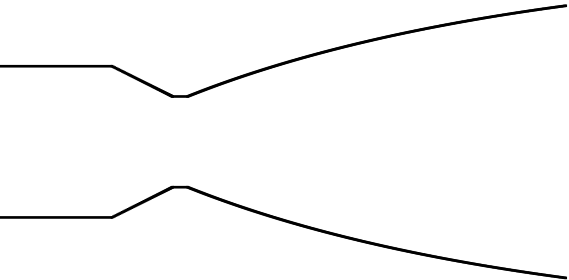
Then the static pressure, static temperature, static density for Viking 5C and Viking 4B engines are

$$\begin{aligned} \mathfrak{M}_{exh} &= 3.312 \\ p &= 68.2\,\text{kPa} \\ T &= 1597.4\,K \\ \rho &= 0.1181\,\text{kg}/m^3 \end{aligned}$$

$$\begin{aligned} \mathfrak{M}_{exh} &= 4.057 \\ p &= 17.0\,\text{kPa} \\ T &= 1265.9\,K \\ \rho &= 0.0372\,\text{kg}/m^3 \end{aligned}$$

- (c)
- (d) Look at Eq. 11 again:

$$\begin{aligned} T &= \dot{m}v_e + (p_e - p_a)A_e = \\ &= \dot{m}\sqrt{\frac{2\gamma RT_0}{\gamma-1} \left(1 - \left(\frac{p_e}{p_0} \right)^{\frac{\gamma-1}{\gamma}} \right)} + (p_e - p_a)A_e \end{aligned}$$



So if $p_e = p_a$, then

$$T = \dot{m}\sqrt{\frac{2\gamma RT_0}{\gamma-1} \left(1 - \left(\frac{p_a}{p_0} \right)^{\frac{\gamma-1}{\gamma}} \right)}$$

- (e) Look at this webpage: <http://www.engapplets.vt.edu/fluids/CDnozzle/cdinfo.html> for a good recap of the physics of shocks, accompanying the lecture by Polk in AE121. Also “Choked flow” wikipedia article https://en.wikipedia.org/wiki/Choked_flow Consider normal shock at the exit of the nozzle. Then before, $\mathfrak{M}_1 > 1$ (supersonic), and right after the shock, $\mathfrak{M}_2 < 1$ (subsonic)

Problem 3: Duct flow with heating.

- (a) In this case, I think that the correct heat capacity to use is C_p because the process occurs at the constant ambient pressure (that ambient pressure, external to the (constant area) duct, remains constant and on the duct). Now

$$Q = C_p d\tau$$

For $c_p := \frac{C_p}{MN}$, $q = mc_p dT$ (experimental physicists’ and engineers’ expression). Thus

$$\dot{q} = \dot{m}c_p dT \implies \int_{\gamma} \dot{q} = \Delta \dot{q} = \dot{m}c_p(T_f - T_i) \text{ or } \frac{\Delta \dot{q}}{\dot{m}c_p} = T_f - T_i$$

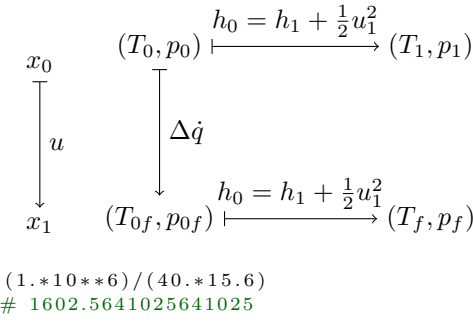
Now this heating raises the temperature of the *stagnation* (enthalpy) temperature, because we’re considering this heating as a (adiabatic?) thermodynamic process, separate from what’s going on with the flow.
For instance, along the flow, at a point, the thermodynamic property is (T, p) . We can always “pull” this “back” to the stagnation properties, whether hypothetical or real:

upstream \longrightarrow downstream

$$(T_0, p_0) \longmapsto (T, p)$$

$$\frac{T_0}{T} = 1 + \frac{\gamma - 1}{2} \mathfrak{M}^2$$
$$\frac{p_0}{p} = \left(1 + \frac{\gamma - 1}{2} \mathfrak{M}^2\right)^{\frac{\gamma}{\gamma - 1}}$$

Thus, when one considers heat addition, the thermodynamic process, $\Delta \dot{q}$ occurs on stagnation temperature, and then can be related to the actual, physical fluid flow and its properties, through formulae we’ve derived:



So

$$\Delta T_0 = 1602.6\,K$$

- (b) Now consider the converging section at the inlet to each 1 cm diameter channel, that “accelerates the flow to a relatively low Mach number and produces a flow rate of 40 g/s”.
Mass continuity (conservation) still holds:

$$\dot{m} = \rho A u = \frac{p}{RT} A \mathfrak{M} \sqrt{\gamma R T} = \frac{p}{\sqrt{RT}} A \mathfrak{M} \sqrt{\gamma} = \frac{p_0 \sqrt{\gamma}}{\sqrt{RT_0}} A \mathfrak{M} \left(1 + \frac{\gamma - 1}{2} \mathfrak{M}^2\right)^{-\frac{\gamma+1}{2(\gamma-1)}}$$

where the “pullback” to the stagnation properties for each point of the flow, before the converging section and after the converging section, was used:

$$\frac{p_0}{p} = \left(1 + \frac{\gamma - 1}{2} \mathfrak{M}^2\right)^{\frac{\gamma}{\gamma - 1}}$$
$$\frac{T_0}{T} = 1 + \frac{\gamma - 1}{2} \mathfrak{M}^2$$

```
massconsEq = Eq(massflow , p_0*sqrt(gamma/(R*T_0))*A*Mach*(1+(gamma-Rat(1))/Rat(2)*Mach**2)**\
((gamma+1)/(-Rat(2)*(gamma-1))))
massconsProb0503 = massconsEq.subs(massflow ,40.*10**(-3)).subs(gamma,1.4).subs(p_0,6.8*10**(6)).\
subs(T_0,673.).subs(A,N(pi)*(10**(-2)/2.))**2).subs(R,k_Boltz.Value/(Decimal(2.0159*10**(-3))/N_Avog.Value))
```

where for R , $R = \frac{k_B}{M}$ and where for M , I used 2.0159g/mol for H₂, and Avogadro’s number, $N_A = 6.022140857 \times 10^{23}$, which is number of particles per mole. So in this example, $R = 4124.4$ for H₂ as

```
>>> k_Boltz.Value/(Decimal(2.0159*10**(-3))/N_Avog.Value)
Decimal('4124.440627733807619868515695')
```

Let’s try to use the derived relation, relating the stagnation temperatures before and after (denoted 1) heat addition:

$$\frac{T_0}{T_{01}} = \left[\frac{1 + \gamma \mathfrak{M}_1^2}{1 + \gamma \mathfrak{M}^2} \left(\frac{\mathfrak{M}}{\mathfrak{M}_1}\right)\right]^2 \left[\frac{1 + \frac{\gamma-1}{2} \mathfrak{M}^2}{1 + \frac{\gamma-1}{2} \mathfrak{M}_1^2}\right]$$

This is implemented in `NozzleTheory.py`:

```
T_01=Symbol('T_01',real=True)
heataddTvsMachEq=Eq(T_0/T_01,((Rat(1)+gamma*Mach_1**2)/(Rat(1)+gamma*Mach**2)*Mach/Mach_1)**2*\
(Rat(1)+(gamma-Rat(1))/Rat(2)*Mach**2)/(1+(gamma-Rat(1))/Rat(2)*Mach_1**2))

heataddTvsMachProb0503=heataddTvsMachEq.subs(gamma,1.4).subs(T_0,675).subs(T_01,675+(1.*10**6)/\
(40.*15.6)).subs(Mach,MachProb0503)
Mach1Prob0503lamb=lambdify(Mach_1,heataddTvsMachProb0503.rhs-heataddTvsMachProb0503.lhs)
plot(heataddTvsMachProb0503.rhs,(Mach_1,0,10))
Mach1Prob0503=scipy.optimize.newton(Mach1Prob0503lamb,0.1)
```

where \mathfrak{M} was obtained for the acceleration by the converging nozzle, T_0 is the stagnation temperature that was given, and T_{01} is obtained from part (a).

Thus,

$$\mathfrak{M}_1 = 0.2024$$

The flow through the channel isn’t thermally choked as \mathfrak{M} doesn’t become 1 by the heat addition.

- (c) After exiting the ducts, the converging-diverging nozzle results in a Mach number given by

$$\frac{A_e}{A^*} = \frac{1}{\mathfrak{M}} \left[\left(\frac{1 + \frac{\gamma-1}{2} \mathfrak{M}^2}{\frac{\gamma+1}{2}} \right)^{\frac{\gamma+1}{\gamma-1}} \right]^{1/2}$$

and from the Mach number definition and energy equation (Bernoulli invariant), we can get the exhaust velocity

$$u_e = \mathfrak{M} a = \mathfrak{M} \sqrt{\gamma R T} = \mathfrak{M} \sqrt{\gamma R \frac{T_0}{1 + \frac{\gamma-1}{2} \mathfrak{M}^2}}$$

Hydrazine monopropellant thrusters

Consider in general 2 reactions of the form

$$\sum_j \nu_j A_j = 0$$

and

$$\xi_K A_K + \sum_k \xi_k B_k = 0$$

with the convention that the stoichiometric coefficients “on the left” (reactants) are negative integers.

Clearly,

$$dN_j = \nu_j d\hat{N}$$
$$dN_k = \xi_k d\hat{N}_2$$

where

$$d\hat{N} \equiv \text{how many times reaction 1 takes place}$$

$$d\hat{N}_2 \equiv \text{how many times reaction 2 takes places}$$

since in chemical reactions, chemical species can’t be “broken down” or “destroyed” into smaller parts (we’re not considering nuclear reactions).

Then for the particular chemical species A_K ,

$$dN_K = \nu_K d\hat{N} \text{ and}$$

$$dN_{K_2} = \xi_{K_2} d\hat{N}_2$$

for chemical species K to occur in the second reaction. Thus, given the fraction X that participates in the second chemical reaction, observe, importantly, that

$$XdN_k = -dN_{K_2} = -\xi_{K_2} d\hat{N}_2$$

$$\frac{-X\nu_K}{\xi_{K_2}} d\hat{N} = d\hat{N}_2$$

and so for the other chemical species (molecules) in the second reaction, their “stoichiometric” coefficients change in the following manner:

$$(12) \quad dN_k = \xi_k \frac{X\nu_k}{-\xi_{K_2}} d\hat{N}$$

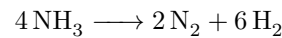
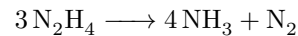
and so writing the first reaction as

$$\sum_j \nu_j A_j = 0 = \sum_{k \neq K} \nu_k A_k + \nu_K A_K$$

then we simply only need to add up reactions 1 and 2, with reaction 2 taking new stoichiometric coefficients according to Eq. 12:

$$(13) \quad \boxed{\sum_{k \neq K} \nu_k A_k + \nu_K (1 - X) A_K + \sum_l \frac{\xi_l X \nu_l}{-\xi_{K_2}} B_l = 0}$$

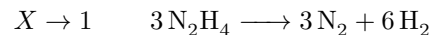
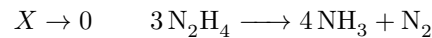
Indeed, consider the (exothermic) decomposition of hydrazine and then the dissociation of ammonia into elevated temperatures:



Then plugging into our formula, Eq. 13,

$$\begin{aligned} -3\text{N}_2\text{H}_4 + 4(1 - X)\text{NH}_3 + \text{N}_2 + \frac{2 \cdot X \cdot 4}{4} \text{N}_2 + \frac{6 \cdot X \cdot 4}{4} \text{H}_2 &= 0 = -3\text{N}_2\text{H}_4 + 4(1 - X)\text{NH}_3 + (1 + 2X)\text{N}_2 + 6X\text{H}_2 \\ \implies \boxed{\text{N}_2\text{H}_4 \longrightarrow 4(1 - X)\text{NH}_3 + (1 + 2X)\text{N}_2 + 6X\text{H}_2} \end{aligned}$$

Indeed, for the following limits,



adiabatic flame temperature

Remember that for adiabatic processes, $Q = 0$, and so $Q = 0 = \Delta H = H_{\text{products}} - H_{\text{reactants}}$ and so for adiabatic processes, $H_{\text{products}} = H_{\text{reactants}}$, given by

$$(14) \quad \begin{aligned} H_P(\tau_P) &= \sum_{i \in \{\text{products}\}} n_i \left[\Delta_f \bar{h}_i^0 + \bar{h}_i(\tau_P) - \bar{h}_i^0 \right] \\ H_R(\tau_R) &= \sum_{i \in \{\text{reactants}\}} n_i \left[\Delta_f \bar{h}_i^0 + \bar{h}_i(\tau_R) - \bar{h}_i^0 \right] \end{aligned}$$

where P denotes products and R denotes reactants.

The upper bound of the adiabatic flame temperature is given by no dissociation ($X = 0$). The lower bound of the adiabatic flame temperature is given by the flame temperature with full dissociation ($X = 1$).

EY : 20160121 My question is this: is this because the dissociation of ammonia is endothermic, as the heat of formation of ammonia is $\Delta H_f^\circ = -45.894 \text{ kJ/mol}$, and so heat is given off in its formation, and so for full dissociation, the temperature goes

down. Is this also because of Le Chatelier’s principle? A clear explanation according to Le Chatelier’s principle would also be appreciated.

Remember that the heat of formation for stable elements is 0.

For the case of $X = 0$, $Q = 0$ leads to

$$3 \cdot \left[\Delta_{f;\text{N}_2\text{H}_4}^\circ + 0 \right] = 4 \cdot \left[\Delta_{f;\text{NH}_3}^\circ + (H - H(T^\circ))_{\text{NH}_3} \right] + 1 \left[0 + (H - H(T^\circ))_{\text{N}_2} \right]$$

For the case of $X = 1$, $Q = 0$ leads to

$$3 \cdot \left[\Delta_{f;\text{N}_2\text{H}_4}^\circ + 0 \right] = 3 \left[0 + (H - H(T^\circ))_{\text{N}_2} \right] + 6 \left[0 + (H - H(T^\circ))_{\text{H}_2} \right]$$

We were given a quadratic polynomial fit to data taken from JANAF (or JANNAF?) tables for $\text{N}_2, \text{H}_2, \text{NH}_3$:

$$\text{N}_2 \quad H - H(T^\circ) = -8.553 + 27.77\theta + 2.317\theta^2$$

$$\text{H}_2 \quad H - H(T^\circ) = -8.292 + 27.39\theta + 1.586\theta^2$$

$$\text{NH}_3 \quad H - H(T^\circ) = -10.37 + 31.32\theta + 11.77\theta^2$$

where $\theta = T/1000$ and T is in Kelvins (K).

Solving for a quadratic function is easy with `scipy` using the module `optimize` and the method `newton`. The code is in `combustion.py`.

The upper bound of the adiabatic flame temperature T_{ad} was found to be

$$T_{ad} = 1644 \text{ K} = 1371 \text{ C} = 2498 \text{ F}$$

and the lower bound of the adiabatic flame temperature was found to be

$$T_{ad} = 866 \text{ K} = 593 \text{ C} = 1099 \text{ F}$$

Gibbs function for a mixture of gases

Let Σ be a 2-dimensional (i.e. $\dim \Sigma = 2$) manifold representing all possible thermodynamic states of the system (in this case, of an ideal gas). Let Σ have coordinates τ, p , representing the temperature τ , and pressure p , of the system, respectively, which could’ve been obtained from (successive) Legendre transformations of (global) coordinates U, V , where U is energy.

Let Q , representing heat applied onto, or heat into, the system, be a 1-form of Σ , i.e. $Q \in \Omega^1(\Sigma)$.

Then Q can be written, in general, as

$$Q = C_p d\tau + \frac{\partial Q}{\partial p} dp$$

Let’s determine $\frac{\partial Q}{\partial p}$. From energy conservation (i.e. “First law” of thermodynamics), $Q = dU + pdV$ (i.e. Q in (global) coordinates of U, V for σ). Next, consider a curve $\gamma : \mathbb{R} \rightarrow \Sigma$ in Σ , representing a thermodynamic process, such that $d\tau(\dot{\gamma}) = 0$ which says that the process occurs at constant temperature τ . Note that $\dot{\gamma}$ is a vector field in Σ , i.e. $\dot{\gamma} \in \mathfrak{X}(\Sigma)$. Then

$$Q(\dot{\gamma}) = C_p d\tau(\dot{\gamma}) + \frac{\partial Q}{\partial p} dp(\dot{\gamma}) = 0 + \frac{\partial Q}{\partial p} dp(\dot{\gamma}) = dU(\dot{\gamma}) + pdV(\dot{\gamma})$$

For an ideal gas, $U = \frac{1}{\gamma-1} N\tau$ and so $dU = \frac{1}{\gamma-1} N d\tau$ and so $dU(\dot{\gamma}) = \frac{1}{\gamma-1} N d\tau(\dot{\gamma}) = 0$. It also sufficed to say that for an ideal gas, $U = U(\tau)$, i.e. U is only a function of temperature, and so U is constant during a process where temperature remains constant.

Then, using $pV = N\tau$ which must always hold for an ideal gas in equilibrium,

$$V = \frac{N\tau}{p} \implies dV = \frac{N\tau}{-p^2} dp$$

and so

$$\frac{\partial Q}{\partial p} dp = 0 + p \left(\frac{N\tau}{-p^2} \right) dp = \frac{-N\tau}{p} dp = -V dp$$

Thus $\frac{\partial Q}{\partial p} = -V$. One can (physically) interpret this as showing how heat, applied to this ideal gas system, can cause our system to be able to do (physical) work.

Then equating $Q = \tau d\sigma$,

$$\tau d\sigma = Q = C_p d\tau - V dp \implies d\sigma = \frac{C_p}{\tau} d\tau - \frac{V}{\tau} dp = \frac{C_p d\tau}{\tau} - \frac{N dp}{p}$$

Taking the integral,

$$\sigma - \sigma_0 = \int_{\tau_0}^{\tau} \frac{C_p d\tau}{\tau} - N \ln \frac{p}{p_0}$$

Let’s talk about, for a mixture of ideal gases, partial pressures p_i for each of the (kinds of) species i . It makes (physical) sense that the contribution to the total pressure p from a single species labeled i would be proportional to the number of particles N_i of the species i (imagine (infinitesimally small) colliding billiard balls). Then $p_i = \frac{N_i}{N}p$ and so

$$N_i \ln \frac{p_i}{p_0} = N_i \ln \frac{N_i p}{N p_0} = N_i \ln \frac{p}{p_0} + N_i \ln \frac{N_i}{N}$$

For a mixture of ideal gases, there is no interaction energy (by definition of *ideal*), and so entropies are *additive* [10]. So

$$\sigma = \sum_i \sigma_i = \sum_i \left(\int_{\tau_0}^{\tau} \frac{C_{pi} d\tau}{\tau} + \sigma_{i0} \right) - N \ln \frac{p}{p_0} - \sum_i N_i \ln \frac{N_i}{N}$$

σ in its current state is unitless. Multiply by the Boltzmann constant k_B and Avogadro’s number N_A as appropriate to obtain the desired units for entropy S .

7.4.1. *ig thermo.xls; Dealing with tables.* Doing a search on Google for `ig_thermo.xls` yields this link:

http://www.google.com/url?q=http://shepherd.caltech.edu/EDL/projects/pde/ig_thermo.xls&sa=U&ved=0ahUKEWjt2ta8x7nKAhVB8GMKHa-tCBYQFggUMAA&usg=AFQjCNGcu2ryEOaIZqwumHooxDjcxTf3gQ

However, ammonia is not to be found in `ig_thermo.xls` *off the Web*.

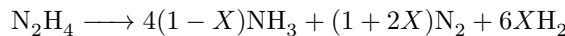
Let’s consider the work-flow or process of scraping the NIST website for JANAF table data. Starting from the **NIST-JANAF Thermochemical Tables** webpage, one can enter, as a string, the CAS number, chemical formula, or compound name, or search by the periodic table. However, using the periodic table, for element N or H, ammonia wasn’t to be found (as of 20160120). However, entering NH₃ or Ammonia (the search text box “field” is case-sensitive), leads to the ammonia gas JANAF table. Then one clicks on the link for “**view**” for the HTML-formatted table and then follow the bottom link for the tab-separated table.

So it would be nice to do the following:

- Obtain a *comprehensive* list of all available compounds of this NIST JANAF database and with that
- scrape all the tab-separated tables.

A comprehensive list appears to be in this pdf link for **Volume 1** of the JANAF tables. Perhaps this pdf could be “scraped” to obtain a list of compounds in text format?

7.4.2. *entropy of a mixture of ammonia, nitrogen, and hydrogen.* Recall that



and that we had derived

$$S(T, P) = \sum_i N_i \left(\int_{T_0}^T c_{pi} \frac{dT}{T} + s_{i0} \right) - NR \ln \frac{P}{P_0} - R \sum_i N_i \ln \frac{N_i}{N}$$

The values we want are for $P = 1$ atm and $P = 10$ atm

Note that

$$N = 4(1 - X) + (1 + 2X) + 6X = 4X + 5$$

and

$$\sum_i N_i \ln \frac{N_i}{N} = 4(1 - X) \ln \left(\frac{4(1 - X)}{4X + 5} \right) + (1 + 2X) \ln \left(\frac{1 + 2X}{4X + 5} \right) + 6X \ln \left(\frac{6X}{4X + 5} \right)$$

8. EQUILIBRIUM FLOW VS. FROZEN FLOW FOR NOZZLE FLOW AND THE EXAMPLE OF THE SPACE SHUTTLE MAIN ENGINE (SSME)

(1) Remember that from Newton’s 2nd law, the so-called momentum equation, is

$$\rho \frac{Du}{Dt} = \text{div}(\mathbf{T}) \text{ or } \rho \left[\frac{\partial \mathbf{u}}{\partial t} + u^j \frac{\partial \mathbf{u}}{\partial x^j} \right] = \text{div}(\mathbf{T})$$

Let

$$k := \frac{1}{2} \rho u^2$$

which is a scalar quantity that depends on time t and spatial coordinates $\mathbf{x} = (x^1 \dots x^n) \in N$ and represents the density of kinetic energy of the bulk fluid at time t and spatial point \mathbf{x} .

