

Supporting Information

Facile C=O Bond Splitting of Carbon Dioxide Induced by Metal-Ligand Cooperativity in a Phosphinine Iron(0) Complex

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S1 Experimental Section

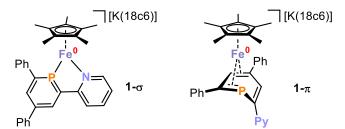
S1.1 General experimental details

All experiments were performed under an atmosphere of dry argon using standard glovebox and Schlenk line techniques. Tetrahydrofuran, toluene and *n*-hexane were purified, dried, and degassed using an MBraun SPS800 solvent purification system. 1,4-Dioxane and 1,2-dimethoxyethane were dried over potassium and distilled under inert gas atmosphere. Deuterated tetrahydrofurane was purchased from Sigma Aldrich and used as received. [K([18]crown-6)][Cp*Fe(C₁₀H₈)] and **L** were synthesized according to literature procedures.^{[1][2]}CO₂ (purity 4.8) was purchased from Linde Gas and used as received.

NMR spectra were recorded on Bruker Avance 300 and Avance 400 spectrometers at 300 K K and a Bruker Avance III HD 600 MHz spectrometer with a fluorine selective TBIF probe at 273 K. ¹H and ¹³C{¹H} spectra were referenced internally to residual solvent resonances, while ³¹P{¹H} and ³¹P spectra were referenced externally to 85% H₃PO_{4 (aq.)}. The assignment of ¹H and ¹³C NMR signals was confirmed by two-dimensional (COSY, HSQC, and HMBC) experiments. Solid state ³¹P MAS-NMR spectra were recorded with a Bruker 400 MHz spectrometer. UV/vis spectra were recorded using a Varian Cary 50 spectrometer. Elemental analyses were determined by the analytical department of the University of Regensburg. IR spectra were recorded using a Bruker ALPHA spectrometer equipped with a diamond ATR unit.

Single-crystal X-ray diffraction data were recorded on an Agilent Technologies SuperNova diffractometer with Cu- K_{α} radiation ($\lambda = 1.54184$ Å). Either semi-empirical multi-scan absorption corrections^[3] or analytical ones^[4] were applied to the data. The structures were solved with SHELXT^[5] and least-square refinements on F^2 were carried out with SHELXL.^[6] The hydrogen atoms were located in idealized positions and refined isotropically with a riding model. CCDC 1942537 (for **1-\sigma**), 1942538 (for **2**), 1942550 (for **3-\sigma**), and 1942542 (for **3-\pi**) contain the supplementary crystallographic data for this paper. These data can be obtained free of charge from the Cambridge Crystallographic Data Centre.

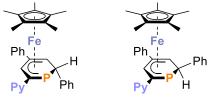
S1.2 Synthesis of 1



A solution of **L** (200 mg, 0.62 mmol, 1 eq.) in 1,2-dimethoxyethane (4 mL) was added dropwise to a solution of [K([18]crown-6)][Cp*Fe($C_{10}H_8$)] (383 mg, 0.62 mmol, 1 eq.) in 1,2-dimethoxyethane (8 mL) at -35 °C. The orange/brown reaction mixture turned into a deep

green suspension. After stirring overnight and warming up to room temperature a deep purple suspension was formed. The suspension was layered with n-hexane (16 mL) and stored at -35 °C over two days. Product **1** could be isolated as dark purple solid after decanting off the supernatant solution and drying under vacuum. Crystals suitable for single-crystal XRD were grown by slow diffusion of n-hexane into a 1,2-dimethoxyethane solution of **1**. Yield 350 mg, 68%. Elemental analysis calcd. for $C_{44}H_{55}FeKNO_6P$ (Mw = 819.84 g·mol⁻¹) C 64.46, H 6.76,

N 1.71; found C 63.27, H 6.69, N 1.54. UV-Vis: (THF, λ_{max} / nm, ϵ_{max} / L·mol⁻¹·cm⁻¹): 310 (sh, 16553), 500 (3708). ¹H NMR (400.13 MHz, 300 K, [D₈]THF): δ = 1.23 (s, 15H, C₅(CH₃)₅), 3.25 (s, 24H, [18]crown-6), 6.20–10.45 (m, 16H signals for **L** of **1-\sigma** and **1-\pi**). ¹³C{¹H} NMR (100.61 MHz, 300 K, [D₈]THF): δ = 11.0 (s, C₅(CH₃)₅), 12.6 (s), 12.8 (s), 70.9 (bs, [18]crown-6), 72.6 (s), 78.7 (s), 80.9 (s, C₅(CH₃)₅), 105.3 (s), 109.6 (s), 111.7 (s), 116.1 (s), 117.0 (s), 122.1 (s), 123.5 (s), 125.00 (s), 125.3 (s), 125.7 (bs), 126.4 (s), 127.4 (s), 127.5 (s), 128.5 (s), 128.8 (s), 130.3 (d, J = 6 Hz), 132.0 (m), 133.8 (s), 144.1 (s), 147.5 (s), 148.7 (s), 158.5 (s).



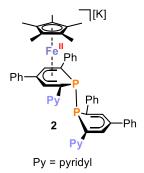
³¹P{¹H} NMR (161.98 MHz, 300 K, [D₈]THF): δ = 131.7 (s, **1-σ**), -44.8 (bs, **1-π**). ³¹P NMR (161.98 MHz, 300 K, [D₈]THF): δ = 130.7 (t, ³J_{PH} = 12 Hz, **1-σ**), -46.2 (bs, **1-π**).

endo-4 exo-4

NMR spectroscopy and elemental analysis consistently indicate the presence of persistent, minor impurities in samples of 1 prepared in the

above manner. Specifically, ³¹P{¹H} NMR analysis indicates the presence of side-products **endo-4** and **exo-4** (typically *ca.* 8%), which are presumed to arise from protonation of **1** by adventitious moisture. ^[7]

S1.3 Synthesis of 2



A solution of **L** (100 mg, 0.31 mmol, 2 eq.) in toluene (2 mL) was added dropwise to a solution of $[K([18]\text{crown-6})][\text{Cp*Fe}(\text{C}_{10}\text{H}_8)]$ (96 mg, 0.15 mmol, 1 eq.) in tetrahydrofuran (0.5 mL) at -35 °C. A colour change from orange brown to wine red was observed. After stirring for 5 hours and warming to room temperature the brown reaction mixture was layered with *n*-hexane (6 mL) and stored at -35 °C overnight. Product **2** was isolated as a dark brown solid after decanting off the supernatant solution and drying under vacuum. Crystals suitable for single-crystal XRD were grown from slow diffusion of *n*-hexane into a 1,2-dimethoxyethane solution of **2**. Yield 126 mg,

72%. Elemental analysis calcd. for $C_{66}H_{71}FeKN_2O_6P_2$ (Mw = $1145.19g\cdot mol^{-1}$) C 69.22, H 6.25, N 2.45; found C 69.60, H 6.26, N 2.28. UV-Vis: (THF, λ_{max} / nm, ε_{max} / L·mol⁻¹·cm⁻¹): 250 (40598), 280 (39399), 510 (8023). ¹H NMR (400.13 MHz, 300 K, [D₈]THF): δ = 0.90 (s, 15H, $C_5(CH_3)_5$), 3.25 (s, 44H, [18]crown-6), 5.87 (m, 1H), 6.17 (m, 2H) 6.51 7.44 (m, 40H), 7.74 8.00 (m, 10H), 8.34 (m, 1H). ¹³C{¹H} NMR (100.61 MHz, 300 K, [D₈]THF): δ = 8.9 (s, $C_5(CH_3)_5$), 71.6 (bs, [18]crown-6), 84.3 (d, 1.66 Hz), 85.2 (m), 85.6 (m), 94.74 (s), 94.7 (s), 112.6 (s) 113.5 (s), 113.9 (s), 117.8 (s), 119.4 (m), 120.6 (s), 120.8 (s), 122.5 (s), 122.6 (s), 123.7 (m), 125.0 (m), 125.2 (m), 125.4 (m), 125.5 (s), 125.9 (s), 126.4 (s), 126.7 (m), 127.5 (s), 127.7 (s), 128.5 (s), 128.5 (s), 128.6 (s), 128.8 (s), 129.3 (s), 129.5 (s), 131.8 (m), 133.3 (s), 133.6 (s), 133.8 (s), 134.5 (s), 134.8 (s), 135.0 (s), 138.3 (s), 143.0 (s), 143.2 (s), 146.1 (s), 148.4 (s), 148,5 (s), 162.5 (s), 162.8 (s), 164.9 (s), 165.2 (s). ³¹P{¹H} NMR (161.98 MHz, 300 K, [D₈]THF): δ = -24.31 (AB, 1 J_{PP} = 262 Hz), -42.09 (AB, 1 J_{PP} = 262 Hz), -25.27 (AB, 1 J_{PP} = 262 Hz), -42.09 (AB, 1 J_{PP} = 262 Hz), -25.27 (AB, 1 J_{PP} = 262 Hz), -42.09 (AB, 1 J_{PP} = 257 Hz).

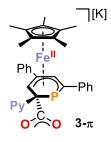
As for compound **1**, samples of **2** prepared in the above manner were found to contain trace impurities, and specifically side-products **endo-4** and **exo-4** (typically *ca.* 1%).^[7]

S1.4 Synthesis of 3- σ and 3- π

Stoichiometric conversion of 1 and CO_2 leads also to 3- σ and 3- π , but leads also to more hydrolysis side products endo-4 and exo-4, due to opening of the reaction vessel and adding CO2 via syringe. It is advisable to keep the reaction mixture in a closed system.

