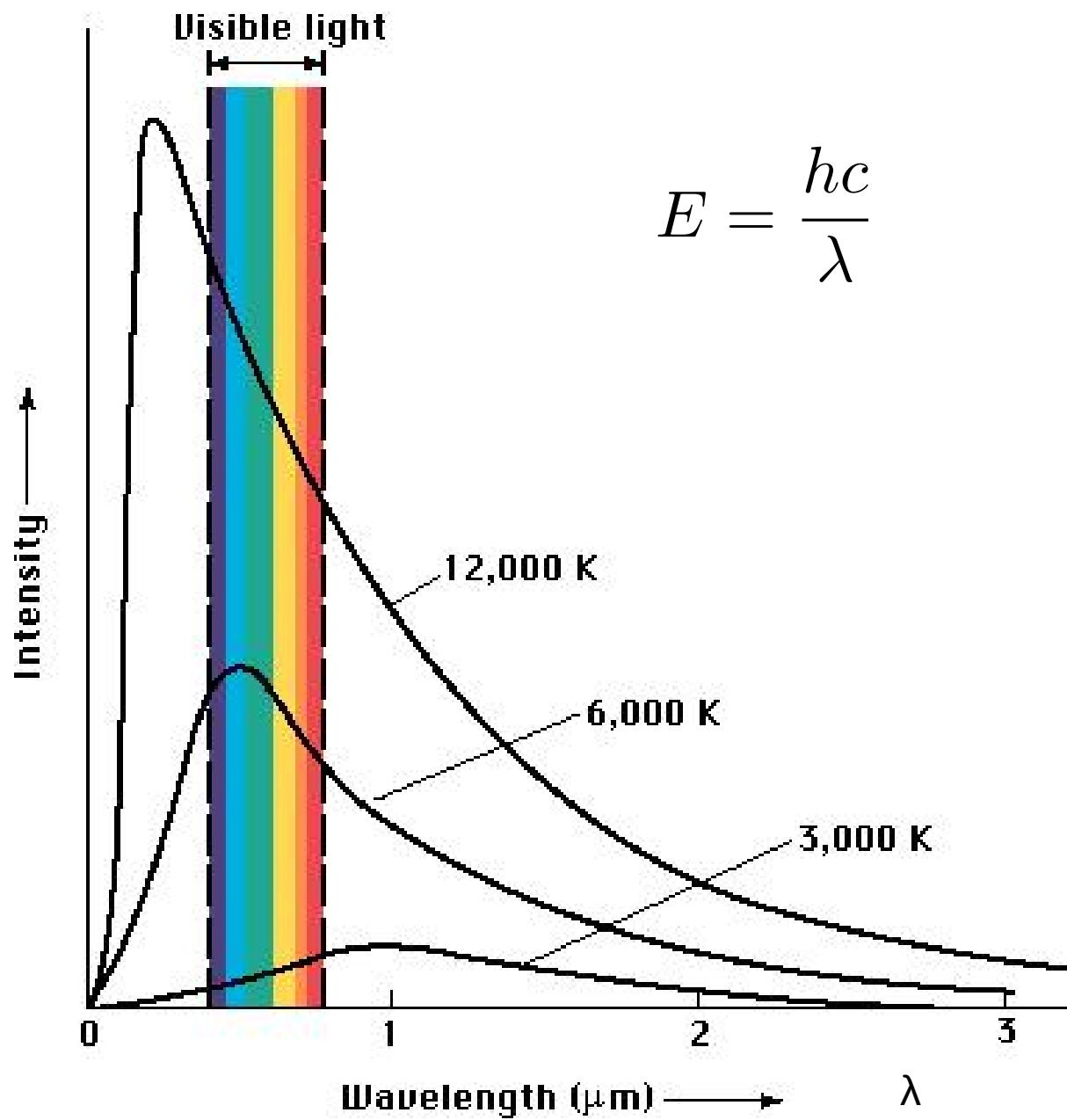
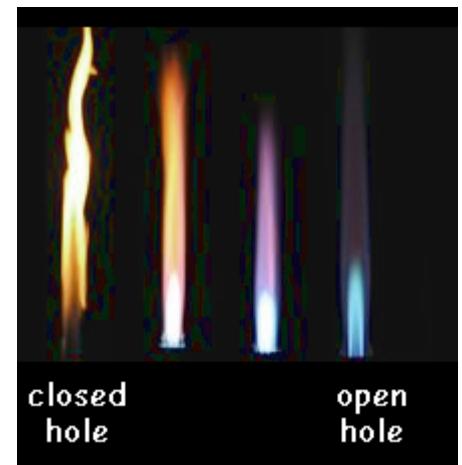