Clearly, product rule still applies when using the material derivative $\frac{D}{Dt} := \frac{\partial}{\partial t} + u^j \frac{\partial}{\partial x^j}$:

$$\frac{D}{Dt} k = \frac{D}{Dt} \frac{1}{2} \rho u^2 = \frac{1}{2} \left[u^2 \frac{D\rho}{Dt} + \rho \left(\frac{D}{Dt} u^2 \right) \right] = \frac{1}{2} \left[u^2 (-\rho \text{div} u) + \rho \left(\frac{D}{Dt} u^2 \right) \right]$$

If, say, $g = 1$, then $u^2 = u_i u^i = (u^i)^2$, and so

$$\rho \frac{D}{Dt} u^2 = 2\rho u^i \frac{D}{Dt} u^i = 2u^i \text{div}(T^i)$$

and so in this case,

$$\frac{D}{Dt} k = \frac{-1}{2} \rho u^2 \text{div}(u) + u^i \text{div}(T^i) = -k \text{div}(u) + u^i \text{div}(T^i)$$

and since

$$u^j \frac{\partial k}{\partial x^j} + k \text{div}(u) = \text{div}(ku)$$

then we could also write

$$\frac{D}{Dt} k + k \text{div}(u) = \frac{\partial k}{\partial t} + \text{div}(ku) = u^i \text{div}(T^i)$$

What about the case where $g \neq 1$ but could be dependent upon time and space?

Invoke the “Fundamental Theorem on Time-Dependent Flows” (Theorem 9.48 in Chapter 9 Integral Curves and Flows, pp. 237 of Lee (2012) [12], which guarantees the existence of a time-dependent flow ϕ for our velocity vector field \mathbf{u} .

So for our time-dependent velocity vector field, $\mathbf{u} \in \mathfrak{X}(\mathbb{R} \times N)$, i.e.

$$\mathbf{u} : \mathbb{R} \times N \rightarrow TN$$

$$\mathbf{u} = \mathbf{u}(t, \mathbf{x}) = u^i(t, \mathbf{x}) \frac{\partial}{\partial x^i}$$

there exists a flow, given initial conditions, $\phi = \phi(t) \in N$, such that

$$\phi(t) = (x^1(t) \dots x^n(t))$$

$$\dot{\phi}(t) = (\dot{x}^1(t) \dots \dot{x}^n(t)) = (u^1 \dots u^n)$$

Thus, again emphasizing on any Riemannian manifold (M, g) , equipped with metric g ,

$$u^2 = g_{ij} u^i u^j = g_{ij} \dot{\phi}^i \dot{\phi}^j$$

Consider any time and space dependent scalar quantitiy, f , $f \in C^\infty(\mathbb{R} \times N)$, i.e. $f : \mathbb{R} \times N \rightarrow \mathbb{R}$. On a flow ϕ for a time-dependent vector field \mathbf{u} generating this flow,

$$\begin{aligned} \frac{d}{dt} f &= \frac{d}{dt} f(t, \mathbf{x}) = \frac{d}{dt} f(t, \mathbf{x}(t)) = \frac{\partial}{\partial t} f + \frac{\partial f}{\partial x^i} \dot{x}^i = \frac{\partial}{\partial t} f + u^i \frac{\partial f}{\partial x^i} = \\ &= \frac{D}{Dt} f \end{aligned}$$

Thus, in this scenario, then $\forall f \in C^\infty(\mathbb{R} \times N)$, the material derivative $\frac{D}{Dt}$ of f is equal to the total time derivative $\frac{d}{dt}$.

Then for scalar quantity $u^2 = u^2(t, \mathbf{x}) = g_{ij}(\phi(t))\dot{\phi}^i\dot{\phi}^j$,

$$\frac{d}{dt}u^2 = \frac{\partial g_{ij}}{\partial x^k}\dot{\phi}^k\dot{\phi}^i\dot{\phi}^j + 2g_{ij}\dot{\phi}^i\ddot{\phi}^j = 2g_{ks}\Gamma_{ij}^k\dot{\phi}^i\dot{\phi}^j\dot{\phi}^s + 2g_{sk}\dot{\phi}^s\dot{\phi}^k = 2g_{ks}\dot{\phi}^s\left[\ddot{\phi}^k + \Gamma_{ij}^k\dot{\phi}^i\dot{\phi}^j\right]$$

where the identity

$$\frac{\partial g_{ij}}{\partial x^k}\dot{\phi}^k\dot{\phi}^i\dot{\phi}^j = 2g_{ks}\Gamma_{ij}^k\dot{\phi}^i\dot{\phi}^j\dot{\phi}^s$$

was used, which is in the proof for Theorem 3.25 on pp. 49 of section 3.8 The natural Lagrangian on manifolds of Calin and Chang (2005) [13]. I want to remark that Calin and Chang (2005) [13] has explicit computations showing this identity to be the case locally, and at least for me, I don't see many other sources that are this thorough with presenting explicit computations.

Noting that

$$\nabla_u u = \nabla_{u^j \frac{\partial}{\partial x^j}} u = u^j \nabla_{\frac{\partial}{\partial x^j}} u = u^j \left(\frac{\partial u^k}{\partial x^j} + \Gamma_{ij}^k u^i \right) \frac{\partial}{\partial x^k} = \left(\ddot{\phi}^k + \Gamma_{ij}^k \dot{\phi}^i \dot{\phi}^j \right) \frac{\partial}{\partial x^k}$$

and so

$$g_{ks}\dot{\phi}^s \left[\ddot{\phi}^k + \Gamma_{ij}^k \dot{\phi}^i \dot{\phi}^j \right] = \langle u, \nabla_u u \rangle \implies \frac{d}{dt}u^2 = 2\langle u, \nabla_u u \rangle \quad (17)$$

Thus, in general (on a Riemannian manifold),

$$(16) \quad \boxed{\frac{\partial k}{\partial t} + \text{div}(ku) = \frac{Dk}{Dt} + k\text{div}(u) = \rho\langle u, \nabla_u u \rangle = \langle u, \text{div}(T) \rangle}$$

Now when following a volume element vol^n , along the flow of the fluid, one should be concerned that in a “lab frame”, the volume itself can change with time:

$$\begin{aligned} \dot{V} &= \frac{d}{dt} \int_V \text{vol}^n = \int_V \mathcal{L}_{\frac{\partial}{\partial t} + \mathbf{u}} \text{vol}^n = \int_V 0 + \mathcal{L}_{\mathbf{u}} \text{vol}^n = \int_V (\mathbf{d}i_{\mathbf{u}} + i_{\mathbf{u}}\mathbf{d})\text{vol}^n = \int_V \mathbf{d}i_{\mathbf{u}}\text{vol}^n + 0 = \\ &= \int_{\partial V} i_{\mathbf{u}}\text{vol}^n = \int_{\partial V} u^j dS_j \end{aligned}$$

Indeed, if the time-dependent velocity vector field \mathbf{u} changes at the boundary ∂V , then so does our volume, in this frame.

The correct way out of this conundrum is to switch to the frame where the bulk velocity of the fluid is 0, so it's at rest in this frame. This is paramount to choosing $\mathbf{u} = 0$ above, and so $\dot{V} = 0$.

The strategy is this: one should recognize that number of particles N , or in general, number of particles for species i , N_i and entropy σ are invariants under Galilean (or Lorentz) transformation: observers, whether in the lab frame, or the frame in which the bulk velocity of the fluid is 0 so the bulk fluid flow is at rest, or any other frame, must agree upon these numbers (that can be objectively measured). So we've fixed our “control volume” that remains constant in time in this fluid-at-rest frame, and count the number of particles that come in, and the entropy.

For only 1 kind of particle, let

$$\check{\sigma} \equiv \sigma/N \text{ entropy per particle}$$

$$\check{e} \equiv E/N \text{ energy per particle}$$

Thus

$$\tau d\sigma = \tau d(N\check{\sigma}) = \tau N d\check{\sigma} + \tau \check{\sigma} dN$$

where

$$\tau N d\check{\sigma} \text{ is entropy change due to change in entropy per particle, i.e. } \quad \mathbf{conduction term}$$

$$\tau \check{\sigma} dN \text{ is entropy change due to change in number of particles, i.e. } \quad \mathbf{convection term}$$

Generalizing this to \mathcal{N} species, indexed by $i = 1 \dots \mathcal{N}$,

$$\check{\sigma}_i \equiv \sigma_i/N_i \text{ entropy per particle of species } i$$

Assuming that entropies are additive (which is a reasonable assumption if the interaction between species is negligible [10]), then for total entropy of the system in volume V in this fluid-at-rest frame (which will be denoted with prime as needed),

$$\sigma = \sum_i \sigma_i = \sum_i \check{\sigma}_i N_i$$

So

$$\tau d\sigma = \tau \sum_i N_i d\check{\sigma}_i + \tau \sum_i \check{\sigma}_i dN_i$$

where

$\tau \sum_i N_i d\check{\sigma}_i$ is entropy change due to change in entropy per particle for each species, i.e. **conduction term**

$\tau \sum_i \check{\sigma}_i dN_i$ is entropy change due to change in number of particles for each species, i.e. **convection term**

Using the above result for the decomposition of $\tau d\sigma$ into a conduction term and convection term, then

$$dE = \tau d\sigma - pdV + \sum_i \mu_i dN_i = \tau \sum_i N_i d\check{\sigma}_i + \tau \sum_i \check{\sigma}_i dN_i - pdV + \sum_i \mu_i dN_i =$$

$$= \tau \sum_i N_i d\check{\sigma}_i + \sum_i (\tau \check{\sigma}_i + \mu_i) dN_i - pdV = \tau \sum_i N_i d\check{\sigma}_i + \sum_i \check{h}_i dN_i - pdV = \tau \sum_i N_i d\check{\sigma}_i + \sum_i \check{h}_i dN_i$$

since $dV = 0$ for any thermodynamic process in this constant volume and that

$$G_i = F_i + p_i V = E_i - \tau \sigma_i + p_i V = H_i - \tau \sigma_i = \mu_i N_i \implies \check{h}_i = \mu_i + \tau \check{\sigma}_i$$

since Gibbs free energy G_i for each species i should be additive, partial pressures p_i for each species i should be additive to the total pressure p , as forces are additive, and from the definition of enthalpy $H_i = E_i + p_i V$ for each species i .

For the case of 1 species, and for this volume at rest in the fluid-at-rest frame, V , then defining

$$\begin{aligned} h &:= \frac{H}{V} \\ s &:= \frac{\sigma}{V} \\ \epsilon &:= \frac{E}{V} \end{aligned}$$

and so from Eq. 17,

$$d\epsilon = \tau \frac{N}{V} d\left(\frac{\sigma}{N}\right) + \frac{H}{N} d\left(\frac{N}{V}\right) = \tau \frac{N}{V} \left(\frac{d\sigma}{N} + -\frac{\sigma dN}{N^2} \right) + \frac{H}{N} \frac{d\sigma}{M} = \tau ds - \frac{\tau s}{\rho} d\rho + \frac{hd\rho}{\rho} = \tau ds + \frac{h - \tau s}{\rho} d\rho$$

We should denote with prime symbols ($'$) the quantities that transform under change of observer's frame, and is specific, in value, to this fluid-at-rest frame:

$$d\epsilon' = \tau ds + \frac{h' - \tau s}{\rho} d\rho$$

For many species, then generalizing the above,

$$d\epsilon' = \tau \sum_i \frac{N_i}{V} d\left(\frac{\sigma_i}{N_i}\right) + \sum_i \frac{H_i}{N_i} d\frac{N_i}{V} = \tau \sum_i \frac{N_i}{V} \left(\frac{d\sigma_i}{N_i} + -\frac{\sigma_i dN_i}{N_i^2} \right) + \sum_i \frac{h'_i}{\rho_i} d\rho_i = \tau ds + \sum_i \frac{h'_i - \tau s_i}{\rho_i} d\rho_i$$

Thus, in this fluid-at-rest frame,

$$(18) \quad \boxed{d\epsilon' = \tau ds + \sum_i \frac{h'_i - \tau s_i}{\rho_i} d\rho_i}$$

Next, use the fact that for a fluid in the “lab frame”, the current density for the total energy E is carried by $(k + h')\mathbf{u}$ and *not* $(k + \epsilon')\mathbf{u}$. This comes from the fact that for compressible flows, the Bernoulli invariant is $k + h' = \frac{1}{2}\rho u^2 + h'$,

and not $k + \epsilon'$. The physical interpretation is that the enthalpy h' is needed to account for all convection terms, to account for the “energy balancing” needed, done by the compression of the volume by $vdN_i := V/N_i dN_i$ to return to the initial volume V in the fluid-at-rest frame, when particles of species i flow through the volume V , convection.

Thus

$$(19) \quad \frac{\partial \epsilon}{\partial t} + \text{div}((k + h')\mathbf{u}) = 0 \text{ or } \frac{\partial \epsilon}{\partial t} + \text{div}(ku) = -\text{div}(h'u)$$

With all these ingredients, begin with the account of all energies in the lab frame, for a fluid:

$$E = KE + U \implies \epsilon = k + \epsilon' \implies \frac{\partial \epsilon}{\partial t} = \frac{\partial k}{\partial t} + \frac{\partial \epsilon'}{\partial t}$$

From Eq. 16, to substitute into $\frac{\partial k}{\partial t}$,

$$\frac{\partial \epsilon}{\partial t} = -\text{div}(ku) + \langle u, \text{div}T \rangle + \frac{\partial \epsilon'}{\partial t}$$

From Eq. 18, we can consider a thermodynamic process where we vary the time, generated by vector $\frac{\partial}{\partial t}$, so that we have

$$\frac{\partial \epsilon'}{\partial t} = \tau \frac{\partial s}{\partial t} + \sum_i \frac{h'_i - \tau s_i}{\rho_i} \frac{\partial \rho_i}{\partial t}$$

and so

$$\frac{\partial \epsilon}{\partial t} = -\text{div}(ku) + \langle u, \text{div}T \rangle + \tau \frac{\partial s}{\partial t} + \sum_i \frac{h'_i - \tau s_i}{\rho_i} \frac{\partial \rho_i}{\partial t} = -\text{div}(ku) - \text{div}(h'u)$$

where in the last equality, we used Eq. 19. Eliminating $-\text{div}(ku)$ from both sides, noting that

$$\text{div}(h'u) = \frac{1}{\sqrt{g}} \frac{\partial(h'u^j \sqrt{g})}{\partial x^j} = u^j \frac{\partial h'}{\partial x^j} + h' \text{div}(u)$$

and using $T = -p1$, the case where the stress tensor is only the isotropic (same in all directions) pressure p , then

$$(20) \quad -u^j \frac{\partial p}{\partial x^j} + \tau \frac{\partial s}{\partial t} + \sum_i \frac{h'_i - \tau s_i}{\rho_i} \frac{\partial \rho_i}{\partial t} = -u^j \frac{\partial h'}{\partial x^j} - h' \text{div}(u)$$

Now we'll need to use a mass conservation law for each of the species i , which I'm not sure is valid (please correct me if I'm wrong):

$$\frac{\partial \rho_i}{\partial t} + \text{div}(\rho_i u) = 0 \implies \frac{\partial \rho_i}{\partial t} = -\rho_i \text{div}u - u^j \frac{\partial \rho_i}{\partial x^j}$$

and so

$$\sum_i \frac{h'_i - \tau s_i}{\rho_i} \frac{\partial \rho_i}{\partial t} = \sum_i -h'_i \text{div}u + \tau s_i \text{div}u - \frac{h'_i - \tau s_i}{\rho_i} u^j \frac{\partial \rho_i}{\partial x^j} = -h' \text{div}u + \tau s \text{div}u - \sum_i \frac{h'_i - \tau s_i}{\rho_i} u^j \frac{\partial \rho_i}{\partial x^j}$$

Plugging this into Eq. 20,

$$-u^j \frac{\partial p}{\partial x^j} + \tau \frac{\partial s}{\partial t} + -h' \text{div}u + \tau s \text{div}u - \sum_i \frac{h'_i - \tau s_i}{\rho_i} u^j \frac{\partial \rho_i}{\partial x^j} = -u^j \frac{\partial h'}{\partial x^j} - h' \text{div}(u)$$

Noting that by definition, $h' := H'/V = \epsilon' + p$ Using Eq. 18 one more time so that

$$\frac{\partial \epsilon'}{\partial x^j} u^j = \left[\tau \frac{\partial s}{\partial x^j} + \sum_i \frac{h'_i - \tau s_i}{\rho_i} \frac{\partial \rho_i}{\partial x^j} \right] u^j$$

and thus

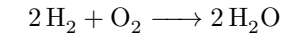
$$\tau \frac{\partial s}{\partial t} + \tau s \text{div}u + \tau u^j \frac{\partial s}{\partial x^j} = 0 \implies \boxed{\frac{\partial s}{\partial t} + \text{div}(su)}$$

Thus, entropy is conserved over time. Indeed

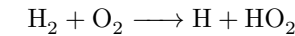
$$\dot{S} = \frac{d}{dt} \int_V s \text{vol}^n = \int_V \left(\frac{\partial s}{\partial t} + \text{div}(su) \right) \text{vol}^n$$

Also notice that I did not assume incompressibility. This, and the statements above, are true for compressible and incompressible flow. Also, note that all the above was proven for any Riemannian manifold N representing space.

(2) I supposed the reaction was



But the reaction listed for `h2o2.cti` or `h2o2_highT.cti` is



9. LIQUID-VAPOR EQUILIBRIUM

From wikipedia's article on “Clausius-Clapeyron relation”, start from the definition of particle diffusion equilibrium, that

$$\mu_v = \mu_l$$

$$d\mu_v = d\mu_l$$

Using the Gibbs-Duhem relation for each side of the above, that $Nd\mu = Vdp - \sigma d\tau$, then

$$d\mu_v = v_v dp - s_v d\tau$$

$$d\mu_l = v_l dp - s_l d\tau$$

Thus

$$(v_v - v_l)dp = (s_v - s_l)d\tau \implies \boxed{\frac{dp}{d\tau} = \frac{s_v - s_l}{v_v - v_l}}$$

Using the definition of enthalpy $H := U + pV$, so that

$$dH = dU + Vdp + pdV = \tau d\sigma + Vdp \text{ or } \frac{1}{\tau}(dH - Vdp) = d\sigma$$

where $\tau d\sigma = dU + pdV$ was used.

Considering the thermodynamic process at constant pressure $dp = 0$ that begins with a single molecule being in a liquid state and changing into vapor state, then

$$\frac{1}{\tau}(\Delta H_v) = \sigma_v - \sigma_l$$

Thus

$$\frac{dp}{dT} = \frac{\Delta H_v}{T(V_v - V_l)} \approx \frac{\Delta H_v}{TV_v}$$

since it's reasonable to assume that the volume of the vapor molecule is much greater than when the molecule is in a liquid state.

Using the ideal gas law for a single molecule, $pV = \tau$, then

$$\frac{dp}{dT} = \frac{p\Delta H_v}{T^2} \text{ or } \frac{dp}{p} = \frac{\Delta H_v dT}{T^2}$$

$$\implies p(T) = p_0 \exp\left(\frac{-\Delta H_v}{T}\right)$$

Part 4. Basic Feeling

10. BOX WITH A HOLE ROCKET; BOTTLED (BOX) ROCKET

Recall Ch. 14: Kinetic Theory, Section “Kinetic Theory of the Ideal Gas Law” of Kittel and Kroemer [4].

Kinetic Theory of the Ideal Gas Law. Consider molecule strike unit area of wall.

Let $v_z \equiv$ velocity component normal to plane of wall.

Suppose molecules, of mass M , reflected specularly (mirror-like) from wall,

$$\Delta p_z = -2M|v_z|$$

Let $a(v_z)dv_z$, number of molecules per unit volume with z -component of velocity between v_z and $v_z + dv_z$.

$$\int a(v_z)dv_z = \frac{N}{V} = n$$

$a(v_z)v_zdv_z$ number of molecules in $(v_z, v_z + dv_z)$ velocity range that strike unit area of wall in (per) unit time

$$\text{pressure } p = \int_0^\infty 2Mv_z a(v_z)v_z dv_z = M \int_{-\infty}^\infty v_z^2 a(v_z)dv_z = Mn\langle v_z^2 \rangle$$

$\frac{1}{2}M\langle v_z^2 \rangle = \frac{1}{2}\tau$ by equipartition of energy (Ch.3)

$$p = nM\langle v_z^2 \rangle = n\tau = \frac{N\tau}{V}$$

Maxwell Distribution of Velocities. cf. Ch.6. distribution function of ideal gas $f(\epsilon_n) = \lambda \exp\left(\frac{-\epsilon_n}{\tau}\right)$

Recall, Ch. 6, Sec. “Classical Limit” of Kittel and Kroemer [4], **an ideal gas is defined as a system of free noninter-acting particles in the classical regime.**

$f(\epsilon) \equiv$ average occupancy of an orbital at energy ϵ

$\epsilon \equiv$ energy of orbital occupied by 1 particle; not energy of system of N particles

Fermi-Dirac and Bose-Einstein distribution $f(\epsilon) = \frac{1}{\exp[(\epsilon - \mu)/\tau] \pm 1}$

In order for $f(\epsilon) \ll 1 \forall$ orbitals, $\exp[(\epsilon - \mu)/\tau] \gg 1 \forall \epsilon$.

$$\implies f(\epsilon) \simeq \exp[(\mu - \epsilon)/\tau] = \lambda \exp(-\epsilon/\tau) \quad \lambda \equiv \exp\left(\frac{\mu}{\tau}\right)$$

$f(\epsilon)$, average occupancy of orbital of energy ϵ , is classical distribution function.

particle in a box: $\epsilon_n = \frac{1}{2M} \left(\frac{\pi n}{L}\right)^2$ (for, recall $\frac{1}{2M} \left(\frac{1}{i}\partial\right)^2 \psi = \frac{-1}{2M} \partial^2 \psi = E\psi$)

number of orbitals in range of quantum number $(n, n + dn)$, probability such orbital is occupied

$$\left(\frac{1}{2}\pi n^2 dn\right) f(\epsilon_n) = \frac{1}{2}\pi n^2 \lambda \exp(-\epsilon_n/\tau) dn$$

$$\frac{1}{2}Mv^2 = \frac{1}{2M} \left(\frac{\pi n}{L}\right)^2 \quad \text{or } n^2 = \frac{(ML)^2}{\pi^2} v^2 \quad \text{or } n = \frac{MLv}{\pi}$$

Consider system of N particles in volume V .

