Ma10: Auger- and Electron Energy Loss Spectroscopy

Jörg Behrmann[*]   Anika Haller[†]

31.10.2011

Contents

---

[*]behrmann@physik.fu-berlin.de
[†]halleran@zedat.fu-berlin.de

1 Introduction

In 1922 Lise Meitner first described the effect that from 1923 on would carry Pierre Augers name. For many years the features of the Auger effect were a nuissance to X-ray spectroscopists, thought to not carry any useful informations. This view has since changed in the last halfcentennial and Auger spectroscopy has thus become a very useful source of information about surface effects, due to the very short mean free path of low-energy electrons in most materials. Thus these electrons cannot penetrate the material very deeply, which makes this spectroscopic technique very surface sensitive. Furthermore, the resulting spectra are specific to the materials analyzed, therefore we will use the Auger effect in the present experiment to measure impurities on the surface of a sample and to later acertain the material of our sample.

Even for energies that cannot lead to Auger excitation electrons do scatter inelastically. The energy and deflection angle of deflected electrons can then be used to gain information about the scatterer, e.g. phonons or plasmons, with which the electron has scattered. This closely related spectroscopic technique is called Electron Energy Loss Spectroscopy (EELS) and will be usedin the our experiment to analyze surface plasmons.

1.1 Universal Curve of Excape Depth

Figure 1 shows the universal curve of electron escape depth, which shows that that the depth from which an incoming electron can escape from a surface again is only dependent on the electron's energy and thus more or less independent of the material. This universal curve can be obtained on theoretical grounds by considering that first for increasing energy (up to $70\,\text{eV}$) more and more interaction modes, e.g. phonon modes, become available as potential scatterers to the incoming electron, which decreases the particles mean free path. Whereas for even higher energies the excess energy can be used to penetrate the material further.

1.2 Inner-Shell Excitations and the Auger Effect

A particle of high enough energy can scatter with an inner shell electron, exciting it enough that it leaves the atom and becomes free, leaving a hole. The atom is now in an unstable state and will relax by either by emitting an X-ray photon or an Auger electron, both possible relaxation processes are schon in 2.

In the X-ray relaxation process is relatively simple. An electron from a higher shell takes the place of the core electron and emits a photon in an energy range between $100\,\text{eV}$ and some MeV, depending on the element. The spectrum of this process is discrete and the remaining atom is singly ionized.

The second possible relaxation process is the emission of an Auger electron. In this relaxation process an electron from a higher shell fills the core hole and the emitted photon is absorbed by another electron in a higher shell that is thus excited in the continuum. The resulting atom is doubly ionized. A schematic for the process is shown
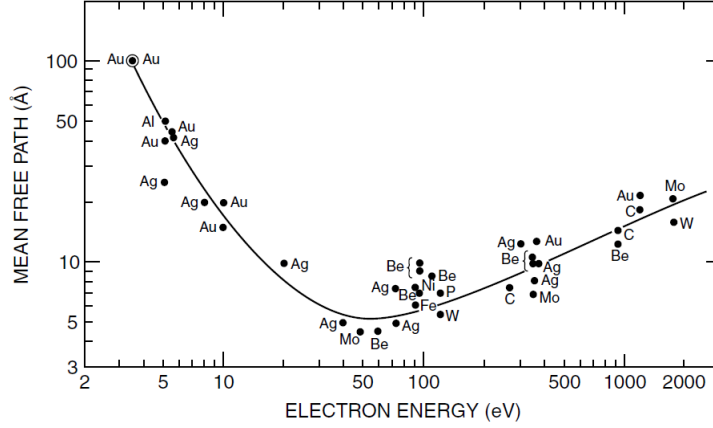
Figure 1: Universal curve of electron escape depth

in 2 using the example of the KL_1K_L3 process. The electron in the K-shell is excited into the continuum, an L_1 electron takes its place and another l_3 is emitted as a result.

