

Physical World 2D Report

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I. MATHEMATICAL MODEL OF THE ALGAE BOTTLE

Based on the Fourier's law of thermal conduction: $R = \frac{\Delta T}{Q}$, and using the substitutions $\dot{Q} = mc\dot{T} = \rho V c \dot{T}$, where ρ , V and c is the density, volume and specific heat capacity of water respectively. we can write a first order differential equation:

$$\Delta T = T - T_{amb} = -R\rho V c \frac{dT}{dt}$$

Solving this differential equation,

$$\begin{aligned} \int_0^t dt &= -R\rho V c \int_{T_{initial}}^T \frac{1}{T - T_{amb}} dT \\ \frac{t}{R\rho V c} &= \ln \left| \frac{T - T_{amb}}{T_i - T_{amb}} \right| \\ (T_{initial} - T_{amb}) \exp\left(-\frac{t}{R\rho V c}\right) &= T - T_{amb} \\ \Delta T &= \Delta T_{initial} \cdot \exp\left(-\frac{t}{R\rho V c}\right) \end{aligned}$$

It can be concluded that a best fit exponential curve of a ΔT against t plot may be used to empirically determine the value of $\lambda_{algae\ bottle}$.

II. METHODOLOGY

Figure 1 shows the experimental set up for determining $\lambda_{algae\ bottle}$.

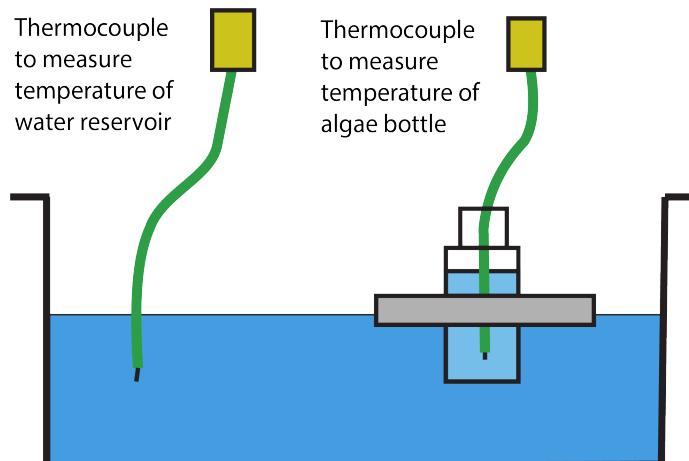


Fig. 1: Experimental set up

The following methodology was used to obtain data:

- 1) Fill the bottle with 30ml of hot water.
- 2) Place thermocouples in the bottle and the tank as shown in Figure 1.

- 3) Use a video camera to record the displays of the digital thermometer and a stopwatch.
- 4) Place the (bottle with floatation device) into the water resevoir and as shown in Figure 1 and start the stopwatch.
- 5) Record for more than 100 seconds.
- 6) Review the footage and record the time displayed by the stopwatch at every temperature decrement ($0.1^\circ C$).

After obtaining data to calculate $\lambda_{algae\ bottle}$, another experiment was carried out to determine λ_{tubing} . Silicone tubing was stuffed into the container, and the cooling curve data was also obtained by the same method as the previous experiment.

III. DATA ANALYSIS

Once the data was obtained, a best fit exponential curve in the form $y = Ae^{-Bx}$, where $A, B \in \mathbb{R}^+$, was plotted in Microsoft Excel. (See figure 2)

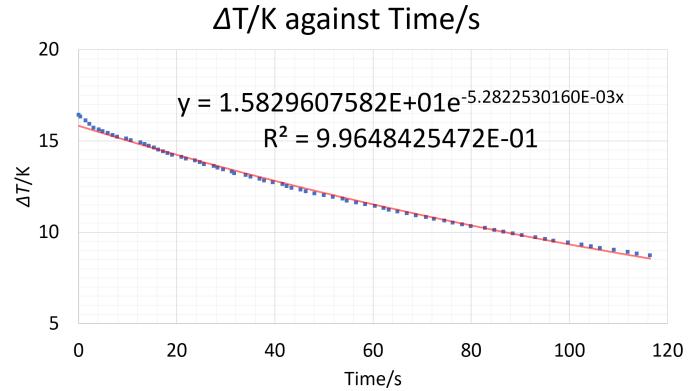


Fig. 2: Graph of $\Delta T/K$ across bottle wall against Time/s

The best fit curve was found to be:

$$y = 1.58296 \times 10^1 \cdot \exp(-5.28225 \times 10^{-3} \cdot t)$$

Given $B = 5.28225 \times 10^{-3}$, we can empirically find R by:

$$R_{algae\ bottle} = \frac{1}{\lambda_{algae\ bottle}} = \frac{1}{B\rho V c}$$

Given $\rho = 0.9881 g \cdot cm^{-3}$ at $50^\circ C$ (see reference [1]), $c = 4.18 J g^{-1} K^{-1}$ (see reference [2]), and $V = 30 cm^{-3}$ (measured), then we can determine using the same methods of data analysis:

$$\lambda_{algae\ bottle} = 0.654512033 WK^{-1}$$

$$\lambda_{tubing} = 0.360332327 WK^{-1}$$

	Lower Limit	Upper Limit
$R_{conv,water}$ at $25^\circ C/KW^{-1}$	1.391873685	0.006959368
$R_{polystyrene}/KW^{-1}$	0.573124458	0.512795568
$R_{conv,water}$ at $35^\circ C/KW^{-1}$	1.391873685	0.023197895
Theoretical λ_{bottle}/WK^{-1}	0.29789639	1.841780616