Let $NP(v)dv$ number of atoms with velocity magnitude in range dv at v

$$NP(v)dv = \frac{1}{2}\pi n^2 \lambda \exp(-\epsilon_n/\tau) \frac{dn}{dv} dv = \frac{1}{2}\pi \lambda \left(\frac{ML}{\pi}\right)^3 v^2 \exp\left(\frac{-Mv^2}{2\tau}\right) dv$$

cf. Ch. 6, $\lambda = \frac{n}{n_Q} = \frac{N}{L^3} \left(\frac{2\pi\hbar^2}{M\tau}\right)^{3/2}$

$$\frac{1}{2}\pi \frac{N}{L^3} \left(\frac{2\pi}{M\tau}\right)^{3/2} \left(\frac{ML}{\pi}\right)^3 = 4\pi N \left(\frac{N}{2\pi\tau}\right)^{3/2}$$

$$(21) \quad \implies P(v) = 4\pi \left(\frac{M}{2\pi\tau}\right)^{3/2} v^2 \exp\left(\frac{-Mv^2}{2\tau}\right)$$

$P(v)$ is **Maxwell velocity distribution**, $P(v)dv$ is probability particle has speed in dv at v

Experimental verification. velocity distribution of atoms which exit from slit of oven.

exit beam weighted in favor of atoms of high velocity at expense of those at low velocity.

weight factor is velocity component $v \cos \theta$ normal to plane of hole.

$$\begin{aligned} \int (\cos \theta) drrd\varphi r \sin \theta d\theta &= \left(\frac{1}{3}\right)(2\pi)R^3 \int \cos \theta \sin \theta d\theta = \left(\frac{2\pi}{3}R^3\right) \int \frac{\sin 2\theta d\theta}{2} = \frac{2\pi R^3}{3} \left(\frac{-\cos 2\theta}{4}\right) \Big|_0^{\pi/2} = \\ &= \frac{2\pi R^3}{3} \left(\frac{1+1}{4}\right) = \frac{2\pi R^3}{3} \left(\frac{1}{2}\right) \end{aligned}$$

Probability atom leaves hole will have velocity in $(v, v + dv)$: $P_{\text{beam}}(v)dv$

$$P_{\text{beam}}(v) \propto v P_{\text{Maxwell}} \propto v^3 \exp\left(\frac{-Mv^2}{2\tau}\right)$$

with, recall $P_{\text{Maxwell}} = 4\pi \left(\frac{M}{2\pi\tau}\right)^{3/2} v^2 \exp\left(\frac{-Mv^2}{2\tau}\right)$

P_{beam} distribution of transmission through a hole is Maxwell transmission distribution.

$$\begin{aligned} (22) \quad \langle v_{\text{out}} \rangle &= \int \int_0^{\pi/2} v \cos \theta \sin \theta d\theta P_{\text{Maxwell}}(v) dv = 4\pi \left(\frac{M}{2\pi\tau}\right)^{3/2} \frac{1}{2} \int_0^\infty v^3 \exp\left(\frac{-M}{2\tau}v^2\right) dv = \\ &= 4\pi \left(\frac{M}{2\pi\tau}\right)^{3/2} \frac{1}{2} \left(\frac{1}{-2\alpha}\right) \left[0 - \frac{1}{\alpha}\right] = 4\pi \left(\frac{M}{2\pi\tau}\right)^{3/2} \frac{1}{4} \left(\frac{4\tau^2}{M^2}\right) = \boxed{\frac{2^{1/2}}{\pi^{1/2}} \left(\frac{\tau}{M}\right)^{1/2}} \end{aligned}$$

for (doing the integration by hand)

$$(e^{-\alpha v^2})' = -2\alpha v e^{-\alpha v^2}$$

$$(v^2 e^{-\alpha v^2})' = -2\alpha v^3 e^{-\alpha v^2} + 2v e^{-\alpha v^2}$$

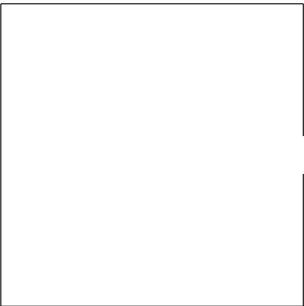
$$(v^2 e^{-\alpha v^2} + \frac{e^{-\alpha v^2}}{\alpha})' = -2\alpha v^3 e^{-\alpha v^2}$$

Armed with this mean velocity $\langle v_{\text{out}} \rangle$ out of a hole of a box, we want the *thrust* that results on a box if we had a box of air, at some pressure, and at some temperature, and then we punch a hole at one end.

What’s happening? Air that’s swirling around the box is then accelerated out of the hole. There’s fluid flow out. For low enough velocities, use *Bernoulli’s equation*, and assume at the starting point of the air’s streamline, the velocity is 0:

$$\begin{aligned} \frac{1}{2}u^2 + \frac{p}{\rho} &= \frac{1}{2}u_f^2 + \frac{p_f}{\rho_f} \\ \implies \frac{p}{\rho} &= \frac{1}{2}u_f^2 + \frac{p_f}{\rho_f} \end{aligned}$$

ρ is really $\frac{MN}{V}$ and by the ideal gas law (still applies), $pV = N\tau$ and $\frac{p}{\tau} = \frac{N}{V}$ but the point is the gas density didn’t change much.



Thus,

$$\frac{p - p_f}{\rho} = \frac{1}{2} u_f^2$$

The thrust is going to be given by the difference in pressure against the walls, the wall in front of the box opposite to the wall with a hole. *You don't need the area of the hole.* This thrust is $(p - p_f)$

$$(p - p_f)A = \rho \frac{1}{2} u_f^2 A = \frac{MN}{V} \frac{1}{2} u_f^2 A = \frac{MN}{L} \frac{1}{2} u_f^2$$

Now from Eq. [22](#), $u_f^2 = \frac{2}{\pi} \frac{\tau}{M}$, and so

$$(p - p_f)A = \frac{N\tau}{\pi L} = \frac{pV}{\pi L} = \frac{pL^2}{\pi}$$

Thus, the thrust on this box is given by

$$F_{\text{thrust}} = \frac{N\tau}{\pi L} = \frac{pL^2}{\pi}$$

Part 5. Combustion

11. MASS FRACTION VS. MOLE FRACTION VS. MOLECULAR MASS I.E. “MOLECULAR WEIGHT”

This follows Powers (2014) [8], Section 2.1. Some general issues, Chapter 2. Gas Mixtures, as some of the following concepts need to be clarified.

Consider a mixture of $\mathcal{N} \in \mathbb{Z}^+$, each “a pure substance”.

Let the total mass of the mixture be M such that (\equiv s.t.)

$$M = \sum_{j=1}^{\mathcal{N}} m_j$$

where $m_j \equiv$ is the mass of the j th substance.

Let N be the total number of *particles* s.t.

$$N = \sum_{j=1}^{\mathcal{N}} N_j$$

where N_j is the (total) number of particles of the j th substance.

Define

$$N_j/N_A =: n_j$$

with n_j be the number of moles of the j th substance, and where N_A is *Avogadro's number*, which really is a conversion *factor* between units of moles to units for number of particles.

Then

$$n = \sum_{j=1}^{\mathcal{N}} n_j$$

is the (total) number of moles of the mixture, or the “total number” (pp. 73, Powers (2014)) [8]. Define

$$(23) \quad Y_j := \frac{m_j}{M} \quad (\text{mass fraction})$$

Y_j to be the mass fraction.

Define

$$(24) \quad X_j := \frac{n_j}{n} = \frac{N_j}{N} \quad (\text{mole fraction or particle fraction})$$

X_j to be the mole fraction or particle fraction or just fraction.

Define the *molecular mass* or “molecular weight” \hat{m}_j to be

$$(25) \quad \hat{m}_j := \frac{m_j}{n_j} = \frac{m_j}{N_j} N_A$$

Then the mass fraction Y_j in terms of molecular mass and in terms of mole fraction, respectively, is

$$Y_j = \frac{m_j}{m} = \frac{\hat{m}_j n_j}{\sum_{\xi=1}^{\mathcal{N}} n_{\xi} \hat{m}_{\xi}} = \frac{\hat{m}_j X_j}{\sum_{\xi=1}^{\mathcal{N}} X_{\xi} \hat{m}_{\xi}}$$

Mole fraction in terms of mass fraction is

$$X_j = \frac{n_j}{n} = \frac{m_j/\hat{m}_j}{\sum_{k=1}^{\mathcal{N}} m_k/\hat{m}_k} = \frac{Y_j/\hat{m}_j}{\sum_{k=1}^{\mathcal{N}} Y_k/\hat{m}_k}$$

The mean molecular mass (of the mixture) is

$$\widehat{M} := \frac{M}{n} = \sum_{j=1}^{\mathcal{N}} \frac{n_j}{n} \hat{m}_j = \sum_{j=1}^{\mathcal{N}} X_j \hat{m}_j$$

Example 2.1 of Powers (2014) [8]

I will now follow the second edition (2000) of Turns [9] (note that there is a third edition for 2011, but I only have the second edition available to me).

On the end of subsection “Ideal-Gas Mixtures” (pp. 15) of Chapter 2 Combustion and Thermochemistry of Turns (2000) [9], recall the Sackur-Tetrode equation for monatomic ideal gas [4]

$$\sigma = N \left[\ln \left(\frac{n_Q}{n} \right) + \frac{d}{2} + 1 \right] \quad \text{where} \quad n \equiv N/V \quad n_Q = \left(\frac{M\tau}{2\pi\hbar^2} \right)^{d/2}$$

Now $\forall j$ th substance, let $n_{Q;j} := \left(\frac{M_j\tau}{2\pi\hbar^2} \right)^{d/2}$, since each substance will have a different mass M_j for each of its molecules, and

let $n = n_j = \frac{N_j}{V} = \frac{p_j}{\tau}$. Then

$$\sigma_j = N_j \left[\ln (n_{Q;j}\tau) - \ln p_j + \frac{d}{2} + 1 \right]$$

where \mathcal{N} is the total number of substances in the mixture.

If we were to add these entropies up for each of the substances, assuming the additivity of entropies, and thus assuming that the interaction between the disparate substances is negligible,

$$\sum_{j=1}^{\mathcal{N}} \sigma_j = \sum_{j=1}^{\mathcal{N}} N_j [\ln n_{Q;j} - \ln p_j] + N_j [\ln \tau + \frac{d}{2} + 1] = \sum_{j=1}^{\mathcal{N}} N_j (\ln n_{Q;j} - \ln p_j) + N (\ln \tau + 1 + \frac{d}{2})$$

If we wanted to also include internal degrees of freedom (e.g. vibrational and rotational degrees of freedom), the expression for the entropy becomes (derived in [thermo.pdf](#), notes on thermodynamics by me)

$$\sigma = N \left[\ln \left(\frac{n_Q}{n} \right) + \left(\frac{d}{2} + 1 \right) \right] + \sigma_{\text{int}}$$

and so

$$\sum_{j=1}^{\mathcal{N}} \sigma_j = \sum_{j=1}^{\mathcal{N}} N_j \ln n_{Q;j} + \sigma_{\text{int};j} - N_j \ln p_j + N (\ln \tau + 1 + \frac{d}{2})$$

with

$$\sigma_j = N_j \ln n_{Q;j} + \sigma_{\text{int};j} - N_j \ln p_j + N_j (\ln \tau + 1 + \frac{d}{2})$$

We can make a *formal* definition, with respect to a reference pressure p_{ref} , of σ_j so that

(26)
$$\sigma_j(\tau, p_j) = \sigma_j(\tau, p_{\text{ref}}) - N_j \ln \frac{p_j}{p_{\text{ref}}}$$

recovering Eq. (2.17a) of Turns (2000) [9].

12. ENTHALPY

12.1. **Stoichiometry.** cf. pp. 18, Chapter 2 Combustion and Thermochemistry, Section Reactant and Product Mixtures, Subsection Stoichiometry of Turns (2000) [9].

(27)
$$C_x H_y + a(O_2 + 3.76 N_2) \longrightarrow x CO_2 + \left(\frac{y}{2}\right) H_2 O + 3.76 a N_2$$

and so $a = x + \frac{y}{4}$.

Therefore, the *stoichiometric air-fuel ratio* in this case is

$$\left(\frac{A}{F}\right)_{\text{stoic}} \equiv \left(\frac{m_{\text{air}}}{m_{\text{fuel}}}\right)_{\text{stoic}} = \frac{4.76 a \, MW_{\text{air}}}{1 \, MW_{\text{fuel}}}$$

where MW_{air} ; MW_{fuel} are molecular weights of air and fuel.

To reproduce Table 2.1 on pp. 19 of Turns (2000) [9], consider the following: reusing Eq. [27](#), then for methane, CH_4 , $x = 1, y = 4$ in this case, and so $a = 2$. For $\text{H}_2 + \text{O}_2$, consider $2\text{H}_2 + \text{O}_2 \longrightarrow 2\text{H}_2\text{O}$. For $\text{C(s)} + \text{air}$, consider a “1 to 1” reaction, of 1 C and 1 “air” reacting with each other.

Then consider how the O/F ratio is **defined**:

(28)
$$\frac{O}{F} := \frac{m_{\text{Ox}}}{m_F} = \frac{\nu_{\text{Ox}} MW_{\text{Ox}}}{MW_F} \equiv \frac{\nu_{\text{Ox}} \hat{m}_{\text{Ox}}}{\nu_F \hat{m}_F}$$

Thus, for example, for CH_4 and air, $\nu_{\text{CH}_4} = 1, \nu_{\text{air}}$, for $\text{H}_2 + \text{O}_2$, $\nu_{\text{H}_2} = 2$ and $\nu_{\text{O}_2} = 1$, and $\nu_{\text{C(s)}} = 1, \nu_{\text{air}} = 1$. Then look up the molecular masses or “molecular weights” for the constituents.

I suggest using the Python library **thermopy3** in this [github repository](#) in the following manner (and the following code snippet is in **combustion.py** of my github repository for [Propulsion](#):

```
import thermopy3
from thermopy3 import nasa9polynomials as nasa9p
nasa9pDB = nasa9p.Database() # this initializes or ‘‘instantiates’’ the database that contains the
# coefficients and other data using NASA 9 polynomials
```

```
# find if a compound is in the database using the list_compound method
# then initialize or ‘‘create’’ the compound by the method set_compound
nasa9pDB.list_compound("methane")
CH4 = nasa9pDB.set_compound("methane")
```

```
nasa9pDB.list_compound("N2")
N2 = nasa9pDB.set_compound('N2')
```

```
nasa9pDB.list_compound("O2")
```

```
O2 = nasa9pDB.set_compound("O2")
```

```
nasa9pDB.list_compound("H2")
H2 = nasa9pDB.set_compound("H2")
```

```
nasa9pDB.list_compound("C")
C = nasa9pDB.set_compound("C")
```

```
CH4.molecular_weight # 0.016042459999999998 kg/mol
O2.molecular_weight # 0.0319988 kg/mol
N2.molecular_weight # 0.0280134 kg/mol
airMW = O2.molecular_weight + 3.76*N2.molecular_weight # 0.137329184 kg/mol
```

```
# O/F for oxidizer air and fuel methane
2*airMW / CH4.molecular_weight # 17.12071390547335
```

```
# O/F for oxidizer oxygen and fuel hydrogen
O2.molecular_weight / (2.*H2.molecular_weight) # 7.936682739051927
```

```
# O/F for oxidizer air and fuel Carbon
airMW / C.molecular_weight # 11.433903436102808
```

equivalence ratio Φ

$$\Phi := \frac{(A/F)_{\text{stoic}}}{(A/F)} = \frac{(F/A)}{(F/A)_{\text{stoic}}}$$

fuel rich $\Phi > 1$

lean $\Phi < 1$

So the immediate previous expressions are in Turns’ notation[9]. In the notation used above, the stoichiometric oxidizer-fuel ratio is

$$\left(\frac{m_{\text{Ox}}}{m_{\text{fuel}}}\right)_{\text{stoic}} = \frac{N_{\text{Ox}} \hat{m}_{\text{Ox}}}{N_{\text{fuel}} \hat{m}_{\text{fuel}}} = \frac{n_{\text{Ox}} \hat{m}_{\text{Ox}}}{n_{\text{fuel}} \hat{m}_{\text{fuel}}}$$

12.2. **Absolute (or standardized) Enthalpy and Enthalpy of Formation.** cf. Turns (2000) [9]

enthalpy of formation h_f - takes into account energy associated iwth chemical bonds

sensible enthalpy change δh_s - enthalpy associated only with temperature

absolute (or standard) enthalpy $\bar{h}_i(\tau) := \bar{h}_{f,i}^0(\tau) + \Delta \bar{h}_{s,i}(\tau_{\text{ref}})$ where

$\bar{h}_i(\tau) \equiv$ absolute enthalpy at temperature τ

$\bar{h}_{f,i}^0(\tau_{\text{ref}}) \equiv$ enthalpy of formation at standard reference state (τ_{ref}, p^0)

$\Delta \bar{h}_{s,i}(\tau_{\text{ref}}) \equiv$ sensible enthalpy change in going from τ_{ref} to τ

$p_{\text{ref}} = p^0 = 1 \text{ atm}$

Example 2.3.[9] gas stream at 1 atm. a mixture of CO , CO_2 , N_2 ; given $X_{\text{CO}} = 0.1$, gas stream temperature $T = 1200 \text{ K}$

$$X_{\text{CO}_2}$$

$$X_{\text{N}_2} = 1 - X_{\text{CO}} - X_{\text{CO}_2} = 0.7$$

13. THERMOCHEMISTRY OF COMBUSTION

cf. 20151203 Dr. Polk AE121a Fall 2015 Lecture

The key is how we define enthalpy. We need to agree on

- (1) how to include enthalpy due to chemical changes in addition to the sensible enthalpy (due to temperature changes)
- (2) Reference state for sensible and chemical enthalpy

Taking into account the above, then

$$H_P(T_P) = \sum_i^{\text{products}} \mathcal{N}_i \left[\underbrace{\Delta_f \bar{h}_i^0}_{\substack{\text{chemical} \\ \text{(enthalpy of formation)}}} + \underbrace{\bar{h}_i(T_P) - \bar{h}_i^0}_{\text{sensible}} \right]$$
$$H_R(T_R) = \sum_i^{\text{reactants}} \mathcal{N}_i \left[\Delta_f \bar{h}_i^0 + \bar{h}_i(T_R) - \bar{h}_i^0 \right]$$

where P denotes products and R denotes reactants, and notation is Polk’s.

To reiterate, in my notation,

$$H_P(\tau_P) = \sum_{i \in \{\text{products}\}} n_i \left[\Delta_f \bar{h}_i^0 + \bar{h}_i(\tau_P) - \bar{h}_i^0 \right]$$
$$H_R(\tau_R) = \sum_{i \in \{\text{reactants}\}} n_i \left[\Delta_f \bar{h}_i^0 + \bar{h}_i(\tau_R) - \bar{h}_i^0 \right]$$

(29)

13.1. Law of Mass Action. Let’s recall the *law of mass action*, derived from Kittel and Kroemer [4] and mentioned for how chemical reactions go in Ch. 3 Chemical Rockets pp. 77 of Oates (1997) [1].

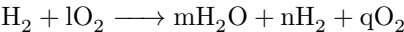
For an ideal gas, $pV = N\tau$ or $p = n\tau$,

$$\prod_j n_j^{\nu_j} = K(\tau) := \prod_j n_{Q_j}^{\nu_j} \exp(-\nu_j F_j(\text{int})/\tau) \xrightarrow{\tau^{\sum_j \nu_j}} \prod_j p_j^{\nu_j} \equiv K_p(\tau) \text{ equilibrium constant}$$
$$\prod_j X_j^{\nu_j} = K_p(\tau) p^{-(\nu_1 + \nu_2 + \cdots + \nu_n)}$$

(30)

where $\frac{p_j}{p} = \frac{n_j}{\sum_i n_i} = X_j$ is the usual mole (molar; particle) fraction for species j .

13.1.1. *Example - H-O Reaction.* Consider



H : $m + n = 1$
O : $m + 2q = 2l$
cf. Chapter 7 Simplified Conservation Equations for Reacting Flows of Turns (2011) [9]
Start from

$$m_i = \int_V \rho_i \text{vol}^n$$

(31)

and so

$$\frac{dm_i}{dt} \equiv \dot{m}_i = \int_V \frac{\partial \rho_i}{\partial t} \text{vol}^n + \int_{\partial V} \rho_i i_{u_i} \text{vol}^n = \int_V \frac{\partial \rho_i}{\partial t} \text{vol}^n + \int_{\partial V} \rho_i u_i^j dS_j = \dot{m}_i = \int_V \dot{m}_i''' \text{vol}^n$$

where we’ve defined

$$\dot{m}_i''' \equiv \text{rate of mass production per unit volume}$$

Thus

$$\implies \frac{\partial \rho_i}{\partial t} + \text{div}(\rho_i u_i) = \dot{m}_i'''$$

where

$$\text{div}(\rho_i u_i) := \frac{1}{\sqrt{g}} \frac{\partial(\rho_i u^j \sqrt{g})}{\partial x^j}$$

Now

$$\rho_i = \frac{M_i N_i}{V}$$
$$\implies \sum_i \rho_i = \rho \equiv \frac{M}{V} \text{ with } \sum_i M_i N_i \equiv M$$

where we assume negligible interaction between species.

Now

$$\sum_i \rho_i \mathbf{u}_i = \rho \mathbf{u} \text{ since}$$
$$\sum_i \rho_i \mathbf{u}_i = \sum_i \frac{M_i N_i \mathbf{u}_i}{V} = \frac{M \mathbf{u}}{V} = \rho \mathbf{u}$$

and so for

$$Y_i := \frac{m_i}{m}$$

Then

$$\mathbf{u} = \sum_i Y_i \mathbf{u}_i = \sum_i \frac{m_i}{m} \mathbf{u}_i = \sum_i \frac{M_i N_i \mathbf{u}_i}{M}$$

(32)

is the **bulk velocity** in usual fluid equations.