The energy of the emitted Auger electron is given by

$$E_{kin} = E(K) - E(L) - E'(L') = E(K) - E(L) - E(L') - C(LL', T) + R \quad (1)$$

where the $E(K)$ is the binding energy of the core electron, $E(L)$ is the binding energy of the electron filling the core hole. $E'(L')$ is the effective binding energy of the electron that is emitted as Auger electron. The effective binding energy consists of the binding energy of the shell $E(L')$, a coupling energy between the holes at L and L' and the final state T and a so-called electronic relaxation R. It is important to note that the kinetic energy of the Auger electron is only dependent on the participating states of the Auger process and not on the energy of the particle that initially excites the core electron.

The dominant relaxation process is dependent on the atomic mass number Z, whereas for lighter atoms the relaxation by emission of an Auger electron dominates over X-ray emission.

1.3 Spin-Orbit Coupling

To fully describe an electron's energy in an atom one has to account for magnetic moments of the electron. Each electron has a magnetic moment due to its orbital movement in the atom and another one due to its spin. All of these magnetic moments of all electrons interact and leed to energy shifts in the spectra. For light atoms the spin-orbit interaction can be approximated by an interaction of the total angular momentum and the total spin (L-S-coupling), which is possible when the coupling of individual spins and orbital angular momenta is neglible. The part of the Hamiltonian
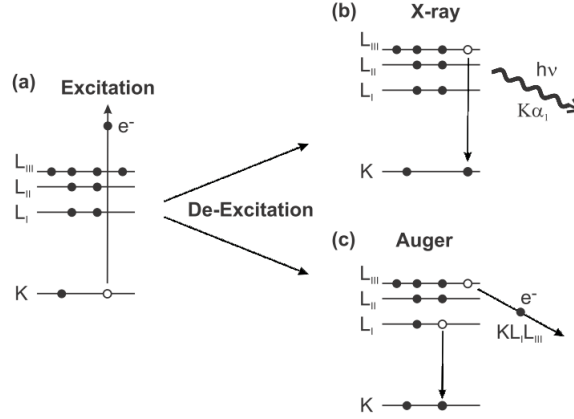
Figure 2: Auger KLM process and emission of an X-ray

describing spin-orbit coupling is then given by

$$H_{SL} = -\frac{\mu_B}{\hbar m_e e c^2} \frac{1}{r} \frac{\partial V(r)}{\partial r} \boldsymbol{L} \cdot \boldsymbol{S}, \tag{2}$$

where μ_B is the Bohr magneton, m_e is the electron mass, e is the elementary charge, c is the speed of light, V is the core potential and r is the distance of the electron to the nucleus.

For heavier atoms the coupling between individual spin and orbital momenta is nonneglible and a better approximation is the couplin of individual total angular moment $J_i = L_i + S_i$, this is called J-J-coupling.

Since Aluminum is a light element it is best described with L-S-coupling. The appropriate selection rules are thus $\Delta S = 0$ and $\Delta L = 0, \pm 1$.

1.4 X-Ray Notation

For historical reasons X-ray notation is used in describen Auger processes. Thus the principal quantum number n is represented by the letters K, L, M... for the numbers $1, 2, 3 ...$ and a subscript to those letters to associate a term symbol to those states. The states $S_{1/2}$ are given the subscript 1, $P_{1/2}$ is given the subscript 2, $P_{3/2}$ is given the subscript 3, and so on. Hence the state L_2 is equivalent to $2P_{1/2}$ in spectroscopic notation.

1.5 Plasmons

A plasmon is a quasi-particle coming up in condensed matter, describing quantized plasma oscillations, i.e. oscillations of the electron gas as a whole. The excitation energy of plasmons is specific to all materials and are given by the materials plasma
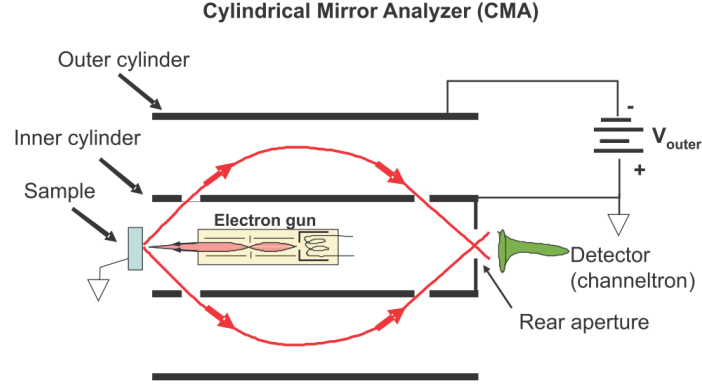
4

