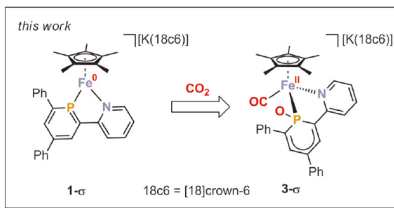


Literature Report: A Facile C=O Bond Splitting of CO₂ Catalyzed by Phosphinine Iron(0) Complex



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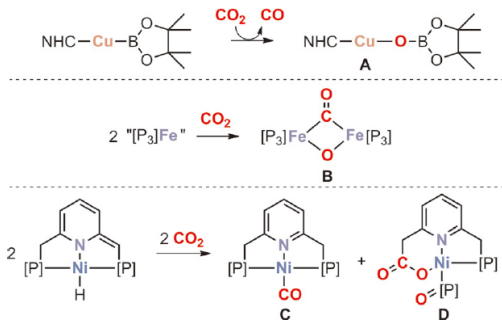
Background

C=O Bond Splitting of CO_2 has become established for early 3d metals and f-block metals.

In comparison, the use of earth-abundant late 3d metals remains surprisingly under-explored, particularly given the role that such metals play in biological CO_2 reduction to CO, mediated by Ni, Fe CO dehydrogenase.

Background

Previous examples of CO₂ cleavage based on 3d metals



Laitar, D. S. et al. *Journal of the American Chemical Society* **2005**, 127, 17196–17197

Sadique, A. R. et al. *Inorganic chemistry* **2008**, 47, 784–786

Oren, D. et al. *Organometallics* **2018**, 37, 2217–2221

This work is the first reported example of C=O cleavage of a CO₂ molecule mediated by a single Fe centre.

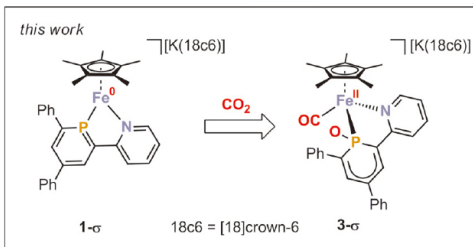


Figure 1: Leitl, J. et al. *Angew. Chem. Int. Ed.* **2019**, 58, 15407–15411

Inspiration

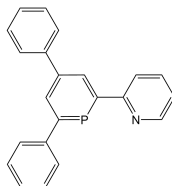
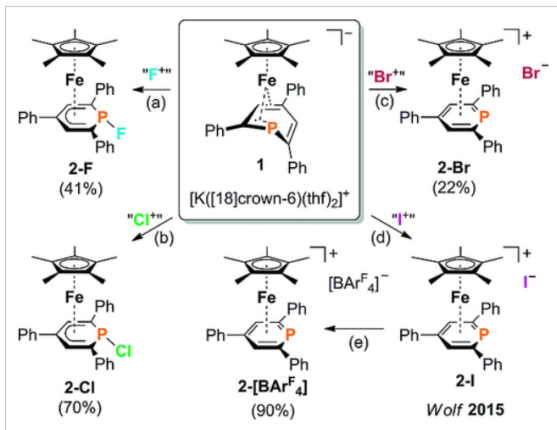


Figure 3: New ligand
2-(2'-pyridyl)-4,6-
diphenylphosphinine

Figure 2: $[\text{Cp}^*\text{Fe}(\eta^4\text{-TPP})]$ and derived compounds
Hoidn, C. M. et al. *Eur. J. Inorg. Chem.* **2019**, *2019*,
1567–1574

Preparation

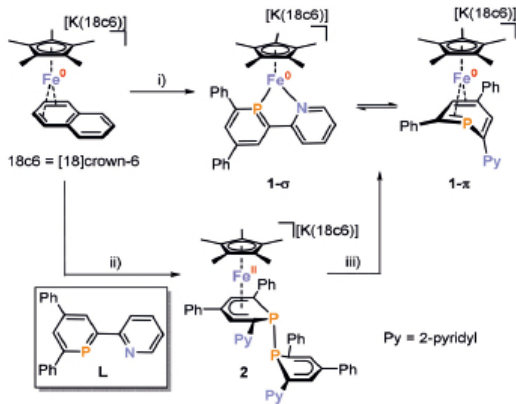
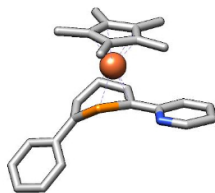


Figure 4: i) L, DME, -35°C to RT, -naphthalene; ii) 1 equiv. $[K([18]crown-6)][Cp^*Fe(C_{10}H_8)]$, 2 equiv. L, toluene/THF, -35°C to RT; iii) 1 equiv. $[K-(18)crown-6][Cp^*Fe(C_{10}H_8)]$, THF.