So taking the sum over species labeled by i , \sum_i ,

$$\sum_i \frac{\partial \rho_i}{\partial t} + \text{div}(\sum_i \rho_i u_i) = \sum_i \dot{m}_i''' \implies \frac{\partial \rho}{\partial t} + \text{div}(\rho u) = \dot{m}''' = 0$$

(33)

Use Galilean invariant, or rather *Galilean transformation*, between “lab frame” (denoted with unprimed notation), and fluid-at-rest frame, or bulk velocity-is-zero frame (primed notation).

$$\mathbf{u}'_i = \mathbf{u}_i - \mathbf{u} \text{ so } \mathbf{u}_i = \mathbf{u}'_i + \mathbf{u}$$

(34)

Thus

$$\frac{\partial \rho_i}{\partial t} + \text{div}(\rho_i(u'_i + u)) = \frac{\partial \rho_i}{\partial t} + \text{div}(\rho_i u'_i) + \text{div}(\rho_i u) = \dot{m}_i'''$$

(35)

where

$$\begin{array}{ll} \text{div}(\rho_i u'_i) & \text{diffusion} \\ \text{div}(\rho_i u) & \text{convection} \end{array}$$

(36)

Define **diffusion flux**

$$\mathbf{j}_i = \rho_i \mathbf{u}'_i$$

(37)

Now, note that

$$Y_i = \frac{m_i}{m} = \frac{M_i N_i}{\sum_j M_j N_j} \text{ for}$$
$$Y_i m = m_i$$
$$Y_i \rho = \rho_i$$

and so

$$\frac{\partial(\rho Y_i)}{\partial t} + \text{div}(\rho Y_i u'_i) + \text{div}(\rho Y_i u) = \dot{m}_i'''$$
$$\implies \rho \frac{\partial Y_i}{\partial t} + Y_i \rho \frac{\partial \rho}{\partial t} + Y_i \text{div}(\rho u) + \rho u^j \frac{\partial Y_i}{\partial x^j} = -\text{div} j_i + \dot{m}_i'''$$
$$\xrightarrow{\frac{\partial \rho}{\partial t} + \text{div}(\rho u) = 0} \rho \frac{\partial Y_i}{\partial t} + \rho u^j \frac{\partial Y_i}{\partial x^j} + \text{div} j_i = \dot{m}_i'''$$

To recap our results, armed with the definitions (and resulting identities) $\rho_i := \frac{M_i N_i}{V} \equiv M_i n_i$, $Y_i := \frac{m_i}{m} = \frac{\rho_i}{\rho}$ and so $Y_i \rho = \rho_i$, where the identity and starting from $m_i = \int_V \rho_i \text{vol}^n$, and using Galilean transformation $u_i = u'_i + u$

$$\begin{aligned} \dot{m}_i &= \int_V \left(\frac{\partial \rho_i}{\partial t} + \text{div}(\rho_i u_i) \right) \text{vol}^n = \int_V \text{vol}^n \left[\frac{\partial \rho Y_i}{\partial t} + \text{div}(\rho_i (u'_i + u)) \right] = \int_V \text{vol}^n \left[Y_i \frac{\partial \rho}{\partial t} + \rho \frac{\partial Y_i}{\partial t} + \text{div}(\rho_i u'_i) + \text{div}(\rho Y_i u) \right] = \\ &= \int_V \text{vol}^n \left[Y_i \frac{\partial \rho}{\partial t} + \rho \frac{\partial Y_i}{\partial t} + \text{div}(\rho_i u'_i) + Y_i \text{div}(\rho u) + \rho u^j \frac{\partial Y_i}{\partial x^j} \right] = \int_V \text{vol}^n \left[\rho \frac{\partial Y_i}{\partial t} + \text{div}(\rho_i u'_i) + \rho u^j \frac{\partial Y_i}{\partial x^j} \right] \end{aligned}$$

where total mass conservation $\dot{m} = 0 \implies \frac{\partial \rho}{\partial t} + \text{div}(\rho u) = 0$ was used in the last equality, which must be true, even if chemical reactions occur.

Then use Fick's law for particle diffusion for each species:

$$\rho_i u'_i = j_{m_i} = -\rho D_{AB} \text{grad} Y_i$$

and so

$$(38) \quad \boxed{\dot{m}_i = \int_V \text{vol}^n \left[\rho \frac{\partial Y_i}{\partial t} + -\text{div}(\rho D_{AB} \text{grad} Y_i) + \rho u^j \frac{\partial Y_i}{\partial x^j} \right]}$$

Compare this to Eq. 3.31, where steady flow is assumed ($\partial(\rho Y_A)/\partial t = 0$) in the section on **Species Conservation** of Chapter 3 Introduction to Mass Transfer of Turns (2011) [9]. I don't agree with his usage of this steady flow assumption because one needs to use the *total* mass conservation first as shown above. Eq. 3.31 of Turns (2011) [9] is, for reference,

$$\dot{m}_A''' - \frac{d}{dx} \left[Y_A \dot{m}'' - \rho D_{AB} \frac{dY_A}{dx} \right] = 0$$

What is useful from Turns (2011) [9] is the physical interpretation he provides: from Eq. 38, \dot{m}_i is the net rate of production of species i by chemical reaction by the entire system. $\rho \frac{\partial Y_i}{\partial t} + -\text{div}(\rho D_{AB} \text{grad} Y_i) + \rho u^j \frac{\partial Y_i}{\partial x^j}$ is the net flow of species i out of the (control) volume V per unit volume.

Then take a look at the section on **Species Mass Conservation (Species Continuity)** on pp. 218 of Chapter 7 “Simplified Conservation Equations” of Turns (2011) [9], namely Eq. (7.8) and (7.10). Further physical interpretation is provided there, and so we can interpret terms in Eq. 38 as follows:

$-\text{div}(\rho D_{AB} \text{grad} Y_i)$	mass flow of species i due to molecular diffusion per unit volume ($\text{kg}/s \cdot m^3$)
$\rho u^j \frac{\partial Y_i}{\partial x^j}$	mass flow of species i due to convection (advection by bulk flow) per unit volume ($\text{kg}/s \cdot m^3$)
\dot{m}_i	net mass production rate of species i by chemical reaction (for the entire system) kg/s

13.2. Energy Conservation applied to Combustion. Following the Section **Energy Conservation** on pp. 334, Chapter 7 Simplified Conservation Equations, of Turns (2011) [9],

$$\dot{E} = - \int_V \text{div}((k + h') \mathbf{u}) \text{vol}^n$$

where we keep in mind the difference in notation and meaning of h and h' : $h' := \frac{H}{V}$ where $H = \int_V h' \text{vol}^n$, is the enthalpy density, whereas $h := \frac{H}{m}$ where $H = \int_V h m = \int_V h \rho \text{vol}^n$ is the so-called *specific enthalpy*, enthalpy per unit mass. So $\rho h = h'$

Noting that

$$\text{div}\left(\left(\frac{\rho u^2}{2} + h'\right)u\right) = \text{div}\left(\left(\frac{\rho u^2}{2} + \rho h\right)u\right) = \text{div}(\rho u)\left(h + \frac{u^2}{2}\right) + \rho u^j \frac{\partial}{\partial x^j} \left(h + \frac{u^2}{2}\right)$$

and that

$$\begin{aligned} u^2 = g_{ij} u^i u^j \implies u^k \frac{\partial}{\partial x^k} u^2 &= 2u^k g_{ij} \frac{\partial u^i}{\partial x^k} u^j + \frac{\partial g_{ij}}{\partial x^k} u^k u^i u^j = 2u^k \left\langle u, \frac{\partial u}{\partial x^k} \right\rangle + 2g_{ks} \Gamma_{ij}^k u^i u^j u^s = \\ &= 2g_{ks} u^s \left[u^j \frac{\partial u^k}{\partial x^j} + \Gamma_{ij}^k u^i u^j \right] = 2g_{ks} u^s (\nabla_u u) = 2\langle u, \nabla_u u \rangle \end{aligned}$$

$$\frac{\partial g_{ij}}{\partial x^k} \dot{\phi}^k \dot{\phi}^i \dot{\phi}^j = 2g_{ks} \Gamma_{ij}^k \dot{\phi}^i \dot{\phi}^j \dot{\phi}^s \text{ with } u^i = \dot{\phi}^i$$

was used, which is in the proof for Theorem 3.25 on pp. 49 of section 3.8 The natural Lagrangian on manifolds of Calin and Chang (2005) [13]

Then

$$\dot{E} = - \int_V \text{vol}^n \left[\text{div}(\rho u) \left(h + \frac{u^2}{2}\right) + \rho u^j \frac{\partial}{\partial x^j} \left(h + \frac{u^2}{2}\right) \right] = \int_V \frac{\partial \rho}{\partial t} \left(h + \frac{u^2}{2}\right) \text{vol}^n - \int_V \text{vol}^n \left[\rho u^j \frac{\partial h}{\partial x^j} + \rho \langle u, \nabla_u u \rangle \right]$$

Now consider the heat $Q \in \Omega^1(\Sigma)$. If we define the heat density $q' \in C^\infty(\mathbb{R} \times N) \times \Omega^1(\Sigma)$, then

$$\dot{Q} = \int_V \left(\frac{\partial q'}{\partial t} + \text{div}(q' u) \right) \text{vol}^n \xrightarrow{\frac{\partial q'}{\partial t} = 0} \int_V \text{vol}^n \left[u^j \frac{\partial q'}{\partial x^j} + q' \text{div} u \right]$$

where we'd have to deal with whether the fluid is compressible $\text{div} u \neq 0$ or incompressible $\text{div} u = 0$.

If we define the “specific heat,” the heat per unit mass, q , defined via $Q = \int \rho q \text{vol}^n$, $q \in C^\infty(\mathbb{R} \times N) \times \Omega^1(\Sigma)$,

$$\begin{aligned} \dot{Q} &= \int_V \left(\rho \frac{\partial q}{\partial t} + q \frac{\partial \rho}{\partial t} \right) \text{vol}^n + \text{div}(\rho q u) \text{vol}^n = \int_V \left(\rho \frac{\partial q}{\partial t} + q \frac{\partial \rho}{\partial t} \right) \text{vol}^n + q \text{div}(\rho u) \text{vol}^n + \rho u^j \frac{\partial q}{\partial x^j} \text{vol}^n = \\ &= \int_V \rho \text{vol}^n \left(\frac{\partial q}{\partial t} + u^j \frac{\partial q}{\partial x^j} \right) \end{aligned}$$

where total mass conservation was used (which should still be valid even if chemical reactions occur; the masses for individual species can change, but the total mass doesn't change for chemical reactions), so that

$$u^j \frac{\partial q}{\partial x^j} = -u^j \frac{\partial h}{\partial x^j} - \langle u, \nabla_u u \rangle \text{ or } -\frac{\partial q}{\partial x^j} = \frac{\partial h}{\partial x^j} + (\nabla_u u)_j$$

Compare this expression we had just derived

$$(39) \quad -\frac{\partial q}{\partial x^j} = \frac{\partial h}{\partial x^j} + (\nabla_u u)_j$$

with Eq. 7.51 on pp. 235 of Chapter 7 of Turns (2011) [9]

$$-\frac{d\dot{Q}_x''}{dx} = \dot{m}'' \left(\frac{dh}{dx} + v_x \frac{dv_x}{dx} \right)$$

14. DROPLET EVAPORATION

14.1. Evaporation Rate. cf. Chapter 3 Introduction to Mass Transfer, Droplet Evaporation section, Evaporation Rate Sub-Section of Turns (2011) [9].

Starting from $m = \int_V \rho \text{vol}^n$,

$$(40) \quad \begin{aligned} \dot{m} &= \int_V \frac{\partial \rho}{\partial t} \text{vol}^n + \int_{\partial V} d(\rho i_u \text{vol}^n) = \int_V \frac{\partial \rho}{\partial t} \text{vol}^n + \int_{\partial V} \rho u^i dS_i \\ &= 0 + \rho u^r(r_s) 4\pi r_s^2 \equiv 4\pi r^2 \dot{m}'' \end{aligned}$$

where we had assumed, first, steady state for the density ρ , and second, spherical symmetry.

Consider $m = \sum_i m_i$. For binary collisions condition, $m = m_A + m_B$. Then $\dot{m} = \sum_i \dot{m}_i$. Assuming the above just like in Eq. 40, steady-state density and spherical symmetry, then $\dot{m}'' = \dot{m}''_A + \dot{m}''_B = \dot{m}''_A$, since $\dot{m}''_B = 0$ (assumed *no flow into droplet*).

14.1.1. *Separating diffusion and bulk flow.* Beginning with $m_i = \int_V \rho_i \text{vol}^n$ (and $m = \sum_i m_i = \int_V \rho \text{vol}^n$, where $\rho = \sum_i \rho_i$, which is allowed if we assume *negligible interaction* between species, until chemical reactions),

$$\dot{m}_i = \int_V \frac{\partial \rho_i}{\partial t} \text{vol}^n + \int_{\partial V} \rho_i u_i^j dS_j = \int_V \frac{\partial \rho_i}{\partial t} \text{vol}^n + \int_{\partial V} \rho_i (u_i')^j dS_j + \int_{\partial V} \rho_i u^j dS_j = \int_V \frac{\partial \rho_i}{\partial t} \text{vol}^n + \int_{\partial V} j_{i,\text{diff}} \cdot dS + \int_{\partial V} \rho_i u \cdot dS \text{ using}$$

$$\int_{\partial V} \rho_i u_i^j dS_j = \int_{\partial V} \rho_i ((u_i')^j + u^j) dS_j$$

Assuming steady state density ρ_i ,
constant physical parameters on a particular choice of the surface boundary ∂V , with surface area A ,

$$\frac{\dot{m}_i}{A} \equiv \dot{m}_i'' = 0 + (j_{i,\text{diff}})^r + \rho_i u^r$$

Now for the total mass of all (i.e. all the species), $m = \sum_i m_i$,

$$\dot{m} = \int \frac{\partial \rho}{\partial t} \text{vol}^n + \int_{\partial V} \rho u^j dS_j$$

Assuming steady state density ρ ,
constant physical parameters on a particular choice of the surface boundary ∂V , with surface area A ,

$$\dot{m}/A \equiv \dot{m}'' = 0 + \rho u^r$$

So one *must define carefully* the choice of volume V (and thus surface ∂V). For we want to consider a choice of volume V s.t. the choice of ∂V results in a *net flux* of the bulk flow (we want a measure of that across the surface ∂V . In this case,

$$\dot{m}'' = \rho u^r \equiv \rho u^n$$

From the definition of $Y_i := \frac{m_i}{m}$, then

$$\rho_i u = Y_i \rho u \equiv Y_i (\dot{m}'')$$

Thus

$$(41) \quad \dot{m}_i'' = Y_i (\dot{m}'') + (j_{i,\text{diff}})$$

which is Eqns. (3.1), (3.5) in Turns (2011) [9].

From Le Bellac, Mortessagne, Batrouni (2004) [11], namely Chapter 6 on Irreversible Processes, for the section on “Particle Diffusion”, for particle diffusion of species A , with the number of particles of species A ,

$$j_\alpha^{NA} = \sum_{j,\beta} L_{NAj}^{\alpha\beta} \partial_\beta \gamma_j = \sum_{j,\beta} L_{NAj}^{\alpha\beta} \partial_\beta \left(\frac{-\mu_j}{\tau} \right) = \sum_{j,\beta} \frac{-L_{NAj}^{\alpha\beta}}{\tau} \sum_i \frac{\partial \mu_j}{\partial n_i} \frac{\partial n_i}{\partial x^\beta}$$

If $\mu_i = \mu_i(n_i)$, so that μ_i doesn’t depend on μ_j , $j \neq i$, i.e. the chemical potential of species i does not depend on the concentrations of the other species,

$$\sum_{j,\beta} \frac{-L_{NAj}^{\alpha\beta}}{\tau} \frac{\partial \mu_j}{\partial n_j} \frac{\partial n_j}{\partial x^\beta} = \sum_{j,\beta} \frac{-L_{NAj}^{\alpha\beta}}{\tau} \frac{1}{\kappa_T n_j^2} \frac{\partial n_j}{\partial x^\beta}$$

and so

$$(42) \quad \mathbf{j}_{NA} = \sum_j \frac{-L_{NAj}}{\tau \kappa_T n_j^2} \nabla n_j$$

Now

$$\frac{N_i}{V} \equiv n_i = \frac{Y_i m / M_i}{V} = \frac{Y_i m}{M_i V} = \frac{Y_i \rho}{M_i} \text{ so } \boxed{\nabla n_i = \nabla \frac{Y_i \rho}{M_i} = \frac{\rho}{M_i} \nabla Y_i} \text{ for}$$

$$\sum_i \rho_i = \rho := \frac{M}{V} \implies \sum_i M_i n_i = \frac{M}{V}$$

So then we can reexpress Eqn. 42 in terms of *mass fractions* Y_i as

$$\mathbf{j}_{NA} = -\rho \sum_j \frac{L_{NAj}}{\tau \kappa_T n_j^2 M_j} \nabla Y_j$$

Note that κ_T is the **coefficient of isothermal compressibility** which was defined in Le Bellac, Mortessagne, Batrouni (2004) [11],

$$\kappa_T = \frac{-1}{V} \left. \frac{\partial V}{\partial p} \right|_\tau$$

Now I guess that $L_{NAj} \propto \delta_{NAj}$ meaning $L_{NAj} \neq 0$ only if $j = N_A$, and so \mathbf{j}_{NA} only depends on the gradient ∇Y_{N_A} , the gradient for what is the concentration of species A , and *not* on the gradient of the concentrations of the other species. I *think* (please contact me if this is wrong) this is because the species are nonreacting, and even when including chemical reactions, the chemical reactions does not affect this diffusion process; the diffusion process is independent of what goes on with the chemical reactions after the chemical reactions occur.

I conjecture that L_{NAj} does not depend on direction (i.e. *isotropic*), which is reasonable, by spatial symmetry.

If so, then

$$(43) \quad \mathbf{j}_{NA} = -\rho \left(\frac{L_{N_A N_A}}{\tau \kappa_T n_A^2 M_A} \right) \nabla Y_A$$

14.1.2. *Particle Diffusion.* cf. pp. 399 of Kittel and Kroemer [4]

Consider a system.

One end in diffusive contact with reservoir at chemical potential μ_1

Other end in diffusive contact with reservoir at chemical potential μ_2

Constant temperature τ .

If $\mu_1 > \mu_2$, particle flow through system from reservoir 1 to reservoir 2; $1 \rightarrow 2$.

$n_i \equiv$ particle concentration in i

Take

$$(44) \quad \mathbf{j}_n = -D \text{grad} n$$

which is **Fick’s law**, and where $D \equiv$ particle diffusion constant or **diffusivity**.

Mean free path l . Particles freely travel over l .

Assume in a collision at z , particles come into local equilibrium at local chemical potential $\mu(z)$, local concentration $n(z)$.

At z , particle flux density in positive z direction $\frac{1}{2} n(z - l_z) \bar{c}_z$

particle flux density in negative z direction $-\frac{1}{2} n(z + l_z) \bar{c}_z$

Note $n(z - l_z)$ is particle concentration at $z - l_z$

$$J_n^z = \frac{1}{2} [n(z - l_z) - n(z + l_z)] \bar{c}_z = -\frac{dn}{dz} \bar{c}_z l_z$$

where $\bar{c}_z = \bar{c} \cos \theta$

$$\bar{l}_z = \bar{l} \cos \theta$$

$$\langle \bar{c}_z l_z \rangle = \bar{c} \bar{l} \frac{\int_{\text{hemisphere}} \cos^2 \theta dS}{\int_{\text{hemisphere}} dS} = \bar{c} \bar{l} \frac{\int_0^{\pi/2} d\theta \int_0^{2\pi} d\varphi \cos^2 \theta \sin \theta d\theta}{\int_0^{\pi/2} d\theta \sin \theta \int_0^{2\pi} d\varphi} = \bar{c} \bar{l} \frac{1}{3}$$

Comparing with Fick’s law,

$$J_n^z = \frac{-1}{3} \bar{c} \bar{l} \frac{dn}{dz} \text{ or } \mathbf{J}_N = -\frac{1}{3} \bar{c} \bar{l} \nabla n$$

For diffusivity D is then $D = \frac{1}{3} \bar{c} \bar{l}$.

Now recall that \bar{l} , the *mean free path*, was derived from kinetic theory:

$$l = \frac{1}{n\pi d^2}$$

where d is the diameter of the particle.

The mean thermal velocity was derived from the Maxwell distribution:

$$\bar{c} = \left(\frac{8\tau}{M\pi} \right)^{1/2}$$

14.2. Binary case for Fick’s law for mass transfer; A and B species only. It seems that for combustion, we only worry about 2 species, A and B , coming together, most of the time.

I am following pp. 84 Chapter 3 Introduction to Mass Transfer, Section **Mass Transfer Rate Laws** of Turns (2011) [9] here, applying Fick’s law of Diffusion from my own development and comparing results from Turns.

Recall the basic quantities for summing up masses and particles (in this case, only two kinds), and mass fraction vs. mole fraction:

$$\begin{aligned} N_A + N_B &= N \\ X_A &:= \frac{N_A}{N} \\ M_A X_A + M_B X_B &= \frac{m}{N} \\ \rho_A &= \frac{M_A N_A}{V} \text{ so } Y_A(\rho_A + \rho_B) = \rho_A \\ m_A + m_B &= m \\ Y_A &:= \frac{m_A}{m} = \frac{M_A N_A}{M_A N_A + M_B N_B} \\ m_A + m_B &= Y_A m + Y_B m = m \end{aligned}$$

Now

$$\dot{m}_A = \frac{d}{dt} \int_V \rho_A \text{vol}^n = \int_V \left[\frac{\partial \rho_A}{\partial t} + \text{div}(\rho_A u_a) \right] \text{vol}^n$$

Using Galilean transformation $\mathbf{u}_A - \mathbf{u} = \mathbf{u}'_A$ or $\mathbf{u}_A = \mathbf{u} + \mathbf{u}'_A$,

$$\text{div}(\rho_A u) = \text{div}(\rho_A (u + u'_A)) = \text{div}(\rho_A u) + \text{div}(\rho_A u'_A)$$

Considering the term $\text{div}(\rho_A u)$,

$$\text{div}(\rho_A u) = \text{div}(Y_A \rho u) \text{ and also note that}$$

$$\int_V \text{div}(Y_A \rho u) = \int_{\partial V} dS \cdot Y_A \rho \mathbf{u}$$

Now for term $\text{div}(\rho_A u'_A)$, using Eq. 44, Fick’s law,

$$\text{div}(\rho_A u'_A) = \text{div}(j_{\rho_A}) = \text{div}(-M_A \mathcal{D}_{AB} \text{grad} n_A) = -\text{div}(\mathcal{D}_{AB} \text{grad} \rho_A) = -\text{div}(\rho \mathcal{D}_{AB} \text{grad} Y_A)$$

Putting everything in terms of a surface area integral,

$$\dot{m}_A = \int_{\partial V} dS \cdot \dot{m}_A'' = \int_{\partial V} dS \cdot \rho u Y_A - \int_{\partial V} dS \cdot \rho \mathcal{D}_{AB} \text{grad} Y_A$$

Then

$$(45) \quad \boxed{\dot{m}_A'' = Y_A \rho u - \rho \mathcal{D}_{AB} \text{grad} Y_A}$$

Note that \mathcal{D}_{AB} is in units of $[m^2/2]$.