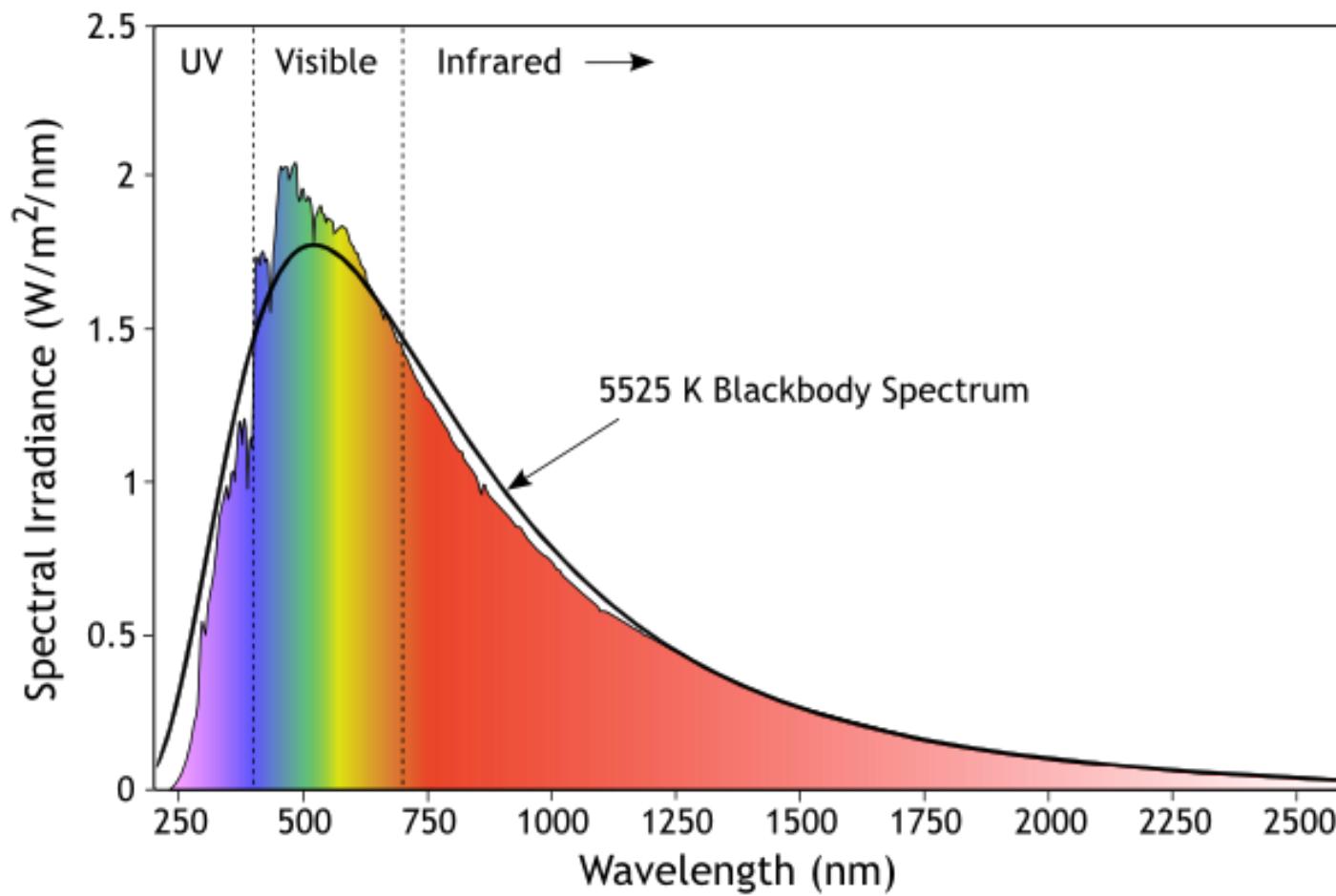
$$E = h c / \lambda$$



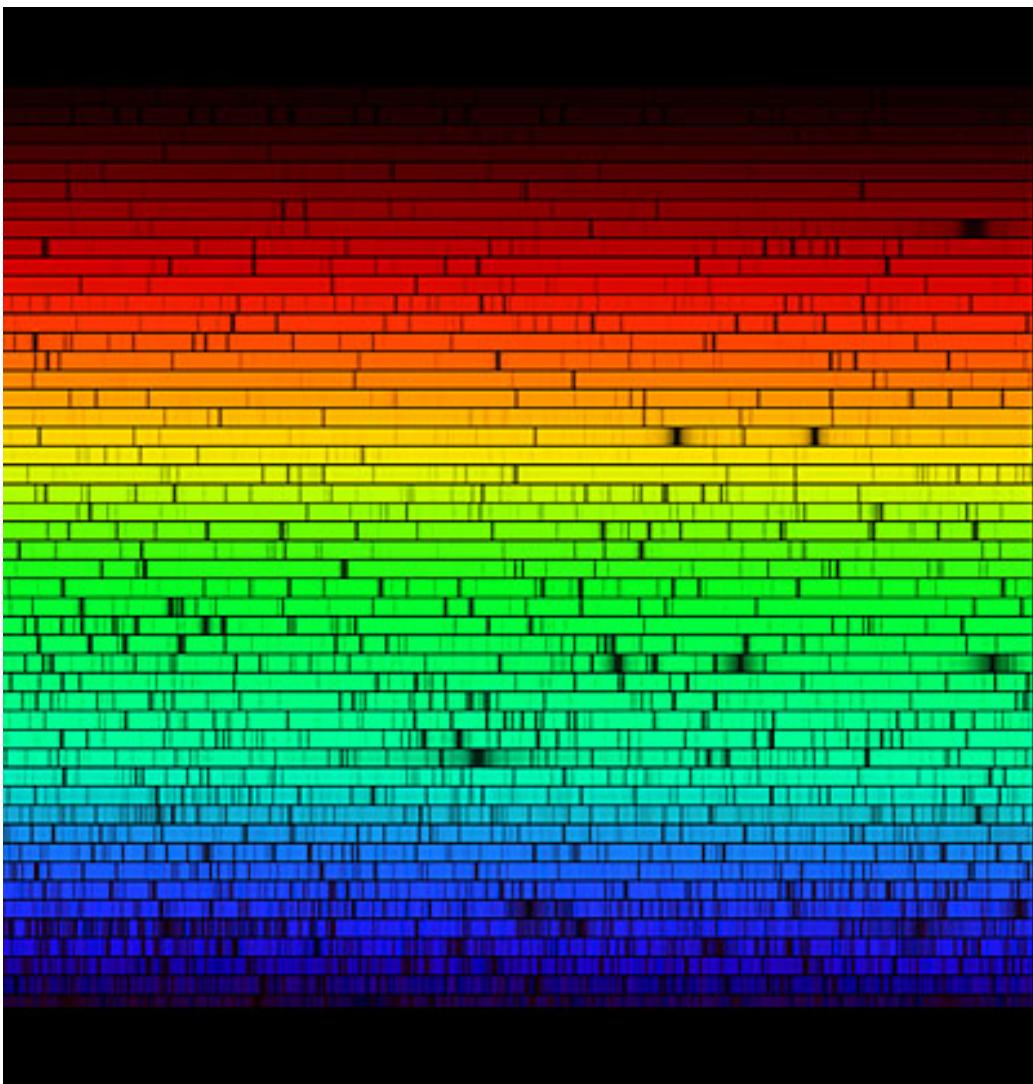