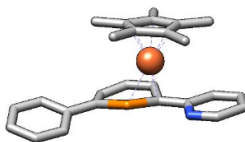
Solution-phase ^{31}P NMR spectrum shows 1- σ and 1- π , while XRD and ^{31}P CP MAS show only 1- σ . Selective crystallization failed.

Calculation for $1-\sigma/1-\pi$

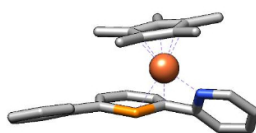
Conversion of $1-\pi$ to $1-\sigma$ is calculated to proceed with a barrier of 27.0 kcal/mol, consistent with an equilibrium at room temperature. NMR indicates an approximately 2:1 ratio of $1-\sigma$: $1-\pi$.



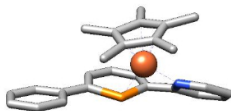
$1-\pi$ ($\Delta E = 0.0$ [0.0])



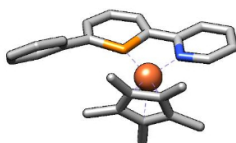
$TS-\pi-int$ ($\Delta E = 25.2$ [27.0])



$1-int$ ($\Delta E = 7.5$ [8.8])



$TS-\sigma-int$ ($\Delta E = 20.2$ [23.0])



$1-\sigma$ ($\Delta E = -1.4$ [-0.9])

Figure S33 - Optimised structures for the isomerisation of $1-\pi$ to $1-\sigma$. Energies are given in kcal·mol⁻¹ relative to the optimised structure of $1-\pi$. Energies in brackets correspond to electronic energies with solvent correction (TPSSH-D3BJ/def2-TZVP CPCM(THF)).

Splitting Reaction

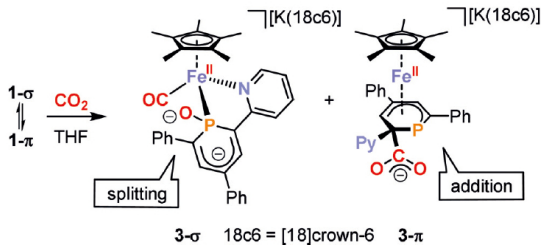
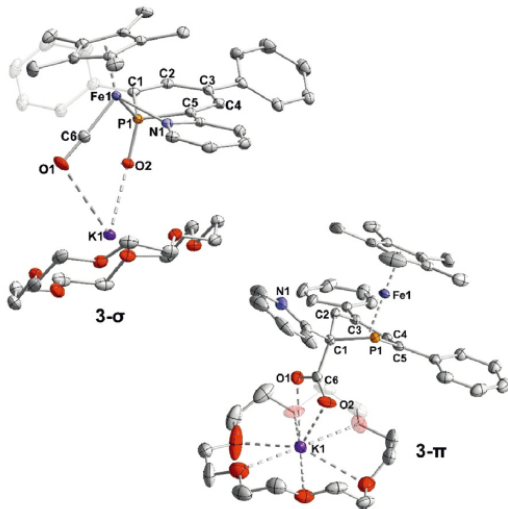


Figure 5: Reaction of $1-\sigma$ and $1-\pi$ with CO_2 (1 atm) in THF at room temperature.

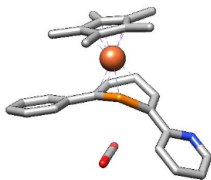
Solid state structures of 3- σ and 3- π



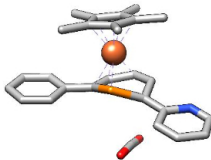
Calculation for $1 \rightarrow 3$ Reaction

Very small energy barriers:

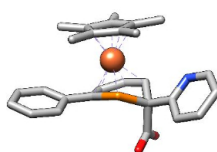
- ▶ $1-\sigma \rightarrow 3-\sigma$: 3.5 kcal/mol
- ▶ $1-\pi \rightarrow 3-\pi$: 5.5 kcal/mol



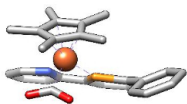
VdW- π -CO₂ ($\Delta E = -3.5$)



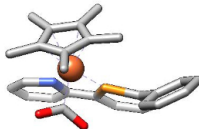
TS- π -CO₂ ($\Delta E = 2.0$)



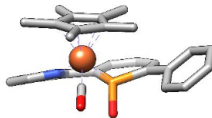
3- π ($\Delta E = -24.0$)



VdW- σ -CO₂ ($\Delta E = -13.3$)



TS- σ -CO₂ ($\Delta E = -9.8$)



3- σ ($\Delta E = -45.4$)

Figure S37 - Optimised key structures for the activation of CO₂ by **1- π** (top) and **1- σ** (bottom). Energies are given in kcal·mol⁻¹ relative to the sum of the electronic energies of CO₂ and the relevant isomer of **1**.

Discussion and Expectations

- ▶ Full catalysation cycle
- ▶ Diversification of $1-\sigma$
- ▶ Orbital analysis