The physical interpretation for the terms is as follows:

\dot{m}_A'' is the mass flow of species A per unit area

$Y_A \rho u$ is the mass flow of species A associated with bulk flow per unit area

$\rho \mathcal{D}_{AB} \text{grad} Y_A$ is the mass flow of species A associated with molecular diffusion per unit area.

Compare this to Eq. 3.1, 3.5 of Turns (2011) [9]:

$$(46) \quad \dot{m}_A'' = Y_A (\dot{m}_A'' + \dot{m}_B'') - \rho \mathcal{D}_{AB} \text{grad} Y_A$$

15. DROPLET EVAPORATION AND BURNING

I am following Chapter 10 Droplet Evaporation and Burning of Turns (2011) [9].

“There are two types of liquid rockets: **pressure-fed**, in which the fule and oxidizer are pushed into the combustion chamber by a high-pressure gas; and **pump-fed**, where turbopumps deliver the propellants.” [9]

16. SHVAB-ZELDOVICH FORMS

Lewis number ($\text{Le} := \frac{K}{\rho c_p D}$) $\text{Le} = 1$ by assumption.

Species flux and Fick’s law.

$$\mathbf{j}_q = -K \text{grad} \tau + \sum_i \dot{m}_i''_{\text{diff}} h_i = -K \text{grad} \tau - \sum_i \rho D (\text{grad} Y_i) h_i = -K \text{grad} \tau - \rho D \text{grad} \sum_i Y_i h_i + \rho D \sum_i Y_i \text{grad} h_i$$

Noting that

$$\sum_i Y_i \text{grad} h_i = \sum_i Y_i c_{pi} \text{grad} \tau = c_p \text{grad} \tau$$

then

$$j_q = -K \text{grad} \tau - \rho D \text{grad} h + \rho D c_p \text{grad} \tau$$

thermal diffusivity $\alpha := \frac{K}{\rho c_p}$.

If $\text{Le} = 1$, $K = \rho D c_p$ and so $j_q = -\rho D \text{grad} h$ or $\dot{Q}'' = -\rho D \text{grad} h$.

16.1. Heat Q . Again $Q \in \Omega^1(\Sigma)$ represents heat on the system.

The conundrum is whether Q is best served with the heat density $q' := Q/V$ or “specific heat” per mass, $q := Q/m$, because it’ll affect whether we have to deal with the compressibility or incompressibility of the fluid or mass conservation, and the form of the “heat flux.”

So

$$Q = \int_V q \rho \text{vol}^n = \int_V q m = \int_V q' \text{vol}^n$$

Then

$$\begin{aligned} \dot{Q} &= \int_V \left(\rho \frac{\partial q}{\partial t} + q \frac{\partial \rho}{\partial t} \right) \text{vol}^n + \int_V \text{div}(\rho q u) \text{vol}^n = \int_V \left[\rho \left(\frac{\partial q}{\partial t} + u^j \frac{\partial q}{\partial x^j} \right) + q \left(\frac{\partial \rho}{\partial t} + \text{div}(\rho u) \right) \right] \text{vol}^n = \\ &= \int_V \left(\frac{\partial q'}{\partial t} + \text{div}(q' u) \right) \text{vol}^n = \int_V \left(\frac{\partial q'}{\partial t} + \text{div}(\mathbf{j}_{q'}) \right) \text{vol}^n \end{aligned}$$

where

$$\text{div}(\mathbf{j}_{q'}) = \frac{1}{\sqrt{g}} \frac{\partial}{\partial x^j} (-\rho D \sqrt{g} \text{grad} h) = \frac{1}{\sqrt{g}} \frac{\partial}{\partial x^j} (-\rho D \frac{\partial h}{\partial x^k} g^{jk} \sqrt{g})$$

Employ the definition of absolute (or standardized) enthalpy in Turns (2011) [9]:

$$h = \sum_i Y_i \Delta_f \bar{h}_i^\circ + \int_{T_{\text{ref}}}^T c_p d\tau$$

The proper way to treat the usage of energy conservation with the Shvab-Zeldovich form is this: starting from $Q = dE - W$, and having no external work done on the system $\dot{W} = 0$

$$(47) \quad Q \left(\frac{d}{dt} \right) = \dot{Q} = \int_V \text{vol}^n \left(\frac{\partial q'}{\partial t} + \text{div}(j_{q'}) \right) = dE \left(\frac{d}{dt} \right) = \dot{E} = - \int_V \text{vol}^n \left[(\rho u^j) \left(\frac{\partial h}{\partial x^j} + (\nabla_u u)_j \right) \right]$$

where on the right hand side (RHS) of Eq. 47 is from the energy conservation of the total energy density of the fluid system,

$$\frac{\partial \epsilon}{\partial t} + \text{div}((h' + k)u) = 0$$

which is from 6.5.2 of Le Bellac, Mortessagne, Batrouni (2004) [11]). Then for steady-state assumption, $\frac{\partial q'}{\partial t} = 0$,

$$\text{div}(j_{q'}) = -(\rho u^j) \left(\frac{\partial h}{\partial x^j} + (\nabla_u u)_j \right)$$

The Shvab-Zeldovich form for $j_{q'}$ is

$$(48) \quad j_q = -\rho D \text{grad} h$$

and so

$$(49) \quad \text{div}(\mathbf{j}_{q'}) = \frac{1}{\sqrt{g}} \frac{\partial}{\partial x^j} (-\rho D \sqrt{g} \text{grad} h) = \frac{1}{\sqrt{g}} \frac{\partial}{\partial x^j} (-\rho D \frac{\partial h}{\partial x^k} g^{jk} \sqrt{g}) = -(\rho u^j) \left(\frac{\partial h}{\partial x^j} + (\nabla_u u)_j \right)$$

Compare Eq. 49 to Eq. (7.61) of Turns (2011) [9].

Employ now Turns' definition of absolute (or standardized) enthalpy [9]:

$$h = \sum_i Y_i \Delta_f \bar{h}_i^0 + \int_{T_{\text{ref}}}^T c_p dT$$

Then applying this definition to Eq. 49,

$$\begin{aligned} \text{div}(\rho D \text{grad} h) &= \text{div} \left(\rho D \sum_i \Delta_f \bar{h}_i^0 \text{grad} Y_i + \rho D \text{grad} \int_{T_{\text{ref}}}^T c_p dT \right) = \\ &= (\rho u^j) \left(\frac{\partial h}{\partial x^j} + (\nabla_u u)_j \right) = (\rho u^j) \left(\sum_i \Delta_f \bar{h}_i^0 \frac{\partial Y_i}{\partial x^j} + \frac{\partial}{\partial x^j} \int_{T_{\text{ref}}}^T c_p dT + (\nabla_u u)_j \right) \text{ or (moving terms from left to right and vice versa)} \\ &(\rho u^j) \frac{\partial}{\partial x^j} \int c_p dT - \text{div}(\rho D \text{grad} \int c_p dT) + \rho u^j (\nabla_u u)_j = \sum_i \text{div}(\rho D \Delta_f \bar{h}_i^0 \text{grad} Y_i) - (\rho u^j) \Delta_f \bar{h}_i^0 \frac{\partial Y_i}{\partial x^j} \end{aligned}$$

Use Eq. 38, the generalized equation for species conservation,

$$\sum_i \Delta_f \bar{h}_i^0 (\text{div}(\rho D \text{grad} Y_i) - (\rho u^j) \frac{\partial Y_i}{\partial x^j}) = \sum_i \Delta_f \bar{h}_i^0 (-\dot{m}_i''')$$

where $\dot{m}_i = \int_V \text{vol}^n \dot{m}_i'''$.

Thus, in general,

$$(50) \quad \boxed{(\rho u^j) \frac{\partial}{\partial x^j} \int c_p dT - \text{div}(\rho D \text{grad} \int c_p dT) + \rho u^j (\nabla_u u)_j = - \sum_i \Delta_f \bar{h}_i^0 \dot{m}_i'''}$$

where Turns (2011) [9] gives the (important) physical interpretations:

$(\rho u^j) \frac{\partial}{\partial x^j} \int c_p dT$	rate of sensible enthalpy transport by convection (advection) per unit volume (W/m^3)
$-\text{div}(\rho D \text{grad} \int c_p dT)$	rate of sensible enthalpy transport by diffusion per unit volume (W/m^3)
$-\sum_i \Delta_f \bar{h}_i^0 \dot{m}_i'''$	rate of formation enthalpy production by chemical reaction per unit volume (W/m^3)

Turns (2011) [9] says that $-\sum_i \Delta_f \bar{h}_i^0 \dot{m}_i'''$ is the “rate of sensible enthalpy production by chemical reaction per unit volume (W/m^3)”, but I think this is wrong because the terms are dealing with the formation enthalpy, the enthalpy locked in the chemical bonds of the constituents.

It will pay to consider the case of spherical symmetry for droplets. Assuming only dependence upon r for all quantities by spherical symmetry, and keeping in mind that total mass conservation holds, i.e.

$$\begin{aligned} \dot{m} = 0 &\implies \frac{\partial \rho}{\partial t} + \text{div}(\rho u) = 0 \xrightarrow{\text{spherical symmetry}} \frac{1}{r^2} \frac{d}{dr} (\rho u^r r^2) = -\frac{\partial \rho}{\partial t} \\ &\xrightarrow{\frac{\partial \rho}{\partial t} = 0 \text{ steady state}} \frac{1}{r^2} \frac{d}{dr} (\rho u^r r^2) = 0 \end{aligned}$$

then

$$\begin{aligned} -\text{div}(\rho D \text{grad} \int c_p dT) &= -\frac{1}{r^2} \frac{d(\rho D r^2 \frac{d \int c_p dT}{dr})}{dr} \\ \frac{d}{dr} (r^2 \rho u^r \int c_p dT) &= \left(\frac{d}{dr} (r^2 \rho u^r) \right) \int c_p dT + r^2 \rho u^r \frac{d}{dr} \int c_p dT \xrightarrow{\dot{m}=0} 0 + r^2 \rho u^r \frac{d}{dr} \int c_p dT \end{aligned}$$

so

$$(51) \quad \frac{1}{r^2} \frac{d}{dr} \left[r^2 \rho u^r \int c_p dT - \rho D r^2 \frac{d \int c_p dT}{dr} \right] = - \sum_i \Delta_f h_i^0 \dot{m}_i'''$$

Compare this to Eq. (7.65) of Turns (2011) [9].

16.1.1. *Temperature distribution in the gas phase for Droplet Evaporation.* From Eq. 51, supposing steady state mass conservation with spherical symmetry:

$$\dot{m} = 4\pi r^2 \rho u^r$$

and supposing that

$$\sum_i \Delta_f \bar{h}_i^0 \dot{m}_i''' = 0$$

which means that “reaction rate term is zero, since no reactions occur for pure evaporation” (cf. pp. 372 Ch. 10 Droplet Evaporation and Burning, Eq. (10.3), of Turns (2011) [9]), and that the Lewis number $Le := \frac{K}{\rho c_p D}$ is 1, then for the remaining terms in Eq. 51

$$\begin{aligned} r^2 \rho u^r \frac{d}{dr} \int c_p dT &= r^2 \rho u^r c_p \frac{dT}{dr} = \frac{\dot{m}}{4\pi} c_p \frac{dT}{dr} \\ \frac{d}{dr} \left[\rho D r^2 \frac{d \int c_p dT}{dr} \right] &= \frac{d}{dr} \left[\rho D r^2 c_p \frac{dT}{dr} \right] = \frac{d}{dr} \left[K r^2 \frac{dT}{dr} \right] \\ &\implies \frac{d}{dr} \left[r^2 \frac{dT}{dr} \right] = \frac{\dot{m} c_{pg}}{4\pi K} \frac{dT}{dr} \equiv Z \dot{m} \frac{dT}{dr} \end{aligned}$$

Let's solve this (easy) ODE (ordinary differential equation):

$$\frac{d(r^2 \frac{dT}{dr})}{dr} = Z \dot{m} \frac{dT}{dr}$$

Then

$$\begin{aligned} \implies 2r\dot{T} + r^2\ddot{T} &= Z\dot{m}\dot{T} \text{ or } \ddot{T} = \frac{(Z\dot{m} - 2r)}{r^2} \dot{T} \text{ so } \ln \dot{T} = \frac{Z\dot{m}}{-r} - 2 \ln r + C \text{ or } \dot{T} = \frac{C_1}{r^2} \exp \left(\frac{-Z\dot{m}}{r} \right) \\ &\xrightarrow{\int dr} T(r) = \frac{C_1 \exp \left(\frac{-Z\dot{m}}{r} \right)}{Z\dot{m}} + C_2 \end{aligned}$$

Let's consider the (important physically) boundary conditions:

$$(52) \quad \begin{aligned} T_b &\equiv T_{\text{boil}} = T(r = r_s) \\ T(r = \infty) &= T_\infty = \frac{C_1}{Z\dot{m}} + C_2 \end{aligned}$$

which is physically interpreted (importantly) that out in $r = \infty$, the system is in thermal equilibrium to some temperature T_∞ , and at the surface of the droplet, $r = r_s$, the liquid making up the droplet is evaporating away, and this occurs at the boiling temperature T_b . Thus, the temperature at $r = r_s$, the surface of the droplet, has to be T_b .

Then

$$\begin{aligned} T_\infty &= \frac{C_1}{Z\dot{m}} + C_2 & T_b &= \frac{C_1 \exp\left(-\frac{Z\dot{m}}{r_s}\right)}{Z\dot{m}} + C_2 \\ C_1 \left(\frac{1 - \exp\left(-\frac{Z\dot{m}}{r_s}\right)}{Z\dot{m}} \right) &= T_\infty - T_b \\ \implies C_2 &= T_\infty - \frac{T_\infty - T_b}{1 - \exp\left(-\frac{Z\dot{m}}{r_s}\right)} = \frac{T_b - T_\infty \exp\left(-\frac{Z\dot{m}}{r_s}\right)}{1 - \exp\left(-\frac{Z\dot{m}}{r_s}\right)} \end{aligned}$$

Therefore

(53)

$$T(r) = \frac{(T_\infty - T_b) \exp\left(-\frac{Z\dot{m}}{r}\right) + T_b - T_\infty \exp\left(-\frac{Z\dot{m}}{r_s}\right)}{1 - \exp\left(-\frac{Z\dot{m}}{r_s}\right)}$$

Note that the dependence on r of $T(r)$, $\frac{dT(r)}{dr}$ is easily given by

(54)

$$\frac{dT(r)}{dr} = \frac{(T_\infty - T_b) \exp\left(-\frac{Z\dot{m}}{r}\right) \left(Z\dot{m} \frac{1}{r^2}\right)}{1 - \exp\left(-\frac{Z\dot{m}}{r_s}\right)}$$

Let’s consider the surface of the droplet. Obviously, if it’s evaporating off the surface, then heat Q is supplied at the surface, raising the temperature of the droplet molecules at the surface to boiling temperature, enough to change (its state) into a vapor. This is the so-called latent heat of vaporization. Denote this heat by $Q_{\text{conduction}}$.

Starting from the general statement that $Q = dH - Vdp$, then the thermodynamic process $\gamma : \mathbb{R} \rightarrow \Sigma$ we want to consider here results in the following:

$$\int Q(\dot{\gamma}) = H_{\text{vap}} - H_{\text{liq}}$$

Then

(55)

$$\dot{Q}_{\text{conduction}} = \frac{d}{dt} \int_\gamma Q(\dot{\gamma}) = \dot{m}(h_{\text{vap}} - h_{\text{liq}}) \equiv \dot{m}h_{fg}$$

Now consider Fourier’s law for heating (cf. pp. 401 Eq. (26) of Kittel and Kroemer (1980) [4]). It says

$$j_u = -K \text{grad} \tau$$

This implies, for the steady-state case, that

$$\dot{Q}_{\text{conduction}} = \int_{\partial V} dS_j j_u^j = K4\pi r_s^2 \left. \frac{dT}{dr} \right|_{r_s}$$

Note that the direction is such that positive heat flows *into* the droplet system, whereas, geometrically, the normal to the surface is outward and *out* of the droplet. Plugging this into Eq. 55,

$$\dot{m}h_{fg} = K4\pi r_s^2 \left. \frac{dT}{dr} \right|_{r_s}$$

Plugging in Eq. 54,

$$\begin{aligned} \frac{(T_\infty - T_b) \exp\left(-\frac{Z\dot{m}}{r_s}\right) \left(Z\dot{m} \frac{1}{r_s^2}\right) \left(\frac{K4\pi r_s^2}{h_{fg}}\right)}{1 - \exp\left(-\frac{Z\dot{m}}{r_s}\right)} &= \dot{m} \implies \frac{(T_\infty - T_b) \exp\left(-\frac{Z\dot{m}}{r_s}\right) \left(Z\frac{1}{r_s^2}\right) \left(\frac{c_{pg} r_s^2}{h_{fg}}\right)}{1 - \exp\left(-\frac{Z\dot{m}}{r_s}\right)} = 1 \\ &\implies \frac{(T_\infty - T_b) \exp\left(-\frac{Z\dot{m}}{r_s}\right)}{1 - \exp\left(-\frac{Z\dot{m}}{r_s}\right)} = \frac{h_{fg}}{c_{pg}} \\ &\implies \ln\left(\frac{c_{pg}}{h_{fg}}(T_\infty - T_b) + 1\right) = \frac{Z\dot{m}}{r_s} \\ &\implies \dot{m} = \frac{4\pi K_g r_s}{c_{pg}} \ln\left(\frac{c_{pg}}{h_{fg}}(T_\infty - T_b) + 1\right) \end{aligned}$$

Noting that $m_d = \rho_l V = \rho_l \frac{4}{3} \pi \left(\frac{D}{2}\right)^3$, then

$$\dot{m}_d = \rho_l 4\pi \frac{D^2}{8} \frac{dD}{dt}$$

Keeping in mind that $\dot{m}_d = -\dot{m}$ because the mass of the droplet decreases with each layer of mass on the surface moving outward and away as a vapor,

$$\begin{aligned} \frac{dD^2}{dt} &= -\frac{8K_g}{c_{pg}\rho_l} \ln\left(1 + \frac{c_{pg}}{h_{fg}}(T_\infty - T_b)\right) \\ \frac{dD^2}{dt} &= -\frac{8K_g}{\rho_l c_{pg}} \ln\left[1 + \frac{c_{pg}(T_\infty - T_b)}{h_{fg}}\right] = -\text{const.} \equiv -k_{\text{evap}} \end{aligned}$$

The decrease in D^2 is constant! Then simply $D^2(t) = D_0^2 - k_{\text{evap}} t$. Thus, the lifetime of the droplet t_d is

$$t_d = \frac{D_0^2}{k_{\text{evap}}} = \frac{D_0^2}{\frac{8k_g}{\rho_l c_{pg}} \ln\left[1 + \frac{c_{pg}(T_\infty - T_b)}{h_{fg}}\right]}$$

16.1.2. Gas Phase Composition.

(56)

$$\frac{d\phi_g}{dx} = \frac{1}{(F/O)_{\phi=1}} \frac{1}{\dot{m}_{\text{Ox}}(0)} \frac{d\dot{m}_g}{dx}$$

17. DROPLET MODEL; BURNING DROPLETS

cf. 20160129 Dr. Jay Polk Ae121b Winter 2016

What we want to solve for:

$\dot{m}_f \implies$ fuel production rate

$D(t) \implies$ applet lifetime

5 unknowns:

- T_f
- T_s
- $Y_{f,s}$
- r_f
- \dot{m}_f

The assumptions for this model is also given on pp. 379 of Turns (2011) [9], with the section **Simple Model of Droplet Burning**.

Note Assumption 3: “The fuel is a single-component liquid with zero solubility for gases.” Thus, there cannot be any net mass flux or mass flow of products inward from the flame to the droplet surface, since products cannot dissolve in the liquid by assumption 3.

Thus, in the inner zone $r \in [r_s, r_f]$, $\dot{m}_{\text{Products}} \equiv \dot{m}_{\text{Pr}}$. However, there is products in the inner zone, that form a stagnant film through which the fuel vapor flows through.

17.1. Species Conservation (Inner Zone). F denotes fuel. We're thinking about the diffusion of fuel vapor, according to Fick's law, and so \mathcal{D} is the diffusivity constant in Fick's law, and is for binary diffusion (fuel vapor and combustion products), and for fuel vapor, in this case. ρ is for the fuel density.

Let $Z_F := \frac{1}{4\pi\rho\mathcal{D}}$.

The boundary conditions (BCs) are

$$\begin{aligned} Y_F(r_s) &= Y_{F,s}(T_s) \\ Y_F(r_f) &= 0 \\ Y_F(r) &= 1 - \frac{(1 - Y_{F,s}) \exp(-Z_F \dot{m}_F / r)}{\exp(-Z_F \dot{m}_F / r_s)} \end{aligned}$$

Applying boundary condition $Y_F(r_f) = 0$,

$$\implies \frac{(1 - Y_{F,s}) \exp(-Z_F \dot{m}_F / r_f)}{\exp(-Z_F \dot{m}_F / r_s)} = 1$$

So

$$(57) \quad \boxed{Y_{F,s}(r_s) = 1 - \frac{\exp(-Z_F \dot{m}_F / r_s)}{\exp(-Z_F \dot{m}_F / r_f)}}$$

17.2. Species Conservation (Outer Zone). Assume stoichiometric combustion.

1 kg fuel + ν kg oxidizer \rightarrow $\underbrace{(\nu + 1)}_{\text{get from stoichiometry}}$ kg of products.