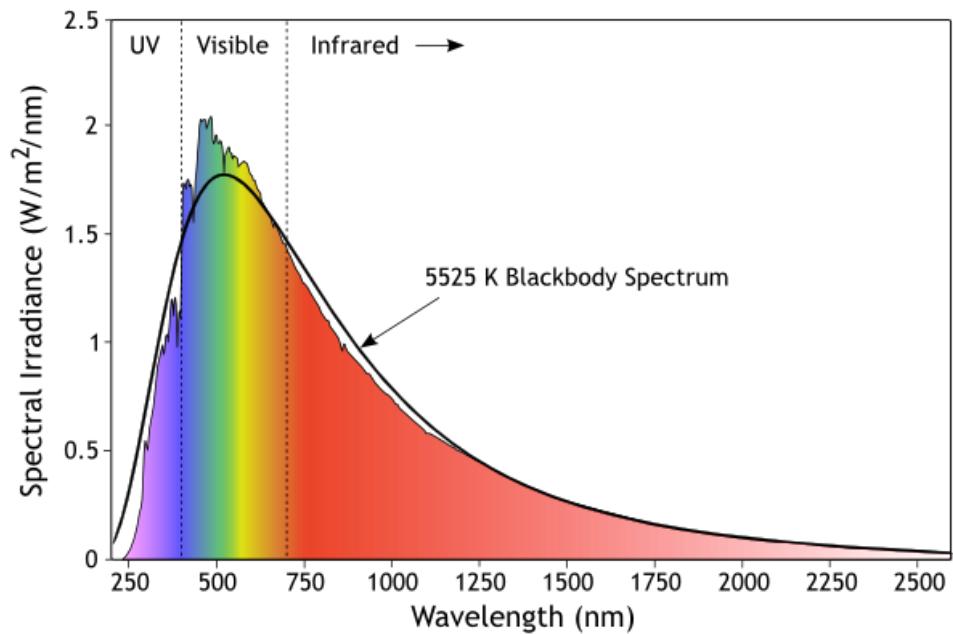
$$E = \frac{hc}{\lambda}$$