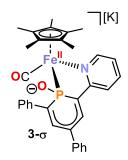
A sealed vessel containing a solution of 1 (100 mg, 0.122 mmol) in tetrahydrofuran (3 mL) was charged with CO₂ (1 atm). An immediate colour change from purple to deep green was observed, and the reaction mixture was stirred for 30 minutes at room temperature. The solvent was removed under vacuum, and the resulting green oil was washed with n-hexane (3 x 2 mL). The remaining residue was extracted with toluene (3 mL), layered with nhexane (6 mL), and stored at room temperature for 7 days. The mother liquor was then decanted from the resulting orange powder.

Isolation of 3-\pi: A small amount of orange powder was isolated and dried under vacuum to give **3-\pi**. This sample was used for spectroscopic characterization The 1 H NMR spectrum of 3- π isolated in this manner shows minor impurities with signals at 1.13 ppm and 3.86 ppm, which is believed to account for deviations in the elemental analysis. Crystals suitable for single-crystal XRD were grown by slow diffusion of n-hexane into a 1,4-dioxane solution of $3-\pi$.



Elemental analysis calcd. for $C_{45}H_{55}FeKNO_8P$ (Mw = 863.85 g·mol⁻¹) C 62.57, H 6.42, N 1.62; found C 59.77, H 6.12, N 1.31. UV-Vis: (THF, λ_{max} / nm , $\varepsilon_{max} / L \cdot mol^{-1} \cdot cm^{-1}$): 334sh (9494). ¹H NMR (400.13 MHz, 300 K, $[D_8]$ THF): $\delta = 1.28$ (s, 15H, $C_5(CH_3)_5$), 3.29 (s, 24H, [18]crown-6), 6.62 (s, 1H, $PC_5H_2Ph_2Py$), 6.96–7.74 (m, 8H, $PC_5H_2Ph_2Py$), 7.91 (m, 2H, $PC_5H_2Ph_2Py$), 8.09 (m, 3H, $PC_5H_2Ph_2Py$), 8.52 (m, 1H, $PC_5H_2Ph_2Py$). $^{13}C\{^{1}H\}$ NMR (100.61 MHz, 300 K, [D₈]THF): $\delta = 9.5$ (s, C₅(CH₃)₅), 70.7 (s, [18]crown-6), 90.9 (s, $C_5(CH_3)_5$, 111.8 (s), 115.7. (s), 121.3 (s), 123.0 (s), 123.7 (s), 125.6 (s), 127.8 (s), 127.8 (s), 128.0 (s), 131.2 (s), 133.6 (s), 157.2 (s). ${}^{31}P{}^{1}H{}^{1}NMR$ (161.98 MHz, 300 K, $[D_8]^{T}HF$): $\delta = -116.6$ (s). ${}^{31}P$ NMR (161.98 MHz, 300 K, $[D_8]$ THF): $\delta = -116.6$ (s).

Isolation of 3-σ: The decanted mother liquor was evaporated to dryness, and the remaining residue was extracted into tetrahydrofuran (3 mL) and layered with n-hexane (6 mL). Dark green crystals of $3-\sigma$ were isolated after standing at room temperature for 6 days, filtering, and drying under vacuum. Crystals suitable for single-crystal XRD were grown from slow diffusion of n-hexane into a 1,4-dioxane solution of $3-\sigma$. Yield: 16 mg, 15%.



Elemental analysis calcd. for $C_{45}H_{55}FeKNO_8P$ (Mw = 863.85 g·mol⁻¹) C 62.57, H 6.42, N 1.62; found C 62.57, H 6.31, N 1.60. UV-Vis: (THF, λ_{max} / nm, ϵ_{max} / L·mol⁻¹·cm⁻¹): 270 (sh, 16070), 340 (13458), 417 (15900), 620 (9617). ATR-IR: $\tilde{\nu}$ (CO) = 1876 cm⁻¹. ¹H NMR (400.13 MHz, 300 K, [D₈]THF): δ = 1.24 (s, 15H, $C_5(CH_3)_5$), 3.38 (s, 24H, [18]crown-6), 6.02 (m, 1H, pyridyl-*H*), 6.75 (m, 1H, pyridyl-*H*), 6.88 (m, 1H, pyridyl-*H*), 6.99 (m, 2H, phenyl-*H*), 7.12 (m, 4H, phenyl-*H*), 7.28 (m, 2H, phenyl-*H*), 7.40 (m, 2H, phenyl-*H*), 8.00 (bd, 1H, pyridyl-*H*, ⁴ $J_{PH'}$ = 5 Hz), 8.23 (bd, 2H, ^{3,5}H of P C_5H_2 ,

 ${}^3J_{\text{PH'}} = 7 \text{ Hz}$). ${}^{13}\text{C}\{{}^1\text{H}\}$ NMR (100.61 MHz, 300 K, [D₈]THF): $\delta = 9.5$ (s, C₅(CH₃)₅), 70.7 (s, [18]crown-6), 90.9 (s, C₅(CH₃)₅), 111.8 (s, pyridyl-CH), 115.7. (s, ${}^{2,6}C$ of PC₅H₂), 121.8 (s, pyridyl-CH), 123.0 (s, pyridyl-CH), 123.7 (s, phenyl-CH), 125.6 (s, phenyl-CH), 127.8 (s, ${}^{3,5}C\text{H}$ of PC₅H₂), 127.8 (s, ${}^{3,5}C\text{H}$ of PC₅H₂), 128.0 (s, phenyl-CH), 128.4 (s, phenyl-CH), 131.2 (s, phenyl-CH), 133.6 (s, phenyl-CH), 146.8 (s, phenyl-C), 147.8 (phenyl-C), 157.2 (s, pyridyl-CH), 170.7 (s, CO). ${}^{31}\text{P}\{{}^{1}\text{H}\}$ NMR (161.98 MHz, 300 K, [D₈]THF): $\delta = 97.9$ (s). ${}^{31}\text{P}$ NMR (161.98 MHz, 300 K, [D₈]THF): $\delta = 96.9$ (t, ${}^{3}J_{\text{PH}} = 17 \text{ Hz}$).

S2 NMR Spectra

S2.1 NMR spectra of compounds $1-\sigma$ and $1-\pi$

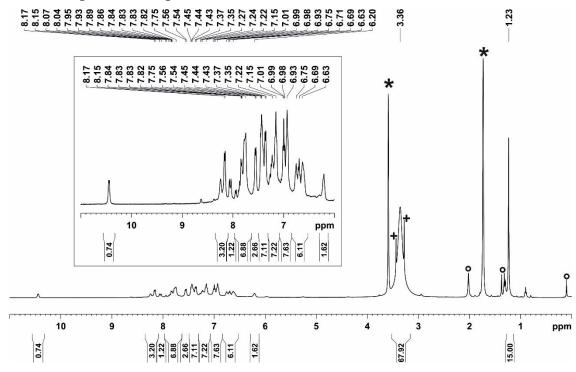


Figure S1 - ¹H NMR spectrum (400.13 MHz, 300 K, [D₈]THF) of **1-\sigma** and **1-\pi** (* = [D₈]THF, + = 1,2-dimethoxy ethane, ° = unidentified impurities).

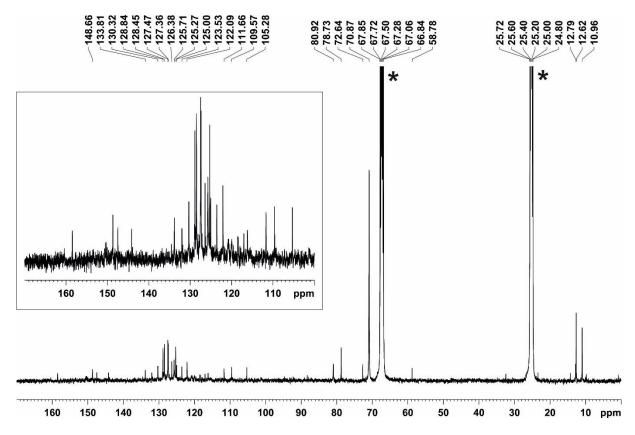


Figure S2 - ${}^{13}C\{{}^{1}H\}$ NMR spectrum (100.61 MHz, 300 K, [D₈]THF) of **1-\sigma** and **1-\pi** (* = [D₈]THF).

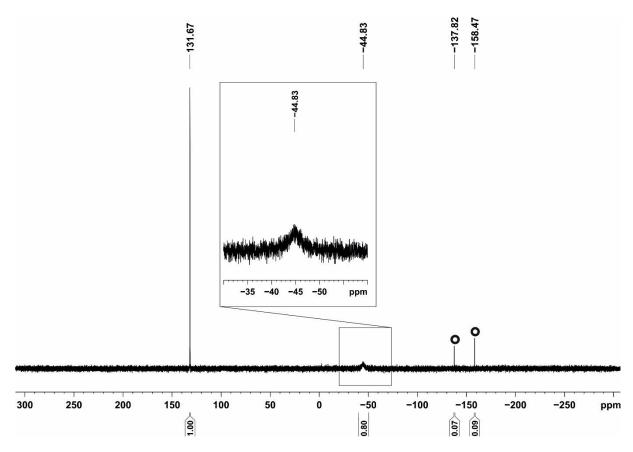


Figure S3 - 31 P{ 1 H} NMR spectrum (161.98 MHz, 300 K, [D₈]THF) of 1-σ and 1-π ($^{\circ}$ = hydrophosphinine iron complexes endo-4 and exo-4).