Figure 3: Cylindrical mirror analyzer

frequency. For a free electron gas the bulk plasma frequency is given by

$$\omega_b = \sqrt{\frac{e^2 n}{m_e \epsilon_0}}, \tag{3}$$

where e is the elementary charge, ϵ_0 the permitivty of free space, m_e the mass of the electron and n the conduction electron density.

Plasmon that appear on the interface between two materials and are confined to a surface are called surface plasmons. Because of their lower dimensionality the plasma frequence is lower than for the three-dimensional case, with $\omega_s = \omega_b/\sqrt{2}$. This excitation can be measured in the EELS part of present experiment as intensity peaks in the energy loss spectra at the energies $E = n\hbar\omega_s$ with $n \in \mathbb{N}$.

2 Experimental Setup

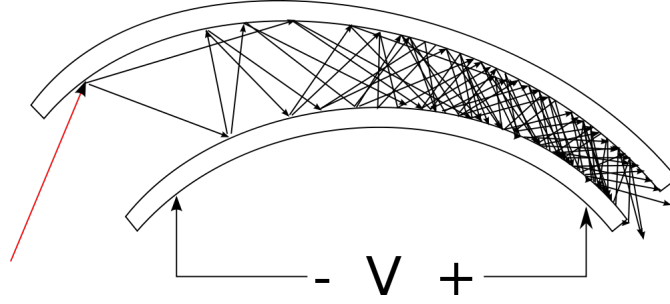In this section we will describe shortly the machinery and techniques used in this experiment.

Figure 4: Channeltron

2.1 Ultra High Vaccuum (UHV)

2.2 Cylindrical Mirror Analyzer

2.3 Channeltron

2.4 Lock-In Amplifier

3 Data Analysis

In this section we will analyze the measured data. We will first look at the cleaning of the sample, will then look at the excitation of bulk and surface plasmons, obtained by Electron Energy Loss spectroscopy, and end this section with an analysis of the Auger spectrum of aluminum.

If not mentioned otherwise we use Gaussian error propagation and least square algorithms for fitting. Also, all spectra were done with an energy step size of $1\,\mathrm{eV}$ and by averaging over 3 measurements.

The errors given for all obtained values are obtained from the uncertainty in fit parameters, these uncertainties usually understate the actual errors, but are used for consistency, since they are the only easily quantifiable errors, because errors and noise in our measurements are nearly eliminated by the averaging process.

3.1 Sample Cleaning

Figure 5 shows the part of the Auger spectrum of the aluminum sample, where peaks resulting from impurities can be found. Using [5] we can identify oxygen, nitrogen, carbon and possibly chlorine as impurities. Nitrogen and oxygen as main parts of air are not surprising but carbon and chlorine are, since they are not part of air.

Having filed on the sample for over half an hour and seeing peaks almost vanish and then reappear shortly thereafter, we think that these peaks result from filing some parts of the sample holder, since it is quite easy to miss the sample with the file.

Furthermore we have to conclude, that the cleaning of the sample is rather useless, since the pressure in the chamber ($p = 4.3 \cdot 10^{-8}$) is too high for the proper conductance
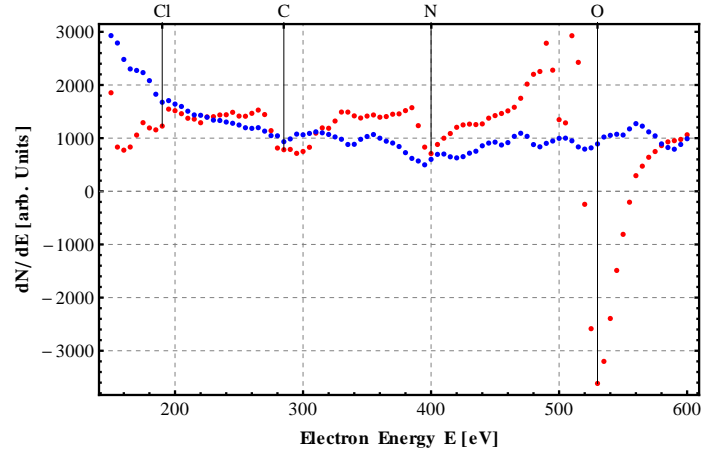
6