Espectro de Emissão do Sol

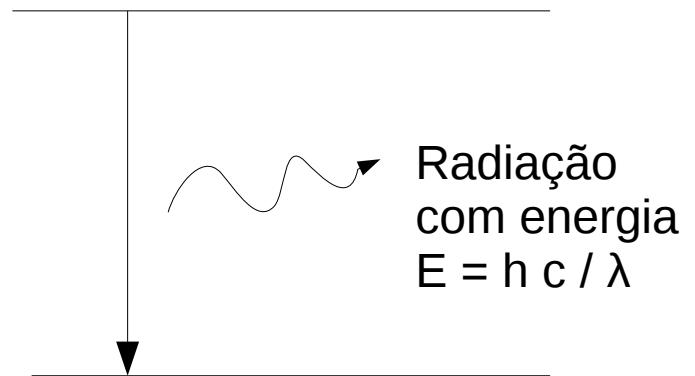
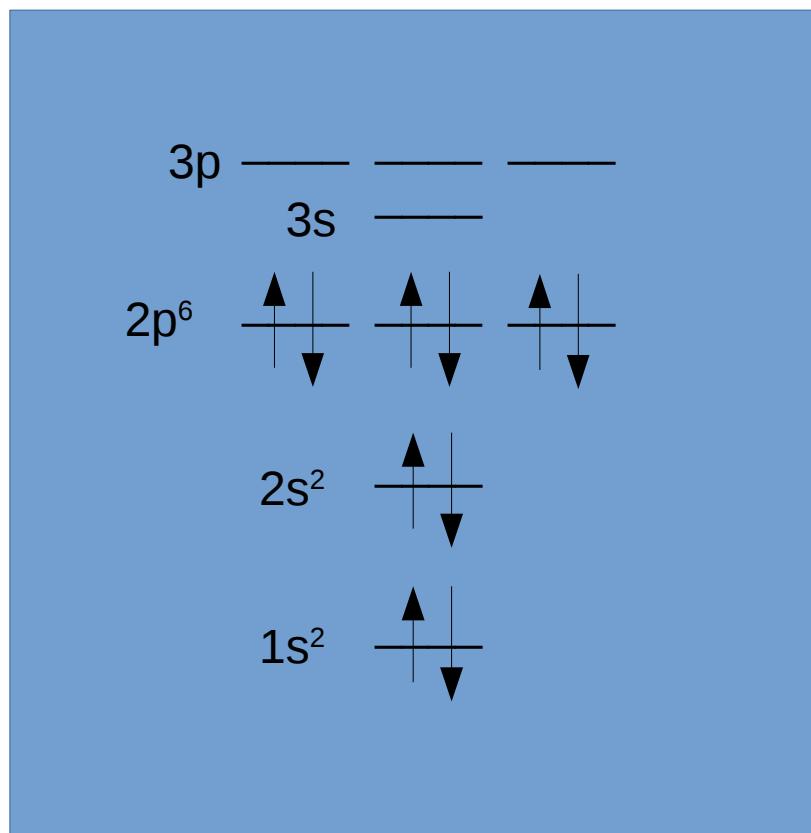


Espectro de Emissão do Sol



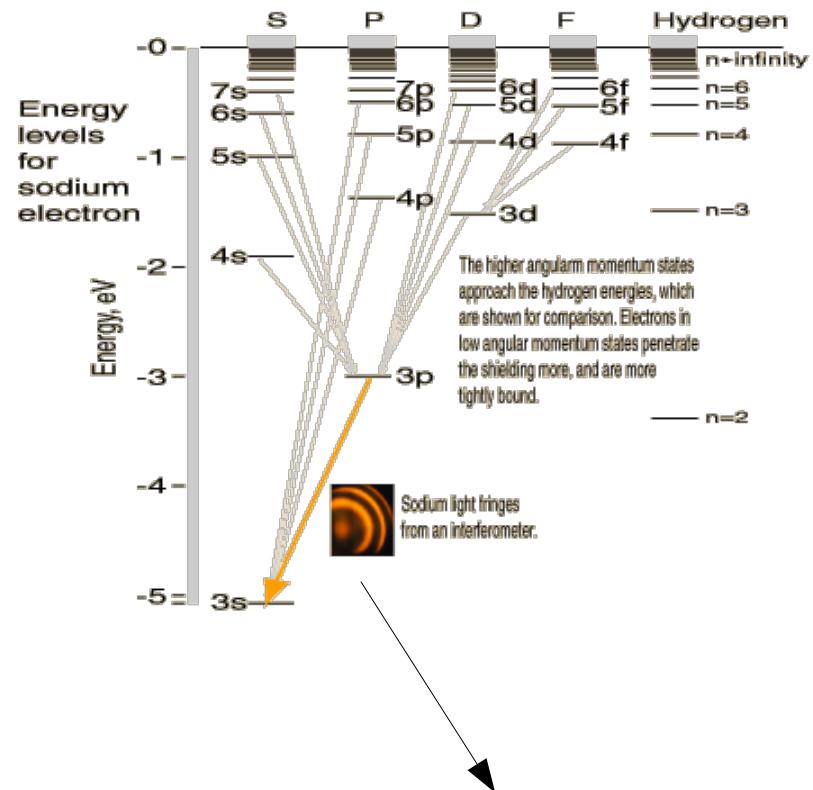
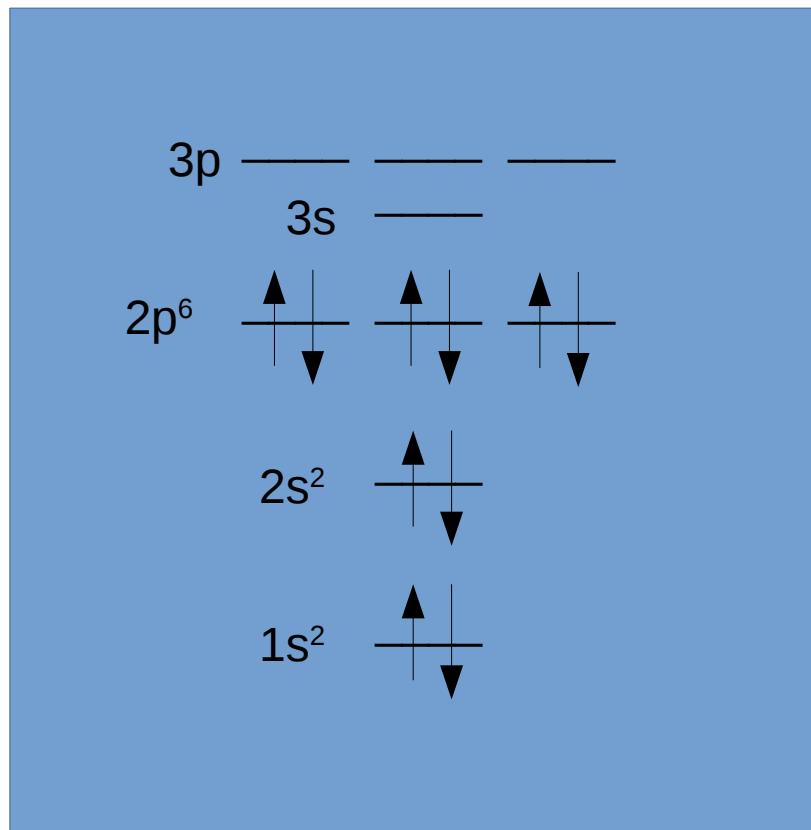
Íon de sódio: Na^+