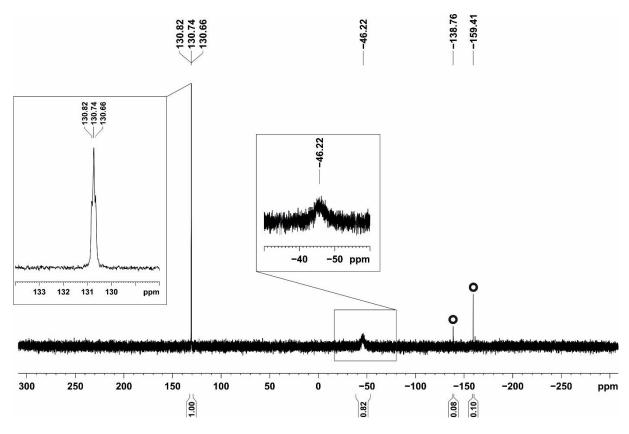


Figure S4 - ³¹P NMR spectrum (161.98 MHz, 300 K, [D₈]THF) of 1-σ and 1-π (° = hydrophosphinine iron complexes endo-4 and exo-4).

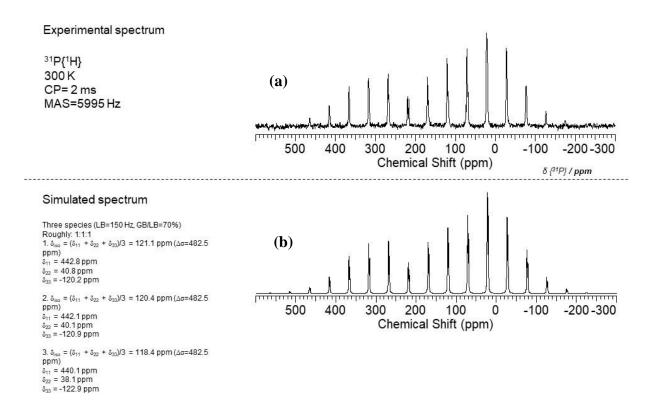


Figure S5. 31 P{ 1 H} CP MAS spectrum (6 kHz, 300 K) of 1-σ; a) experimental spectrum (δ_{iso} = 121.1 ppm), b) simulated spectrum. The spectrum of 1-σ indicates that there are three structurally similar species. These might arise from crystallographically different molecules.

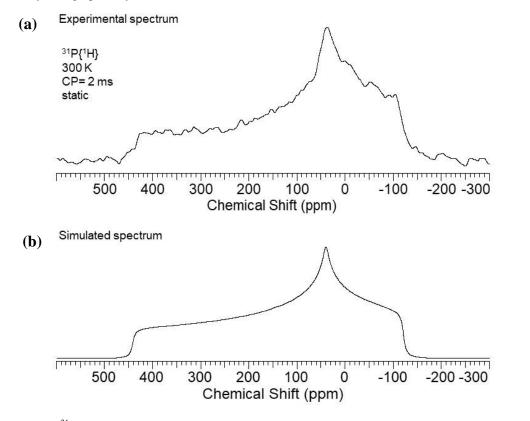


Figure S6 - ³¹P CP MAS static spectrum (6 kHz, 300 K) of **1-σ**; a) experimental spectrum, b) simulated spectrum.

S2.2 NMR spectra of compound 2

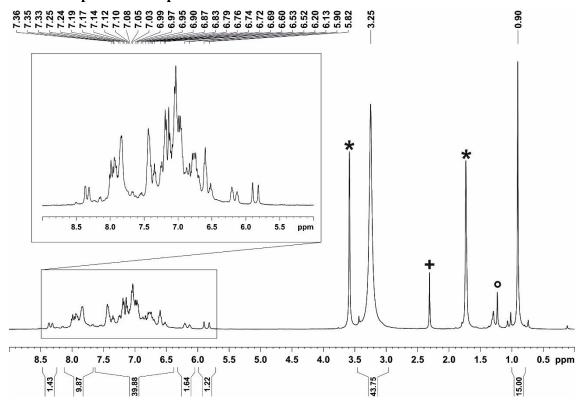


Figure S7 - 1 H NMR spectrum (400.13 MHz, 300 K, [D₈]THF) of 2 (* = [D₈]THF, + = toluene, $^{\circ}$ = unidentified impurity).

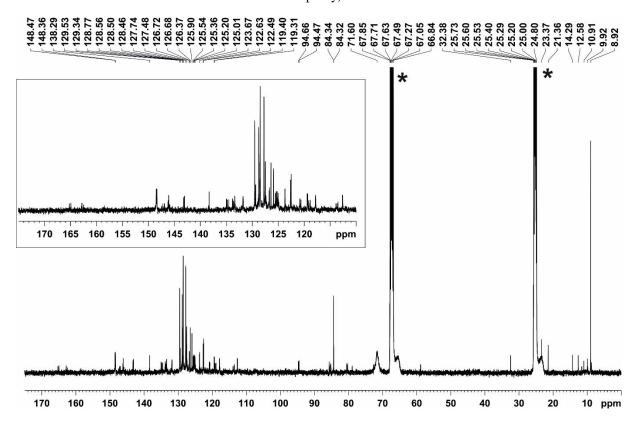


Figure S8 - ${}^{13}C\{{}^{1}H\}$ NMR spectrum (100.61 MHz, 300 K, [D₈]THF) of **2** (* = [D₈]THF).

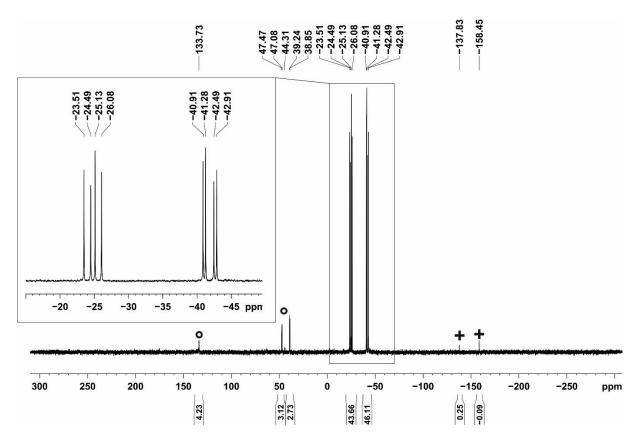


Figure S9 - ${}^{31}P\{{}^{1}H\}$ NMR spectrum (161.98 MHz, 300 K, [D₈]THF) of 2 (° = unidentified impurities, + = hydrophosphinines endo-4 and exo-4).

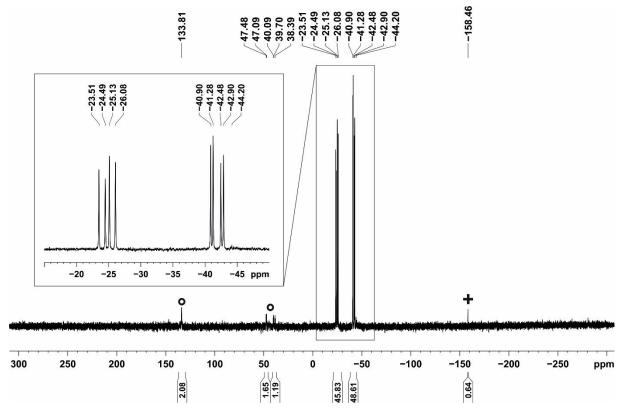


Figure S10 - 31 P NMR spectrum (161.98 MHz, 300 K, [D₈]THF) of 2 ($^{\circ}$ = impurities, + = hydrophosphinines endo-4 and exo-4).

S2.3 NMR spectra of compound $3-\sigma$

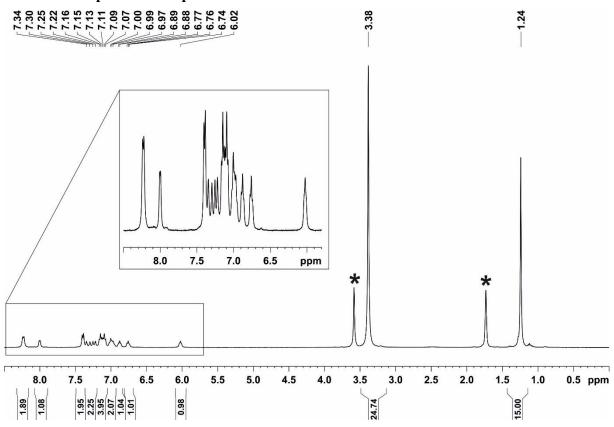


Figure S11 - 1H NMR spectrum (400.13 MHz, 300 K, $[D_8]THF)$ of 3- σ (* = $[D_8]THF).$

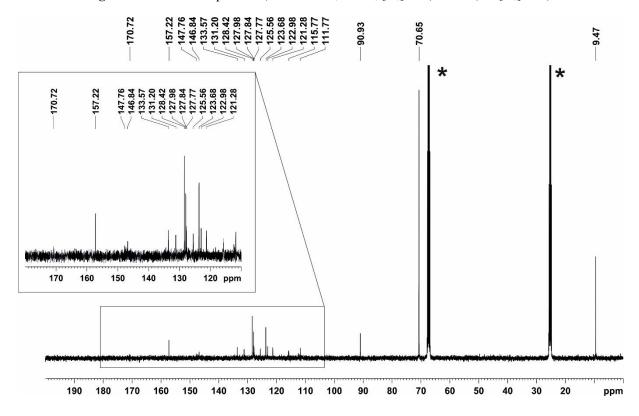


Figure S12 - ${}^{13}C\{{}^{1}H\}$ NMR spectrum (100.61 MHz, 300 K, [D₈]THF) of **3-\sigma** (* = [D₈]THF).

