

Document Title

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ABSTRACT

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INTRODUCTION

For molar mass, use \mathcal{M} .

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Stoichiometry The relationship between the relative quantities of substances taking part in a reaction or forming a compound, typically a ratio of whole integers.

Atomic mass The mass of an atom of a chemical element expressed in atomic mass units. It is approximately equivalent to the number of protons and neutrons in the atom (the mass number) or to the average number allowing for the relative abundances of different isotopes.

Law of Definite Proportions: A given compound always contains exactly the same proportion of elements by mass.

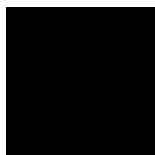


FIGURE 1: The pairwise nature of the Born radius.

See Figure 1.

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TABLE 1: Gravimetric analysis of silver halides in a 1.27-mL sample of Salton Sea water.

Qty of Sample	Test Tubes				Avg
	A	B	C	D	
Mass (g)	1.399	1.32	1.328	1.408	1.364
Density (g/mL)	1.10	1.04	1.05	1.109	1.07
Mass w/ Precipitate (g)	13.443	13.401	13.348	—	13.397
Mass AgCl (10^{-2} g)	9.0	9.2	8.7	—	8.9
Moles AgCl (10^{-4} mol)	6.28	6.42	6.08	—	6.50

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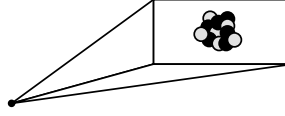


FIGURE 2: The nucleus.

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0.1. Procedure

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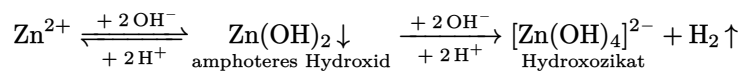
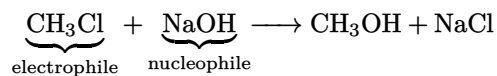
0.2. Materials

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