Configuração eletrônica: $1s^2 2s^2 2p^6$

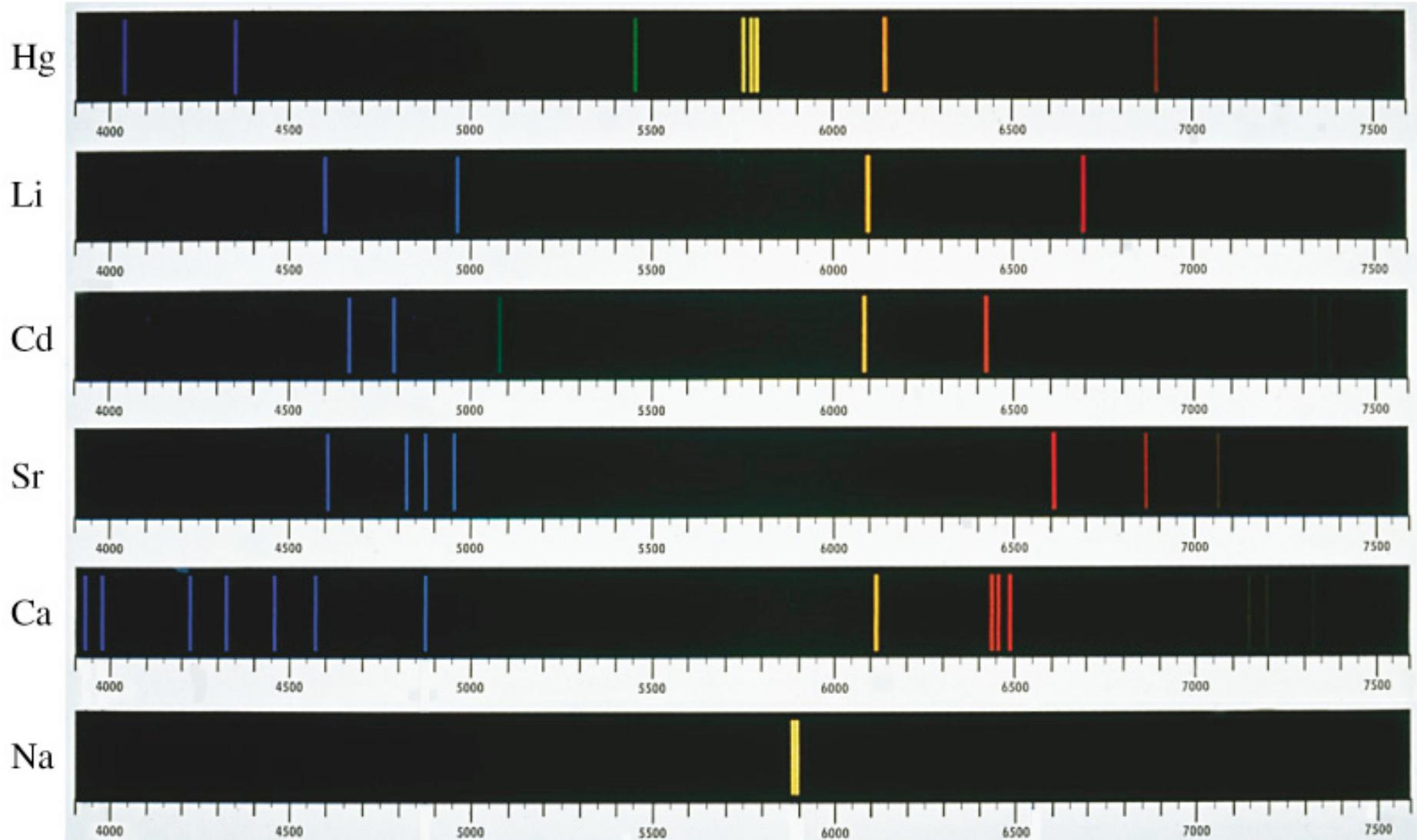


Íon de sódio: Na^+

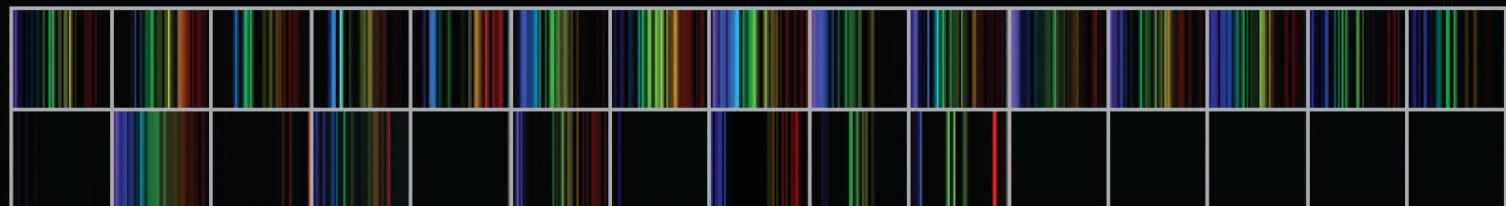
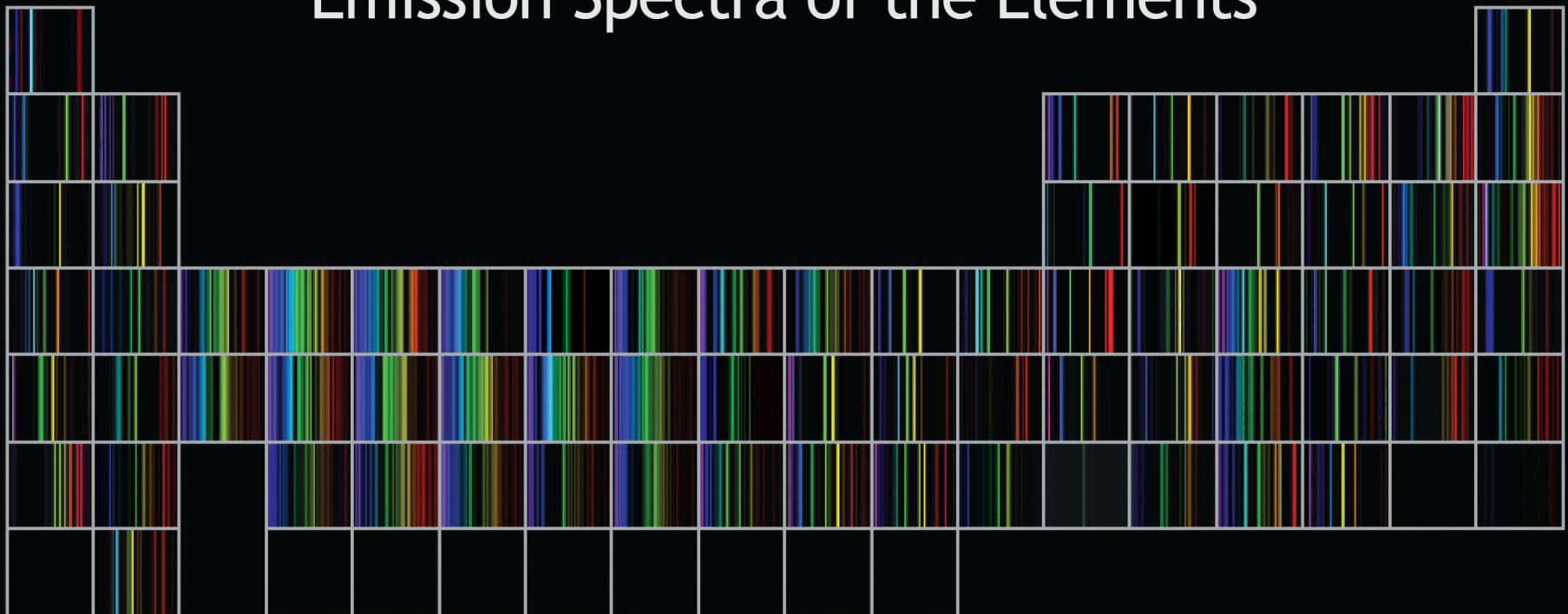
Configuração eletrônica: $1s^2 2s^2 2p^6$