EY : 20160216 Note that by mass conservation,

$$\begin{aligned} m_f + m_{\text{Ox}} &= m_p \implies 1 + \frac{m_{\text{Ox}}}{m_f} = \frac{m_p}{m_f} \\ \nu_f M_f + \nu_{\text{Ox}} M_{\text{Ox}} &= m_p \implies 1 + \frac{\nu_{\text{Ox}} M_{\text{Ox}}}{\nu_f M_f} = \frac{m_p}{\nu_f M_f} \end{aligned}$$

The boundary conditions (BCs) are

$$\begin{aligned} Y_{\text{Ox}}(r_f) &= 0 \\ Y_{\text{Ox}}(\infty) &= 1 \end{aligned}$$

With

$$Y_{\text{Ox}}(r) = \nu \left[\frac{\exp(-Z_F \dot{m}_F / r)}{\exp(-Z_F \dot{m}_F / r_f)} - 1 \right]$$

then, from $Y_{\text{Ox}}(\infty) = 1$,

$$(58) \quad \boxed{\exp(Z_F \dot{m}_F / r_f) = \frac{\nu + 1}{\nu}}$$

One needs to take care in using the above. For instance, for the *outer zone*, by Assumption 5, *only* the *oxidizer and products* are there. So in

$$\dot{m}_F = 4\pi r^2 \frac{\rho \mathcal{D}}{\nu + Y_{\text{Ox}}} \frac{dY_{\text{Ox}}}{dr}$$

and in solving it

$$\begin{aligned} \implies \frac{dr}{r^2} &= \frac{4\pi\rho\mathcal{D}\dot{m}_F}{\nu + Y_{\text{Ox}}} dY_{\text{Ox}} \\ \implies \frac{-1}{r} + \frac{1}{r_f} &= \frac{4\pi\rho\mathcal{D}}{\dot{m}_F} (\ln(\nu + Y_{\text{Ox}}) - \ln(\nu + Y_{\text{Ox}}(r_f))) \end{aligned}$$

then $Z_f = \frac{1}{4\pi\rho\mathcal{D}}$ was used, but in this case, with binary diffusion between species A , which is the oxidizer, Ox, in this case, and species B which is the products, Pr, in this case, then $\rho = \rho_A + \rho_B$ is for the total density of Ox and Pr in this case, not fuel F and Pr. Also, \mathcal{D} in this case refers to the binary diffusion of Ox and Pr, not to F and Pr, as in the inner zone, previously.

17.3. Energy conservation at droplet. Chemical reactions confined to occur only at the boundary i.e. the flame sheet (flame front). The reaction rate term is 0 both inside and outside the flame.

Thus, the energy equation for droplet evaporation is used.

$Z_T = \frac{c_{pg}}{4\pi k_g} = Z_F$ if Le = 1, Lewis number is 1.

$$(59) \quad \underbrace{\dot{q}_{i-l}}_{\text{interface into liquid}} \quad \underbrace{\dot{q}_{g-i}}_{\text{gas to interface}} = - \left[-k_g 4\pi r^2 \frac{dT}{dr} \Big|_{r_s} \right]$$

$$\boxed{\frac{c_{pg}(T_f - T_s)}{(q_{i-l} + h_{fg})} \frac{\exp(-Z_T \dot{m}_F / r_s)}{[\exp(-Z_T \dot{m}_F / r_s) - \exp(-Z_T \dot{m}_F / r_f)]} + 1 = 0}$$

17.4. Energy Conservation at Flame Front. Chemical energy released at flame taken account by absolute enthalpy fluxes for fuel, oxidizer, and products:

$$\dot{m}_F h_F + \dot{m}_{\text{Ox}} h_{\text{Ox}} - \dot{m}_{\text{Pr}} h_{\text{Pr}} = \dot{Q}_{f-i} + \dot{Q}_{f-\infty}$$

Enthalpies are defined as

$$h_F := h_{f,F}^\circ + c_{pg}(T - T_{\text{ref}})$$

$$h_{\text{Ox}} := h_{f,\text{Ox}}^\circ + c_{pg}(T - T_{\text{ref}})$$

$$h_{\text{Pr}} := h_{f,\text{Pr}}^\circ + c_{pg}(T - T_{\text{ref}})$$

Heat of combustion Δh_c per unit mass of fuel given by

$$\Delta h_c(T_{\text{ref}}) = (1)h_{f,F}^\circ + (\nu)h_{f,\text{Ox}}^\circ - (1 + \nu)h_{f,\text{Pr}}^\circ$$

Note that although products exist in inner region, \nexists net flow of products between droplet surface and flame; thus all products flow radially outward away from flame, i.e.

$$\begin{array}{ccc} r_s & \xrightarrow{\dot{m}_F} & r_f \xleftarrow{\dot{m}_{\text{Ox}}} \\ & & \searrow \dot{m}_{\text{Pr}} \end{array} \quad \text{Thus } \dot{m}_F + \dot{m}_{\text{Ox}} = -\dot{m}_{\text{Pr}}, \text{ or } \dot{m}_F + \nu \dot{m}_F = -\dot{m}_{\text{Pr}}$$

or thus $\dot{m}_F(1 + \nu) = -\dot{m}_{\text{Pr}}$

Hence

$$\begin{aligned} \dot{m}_F [h_F + \nu h_{\text{Ox}} - (\nu + 1)h_{\text{Pr}}] &= \dot{Q}_{f-i} + \dot{Q}_{f-\infty} \\ \implies \dot{m}_F \Delta h_c + \dot{m}_F c_{pg} [(T_f - T_{\text{ref}}) + \nu(T_f - T_{\text{ref}}) - (\nu + 1)(T_f - T_{\text{ref}})] &= \dot{Q}_{f-i} + \dot{Q}_{f-\infty} \end{aligned}$$

Since we assume c_{pg} constant, then Δh_c independent of temperature; thus choose flame temperature as reference state,

$$\implies \dot{m}_F \Delta h_c = \dot{Q}_{f-i} + \dot{Q}_{f-\infty}$$

Therefore, the trick that we can employ is to simply only consider the Δh_c at the flame temperature.

$$\begin{array}{ccc} \text{interface} & \xleftarrow{\dot{q}_{f-i}} & \parallel \text{flame} \parallel \xrightarrow{\dot{q}_{f-\infty}} \infty \\ \text{i} & \xrightarrow{\dot{m}_f h_f} & \parallel f \parallel \xleftarrow{\dot{m}_{\text{Ox}} h_{\text{Ox}} = \nu \dot{m}_f h_{\text{Ox}}} \infty \\ & & \parallel f \parallel \xrightarrow{\dot{m}_p h_p = (\nu + 1) \dot{m}_f h_p} \infty \end{array}$$

Thus

$$\begin{aligned}\dot{m}_f[h_f + \nu h_{\text{Ox}} - (\nu + 1)h_f] &= \dot{q}_{f-i} + \dot{q}_{f-\infty} \\ \Delta h_c(T_{\text{ref}}) &= h_{f,f}^\circ + \nu h_{\text{Ox},f}^\circ - (\nu + 1)h_{f,f}^\circ \\ \dot{m}_f \Delta h_c &= \dot{q}_{f-i} + \dot{q}_{f-\infty}\end{aligned}$$

(66)

Using $\left.\frac{dT}{dr}\right|_{\text{rs}}$ from inner+outer energy conservation.

(60)

$$\frac{c_{pg}}{\Delta h_c} \left[\frac{(T_s - T_f) \exp(-Z_T \dot{m}_F / r_f)}{\exp(-Z_T \dot{m}_F / r_s) - \exp(-Z_T \dot{m}_F / r_f)} - \frac{(T_\infty - T_f) \exp(-Z_T \dot{m}_F / r_f)}{[1 - \exp(-Z_T \dot{m}_F / r_f)]} \right] - 1 = 0$$

17.5. **Liquid-Vapor Equilibrium at Droplet Surface.** Clausius-Clapeyron Eqn.

$$p_v = p_0 \exp \left[\frac{h_{fg}}{R} \left(\frac{1}{T_0} - \frac{1}{T} \right) \right]$$

can rewrite as

$$p_v = A \exp \left(\frac{-B}{T_s} \right)$$

A, B are constants for a given liquid.

So fuel partial pressure at surface \simeq equilibrium vapor.

pressure $p_{F,s} = A \exp \left(\frac{-B}{T_s} \right)$

Fuel mole fraction $X_{Fs} = \frac{p_{Fs}}{p}$ and

$$Y_{F,s} = \frac{N_F m_F}{N_F m_F + N_P m_P} = X_F \frac{m_F}{X_F m_F + (1 - X_F) m_P}$$

(61)

$$Y_{F,s} = \frac{A \exp(-B/T_s) m_F}{A \exp(-B/T_s) m_F + [p - A \exp(-B/T_s)] m_P}$$

17.5.1. *Empirical relations.* Turns (2011) [9] quotes Law and Williams (1972) (cf. Law, C.K., and Williams, F.A., “Kinetics and Convection in the Combustion of Alkane Droplets,” *Combustion and Flame*, 19(3): 393-406 (1972)) for empirical relations to use for k_g and c_{pg} (EY : 20160227 I need to check this article out myself).

Therefore,

(62)

$$\begin{aligned}c_{pg} &= c_{pF}(\overline{T}) \\ k_g &= 0.4k_F(\overline{T}) + 0.6k_{\text{Ox}}(\overline{T}) \\ \rho_l &= \rho_l(T_s)\end{aligned}$$

Also, from 20160129 Dr. Polk, Ae121b Winter 2016, $\overline{T} = (T_s + T_f)/2$, and a good initial guess for T_s and T_f are $T_b(P)$ and T_{ad} , the adiabatic flame temperature for stoichiometric mixture.

After algebraic manipulations (which can be done in `BurningDroplet.py` with `sympy` in Python, instead of by hand), then we have 5 equations

(63)

$$\dot{m}_F = \frac{4\pi k_g r_s}{c_{pg}} \ln[1 + B_{oq}]$$

(64)

$$T_f = \frac{q_{i-l} + h_{fg}}{c_{pg}(1 + \nu)} [\nu B_{oq} - 1] + T_s$$

(65)

$$r_f = r_s \frac{\ln(1 + B_{oq})}{\ln[(\nu + 1)/\nu]}$$

•

$$Y_{F,s} = \frac{B_{oq} - 1/\nu}{B_{oq} + 1}$$

•

(67)

$$T_s = \frac{-B}{\ln \left[\frac{-Y_{F,s} p m_P}{A(Y_{F,s} m_F - Y_{F,s} m_P - m_F)} \right]}$$

with

$$B_{oq} := \frac{\Delta h_c / \nu + C_p (T_\infty - T_s)}{q_{i-l} + h_{fg}}$$

Keep in mind the constant decrease in the size of the droplet, measured by diameter squared, or D^2 ,

$$\frac{dD^2}{dt} = -\kappa$$

with

$$\kappa := \frac{8k_g}{\rho_l c_{pg}} \ln(1 + B_{oq})$$

I will compare with Turns (2011) [9], from pp. 378, Section **Simple Model of Droplet Burning** of Chapter 10 Droplet Evaporation and Burning.

Turns (2011) [9] lists 10 assumptions on pp. 379 for the Simple Model of Droplet Burning. Beginning with his fifth assumption,

The gas phase consists of only 3 species: fuel vapor, oxidizer, and combustion products. The gas phase region is divided into 2 zones, inner zone, and outer zone. So for this spherically symmetric problem, parametrize space by radius $r \in \mathbb{R}^+$ from the center of the (fuel) droplet.

Assumption 3 of Turns (2011) [9]: Fuel is single-component liquid with 0 solubility for gases. Phase equilibrium at liquid-vapor interface (which I believe is at the surface of the droplet, r_s).

Assumption 10 of Turns (2011) [9]: “Liquid fuel droplet is the only condensed phase; no soot or liquid water is present.”

$Y_F \equiv$ fuel vapor mass fraction.

17.5.2. *Inner zone.* Define the inner zone $r \in [r_s, r_f]$ between droplet surface r_s and flame front r_f .

Inner zone only contains fuel vapor and combustion products: binary diffusion prevails.

$T_s \equiv$ droplet surface temperature.

$T_f \equiv$ flame temperature.

$Y_{F,s} = Y_F(r = r_s) \equiv Y_F(r_s) =$ fuel vapor mass fraction at the droplet surface

$Y_F(r_f) = 0$ means that all the fuel is consumed by the time we reach the flame front.

Turns mentions that “a more elegant approach” is described in Kuo’s **Principles of Combustion**, that “combines the species and energy equations to create a conserved-scalar variable.

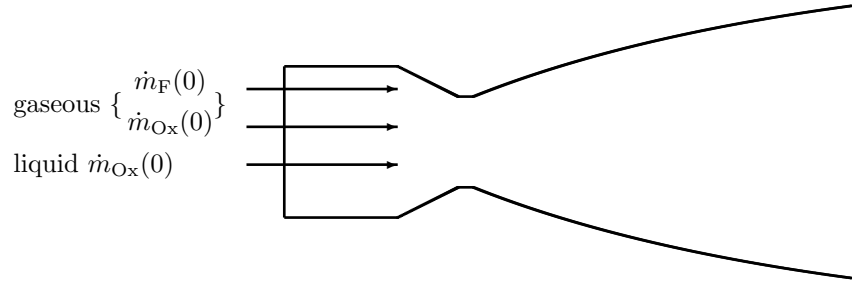
18. COMBUSTION CHAMBER FLOW MODEL

cf. 20160202 Ae121b Dr. Polk.

Parametrize the axisymmetric axis $x \in \mathbb{R}$, with $x = 0$ being the inlet exit into the combustion chamber.

Consider the (mass) inlet flow of the oxidizer, $\dot{m}_{\text{Ox}}(x = 0) \equiv \dot{m}_{\text{Ox}}(0)$ and the (mass) inlet flow of the fuel F , $\dot{m}_{\text{F}}(x = 0) \equiv \dot{m}_{\text{F}}(0)$. Here, fuel is a vapor (i.e. gas).

Suppose some of the fuel from the inlet will come in as a liquid l ; denote the inlet flow of liquid fuel as $\dot{m}_l(0)$.



Consider the following assumptions:

- There are only **2 phases**: the gaseous phase of fuel + oxidizer, and the liquid phase of liquid fuel.
- Assume 1-dimensional flow, with no diffusion in the bulk fluid.
- Assume constant pressure in the combustion chamber, and adiabatic flow, with no work on the flow.
- Fuel is injected as a *monodisperse* spray, meaning that all the droplets have the same initial diameter.
- Gas phase is in equilibrium at even the axial position x .

The parameters of interest along the axis x are the following:

$$\begin{array}{ll} \dot{m}_g(x) & \dot{m}_l(x) \\ T_g(x) & D(x) \\ \phi_g(x) & v_d(x) \\ v_g(x) & \end{array}$$

18.1. **Inlet conditions.** $\dot{m}_g(0)$ refers to the inlet.

$\dot{m}_f(0)$ refers to flow rate of fuel injected as gas.

Then by accounting for the total mass of the gas injected at the inlet, being part gaseous fuel, f , and gaseous oxidizer, Ox,

$$\dot{m}_g(0) = \dot{m}_f(0) + \dot{m}_{\text{Ox}}(0)$$

By definition,

$$\left(\frac{F}{O}\right)_x = \frac{\dot{m}_F(x)}{\dot{m}_{\text{Ox}}(x)}$$

Now

$$\phi_g(x = 0) = \phi_g(0) = \frac{\dot{m}_f(0)/\dot{m}_{\text{Ox}}(0)}{(F/O)_{\Phi=1}} = 0.45$$

by definition of an *equivalence ratio*, ϕ_g for the gas g , and so we can determine $\dot{m}_f(0)$ after doing some algebra

$$(68) \quad \dot{m}_f(0) = \dot{m}_{\text{Ox}}(0) \left(\frac{F}{O}\right)_{\Phi=1} \phi_g(0)$$

Now, if given an overall equivalence ratio, ϕ_{overall} at a point x along the axis,

$$(69) \quad \phi_{\text{overall}}(x = 0) = \frac{\dot{m}_{\text{fuel}}(0)/\dot{m}_{\text{Ox}}(0)}{(F/O)_{\Phi=1}} = \frac{\frac{\dot{m}_f(0)}{\dot{m}_{\text{Ox}}(0)} + \frac{\dot{m}_l(0)}{\dot{m}_{\text{Ox}}(0)}}{(F/O)_{\Phi=1}} = \phi_g(0) + \frac{\frac{\dot{m}_l(0)}{\dot{m}_{\text{Ox}}(0)}}{(F/O)_{\Phi=1}}$$

Clearly, the total mass of fuel inputted in is the mass of the fuel as a gas and mass of the fuel as a liquid (droplets):

$$(70) \quad \dot{m}_{\text{Fuel}}(0) = \dot{m}_f(0) + \dot{m}_l(0)$$

By the physical parameters of a given setup, such as injection velocity of the gas $v_g(x = 0) \equiv v_g(0)$ and total fuel injector cross-sectional area, $A_{\text{inlet Fuel}}$

$$(71) \quad \dot{m}_g(0) = \rho_g v_f(0) A_{\text{inlet Fuel}} = \rho_g(T_{\text{inlet}}, P) v_f(0) A_{\text{inlet Fuel}}$$

where we remember that ρ_f is a function of inlet temperature T_{inlet} and pressure P of the combustion chamber. ρ_g can be obtained by setting the state of the gas with *Cantera*, with the appropriate equivalence ratio for the gases $\phi_g(0)$.

From Eq. 68, we can obtain $\dot{m}_{\text{Ox}}(0)$ with the help of Eq. 71, namely

$$\dot{m}_g(0) = \dot{m}_{\text{Ox}}(0) \left(1 + \left(\frac{F}{O}\right)_{\Phi=1} \phi_g(0)\right)$$

And then, from the first equality of Eq. 69, and Eq. 70, then $\dot{m}_l(0)$ can be obtained, namely

$$\phi_{\text{overall}}(0)(F/O)_{\Phi=1} = \frac{\dot{m}_f(0) + \dot{m}_l(0)}{\dot{m}_{\text{Ox}}(0)}$$

18.2. **Gas phase continuity.** Since for a gas, $pV = N\tau$, so

$$p = \frac{N}{V}\tau = \frac{MN}{V} \frac{1}{M} \frac{T}{k_B} = \rho \frac{1}{M k_B} T = \rho R_M T$$

$$(72) \quad \boxed{v_g = \frac{\dot{m}_g}{\rho_g A} = \frac{\dot{m}_g R_U T_g}{m_g P A}}$$

$$\text{i.e. } v_g = \frac{\dot{m}_g \tau_g}{M_g p A}$$

18.3. **Gas phase Energy Conservation.** For $H_{\text{tot}} = H_{\text{tot}}(\tau, p, \{N_i\}, x) \equiv H_{\text{tot}}(x) \in C^\infty(\Sigma \times N) = C^\infty(\Sigma \times \mathbb{R})$, $H_{\text{tot}} = m_g h_g + m_l h_l$.

For $\dot{H}_{\text{tot}} = \dot{m}_g h_g + \dot{m}_l h_l$ where I take into consideration the steady state assumption, then by

$$Q = 0, dH = 0, \text{ for } dp = 0 \implies \frac{d}{dx}(\dot{m}_g h_g) + \frac{d}{dx}(\dot{m}_l h_l) = 0$$

Assume constant T droplets, so $h_l = \text{constant}$

$$(73) \quad \frac{dh_g}{dx} = \frac{-1}{\dot{m}_g} \left[h_g \frac{d\dot{m}_g}{dx} + h_l \frac{d\dot{m}_l}{dx} \right]$$

Since $h_g = h_g(T_g, p, \phi_g) \in C^\infty(\Sigma)$,

$$(74) \quad \frac{dh_g}{dx} = \frac{\partial h_g}{\partial T} \frac{dT}{dx} + \frac{\partial h_g}{\partial \phi_g} \frac{d\phi_g}{dx} \quad (p \text{ is const.})$$

Equating Eqns. 73, 74, then

$$\frac{dT_g}{dx} = \left[\frac{-1}{\dot{m}_g} \left(h_g \frac{d\dot{m}_g}{dx} + h_l \frac{d\dot{m}_l}{dx} \right) - \frac{\partial h_g}{\partial \phi_g} \frac{d\phi_g}{dx} \right] / \frac{\partial h_g}{\partial T_g}$$

Using mass conservation on fuel, $\frac{d\dot{m}_g}{dx} = -\frac{d\dot{m}_l}{dx}$, so

$$(75) \quad \boxed{\frac{dT_g}{dx} = \left[\frac{-1}{\dot{m}_g} \frac{d\dot{m}_g}{dx} (h_g - h_l) - \frac{\partial h_g}{\partial \phi_g} \frac{d\phi_g}{dx} \right] / \frac{\partial h_g}{\partial T_g}}$$

18.4. Droplet momentum conservation.

$$F_d = m_d \frac{dv_d}{dt} = m_d \frac{dx}{dt} \frac{dv_d}{dx} = m_d v_d \frac{dv_d}{dx}$$

$$F_d \text{ for a sphere} = C_d \rho_g \frac{v_{\text{rel}}^2}{2} \frac{\pi D^2}{4}$$

Recall the Reynolds number for a sphere moving through a fluid,

$$\text{Re} = \frac{D_{\text{droplet}} u_{\text{rel}} \rho_{\text{fluid}}}{\mu_{\text{fluid}}}$$

where μ viscosity of the fluid, $\rho_{\text{fluid}} = \rho_g$, $\mu_{\text{fluid}} = \mu_g$. Keep in mind that $[D_{\text{droplet}} u_{\text{rel}} \rho_{\text{fluid}}] = m \cdot \frac{m}{s} \cdot \frac{\text{kg}}{m^3} = \frac{kg}{m \cdot s}$.

$$(76) \quad \begin{aligned} \frac{dD^2}{dx} &= \frac{-K}{v_d} \\ \frac{dv_d}{dx} &= \frac{3C_D \rho_g (v_g - v_d) |v_g - v_d|}{4\rho_l D v_d} \\ \frac{dT_g}{dx} &= \left[\frac{-1}{\dot{m}_g} \frac{d\dot{m}_g}{dx} (h_g - h_l) - \frac{\partial h_g}{\partial \phi_g} \frac{d\phi_g}{dx} \right] / \frac{\partial h_g}{\partial T_g} \end{aligned}$$

droplet $\in \Sigma \times \Gamma(TN)$

$(\rho_l, p_l) \times (x, v_d(x))$

assume constant T droplets.

So I claim by clausius Clapeyron $T_l = T_{\text{boil}}$

Let's take a look at Eq. 76.

- Looking at $\frac{dD^2}{dx} = \frac{-K}{v_d}$, and looking at what each of the factor's respective formulae,

$$\frac{dD^2}{dx} = \frac{-K}{v_d} \implies \begin{aligned} K &= \frac{8k_g}{\rho_l C_{pg}} \ln(1 + B_{oq}) \\ B_{oq} &= \frac{\Delta h_c / \nu + C_{pg}(T_\infty - T_s)}{q_{i-l} + h_{fg}} \end{aligned}$$

Then, examining the terms *related to gases* that depend on Σ and spatial manifold N , parametrize by $x \in \mathbb{R}$,

$$\begin{aligned} k_g &= k_g(T_g, P, \phi_g) \in C^\infty(\Sigma) \\ k_g(T_g(x), P, \phi_g(x)) &\in C^\infty(\Sigma_x) \implies k_g(x) \in C^\infty(N) \end{aligned}$$

with $\Sigma \rightarrow N$ being a fibered bundle, with fibers Σ_x .