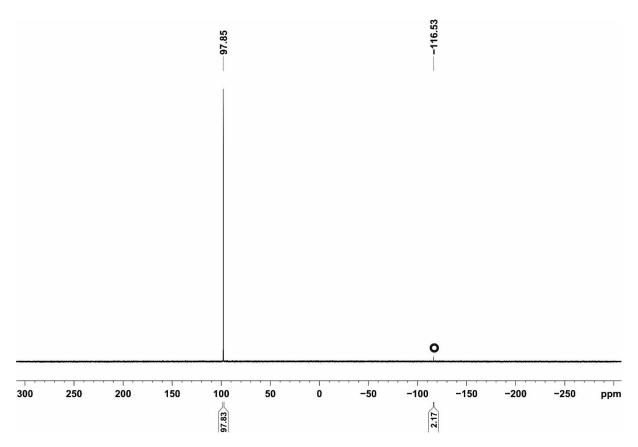


Figure S13 - ${}^{31}P\{{}^{1}H\}$ NMR spectrum (161.98 MHz, 300 K, [D₈]THF) of **3-** σ (° = **3-** π).

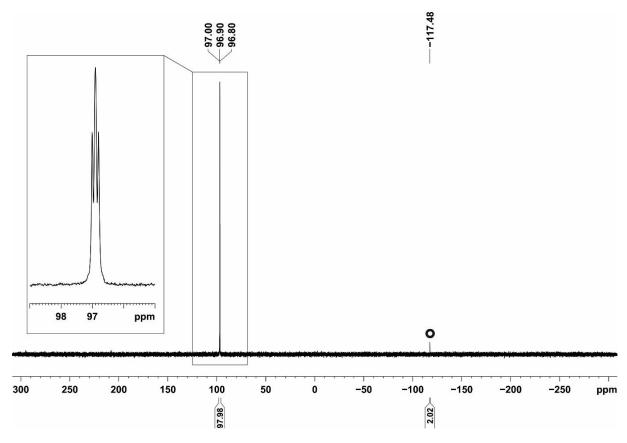


Figure S14 - 31 P NMR spectrum (161.98 MHz, 300 K, [D₈]THF) of **3-\sigma** (° = **3-\pi**).

S2.4 NMR Spectra of compound $3-\pi$

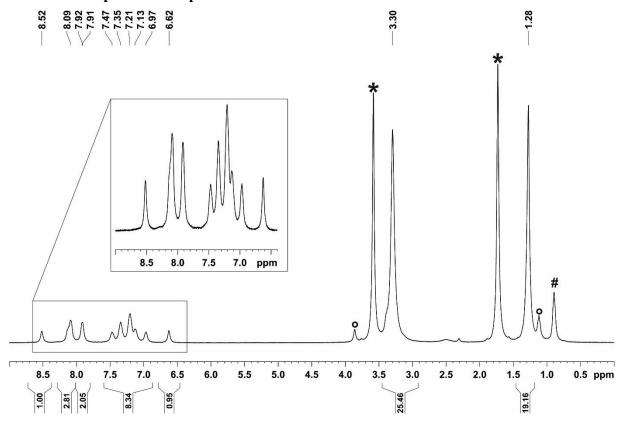


Figure S15 ¹H NMR spectrum (400.13 MHz, 300 K, [D₈]THF) of **3-\pi** (* = [D₈]THF, # = n-hexane, ° = unidentified impurities).

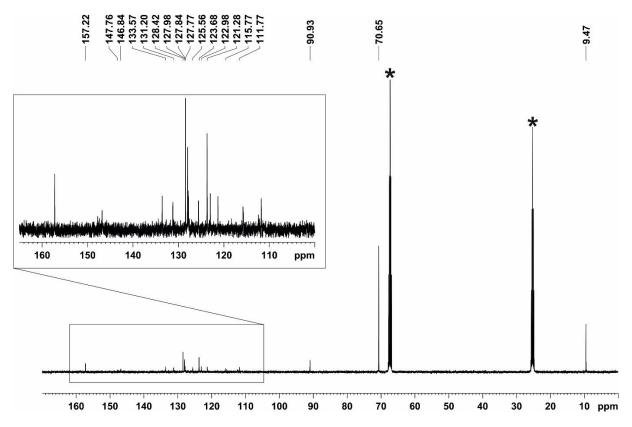
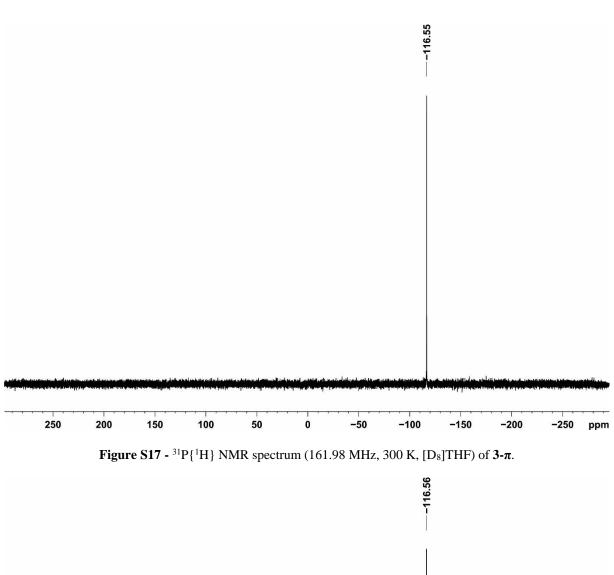


Figure S16 - ${}^{13}\text{C}\{{}^{1}\text{H}\}$ NMR spectrum (100.61 MHz, 300 K, [D₈]THF) of **3-** π (* = [D₈]THF).



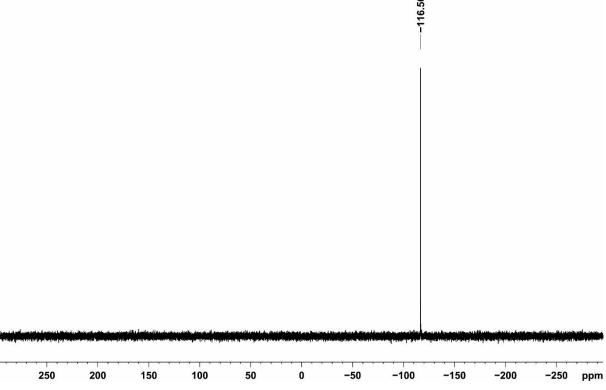


Figure S18 - 31 P NMR spectrum (161.98 MHz, 300 K, [D₈]THF) of **3-\pi**.

S2.5 Variable temperature $^{31}P\{^{1}H\}$ NMR spectra of compounds 1- σ and 1- π

[K([18]crown-6)][Cp*Fe($C_{10}H_8$)] (38 mg, 0.061 mmol) and **L** (20 mg, 0.061 mmol) were dissolved in [D₈]THF (0.5 mL) at room temperature. ³¹P{¹H} NMR spectra were recorded at 300 K, 243 K and 333 K.

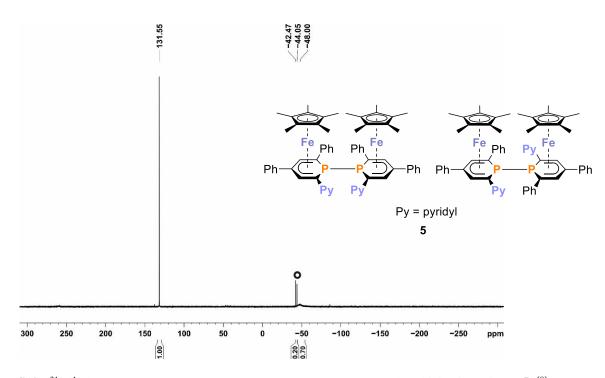


Figure S19 - ${}^{31}P\{{}^{1}H\}$ NMR spectrum (161.98 MHz, [D₈]THF) at 300 K (° = phosphinine iron dimers **5**). [8]

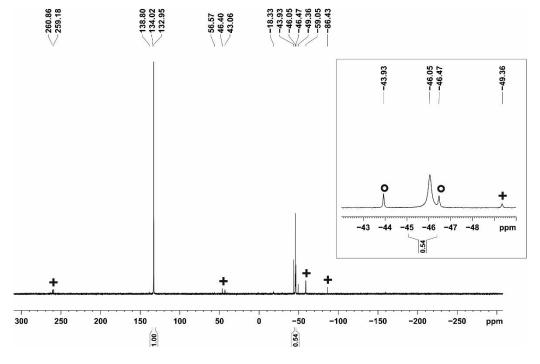


Figure S20 - $^{31}P\{^{1}H\}$ NMR spectrum (161.98 MHz, [D₈]THF) at 243 K (° = phosphinine iron dimers 5, + = unidentified impurities).