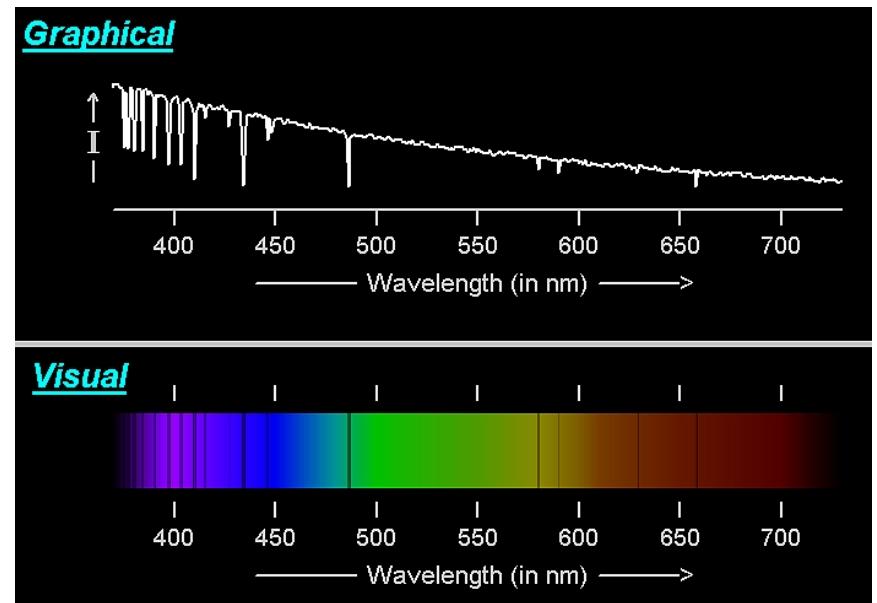
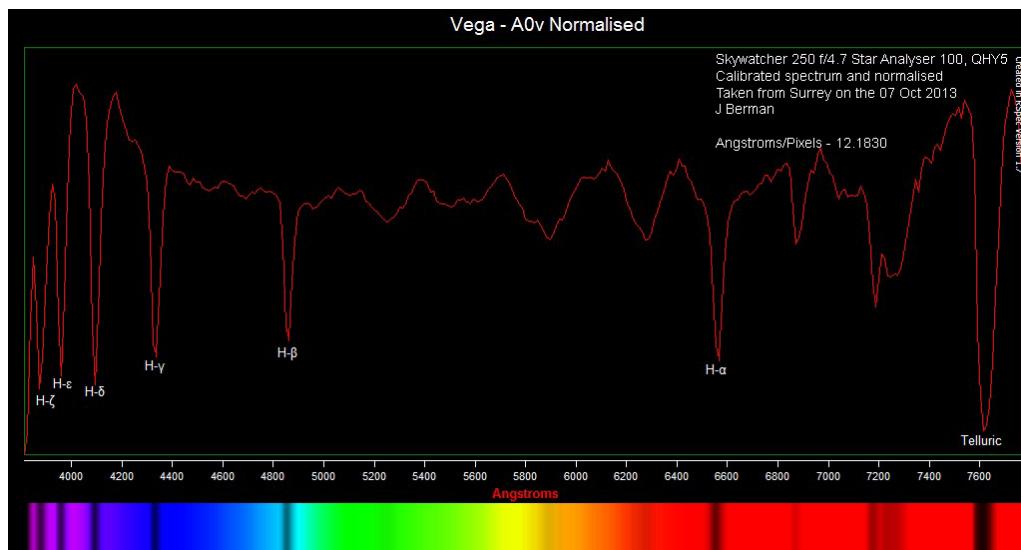