Likewise,

$$\begin{aligned} C_{Pg} &= C_{Pg}(T_g, P, \phi_g) \in C^\infty(\Sigma) \\ C_{Pg}(T_g(x), P, \phi_g(x)) &\in C^\infty(\Sigma_x) \text{ and } C_{Pg}(x) \in C^\infty(N) \end{aligned}$$

In B_{oq} , for term h_{fg} , for the heat of formation,

$$\begin{aligned} h_{fg} &= h_{fg}(T^0, P, \phi_g) \in C^\infty(\Sigma) \\ h_{fg}(T^0, P, \phi_g(x)) &\in C^\infty(\Sigma_x) \text{ and } h_{fg}(x) \in C^\infty(N) \end{aligned}$$

and since assuming $C_{PF} = C_{POx} = C_{pg}$ then

$$\Delta h_c = \Delta h_c(T) \in C^\infty(\Sigma) \text{ and } \Delta h_c(T = T_{\text{ref}}) \equiv \Delta h_c(T_{\text{ref}}) = \Delta h_c(T)$$

with the upshot that the “heat of combustion” is the same at $T = T_{\text{ref}}$ and at arbitrary T .

For terms involving the droplet, the “*liquid*”,

$$\rho_l = \rho_l(T) \in C^\infty(\Sigma) \rho_l(T(x)) \in C^\infty(\Sigma_x) \text{ and } \rho_l(x) \in C^\infty(N)$$

In our particular case, we assume the droplet has uniform temperature, which happens to be the temperature at the surface, which then happens to be dependent on pressure, by the Clausius-Clapeyron relation. Keep in mind that the relation $\rho_l = \rho_l(T)$ is from the relation involving the volumnic thermal expansion α of a *liquid*: $\rho = \frac{\rho_{T_0}}{1 + \alpha(T - T_0)}$, which, I think, is an empirical relation, as α needs to be measured and is a parameter to input in.

$$\rho_l = \rho_l(T_s) = \rho_l(T_b) \in C^\infty(\Sigma)$$

- Looking at $\frac{dv_d}{dx} = \frac{3C_D \rho_g (v_g - v_d) |v_g - v_d|}{4\rho_l D v_d}$, and taking a look at terms involving the *gaseous* phase,

$$v_g \in TN(x, v_g) = v_g(x) \in T_x N$$

and

$$\begin{aligned} \rho_g &= \rho_g(T_g, p, \phi_g) \in C^\infty(\Sigma) \\ \rho_g(T_g(x), p, \phi_g(x)) &\in C^\infty(\Sigma_x) \text{ and } \rho_g(x) \in C^\infty(N) \end{aligned}$$

Looking at terms involving the *liquid* phase, the **droplet**,

$$\begin{aligned} v_d &\in TN \\ (x, v_d) &= v_d(x) \in T_x N \\ D &= D(x) \in C^\infty(N) \end{aligned}$$

Then, examining the formulae for the Reynolds number Re and C_D , affecting the drag on the droplet,

$$\text{Re}_{D,\text{rel}} = \frac{D |v_g - v_d| \rho_g}{\mu}$$

with viscosity μ being a transport property of the *gas* that the droplet is moving around in, and

$$\begin{aligned} \mu &= \mu(T_g, P, \phi_g) \in C^\infty(\Sigma) \\ \mu(T_g(x), P, \phi_g(x)) &\in C^\infty(\Sigma_x) \text{ and } \mu(x) \in C^\infty(N) \end{aligned}$$

C_D drag coefficient is fairly straightforward, depending upon $x \in N$, but through $\text{Re}_{D,\text{rel}}$, *only*:

$$C_D \simeq \frac{24}{\text{Re}_{D,\text{rel}}} + \frac{6}{1 + \sqrt{\text{Re}_{D,\text{rel}}}} + 0.4$$

- Looking at $\frac{dT_g}{dx} = \left[\frac{-1}{\dot{m}_g} \frac{d\dot{m}_g}{dx} (h_g - h_l) - \frac{\partial h_g}{\partial \phi_g} \frac{d\phi_g}{dx} \right] / \frac{\partial h_g}{\partial T_g}$, the terms involving the *gas* phase are the following:

$$\begin{aligned} T_g &\in \Sigma \text{ and } T_g(x) \in \Sigma(x) \equiv \Sigma_x \text{ and} \\ T_g(x) &\in C^\infty(N) \\ \phi_g &= \phi_g(x) \in C^\sigma(\Sigma_x) \text{ and} \\ \phi_g(x) &\in C^\infty(N) \\ h_g &= h_g(T_g, P, \phi_g) \in C^\infty(\Sigma) \\ h_g(T_g(x), P, \phi_g(x)) &\in C^\infty(\Sigma_x) \text{ and } h_g(x) \in C^\infty(N) \end{aligned}$$

and the (physical) quantity that doesn't depend upon Σ for the *gaseous* phase is

$$\dot{m}_g = \dot{m}_g(x) \in C^\infty(N)$$

For the terms involving the droplet, or the *liquid* phase,

$$\begin{aligned} h_l &= h_l(T_l, P) \in C^\infty(\Sigma) \\ h_l(T_s, P) &= h_l(T_{\text{boil}}, P) \in C^\infty(\Sigma) \end{aligned}$$

18.4.1. *Considerations of the physical parameters involved.* Start with the “mass flow” rates that are smooth functions of space and N . There are 2 reactant *species* involved, the fuel and oxidizer, Ox, amongst 2 phases, gas and liquid, considered.

$$\dot{m}_g = \dot{m}_g(x), \dot{m}_f = \dot{m}_f(x), \dot{m}_{\text{Ox}} = \dot{m}_{\text{Ox}}(x), \dot{m}_l = \dot{m}_l(x) \in C^\infty(N)$$

Clearly, we have

$$(77) \quad \dot{m}_g = \dot{m}_f + \dot{m}_{\text{Ox}}$$

as gaseous oxidizer is injected in, and gaseous fuel is from the evaporation off the liquid droplet.

Modelling the liquid fuel being injected into the combustion chamber through an injection plate of total fuel injection cross-section $A_{\text{tot fuel inj}}$, at injection speed $v_d(0)$, then clearly

$$\dot{m}_l(x=0) = \dot{m}_l(0) = \rho_l v_d(0) A_{\text{tot fuel inj}} = \rho_l(T_b, P) v_d(0) A_{\text{tot fuel inj}}$$

I would argue that for the liquid density $\rho_l = \rho_l(T_b, P)$, which is completely specified by 2 thermodynamic quantities in Σ^l , $(T, P) \in \Sigma^l$, the temperature is T_b , the boiling temperature specified by the Clausius-Clapeyron relation for a given combustion chamber pressure P . The inlet temperature is usually much higher than the boiling temperature of our fuel: we wouldn’t have any liquid, or droplets, if its local temperature is this inlet temperature. I will also argue that the liquid droplet is entirely in thermal equilibrium and at its surface, fuel molecules are evaporating away at T_b : so the entire liquid is at $T_b = T_b(P)$. This remains the case as the liquid droplet travels throughout the combustion chamber.

Out of this injection plate, for a droplet of initial size D_0 , then for number of droplets per unit time emerging out of the injection plate, \dot{N} , clearly

$$\dot{m}_l(0) = \frac{\dot{N} \rho_l \pi D_0^3}{6}$$

Consider the so-called equivalence ratio; in this case the overall equivalence ratio ϕ_{overall} . It is an interesting quantity that relates purely physical quantities, masses or number of particles (i.e. moles), to thermodynamic variables $\{N_i\}_i \in \Sigma$, i.e. $\phi_{\text{overall}} \in C^\infty(\Sigma)$. By definition of ϕ_{overall} ,

$$(78) \quad \phi_{\text{overall}} := \frac{\frac{\dot{m}_f + \dot{m}_l}{\dot{m}_{\text{Ox}}}}{(F/O)_{\Phi=1}} = \frac{\frac{\dot{m}_g - \dot{m}_{\text{Ox}} + \dot{m}_l}{\dot{m}_{\text{Ox}}}}{(F/O)_{\Phi=1}} \implies \dot{m}_g = \left(\phi_{\text{overall}} \left(\frac{F}{O} \right)_{\Phi=1} + 1 \right) \dot{m}_{\text{Ox}} - \dot{m}_l$$

Also, by definition of the equivalence ratio of gaseous fuel to oxidizer (which is assumed to have completely vaporized into gas), $\phi_g \in C^\infty(\Sigma)$,

$$(79) \quad \phi_g := \frac{\frac{\dot{m}_f}{\dot{m}_{\text{Ox}}}}{\left(\frac{F}{O} \right)_{\Phi=1}} = \frac{\frac{\dot{m}_g - \dot{m}_{\text{Ox}}}{\dot{m}_{\text{Ox}}}}{\left(\frac{F}{O} \right)_{\Phi=1}} \implies \dot{m}_g = \left(\left(\frac{F}{O} \right)_{\Phi=1} \phi_g + 1 \right) \dot{m}_{\text{Ox}}$$

By equating the expressions for \dot{m}_g in Eqns. 78, 79, then

$$\dot{m}_{\text{Ox}} = \frac{\dot{m}_l}{(\phi_{\text{overall}} - \phi_g) \left(\frac{F}{O} \right)_{\Phi=1}}$$

Notice that $\phi_{\text{overall}} > \phi_g \geq 0$.

I haven’t mentioned where this is all occurring. Here is one important fact about the physical setup:

$$\dot{m}_{\text{Ox}}(x) = \dot{m}_{\text{Ox}}(0) \quad \forall x \in N$$

i.e. the oxidizer is constantly being injected in and flowing uniformly through the chamber.

At $x = 0$, the start of the combustion chamber,

$$\dot{m}_{\text{Ox}}(0) = \frac{\dot{m}_l(0)}{(\phi_{\text{overall}}(x=0) - \phi_g(x=0)) \left(\frac{F}{O} \right)_{\Phi=1}}$$

Let’s take stock of the quantities and its properties that we’ve discussed so far:

$$\begin{aligned} \dot{N}(x) &= \dot{N}(0) & \forall x \in N & \quad (\text{droplets don't split or breakup, and don't combine together}) \\ \dot{m}_{\text{Ox}}(x) &= \dot{m}_{\text{Ox}}(0) & \forall x \in N & \quad (\text{constant injection of oxidizer}) \end{aligned}$$

$$\dot{m}_g(x) = \left(\left(\frac{F}{O} \right)_{\Phi=1} \phi_g(x) + 1 \right) \dot{m}_{\text{Ox}}(0) = \dot{m}_g(\phi_g(x)) \in C^\infty(\Sigma_x)$$

where we notice that \dot{m}_g is a function of space $x \in N$ or a function of thermodynamic property $\phi_g(x)$. Also note that the quantity $\left(\frac{F}{O} \right)_{\Phi=1}$, the so-called *stoichiometric* ratio between the *mass* of fuel to *mass* of oxidizer that combines together in a reaction that goes to 100% completion (hence “stoichiometric”); once the chemical reaction is chosen, this quantity remains *fixed* throughout $x \in N$.

Thus, so far, we should carry forward these 3 physical quantities, that are completely determined by inlet conditions:

$$(\dot{m}_l(0), \dot{N}(0), \dot{m}_{\text{Ox}}(0)) \in C^\infty(N_{x=0}) \times C^\infty(N_{x=0}) \times C^\infty(N_{x=0})$$

For $x \in N$ (i.e. along the rest of the combustion chamber), and recalling this derivation,

$$\frac{d}{dx} \dot{m}_l = \frac{d}{dx} (\dot{N} \rho_l \frac{\pi D^3}{6}) = \frac{3}{2} \frac{\dot{m}_l(0)}{D_0^3} D \frac{dD^2}{dx} \xrightarrow{\frac{dD^2}{dx} = \frac{-K}{v_d}} \frac{-3}{2} \frac{\dot{m}_l(0)}{D_0^3} D \frac{K}{v_d}$$

Droplet evaporation off surface is expressed as:

$$\frac{d\dot{m}_g}{dx} = -\frac{d\dot{m}_g}{dx} \implies \frac{d\dot{m}_g}{dx} = \frac{3}{2} \frac{\dot{m}_l(0)}{D_0^3} D \frac{K}{v_d}$$

and thus, taking the spatial derivative of both sides of Eq. 79

$$\frac{d\phi_g}{dx} = \frac{1}{\left(\frac{F}{O} \right)_{\Phi=1} \dot{m}_{\text{Ox}}(0)} \frac{d\dot{m}_g}{dx} = \frac{1}{\left(\frac{F}{O} \right)_{\Phi=1} \dot{m}_{\text{Ox}}(0)} \frac{3}{2} \frac{\dot{m}_l(0)}{D_0^3} D \frac{K}{v_d}$$

By gas phase continuity,

$$v_g = v_g(x) = \frac{m_g}{\rho_g A_{cc}} = \frac{\dot{m}_g(X) R_U T_g(x)}{MW_g P A_{cc}} = \frac{\left(\left(\frac{F}{O} \right)_{\Phi=1} \phi_g(x) + 1 \right) \dot{m}_{\text{Ox}}(0) R_U T_g(x)}{MW_g P A_{cc}}$$

We can obtain $\dot{m}_l = \dot{m}_l(x)$ (i.e. as a function of $x \in N$):

$$\dot{m}_l(x) = \frac{\dot{N} \rho_l(T_b, P) \pi D^3}{6}$$

Notice that it is a function of $D = D(x) \in C^\infty(N)$.

Now consider that

$$\frac{d\phi_{\text{overall}}}{dx} = \frac{\frac{d\dot{m}_g}{dx} - 0 + \frac{d\dot{m}_l}{dx}}{\dot{m}_{\text{Ox}} \left(\frac{F}{O} \right)_{\Phi=1}} = 0$$

i.e. ϕ_{overall} is constant throughout the chamber! This could be attributed to overall mass conservation, but it’s explicitly shown above.

Therefore

$$(80) \quad \begin{aligned} \dot{m}_g(x) &= \left(\phi_{\text{overall}} \left(\frac{F}{O} \right)_{\Phi=1} + 1 \right) \dot{m}_{\text{Ox}}(0) - \dot{m}_l(x) \\ \implies \phi_g(x) &= \left(\frac{\dot{m}_g(x)}{\dot{m}_{\text{Ox}}(0)} - 1 \right) \frac{1}{\left(\frac{F}{O} \right)_{\Phi=1}} \end{aligned}$$

and so $\phi_g(x)$ can be obtained from this formula above, $\forall x \in N$.

It may be interesting to notice that, and to take stock of what we’ve uncovered so far, that the dynamic quantities (dynamic in that they’ll be involved in our system of ordinary differential equations (ODEs))

$$(T_g(x), D(x), v_d(x)) \in C^\infty(\Sigma^g) \times C^\infty(N) \times TN$$

along with initial (inlet) conditions

$$(\dot{m}_{\text{Ox}}(0), \dot{N}, \dot{m}_l(0)) \in C^\infty(N_0) \times C^\infty(N_0) \times C^\infty(N_0)$$

completely determines, algebraically, the quantities

$$(\phi_g(x), \dot{m}_g(x), v_g(x)) \in C^\infty(\Sigma^g) \times C^\infty(N) \times TN, \quad \dot{m}_l(x) \in C^\infty(N)$$

all throughout $x \in N$.

Part 6. Numerical Computation; Scientific Computation

19. SYSTEMS OF ORDINARY DIFFERENTIAL EQUATIONS ODES

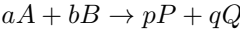
EY : 20160228 There’s a lack of fully general and useful examples of solving a system of ordinary differential equations (ODEs) in `scipy` if one does a search on Google. There is the [scipy cookbook](#), which has toy examples for the [Coupled spring-mass system](#), [Korteweg de Vries equation](#), [Matplotlib: lotka volterra tutorial](#), [Modeling a Zombie Apocalypse](#), [Theoretical ecology: Hastings and Powell](#). Then, there’s Kevin Dunn’s material for Process Model Formulation and Solution, which is *excellent* for its layout, thoroughness, side-by-side comparison between Matlab and Python (scipy), and *real-world* examples, though related to *chemical engineering* (but which is an interesting and useful subject in its own right; so we’d might as well learn some chemical engineering).

There was some [MIT OCW material on CSTR](#), but it only solved for the steady-state solution.

19.1. Continuously Stirred Tank Reactor (CSTR). Consider the *continuously stirred tank reactor* (CSTR) or vat or backmix reactor.

The main assumption is of perfect mixing - perfect mixing throughout, with each locality’s mixing the same as another locality, so reaction rate is equal everywhere throughout a volume V .

From [wikipedia’s Reaction Rate](#) article, for a general chemical reaction of the form



the reaction rate r is given by

$$r = \frac{-1}{a} \frac{d[A]}{dt} = -\frac{1}{b} \frac{d[B]}{dt} = \frac{1}{P} \frac{d[P]}{dt} = \frac{1}{q} \frac{d[Q]}{dt}$$

where $[X] \equiv$ concentration of substance X .

19.1.1. Second-Order Reaction Rate. A second order reaction rate would be of the form

$$r = kC_A^2$$

and if $[C_A] = \text{mol}/L$, then the units for k are $[k] = \frac{1/s}{\text{mol}/L} = \frac{L}{\text{mol}\cdot s}$.

Let’s count (or account) for number of particles N . In general, for concentration $n := \frac{N}{V}$, and $N = N(t)$, dependent upon time, then

$$N = \int_V n \text{vol}^n$$
$$\frac{dN}{dt} \equiv \dot{N} = \int_V \frac{\partial n}{\partial t} \text{vol}^n + \int_V \text{div}(nu) \text{vol}^n = \int_V \frac{\partial n}{\partial t} \text{vol}^n + \int_{\partial V} nu^j dS_j$$

One can interpretation the terms of the last equation immediately above as [accumulation] = [generation] + ([in] – [out]), respectively.

Assuming perfect mixing, so the reaction rate equal everywhere throughout volume V , so $\frac{\partial n}{\partial t}$ the same throughout V , and that $\partial V = S_{\text{in}} \amalg S_{\text{out}}$, then

$$\frac{dN}{dt} = \nu k C^2 V + \int_{S_{\text{in}}} nu^j dS_j + \int_{S_{\text{out}}} nu^j dS_j$$

Now for $F(t)$ being the inlet flow of volume per second or volume per minute, then

$$\int_{S_{\text{in}}} nu^j dS_j = F(t) C_{A,\text{in}}$$

Supposing the flow rate out, *in terms of volume*, is the same as flow rate in,

$$\int_{S_{\text{out}}} nu^j dS_j = -F(t) C_A$$

i.e., after perfect mixing, obtaining a concentration $n = C_A$, then multiply it by the volume per second or volume per minute flowing out, $-F(t)$ to obtain the number of particles rushing out.

I think the flow rate out being the same as flow rate in, *in terms of volume*, is what Kevin Dunn meant in the [first example problem for ODE integration](#).

Thus, after dividing by fixed V (the volume of the tank),

$$\xrightarrow{\frac{1}{V}} \frac{dn_A}{dt} \equiv \frac{dC_A}{dt} = \frac{F(t)C_{A,\text{in}}}{V} - \frac{F(t)C_A}{V} + \nu_A k C_A^2$$

For the “right-hand side” (RHS) of an ODE with constant coefficients, and a RHS that doesn’t depend on time, then solving this problem is pretty easy. In the language of flows on a manifold, letting curve, or our solution, be $\gamma : \mathbb{R} \rightarrow M$, so $\dot{\gamma} \in \mathfrak{X}(M)$, a vector field on smooth manifold M , then

$$\frac{d\gamma}{dt} = f(\gamma, t) = f(\gamma) \implies \frac{d\gamma}{f(\gamma)} = dt$$

A copy of Dunn’s implementation is in `CSTRconstant.py` and `CSTRdynamic.py`, copied from verbatim.

EY : 20160229 Things I need to understand further are the `set_integrator` method of `scipy.integrate.ode`, with option ‘vode’ and method ‘bdf’ (i.e. `method=’bdf’`) and what a [Stiff Equation](#) is.

Part 7. Cantera

20. PROFESSOR DAVID GOODWIN, 1957-2012

“Keep on working to make this world a better place. It is important to get your thoughts away from your problems and focused on those who are less fortunate than you are.”

“We can’t control the hand of cards we are dealt, but we sure can determine how we play them.”

- Open-Source, Berkeley license,
- Object-Oriented Structure, multi-interface structure,
- formalism for collision integrals in evaluating transport parameters of dilute ideal gases

cf. <https://code.google.com/archive/p/cantera/wikis/DaveGoodwin.wiki>

The following includes (disorganized) notes on using, implementing, and developing for Cantera.

EY : 20160125 I wanted to obtain the *equation of state* but in Python. This was implemented in Matlab as `eosType.m` here <https://github.com/Cantera/cantera/blob/master/interfaces/matlab/toolbox/%40ThermoPhase/eosType.m> It leads to `thermo_get` for the code says

```
e = thermo_get(tp.tp_id , 18);
```

I don’t think the `tp.tp_id` equivalent in Python is `tp.ID`; searching for `tp_id` led to [ThermoPhase.m](#) and

```
if nargin == 1
    if isa(r, 'ThermoPhase')
        % create a copy
        t = r;
        return
    elseif isa(r, 'XMLNode')
        t.owner = 1;
        hr = xml_hndl(r);
        t.tp_id = thermo_get(hr, 0);
        if t.tp_id < 0
            error(geterr);
        end
    else
        t.owner = 0;
        t.tp_id = r;
```

```
end
t = class(t, 'ThermoPhase');
```

thermo_get of MatLab is here, in the private folder: https://github.com/Cantera/cantera/blob/master/interfaces/matlab/toolbox/%40ThermoPhase/private/thermo_get.m which leads to the MatLab command `ctmethods` and the following inputs/parameters

```
function i = thermo_get(n, job, a, b)
if nargin == 2
    i = ctmethods(20, n, job);
```

I was only able to find the “definition” of `ctmethods` for *MatLab* here: <https://github.com/Cantera/cantera/blob/5926d2db7c0d4919b75ee50828b0adab4e691a51/src/matlab/ctmethods.cpp> which was essentially a so-called “mex function” which was a bunch of different cases; the case I was interested in was for Thermo, leading me to <https://github.com/Cantera/cantera/blob/5926d2db7c0d4919b75ee50828b0adab4e691a51/src/matlab/thermomethods.cpp> with its `thermo_get` and `thermo_set` as above, the so-called “job” I wanted was `job=18`, which lead me to the command `th_eosType`

```
case 18:
    vv = double(th_eosType(n));
    break;
```

In `soundspeed.m` for the Matlab toolbox ThermoPhase, notice the line `if isIdealGas(tp)`. This is implemented in `isIdealGas.m` by returning `v` in

```
if eosType(tp) == 1
    v = 1;
else
    v = 0;
```

21. SPEED OF SOUND AND SOUND SPEED.PY

Take a look at `sound_speed.py` in the github repository for [cantera](#). Take a look at `afrozen` in the code. This is an exact implementation of the usual speed of sound:

$$a_{\text{frozen}}^2 = \left(\frac{\partial p}{\partial \rho}\right)_\sigma$$

Take a look at `afrozen2` in the code. This is an exact implementation of the speed of sound for an ideal gas. Let’s review its derivation.