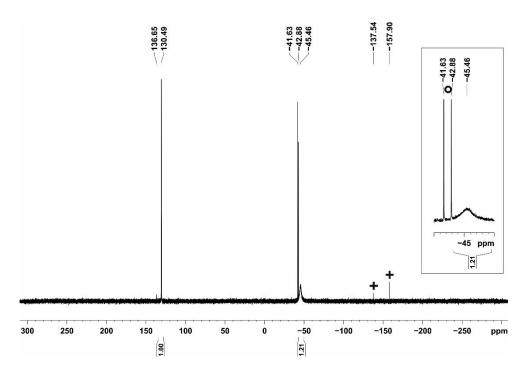


Figure S21 - $^{31}P\{^{1}H\}$ NMR spectrum (161.98 MHz, [D₈]THF) at 333 K ($^{\circ}$ = phosphinine iron dimers 5, + = hydrophosphinine iron complexes **endo-4** and **exo-4**).

S2.6 ${}^{31}P{}^{1}H}$ NMR monitoring of the formation of 1- σ and 1- π at 273 K

[K([18]crown-6)][Cp*Fe($C_{10}H_8$)] (38 mg, 0.061 mmol) and **L** (20 mg, 0.061 mmol) were dissolved in [D₈]THF (0.8 mL) cooled to -35 °C inside an NMR tube fitted with a screw cap. ³¹P{¹H} NMR spectra were immediately recorded at a controlled temperature of 273 K, and recorded periodically while being maintained at the same temperature.

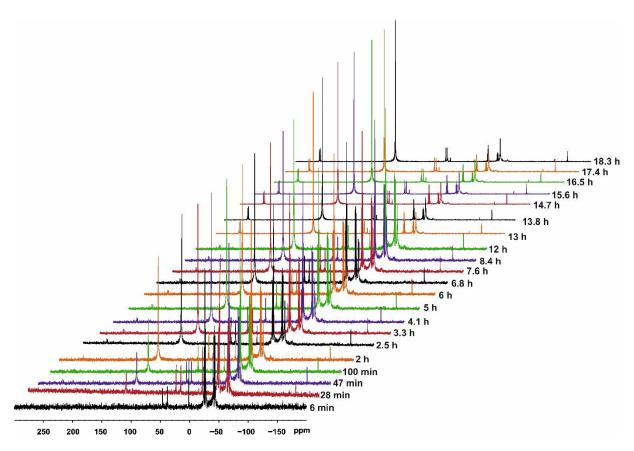


Figure S22 - $^{31}P\{^{1}H\}$ NMR monitoring (242.87 MHz, [D₈]THF) of the reaction between [K([18]crown-6)][Cp*Fe(C₁₀H₈)] and **L** at 273 K; range of 300 ppm $\geq \delta \geq -200$ ppm.

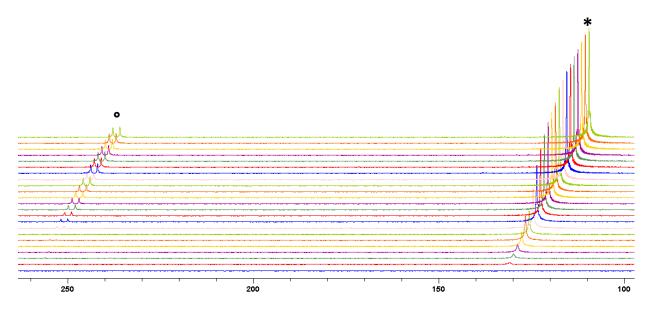


Figure S23 - $^{31}P\{^{1}H\}$ NMR monitoring (242.87 MHz, [D₈]THF) of the reaction between [K([18]crown-6)][Cp*Fe(C₁₀H₈)] and **L** at 273 K; range of 260 ppm $\geq \delta \geq 130$ ppm. * = **1-\sigma**, ° = unknown species.

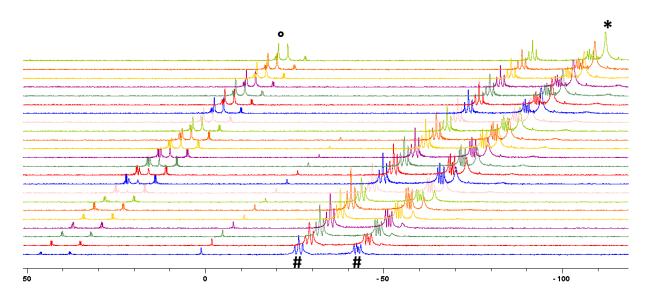


Figure S24 - $^{31}P\{^{1}H\}$ NMR monitoring (242.87 MHz, [D₈]THF) of the reaction between [K([18]crown-6)][Cp*Fe(C₁₀H₈)] and **L** at 273 K; range of 50 ppm $\geq \delta \geq -50$ ppm. * = **1-** π , ° = unknown species, # = **2**.

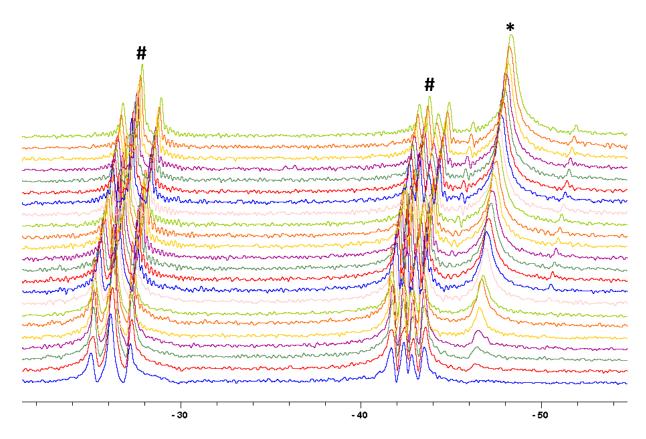


Figure S25 - $^{31}P\{^{1}H\}$ NMR monitoring (242.87 MHz, [D₈]THF) of the reaction between [K([18]crown-6)][Cp*Fe(C₁₀H₈)] and **L** at 273 K; range of -20 ppm $\geq \delta \geq -50$ ppm. * = **1-** π , # = **2**.

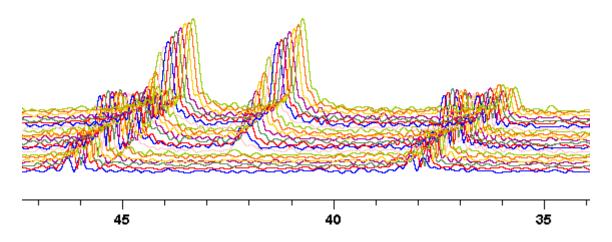


Figure S26 - $^{31}P\{^{1}H\}$ NMR monitoring (242.87 MHz, [D₈]THF) of the reaction between [K([18]crown-6)][Cp*Fe(C₁₀H₈)] and **L** at 273 K; range of 50 ppm $\geq \delta \geq$ 35 ppm; unknown species.

S3 UV-vis Spectra

S3.1 UV-vis spectrum of $1-\sigma$ and $1-\pi$

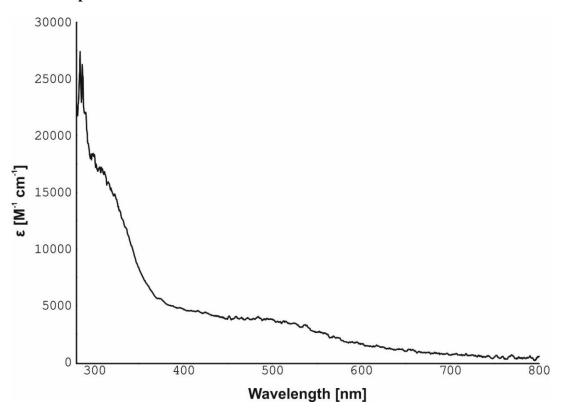


Figure S27 - UV/vis spectrum of 1- σ and 1- π in THF.

S3.2 UV-vis spectrum of 2

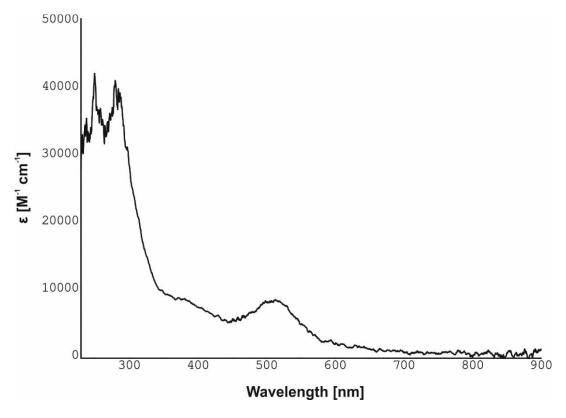


Figure S28 - UV/vis spectrum of 2 in THF.

S3.3 UV-vis spectrum of $3-\sigma$

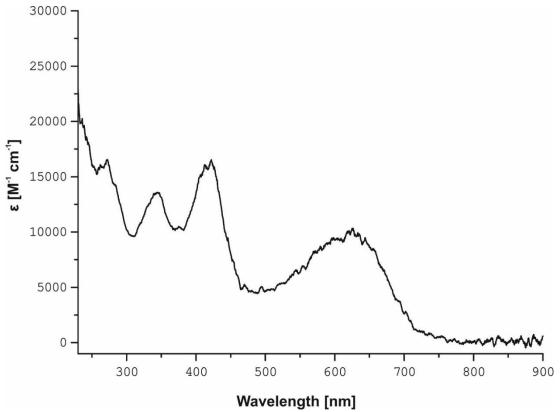


Figure S29 - UV/vis spectrum of 3- σ in THF.