TABLE I: Upper and Lower Limits of λ_{bottle}

Using the value of $\lambda_{algae,bottle}$ and the dimensions of the algae bottle, we are able to find the thermal conductivity k of the polystyrene [3] algae bottle. $k_{polystyrene} = 1.7651032 \times 10^{-1} Wm^{-1} K^{-1}$. The $\lambda_{algae,bottle}$ calculated falls within upper and lower limits as calculated in Table I (see attached excel files, reference: [4], [5]).

Based on research [6], the conductivity of (non-expanded) high impact polystyrene is approximately $1.72 \times 10^{-1} Wm^{-1} K^{-1}$, indicating that the result attained is accurate.

IV. DESIGN OF HEAT EXCHANGER

Since the calculated value of $\lambda_{algae,bottle}$ is 82% higher than the calculated λ_{tubing} , a heat exchanger was designed to exploit this higher heat conductance to remove heat more rapidly from the bottle. This allows for a larger heat load (such as from the sun) to be placed on the algae bottle, both increasing the safety margin of acceptable heat load before actuator saturation (i.e. peristaltic pump running at full power) and also energy efficiency at a constant heat load.

A “water jacket” design was created in Fusion360 Computer Aided Design (CAD) software and subsequently 3D printed. Figure 3 below is rendered in a transparent material to show the internal structure of the heat exchanger. The inlet is situated at the bottom to allow the pump to fill the container fully; this will also allow the colder water to remain at the bottom while hotter water rises to the top of the heat exchanger and then expelled from the outlet.



Fig. 3: Internal Structure of the Heat Exchanger

Figure 4 below clearly shows the triangular protrusions on the inner wall of the heat exchanger which creates channels for the cooling water to flow around the algae bottle while minimising contact area.

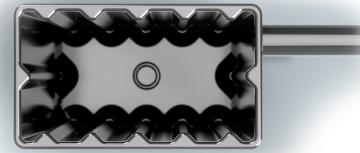


Fig. 4: Top View of the Heat Exchanger

Figure 5 below shows air gaps in the wall of an incompletely 3D printed heat exchanger which minimises heat gain or heat loss to the surroundings when the pump is not active so as to maintain temperature small temperature range for a long period.



Fig. 5: Incomplete 3D printed heat exchanger showing air gaps

Figure 6 below shows the actual 3D printed heat exchanger with an algae bottle fitted.



Fig. 6: Actual 3D Printed Heat Exchanger with an Algae Bottle

V. CONSIDERING THE TUBING AS AN OPEN SYSTEM

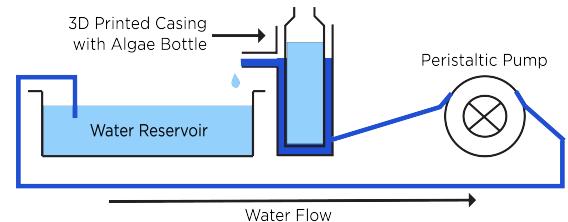


Fig. 7: Final Demonstration Setup Schematic Diagram

Considering the silicone tubing, pump and heat exchanger (excluding the algae bottle) shown in Figure 7 as an open system, the rate of change of energy in the control volume \dot{E}_{cv} is given by:

$$\dot{Q}_{cv} - \dot{W}_{cv} + \dot{m}_i \left(h_i + \frac{V_i^2}{2} + gz_i \right) + \dot{m}_e \left(h_e + \frac{V_e^2}{2} + gz_e \right)$$

Assuming steady state, $\dot{m}_i = \dot{m}_e = \dot{m}_w$ and $\dot{E}_{cv} = 0$ and ignoring gravitational potential energy and kinetic energy effects, the equation may be simplified to:

$$\frac{dE_{cv}}{dt} = \dot{Q}_{cv} - \dot{W}_{cv} + \dot{m}_w (h_1 - h_2)$$

Expanding \dot{Q}_{cv} ,

$$\dot{Q}_{cv} = \dot{Q}_{exchanged} + \dot{Q}_{pump,water}$$

However, in this case, the pump can be assumed to impart negligible heat on the water flowing through it. Thus, the heat transferred from the pump to the cooling water $\dot{Q}_{pump,water} = 0$