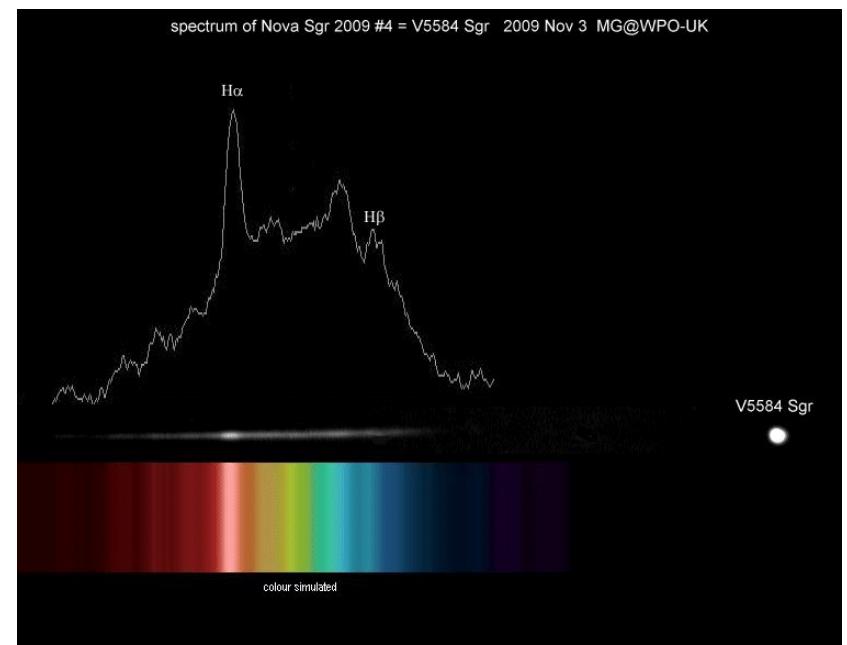
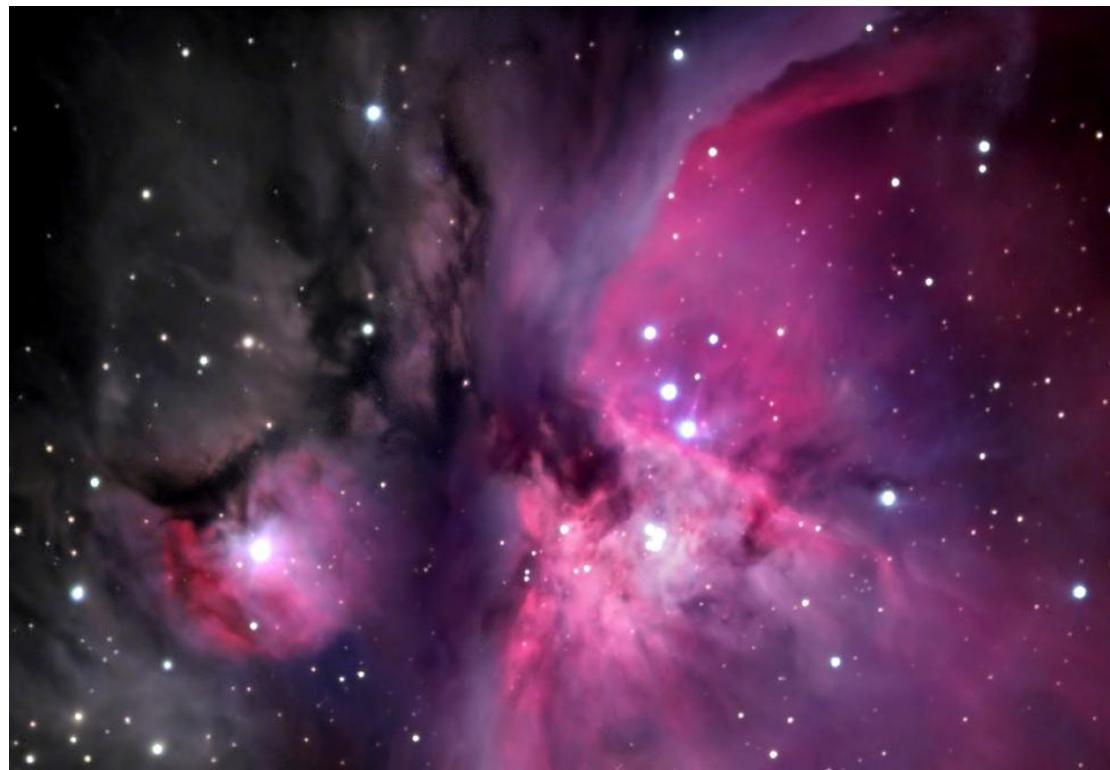