S3.4 UV-vis spectrum of $3-\pi$

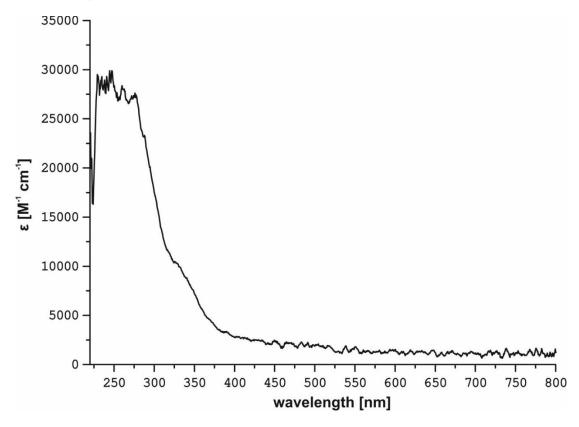


Figure S30 UV/vis spectrum of $3-\pi$ in THF.

S4 IR Spectra

S4.1 IR spectrum of 3-σ

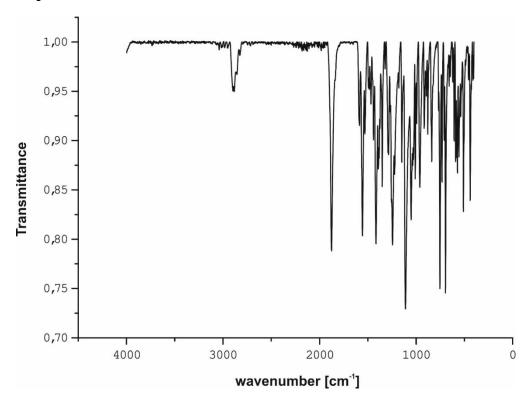


Figure S31 Solid state IR spectrum of $3-\sigma$.

S4.2 IR spectrum of $3-\pi$

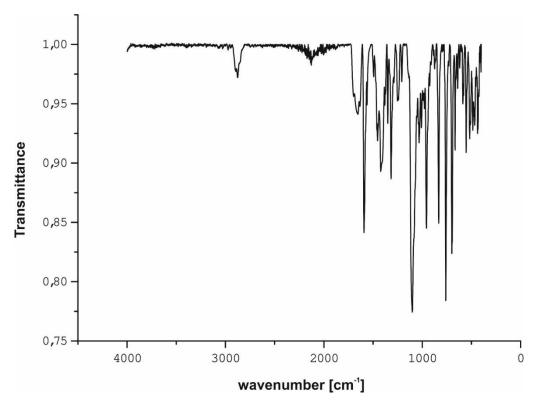


Figure S32 Solid state IR spectrum spectrum of $3-\pi$.

S5 X-ray Crystallographic Data

Table S1. Crystallographic data and structure refinement of $1-\sigma$, 2, $3-\sigma$ and $3-\pi$.

	1-σ	2	3-σ	3-π
Empirical formula	C ₄₄ H ₅₅ FeKNO ₆ P	$C_{66}H_{71}FeKN_2O_6P_2$	C ₄₅ H ₅₅ FeKNO ₈ P	C ₄₅ H ₅₅ FeKNO ₈ P
Formula weight / g·mol ⁻¹	819.81	1145.13	863.82	863.82
Temperature / K	123.01(10)	123.00(10)	123.00(10)	123.00(10)
Crystal system	monoclinic	orthorhombic	triclinic	monoclinic
Space group	$P2_1/n$	Pbcn	P-1	P21/n
a / Å	15.3111(2)	13.8522(4)	12.0465(4)	16.9137(3)
b / Å	17.0781(2)	38.2033(10)	12.2446(5)	11.6932(2)
c / Å	16.7872(3)	23.0082(6)	15.7693(5)	21.2321(4)
α /°	90	90	108.340(3)	90
β/°	106.967(2)	90	95.737(3)	99.828(2)
γ /°	90	90	92.223(3)	90
V / Å ³	4198.53(11)	12175.9(6)	2190.75(14)	4137.56(13)
Z	4	8	2	4
$ ho_{ m calc}$ / g cm ⁻¹	1.297	1.249	1.310	1.387
μ / mm ⁻¹	4.503	3.503	4.379	4.637
F(000)	1736.0	4832.0	912.0	1824.0
Crystal size / mm ³	$0.369 \times 0.172 \times$	$0.376 \times 0.305 \times$	$0.385 \times 0.243 \times$	$0.739 \times 0.222 \times$
•	0.122	0.186	0.125	0.072
Radiation / Å	$CuK\alpha (\lambda =$	$CuK\alpha (\lambda =$	$CuK\alpha (\lambda =$	$CuK\alpha (\lambda =$
	1.54184)	1.54184)	1.54184)	1.54184)
2Θ range for data collection /°	7.558 to 147.676	6.788 to 147.266	7.396 to 147.044	7.324 to 152.886
Diffractometer	SuperNova	SuperNova	SuperNova	SuperNova
Index ranges	$-19 \le h \le 18, -20$ $\le k \le 21, -20 \le 1$ ≤ 19	$-15 \le h \le 16, -38 \le k \le 46, -19 \le 1 \le 28$	$-14 \le h \le 14, -15$ $\le k \le 15, -19 \le 1$ ≤ 17	$-21 \le h \le 21, -12$ $\le k \le 14, -26 \le 1$ ≤ 24
Reflections collected	33166	32483	16728	27823
Independent reflections	$8388 \; [R_{int} =$	$11993 [R_{int} =$	$8612 \; [R_{int} =$	8583 [Rint =
	$0.0402,R_{sigma}=$	$0.0520, R_{sigma} =$	$0.0404,R_{sigma}=$	0.0310, Rsigma =
	0.0289]	0.0493]	0.0481]	0.0267]
Data/restraints/parameters	8388/390/655	11993/1623/1076	8612/0/519	8583/46/538
Goodness-of-fit on F ²	1.029	1.052	1.037	1.093
Final R indexes [I>= 2σ	$R_1 = 0.0303, wR_2$	$R_1 = 0.0850, wR_2$	$R_1 = 0.0411, wR_2$	R1 = 0.0782,
(I)]	= 0.0796	= 0.2172	= 0.1034	wR2 = 0.1713
Final R indexes [all data]	$R_1 = 0.0322, wR_2$	$R_1 = 0.0998, wR_2$	$R_1 = 0.0427, wR_2$	R1 = 0.0799,
	= 0.0812	=0.2290	= 0.1051	wR2 = 0.1722
Largest diff. peak/hole / e Å-3	0.46/-0.40	0.78/-0.54	0.87/-0.51	1.23/-0.49

S6 DFT Calculations

S6.1 General methods

All calculations were carried out with the ORCA program package. [9,10] All geometry optimisations were performed at the BP86-D3BJ/def2-TZVP^[11–15] level of theory in the gas phase. Frequency calculations were carried out to confirm the nature of stationary points found by geometry optimisations. Density fitting techniques, also called resolution-of-identity approximation (RI), [16] were used for GGA calculations, whereas the RIJCOSX^[17] approximation was used for TPSSh calculations. To save computational cost the phenyl groups at the 4-position of the phosphinine moiety of ligand **L** were replaced by hydrogen atoms, and [K([18]crown-6)]⁺ countercations were omitted. Approximate transition states were generated using the nudged elastic band (NEB) method implemented in ORCA, followed by a saddle-point optimisation.

S6.2 Isomerisation of complex 1

Final single-point calculations on the BP86 geometries were conducted at the TPSSh-D3BJ/def2-TZVP level of theory and zero-point energies and thermal corrections at 298 K were added from the BP86 calculations. Additionally, final single-point calculations of the calculated minima and transition states were carried out with the CPCM^[19] model for THF at the TPSSh-D3BJ/def2-TZVP level.

The isomerisation of $1-\pi$ to $1-\sigma$ proceeds via a two-step mechanism. First, the intermediate 1-int is formed, in which the Cp*Fe moiety 'slips' from η^4 to η^2 coordination to the phosphinine moiety, with concomitant formation of a new Fe—N interaction. Subsequent reorientation of the phosphinine switches it from a π - to a σ -coordination mode, providing the final isomerised complex (Figure S33).

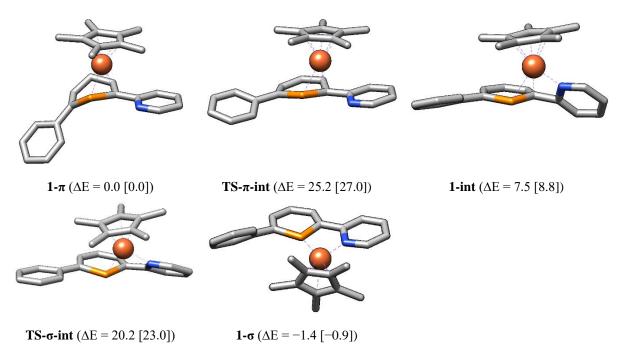


Figure S33 - Optimised structures for the isomerisation of **1-π to 1-σ**. Energies are given in kcal·mol⁻¹ relative to the optimised structure of **1-π**. Energies in brackets correspond to electronic energies with solvent correction (TPSSh-D3BJ/def2-TZVP CPCM(THF)).