$$\dot{Q}_{cv} = m_a c_a \dot{T}_a - \dot{Q}_{ambient}$$

$$\dot{Q}_{cv} = m_a c_a \dot{T}_a - \lambda_{ambient} (T_a - T_{amb})$$

However, since the algae bottle is largely in direct contact with the cooling water and due to the high thermal conductance measured ($\lambda_{algae\ bottle} = 0.654512033 W K^{-1}$) we can assume that a negligible proportion of heat escapes into the air via convection. Thus, $\lambda_{ambient} (T_a - T_{amb}) = 0$, and the energy balance equation can be written:

$$m_a c_a \dot{T}_a - \dot{W}_{cv} + \dot{m}_w c_w (T_i - T_e) = 0$$

The simplified equation for \dot{E}_{cv} can be thus written (part b4):

$$\dot{E}_{cv} = m_a c_a \dot{T}_a - \dot{W}_{cv} + \dot{m}_w c_w (T_i - T_e)$$

To express $(\dot{W})_{cv}$ as a function of T_a , we assume $\dot{E}_{cv} = 0$.

$$m_a c_a \dot{T}_a - \lambda_{amb} (T_a - T_{amb}) - \dot{W}_{cv} + \dot{m}_w c_w (T_i - T_e) = 0$$

$$\int dt = \int \frac{m_a c_a}{\lambda_{amb} (T_a - T_{amb}) + \dot{W}_{cv} - \dot{m}_w c_w (T_i - T_e)} dT_a$$

$$t + c = \frac{m_a c_a}{\lambda_{amb}} \ln(\lambda_{amb} (T_a - T_{amb}) + \dot{W}_{cv} - \dot{m}_w c_w (T_i - T_e))$$

$$\frac{\lambda_{amb} (t + c)}{m_a c_a}$$

$$= \ln(\lambda_{amb} (T_a - T_{amb}) + \dot{W}_{cv} - \dot{m}_w c_w (T_i - T_e))$$

$$A e^{\lambda_{amb} t / (m_a c_a)} = \lambda_{amb} (T_a - T_{amb}) + \dot{W}_{cv} - \dot{m}_w c_w (T_i - T_e)$$

$$\dot{W}_{cv} = A e^{\lambda_{amb} t / (m_a c_a)} - \lambda_{amb} (T_a - T_{amb}) + \dot{m}_w c_w (T_i - T_e)$$

where $A \in \mathbb{R}$

VI. ANALYSIS

|W_cv|/W against t/s, Pump Voltage = 6V

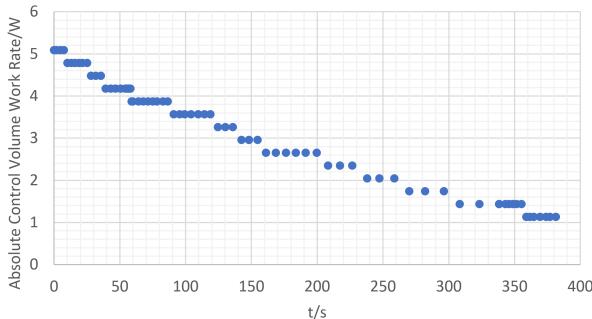


Fig. 8: Graph of Calculated $|\dot{W}_{cv}|$ against Pump Voltage

Based on the formula derived to calculate \dot{W}_{cv} , the graph of the calculated $|\dot{W}_{cv}|$ against pump voltage was plotted in Figure 8. The value of A was experimentally determined to be $-1.32 + 0.654(308.6 - 298.1) - 0.728 \cdot 4.18(298.6 - 295.4) = 4.94W$. Further experiments were conducted to measure the flow rate of the pump against the voltage supplied by plotting a volume against time graph for each voltage and finding the gradient of the graph. Thus, the actual \dot{W}_{cv} done by the pump can be calculated. The results are shown in Figure 9.

Actual Wcv against Pump Voltage

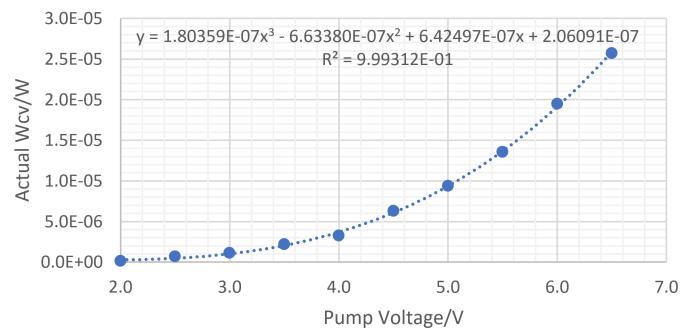


Fig. 9: Graph of Actual $|\dot{W}_{cv}|$ against Pump Voltage

The data collected indicates that when the $V_{pump} = 6V$ the actual \dot{W}_{cv} is constant at $-1.94653 \times 10^{-5} W$. The discrepancy between the actual and calculated values can be attributed to heat loss to the surroundings. This unaccounted heat loss could be attributed to:

- 1) Convection to air
- 2) Non-zero thermal mass of the system

REFERENCES

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- [6] Polystyrene (ps) typical properties generic ps (hips) — ul prospector. [Online]. Available: <https://plastics.ulprospector.com/generics/43/c/t/polystyrene-ps-properties-processing/sp/4>