A intensa cor amarela está associada às transições transição $3p \rightarrow 3s$



Emission Spectra of the Elements





Espectrometria de absorção ou emissão atômica:

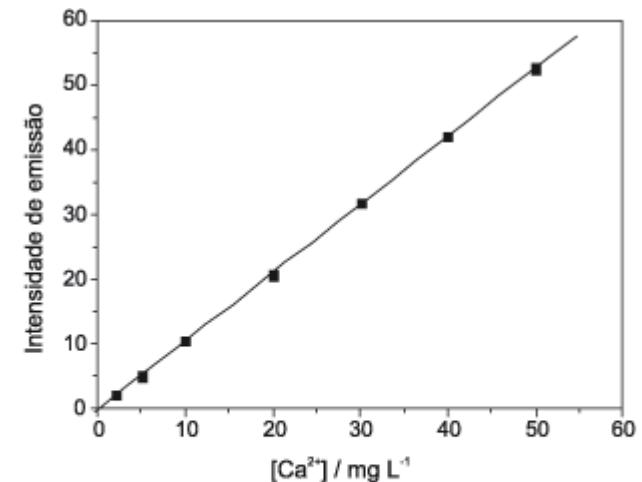
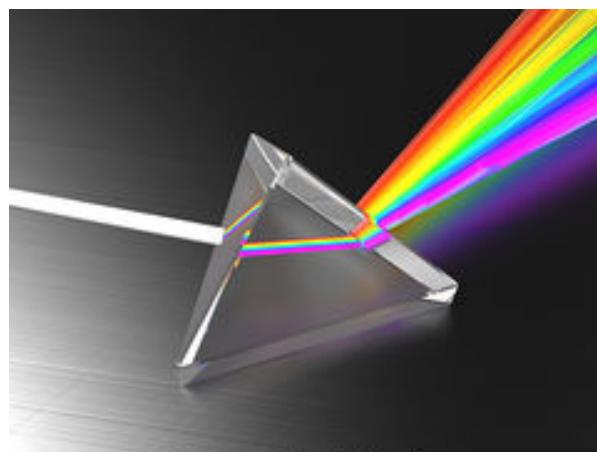


Figura 5. Curva analítica para cálcio, obtida usando o fotômetro de chama. Cada ponto é a média de cinco determinações. Neste intervalo de concentração obteve-se uma relação linear representada pela Equação 3, do texto



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