S6.3 Formation and electronic structure of compound 2

The reaction between $1-\pi$ and L in the gas phase involves two steps (energies were obtained at the TPSSh-D3BJ/def2-TZVP^[18] level of theory): the activation barrier-free formation of the Van-der-Waals complex **VdW-** π -**L**, and subsequent P-P bond formation yielding **2**. The respective transition state **TS-** π -**L** has a low energy (Figure S35), consistent with the experimental observation of rapid formation of **2** at room temperature.

The highest occupied molecular orbital (HOMO) of 2 strongly resembles the HOMO of the phosphacyclohexadienyl anion (Figure S35). Additionally, inspection of the molecular orbitals of 2 (Figure S36) revealed a $3d^6$ configuration at the iron center. Therefore, 2 can be described as an iron(II) complex with an η^6 -coordinating, dianionic di(phosphacyclohexadienyl) ligand. Thus, the formation of 2 could formally be regarded as a redox reaction.

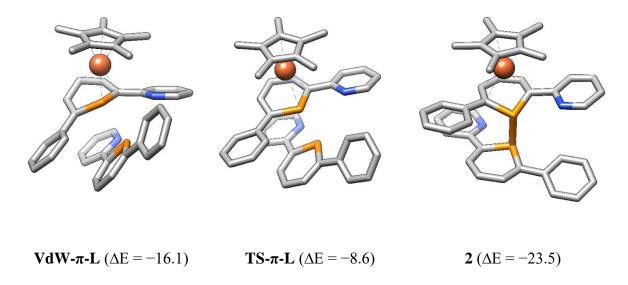


Figure S34 - Optimised key structures for the formation of **2** from **1-** π and **L**. Energies are given in kcal·mol⁻¹ relative to the sum of the electronic energies of **1-** π and **L**.

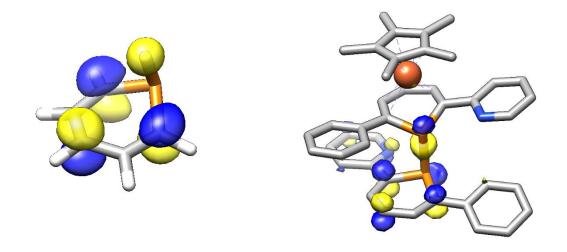


Figure S35 - Comparison of the HOMOs of the phosphacyclohexadienyl anion (left) and **2** (right), obtained at the BP86-D3BJ/def2-TZVP level. Hydrogen atoms of **2** have been omitted for claritiy. Surface isovalue = 0.06.

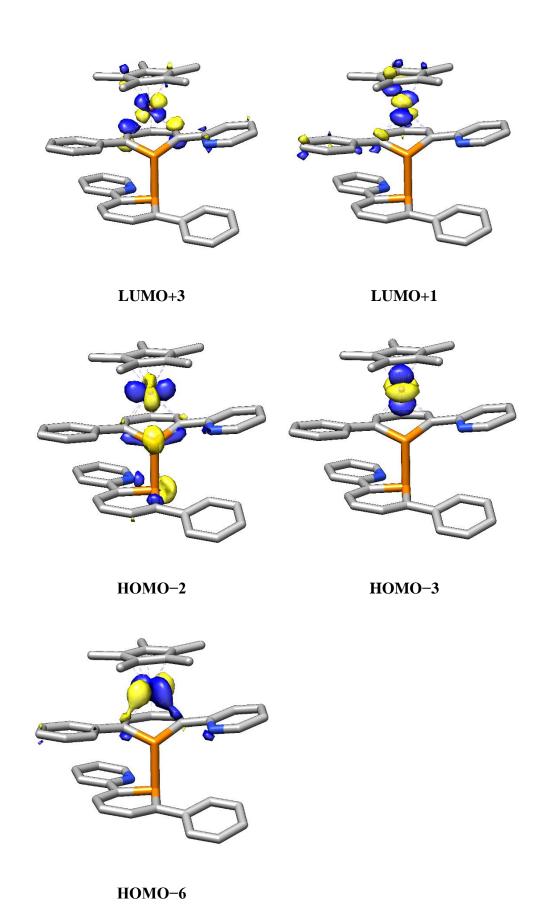


Figure S36 - Kohn-Sham orbitals of **2** with significant 3d character at iron (BP86-D3BJ/def2-TZVP level). Hydrogen atoms have been omitted for claritiy. Surface isovalue = 0.06.

S6.4 Reactions of $1-\pi$ and $1-\sigma$ with CO₂

To account for solvent effects, final single point calculations of the calculated minima and transition states were carried out with the CPCM^[19] model for THF at the TPSSh-D3BJ/def2-TZVP level.

In the case of $1-\pi$, the first step of the mechanism involves the activation barrier-free formation of the van-der-Waals complex VdW- π -CO₂, followed by the exothermic formation of $3-\pi$ *via* an energetically low-lying transition state TS- π -CO₂ (+5.5 kcal·mol⁻¹, Figure S37). Note that for these calculations the [K([18]crown-6)]⁺ countercation has been omitted for the sake of computational efficiency; they therefore do not account for any addition electrostatic stabilisation as a result of K···O—C interactions (as observed in the solid state for $3-\pi$).

In the case of **1-** σ , the first step involves the activation barrier-free coordination of CO₂ at the iron atom forming intermediate **VdW-\sigma-CO₂**. Subsequently, **VdW-\sigma-CO₂** undergoes CO₂ cleavage *via* transition state **TS-** σ -CO₂ to form **3-** σ in a very strongly exothermic reaction. As for **1-** π , the activation barrier for this process is very small (+3.5 kcal·mol⁻¹, Figure S37), in agreement with the instantaneous reaction observed experimentally.

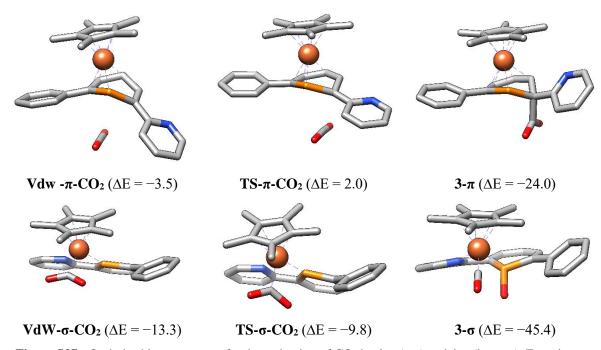


Figure S37 - Optimised key structures for the activation of CO_2 by $1-\pi$ (top) and $1-\sigma$ (bottom). Energies are given in kcal·mol⁻¹ relative to the sum of the electronic energies of CO_2 and the relevant isomer of 1.

S6.5 Electronic structure of 3-σ

Because the analysis of the Kohn-Sham orbitals of $3-\sigma$ was not as straightforward as for 2, CASSCF/def2-TZVP calculations with ten electrons in seven orbitals were carried out in order to obtain an insight into the electronic structure and bonding in $3-\sigma$. To aid convergence, the CASSCF calculation was carried out including six roots. The ground-state of $3-\sigma$ is of dominant single-reference character (94% contribution of the ground-state configuration state function). Inspection of the natural orbitals of the active space reveals three lone pairs of electrons at the iron center (HOMO, HOMO-1 and HOMO-2, Figure S39) accounting for a $3d^6$ configuration and two metal-ligand bonding orbitals (HOMO-3 and HOMO-4) and their respective antibonding counterparts

(LUMO and LUMO+1). Additionally, the HOMO-5 resembles the HOMO of the phosphacyclohexadienyl anion, thus indicating the presence of a negative charge within the C_5P ring in **3-\sigma**. The bonding situation can therefore be described as an interaction of a cationic $[Fe^{II}(CO)(Cp^*)]^+$ fragment with a dianionic oxophosphacyclohexadiene ligand.

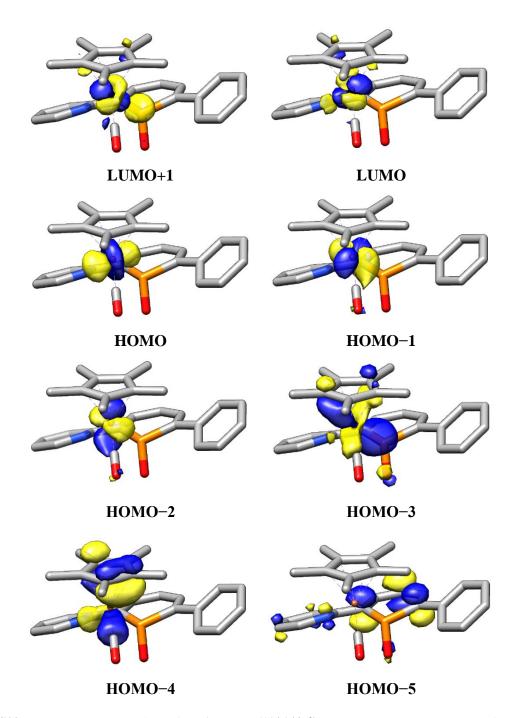


Figure S38 Selected molecular orbitals of $4-\sigma$ (CASSCF(10/7)/def2-TZVP level). Hydrogen atoms of the $4-\sigma$ have been omitted for claritiy. Surface isovalue = 0.06.