From $pV = N\tau$, which is how ideal gases behave at any time, then $p = \frac{\rho\tau}{M}$ where $\rho := \frac{MN}{V}$ and M is the mass of a single particle of the species. Then

$$dp = \frac{\tau d\rho + \rho d\tau}{M}$$

and so

$$C_p \frac{d\tau}{\tau} = C_p \left(\frac{dp}{p} - \frac{d\rho}{\rho}\right) = \frac{V dp}{\tau}$$

where the last equality comes from the following derivation

$$\begin{aligned} Q = C_p d\tau - V dp = \tau d\sigma = 0 \text{ since } d\sigma = 0 \text{ for an isentropic process} \\ \implies \frac{C_p}{\tau} d\tau = \frac{V}{\tau} dp \end{aligned}$$

Thus

$$C_p \frac{d\rho}{\rho} = \left(\frac{C_p}{p} - \frac{V}{\tau}\right) dp = \left(\frac{C_p - N}{p}\right) dp = C_V \frac{dp}{p} \implies \left(\frac{\partial p}{\partial \rho}\right)_\sigma = \frac{\gamma\tau}{M}$$

recovering the usual speed of sound for ideal gases.

The derivation on pp. 3-4 of Lecture 14 of MIT OCW 16.512 by Martinez-Sanchez (2005) [16] is fallacious because M is constant always and even in chemical reactions, masses for each species don’t get created or destroyed. The changes due to

chemical reactions show up in dN_i . And even then, we know that the changes in N_i , the number of particles for each species i , dN_i , is governed by the chemical reaction and the number of times the chemical reaction occurs.

For instance, surely the total mass m of the system is thus given by

$$m \equiv \sum_j m_j := \sum_j N_j M_j$$

and so

$$dm = \sum_j dm_j = \sum_j M_j dN_j = \sum_j M_j \nu_j d\hat{N}$$

where $d\hat{N} \in \mathbb{Z}$ indicates how many times a reaction occurs, and $dN_j = \nu_j d\hat{N}$, which tells us how the number of particles in each species changes everytime a chemical reaction occurs, as the chemical reaction is given by $\sum_j \nu_j A_j = 0$ (see pp. 267 Section **Equilibrium in Reactions**, Ch. 9 Gibbs Free Energy and Chemical Reactions of Kittel and Kroemer [4]).

Considering the total number of particles $N = \sum_j N_j$, the total pressure $p = \sum_j p_j$ as a sum of partial pressure p_j , and taking the system as a whole to always obey ideal gas behavior, $pV = N\tau$, then

$$p_j V = N_j \tau$$

so that each species, at any time, behaves as an ideal gas.

Considering the (mass) density of species i , $\rho_j := \frac{N_j M_j}{V}$, and so $\sum_j \rho_j = \frac{m}{V} = \rho$, then rewrite the ideal gas law as $p_j = \frac{N_j \tau}{V} = \frac{\rho_j \tau}{M_j}$.

If one takes the differential as does Martinez-Sanchez (2005) [16] in Lecture 14, pp. 3-4,

$$dp_j = \frac{\tau d\rho_j + \rho_j d\tau}{M_j} \implies \frac{dp_j}{p_j} = \frac{d\rho_j}{\rho_j} + \frac{d\tau}{\tau}$$

and the derivation for speed of sound $a^2 = \left(\frac{\partial p_i}{\partial \rho_j}\right)_\sigma$ goes as above. It is fallacious to say that M_j changes as it is a constant for each species j . The changes in N_j due to chemical reactions are already accounted for in ρ_j .

Instead, one must go back to the general form of heat Q :

$$Q = Q(\tau, p, \{N_i\}) = C_p d\tau + \left(\frac{\partial \tau \sigma}{\partial p}\right)_{\tau, \{N_i\}} + \sum_i \frac{\partial Q}{\partial N_i} dN_i = \tau d\sigma$$

and realize that the implementation of `aequil` in `sound_speed.py` is this general implementation of

$$a^2 = \left(\frac{\partial p}{\partial \rho}\right)_{\sigma, \{N_i\}}$$

that does not assume ideal gas behavior, and also equilibrates the changes in species, dN_i , and does not presume that heat capacities C_V, C_p have the nice form that arises from the assumption of ideal gas behavior.

Part 8. Compressible Flow

I’ll review or verify aspects of *compressible flow* for gas dynamics.

I want to verify the so-called *Bernoulli constant* or Bernoulli invariant $h + u^2/2 + \Phi$ on pp. 38 of Thompson (1988) [15] for *compressible* flow. Note that Thompson uses $h := H/m$ to denote the *specific enthalpy*, whereas I used the notation of $h := H/V$ for the enthalpy density. I’ll use the same notation and one should be able to understand which one is being used from context and/or from the units.

Now for the (internal) energy density ϵ and kinetic energy density k , defined such that

$$\begin{aligned} \epsilon &:= \frac{E}{V} \\ k &= \frac{1}{2} \rho u^2 \implies E_{\text{tot}} = \int_V (\epsilon + k) \text{vol}^n \end{aligned}$$

(which is the notation used by Le Bellac, Mortessagne, Batrouni (2004) [11]), then from the physical principle that the change in total energy of a system is given by the work done on it by the sum of all external forces on the system,

$$\frac{d}{dt} \int_V (\epsilon + k) \text{vol}^n = \int_V \left[\frac{\partial}{\partial t} (\epsilon + k) + \text{div}((\epsilon + k)\mathbf{u}) \right] \text{vol}^n = \int_V \rho \langle u, b \rangle \text{vol}^n + \int_{\partial V} \langle u, T^j \rangle dS_j - \int_{\partial V} q^j dS_j$$

where b is the density associated with (external) body forces on the system and q is the outward-directed heat flux representing the system's heat loss, i.e.

$$q^j dS_j = q \cdot n > 0 \text{ for outward-directed heat flux,} \quad - \int_{\partial V} q^j dS_j \text{ represents heat loss}$$

The *specific energy* (unfortunately denoted as e in Thompson (1988) [15], as it looks like the exponential), the energy per unit mass, or specific internal energy, a thermodynamic quantity, is defined as such:

$$\rho e = \epsilon = \frac{MN}{V} e = \frac{E}{V} \implies e := \frac{E}{m}$$

Then

$$\frac{d}{dt} \int_V (\rho e + k) \text{vol}^n = \int_V \left[\frac{\partial}{\partial t} (\rho e + k) + \text{div}((\rho e + k)\mathbf{u}) \right] \text{vol}^n \implies \frac{D}{Dt} (\rho e + k) + (\rho e + k) \text{div} u = \text{div} \langle u, T \rangle + \rho \langle u, b \rangle - \text{div} q$$

If one can use mass conservation for the flow

$$\frac{\partial \rho}{\partial t} + \text{div}(\rho u) = \frac{D\rho}{Dt} + \rho \text{div} u = 0 \text{ for } \dot{m} = 0$$

then

$$\rho \frac{D}{Dt} \left(e + \frac{k}{\rho} \right) = \text{div} \langle u, T \rangle + \rho \langle u, b \rangle - \text{div} q$$

since, from Newton's second law and using that on $\frac{Du^2}{Dt}$,

$$\begin{aligned} \rho \frac{Du}{Dt} &= \text{div}(T) \\ \rho \frac{Du^2}{Dt} &= \langle u, \text{div} T \rangle \end{aligned}$$

Working the following out,

$$\text{div} \langle u, T \rangle = \frac{1}{\sqrt{g}} \frac{\partial}{\partial x^k} (g_{ij} u^i T^{jk} \sqrt{g}) = \langle u, \text{div} T \rangle + T^{jk} \frac{\partial}{\partial x^k} (g_{ij} u^i)$$

For $T = -pg$,

$$\text{div} \langle u, T \rangle = \text{div}(g_{ij} u^i T^j) = \frac{1}{\sqrt{g}} \frac{\partial (g_{ij} u^i T^{jk} \sqrt{g})}{\partial x^k} = \frac{1}{\sqrt{g}} \frac{\partial (g_{ij} u^i (-p) g^{jk} \sqrt{g})}{\partial x^k} = \frac{-1}{\sqrt{g}} \frac{\partial (u^k p \sqrt{g})}{\partial x^k} = -u^j \frac{\partial p}{\partial x^j} - p \text{div} u$$

and

$$\frac{D}{Dt} \left(\frac{p}{\rho} \right) = \frac{1}{\rho} \frac{D}{Dt} p + \frac{-p}{\rho^2} \frac{D}{Dt} (\rho) = \frac{1}{\rho} \frac{Dp}{Dt} - \frac{p}{\rho} \text{div} u$$

since, again, using mass conservation,

$$\frac{\partial \rho}{\partial t} + \text{div}(\rho u) = \frac{D\rho}{Dt} + \rho \text{div} u = 0$$

Then

$$\rho \frac{D}{Dt} \left(\frac{p}{\rho} \right) = \frac{Dp}{Dt} - p \text{div} u$$

and so

$$\implies \rho \frac{D}{Dt} \left(e + \frac{p}{\rho} + \frac{k}{\rho} \right) = \frac{\partial p}{\partial t} + -q$$

in the case of $b = 0$ (no external body forces).

22. NOZZLE FLOW OF A REACTING GAS

This follows from Section 3.7 **Nozzle Flow of a Reacting Gas** of Oates (1997) [1]. Also see Lecture 14 of Martinez-Sanchez (2005) [16].

Investigate 2 limits to flow within nozzles, “equilibrium” and “frozen.”

22.1. Equilibrium flow. Suppose all chemical reactions occur in times very short compared to the time of fluid passage through the nozzle, so that the fluid will be at all times (almost) in a state of chemical equilibrium. Clearly, this is an approximation as we can imagine that during supersonic flow, the reactants could've traveled far, already, before the chemical reaction goes to completion.

The reactions occur continuously throughout the flow, leading to a continuous passage of energy from chemical binding and excitation modes to translational modes. Since the fluid is kept at equilibrium (locally) at all times, the “equivalent” temperatures of all such modes of energy storage are equal and as a result the total entropy of the fluid remains constant.” (cf. pp. 86 of Oates (1997) [1]).

From Oates (1997) [1], starting from

$$\tau d\sigma_j = C_{p_j} d\tau - V dp_j \text{ or } d\sigma_j = \frac{C_{p_j} d\tau}{\tau} - \frac{N_j dp_j}{p_j}$$

$$\implies \tau d\sigma_j = C_{p_j} d\tau - \frac{N_j \tau}{p_j} dp_j \text{ or } \hat{\sigma}_j - \hat{\sigma}_j(p_0, \tau_0) = \int_{\tau_0}^{\tau} \frac{\hat{C}_{p_j} d\tau}{\tau} - \ln \frac{p_j}{p_{j0}}$$

$$\hat{\sigma}_j - \hat{\sigma}_j(p_0, \tau_0) = \int_{\tau_0}^{\tau} \frac{\hat{C}_{p_j} d\tau}{\tau} - \ln \frac{p_j}{p_{j0}}$$

$$\sigma = \sum_{j=1}^{\mathcal{N}} N_j \hat{\sigma}_j = \sum_{j=1}^{\mathcal{N}} N_j \left[\left(\int_{\tau_0}^{\tau} \frac{\hat{C}_{p_j} d\tau}{\tau} + \hat{\sigma}_j(p_0, \tau_0) \right) - \ln \frac{p_j}{p_{j0}} \right]$$

Recall that a chemical reaction, labeled by I , with the convention that reactant (stoichiometric) coefficients are negative integers, is given by $\sum_j \nu_{jI} A_j = 0$ and so

$$\sum_j \nu_{jI} A_j = 0 \implies dN_{jI} = \nu_{jI} d\hat{N}_I$$

which informs us that a chemical reaction, every time it occurs, adds in and subtracts out particles of each species, represented by dN_{jI} , each time a chemical reaction is run, $d\hat{N}$.

Recalling the definition of enthalpy $H = H(\sigma, p; \{N_i\}_i) := U + pV$, then the total enthalpy H of a system is completely specified by total entropy σ , pressure p , and \mathcal{N} numbers, $\{N_i\}_i \equiv \{N_i\}_{i=1 \dots \mathcal{N}}$ for the number of particles in each species, with \mathcal{N} being the total number of (different) species to consider in all possible chemical reactions I . Effectively, H is a smooth scalar function on a $2 + \mathcal{N}$ manifold Σ , with Σ representing all possible thermodynamic processes. Also, recall that

$$dH = \tau d\sigma + V dp + \sum_i \mu_i dN_i$$

Consider a curve on Σ , $\gamma : \mathbb{R} \rightarrow \Sigma$, representing a thermodynamic process such that σ is constant (isentropic) and p is constant (isobaric). This means that $d\sigma(\dot{\gamma}) = 0$ and $dp(\dot{\gamma}) = 0$, respectively, and so we have $d\sigma = dp = 0$ in this case. Then, this case,

$$dH = \sum_i \mu_i dN_i$$

$$\implies dH(\dot{\gamma}) = \sum_i \mu_i \Delta N_i = H_P(\tau_P) - H_R(\tau_0)$$

where H_P , H_R refer to the total enthalpies of the products and reactants, respectively, as defined in Eq. 29.

Recall that $dG = dH - d(\tau\sigma)$ and $dF = dU - d(\tau\sigma) = 0$. Suppose $dH = 0$ and $dG = 0$. This implies that $d(\tau\sigma) = 0$ so therefore $dU = 0$ and $dF = 0$ and so we're in Helmholtz free energy equilibrium. One can imagine that both σ and τ change during the process in such a way that $d(\tau\sigma) = 0$.

If we suppose we want Gibbs equilibrium $dG = 0$, and so for an isentropic process $d\sigma = 0$, then $dH = \sigma d\tau$ and so the so-called *stagnation* enthalpy can drop (or at least change) with a drop (or change) in temperature.

Thus, what's going on is that there's a transformation, or mapping, via a thermodynamic process represented by curve $\gamma : \mathbb{R} \rightarrow \Sigma$ in Σ , a $2 + \mathcal{N}$ -dim. manifold:

$$\begin{array}{ccc} \Sigma & \xrightarrow{\gamma} & \Sigma \\ & \gamma \text{ s.t.} & \\ & dG(\dot{\gamma})=0 & \\ & d\sigma(\dot{\gamma})=0 & \\ & dp(\dot{\gamma})=0 & \\ (\tau_0, p_0, \{N_{i0}\}) & \longmapsto & (\tau_p, p_0, \{N_{ieq}\}) \end{array}$$

and if we wanted enthalpy to remain constant,

$$\begin{array}{ccc} \Sigma & \xrightarrow{\gamma} & \Sigma \\ & \gamma \text{ s.t.} & \\ & dG(\dot{\gamma})=0 & \\ & dH(\dot{\gamma})=0 & \\ & d(\tau\sigma)(\dot{\gamma})=0 & \\ (\tau_0, p_0, \{N_{i0}\}) & \longmapsto & (\tau_p, p_0, \{N_{ieq}\}) \end{array}$$

I think what we want, since the flow was shown to be isentropic, and, again, as Oates [1] argued, that local equilibrium leads to equilibrium in all modes of energy storage, chemical bonds, vibrational modes, to translational modes, that the total entropy of the fluid remains constant, this:

$$\begin{array}{ccc} \Sigma & \xrightarrow{\gamma} & \Sigma \\ & \gamma \text{ s.t.} & \\ & dG(\dot{\gamma})=0 & \\ & d\sigma(\dot{\gamma})=0 & \\ & dp(\dot{\gamma})=0 & \\ (\tau_0, p_0, \{N_{i0}\}) & \longmapsto & (\tau_p, p_0, \{N_{ieq}\}) \end{array}$$

22.1.1. *Continuity equations for equilibrium flow.* The continuity equations to use for equilibrium flow should be (correct me if I'm wrong) are

$$\dot{m} = \rho u A$$

for which \dot{m} should be constant at each point along the 1-dimensional flow (out) by mass conservation (which should still should hold for chemical reactions).

One possible form of mass conservation that could prove useful is

$$\rho_2 u_2 A_2 = \rho_1 u_1 A_1 \text{ or } \frac{A_2}{A_1} = \frac{\rho_1 u_1}{\rho_2 u_2}$$

Also, throughout the flow, the Bernoulli invariant is $h + k$, i.e.

$$\frac{u_2^2}{2} + h_2 = \frac{u_1^2}{2} + h_1 \text{ or } u_2^2 = u_1^2 + 2(h_1 - h_2)$$

Note that the speed of sound $a^2 = \left(\frac{\partial p}{\partial \rho}\right)_{\sigma, \{N_i\}}$ can be calculated at each instance, and so the Mach number can be calculated:

$$\mathfrak{M} = \frac{u}{a}$$

22.2. **Frozen flow.** Note that there are a number of critical typos in Martinez-Sanchez (2005) [16] that makes it difficult to understand, in this case, Lecture 14.

Given nozzle exit (i.e. entering combustion chamber), pressure p_c , and consider chamber entropy σ_c . Then (σ_c, p_c) specifies completely the thermodynamic state, and so enthalpy (or specific enthalpy) h_c is specified.

$$(\sigma_c, p_c) \mapsto h_c(\sigma_c, p_c)$$

For adiabatic flows, the energy equation (for steady state) is

$$\frac{u_2^2}{2} = h_c - h_2$$

In general, for multiple number of species,

$$(\sigma_c, p_c; \{N_i\}_{i=1\dots\mathcal{N}}) \mapsto h_c(\sigma_c, p_c; \{N_i\})$$

For throat conditions, consider mass flow $\dot{m} = \rho u A$.

Assuming steady state (EY: 20160210 check if throat conditions necessitates $\dot{m} = 0$), then

$$\frac{d\rho}{\rho} + \frac{du}{u} + \frac{dA}{A} = 0 \implies \frac{du}{u} = -\frac{dh}{2(h_c - h)} = \frac{-dp}{2\rho(h_c - h)} = \frac{-dp}{\rho u^2} \text{ since } u du = -dh \text{ and}$$

$$dh = \tau ds + \frac{1}{\rho} dp + \frac{h}{\rho} d\rho \text{ and for this thermodynamic process } ds = d\rho = 0 \text{ in the fluid-at-rest frame}$$

At the throat, $dA = 0$ (by definition of a throat), and so

$$\frac{d\rho}{\rho} = -\frac{du}{u} = \frac{dp}{\rho u^2} \implies (u^*)^2 = \left(\frac{\partial p}{\partial \rho}\right)_\sigma$$

Also, consider from mass conservation, that, at the throat, $dA = 0$,

$$\begin{aligned} d(\rho u)A + \rho u dA &= 0 \\ \implies \frac{d(\rho u)}{\rho u} + \frac{dA}{A} &= 0 \xrightarrow{dA=0} d(\rho u) = 0 \end{aligned}$$

ρu is at a *maximum* at the throat.

22.3. **Implementation in Cantera of equilibrium flow and frozen flow.** The first thing that happens is *combustion*! Fule and oxidizer enters through inlets into combustion chamber. Then combustion occurs in the (combustion) chamber!

$$\begin{array}{ccc} \Sigma & \xrightarrow{\gamma} & \Sigma \\ & \gamma \text{ s.t.} & \\ & dH(\dot{\gamma})=0 & \\ & dp(\dot{\gamma})=0 & \\ (\tau_0, p_0, \{N_{i0}\}) & \longmapsto & (\tau_c, p_0, \{N_{ic}\}) \end{array}$$

and $dH(\dot{\gamma}) = 0$ since $H_P(\tau_c) = H_R(\tau_0)$ so $\Delta H \equiv H_p(\tau_c) - H_R(\tau_0) = 0$

$h(\sigma_c, p_0, \{N_{ic}\}_i) \equiv h_0 \in C^\infty(\Sigma)$ and h_0 is the stagnation enthalpy, used as the Bernoulli invariant in the energy equation (i.e. $\frac{u^2}{2} = (h_0 - h)$).

Then consider isentropic flow γ_{isent} :

$$\begin{aligned} \Sigma &\xrightarrow{\gamma_{\text{isent}}} \Sigma \\ (\tau_c, p_0, \{N_{i\,c}\}) &\xrightarrow[\gamma_{\text{isent}} \text{ s.t. } d\sigma(\gamma_{\text{isent}})=0]{} (\tau, p, \{N_{i\,f}\}) \\ C^\infty(\Sigma) &\xrightarrow{\gamma_{\text{isent}}} C^\infty(\Sigma) \\ h(\sigma_c, p_0, \{N_{ic}\}) \equiv h_0 &\xrightarrow{\gamma_{\text{isent}}} h(\sigma_c, p, \{N_{if}\}) \end{aligned}$$

How I implemented this in Cantera is by setting the state of the gas to the desired p , and then equilibrating the gas under constant entropy σ_c and constant pressure p .

22.3.1. *Getting the results for rocket (and nozzle) performance (i.e. collecting the dividends).* You want to calculate the specific impulse (for the rocket). Use this handy relation:

$$gI_{\text{sp}} = \frac{F}{\dot{m}} = \frac{\dot{m}u_{\text{exh}} + (p_{\text{exh}} - p_a)A_e}{\rho_{\text{exh}}u_{\text{exh}}A_e} = u_{\text{exh}} + \frac{p_{\text{exh}} - p_a}{\rho_{\text{exh}}u_{\text{exh}}}$$

and $\dot{m} = \rho^*u^*A^*$.

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There’s an 8th edition of Biblarz and Sutton [2] for 2010 that I would like to have. If you find any of this material useful or if you’d like to help, email me or visit my Open/Tilt page ernestyalumni.tilt.com and donate to the crowdfunding campaign or click on the PayPal donate button.