S6.6 Cartesian coordinates of optimised structures

[C₅H₅P]⁻

C	-7.61943319559269	1.22409425547681	0.13533784859837
C	-7.64465161413261	2.59587319386719	0.30655806150595
C	-6.46424418962032	0.47009145181158	-0.19230249875865
C	-5.17738486014462	1.04326799732260	-0.03165911543904
C	-4.92836447320168	2.39473921664987	0.12075910803984
P	-6.23079634688711	3.61428255628433	-0.18313646659172
Η	-8.53726116116016	0.65629539320926	0.35526367967173
Η	-8.52398638132399	3.07923019770116	0.74215399853911
Η	-4.33148673543528	0.34486561603294	0.06765676734310
Η	-6.55098221652423	-0.61550683348109	-0.28501854811743
Η	-3.93687915667162	2.73953775574318	0.42846629093591
Η	-6.34139966930567	3.47491919938217	-1.64968912572718

L

P	6.04122337926595	24.63408168499444	10.97933198060370
C	6.08960118303451	26.38102873225077	10.75022889366670
C	5.86394350235381	27.01786074063859	9.52198000089409
Н	5.94173500748529	28.10846159769303	9.48776660296668
C	5.58319435084327	26.34316947281555	8.33494039956308
C	5.49794456569007	24.95320382160534	8.25123074122980
Η	5.30488939279188	24.51332433392341	7.26873401066420
C	5.68871142267727	24.09357247778719	9.33797204441012
C	5.57916845761728	22.62759068525408	9.15123071191018
C	4.72646922398813	22.05698178249720	8.18793551438395
Н	4.09772039713966	22.69146888071547	7.56349075890103
C	4.67201228383245	20.67090387157797	8.06514324487574
Η	4.01602743110801	20.20974644533032	7.32525654825285
C	5.45558524081305	19.88503800485405	8.91149915598559
Н	5.44375709874828	18.79671836851895	8.85008926594723
C	6.25604186692401	20.53456912567334	9.85513931531024
C	6.36810393266843	27.20464479007647	11.95036446841172
Η	5.44717675621914	26.92807203796902	7.42378818398789
C	5.66160778644344	28.39942471765804	12.18383860842483
C	5.91647771486721	29.17118253374181	13.31686695234687
C	6.88192929376833	28.76510071451661	14.24259996517816
C	7.58856340350576	27.57932192690138	14.02585573436544
C	7.33467792257424	26.80721372213552	12.89250019850521
Η	7.90164320401353	25.89157814763470	12.71613940404032
N	6.32603590454227	21.86411741452478	9.97879964270533
Η	6.87779545221354	19.95805821272706	10.54645969983165
Η	4.88436086533311	28.70664610395373	11.48235146563765
Н	5.35129726149673	30.08954095601434	13.48311094968027
Н	7.08159905106075	29.36934419713680	15.12843577416666
Η	8.34813664698058	27.25666449888003	14.73934976315279

_	2 5250240405220	10 1010 (00 (00 (00)	10.05011010000000
	3.72592401196239	13.42136286396835	13.05841048003086
P	5.50741856996626	12.40791601565484	12.77336220237265
N	4.45241524290449	14.58019982644907	11.69500525801739
C	7.96462190229615	11.14102547354741	12.95606485756925
Н	8.56915206010328	10.28075232362786	13.26636731763079
C	5.80641481923798	14.54496225852838	11.26442046771582
C	6.30886722935948	15.51378447711643	10.37749995200805
Η	7.36328550097966	15.45172990411800	10.10107369759245
C	8.61257994180999	12.14895602150099	12.21596825323474
C	7.94006245125282	13.23602191264865	11.64947452229262
Η	8.53515965993015	13.95662339415390	11.07573608031956
C	6.04969205736462	9.90197877564039	13.92441084653583
C	5.50499492078208	16.50825541721105	9.84398734262572
Н	5.90390161046380	17.24814981645005	9.14910467590880
C	6.60301295011311	11.08955062401833	13.26070281068218
Č	2.06852235235543	12.38475414999738	13.72796025904286
Č	6.56911992103160	13.45528404035571	11.78893669337016
Č	2.98081276905463	12.65666820618897	14.81297457402019
C	3.15405039715675	14.08461905644085	14.90286003940619
C	4.83456319370247	9.33506960098914	13.48305128108727
Н	4.31361889190257	9.81972260698764	12.65523824027235
C	3.67656117724735	15.53917517058476	11.08757962846528
Н	2.62314082720625	15.50252095239016	11.34565220829595
C	4.31144986189041	8.18986829728283	14.07885684681174
H	3.37135936815636	7.77312205640991	13.71040867036698
C	2.30985713851632	14.68902834537661	13.90626746829831
C	1.64324555980373	13.64169246323483	13.17061768004312
C	4.13596313835306	16.49196859562335	10.20490336036189
Η	3.42638997780370	17.20825812544575	9.78736340357240
C	6.71017006174308	9.26860520477863	14.99909926502537
Н	7.63453278865358	9.70764137059168	15.37903089264052
C	4.97964493548110	7.57430956223483	15.14498562727098
Н	4.56431678037425	6.68346009548213	15.61936254361677
C	2.11383635004085	16.16705541338921	13.76137267678583
Η	3.06235146947623	16.69551462891634	13.57668379258487
Η	1.67072971472238	16.59054152389485	14.67931428666458
Η	1.43629224333642	16.41405881993866	12.93252660351801
C	3.58157681527615	11.66108258655962	15.75583659653268
Η	3.58984112836042	10.64899607456343	15.33469526972583
Η	3.01723092338621	11.62411800589632	16.70577162028925
Н	4.62376025244625	11.91323213049187	15.99862602858001
C	6.18396301654359	8.12388323293657	15.59808025445408
Н	6.71052148874602	7.66448735273890	16.43799462637624
C	3.96119197469626	14.81723702801862	15.92748433455901
Н	4.82076387435136	14.21610210308666	16.25397150416272
Н	3.36429959773181	15.06391193851799	16.82811354696823
Н	4.35593206491193	15.76005496209445	15.52276601974220
C	1.55142316545978	11.03854484229403	13.32953550274087
Н	1.40274959551542	10.97051456401285	12.24164636164965
Н	0.58195927042286	10.81344541188779	13.81368058140783
Н	2.25163750959263	10.24466763738927	13.61942752165738
C	0.59863817538809	13.81063038573668	12.11381759495715
Н	0.72122631804818	14.74995394836136	11.55356499138391
Н	-0.42320135351064	13.82443949026823	12.53957132469791
Н	0.63885161773648	12.99147387278839	11.38086274489048
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C
C
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CO_2

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VdW-σ-CO₂

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C 4.69185285488964	14.35965477588063	12.11218145065662
C 4.95400152694683	15.73745226523233	12.11167264842955
H 5.81913485359461	16.10233045040087	12.66730357574632
C 7.29613567266747	12.54467393568061	14.23035921630320
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C 5.75415859857014	9.22142107951128	13.12328098317932
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H 4.34279337611113	17.69194285421748	11.41424537622335
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5.32527428530807 8.69456298135441 \mathbf{C} 11 88626712304251 5.05815094490507 H 9.37151911694905 11.06756449656322 2.86568241249067 14.71082415048875 10.67050766612572 Η 2.13609236398676 14.23850113657486 10.02059737949864 C 5.20424020007035 7.31792064038975 11.70643573855699 4.87101184284879 6.93895608290193 10.73845198480673 1.32544436924973 12.56812927947231 12.93165456402186 \mathbf{C} 0.79551016943479 12.34978271029655 11.60255177332506 \mathbf{C} 3.08826165244912 16.07761320174674 10.65141791618134 2.46128665598992 H 16.70629876494466 10.01756232954228 \mathbf{C} 6.02193914603173 8.31811686614762 14.17295062813469 Η 6.31120977075130 8.70904197492446 15.15048013516069 C 5.48948482514096 6.43086039480131 12.75014660950437 5.38536688744098 H 5.35374543548549 12.60556280811654 1.26478887981874 13.85226976550800 13.69603597977122 2.07767461124135 13.91141481686829 14.43329471350814 0.30872769218668 14.23845027269546 13.96246314877646 1.36792594935871 14.71924684482256 13.02665086445770 \boldsymbol{C} 2.10790719256539 8.91469132761268 12.38526270679506 Н 2.50420731980277 8.57509377145833 11.41967654577394 1.22649211035141 8.29467672417743 12.63095517546906 Н 13.14116515142021 2.87771052483301 8.71626716294510 Н 5.89293058573302C 13.98852471217716 6.94128307666253 Η 6.09641890102428 6.26309506362679 14.82049197655864 C 2.47695132691017 11.06329137309089 14.73947466750421 Η 3.27358617158084 10.30781736879777 14.69357077144267 1.70289524277102 10.69114422009256 15.43441880753673 2.91511985255239 11.97052953510036 15.17827121507038 \mathbf{C} 0.52502915135889 10.29692591394056 10.02835285493803 Η 0.47299142138064 10.98332128441680 9.17391258160363 H -0.48947588754974 9.90089553105803 10.21881318528188 9.73991560655683 H 9.46176641535396 1.17454906987997 \mathbf{C} -0.04075989678452 13.29562421371610 10.80104085311612 H 0.07076762371514 14.33153038219108 11.14742259228973 H -1.11211957380265 13.03446049382925 10.87616006351614 0.23405671082369 13.25987837825828 9.73619977226587 8.15253505360587 12.79729971192003 14.85825359188132 3.20091851853001 11.47585559730916 9.66862498558332 2.80957746600907 12.35079709671019 8.88345476598036 O 3.68264249928746 10.34064242437345 9.54261506235651

ΤS-σ-CO2

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3-σ

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VdW-π-CO₂

Fe 8.21324871075741	15.39579092730344	2.44799070725311
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3-π

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