Figure 5: Exemplary differential Auger spectra before (red) and after (blue) cleaning.

of the experiment, since after two minutes, this leads to the adsorption of nearly four monolayers of oxygen:

$$4.3 \cdot 10^{-8} \, \text{mbar} \times 120 \, \text{s} = 3.87 \, \text{L}, \tag{4}$$

which we could verify in so far, as the state of the oxygen peak at the end of the cleaning process (blue in figure 5) had changed back to state before the cleaning (red in figure 5).

Neverthless the experiment could be conducted and lead to qualitatively good results, as will be shown in the following section, but we see the described problems as a major error source.

3.2 Plasmons

The Electron Energy Loss Spectroscopy experiments were done after taking the Auger spectrum of the sample.

We think, that the excitation of bulk plasmons is needed to understand the results of the Auger spectrum, which is why we will discuss the plasmon excitations first.

The electron gun energy was adjusted to $1 \, \text{kV}$ for the spectroscopy of the bulk plasmons and to $0.2 \, \text{kV}$ for the surface plasmon spectroscopy. Both energies are too low to excite Auger electrons, so that the Auger effect can be excluded for the measured phenomena and thez have to be explained by plasmon excitation.

3.2.1 Bulk Plasmons

Figure 6 shows the differential and the integrated spectrum of the EELS of bulk plasmon excitation.
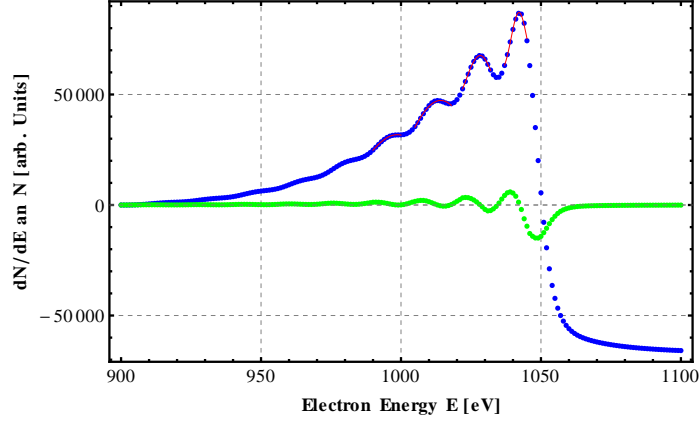
7

Figure 6: Integrated (blue) and differential (green) electron energy loss spectrum for the bulk plasmon measurement. Fits for peak positions are plotted on top of the integrated spectrum in red.

Peak Position E [eV]
998.66 ± 0.17
1014.17 ± 0.19
1028.27 ± 0.02
1042.15 ± 0.06

Table 1: Peak positions of bulk plasmons.

To obtain the bulk plasmon excitation energy we fitted the first four peaks from the right with gaussians, since they were the only peaks, that exhibited a definite shape, i.e. had a discernible midpoint. The peak positions can be found in table 1.

Using the gaussian midpoint positions, we then obtained the following value for the excitation energy by linear regression:

$$\Delta E_{bulk} = (14.19 \pm 0.05)\,\text{eV}. \tag{5}$$

This value does not agree with the literature value of $15.3\,\text{eV}$ [4], but is at least qualitatively in the vicinity of the literature value. It is possible that this is due to the contamination of the surface with oxygen, although this seems unlikely for the excitation of bulk plasmons. Another possible source of error can be in our energy detector. This will be discussed later in the section on the Auger spectrum.
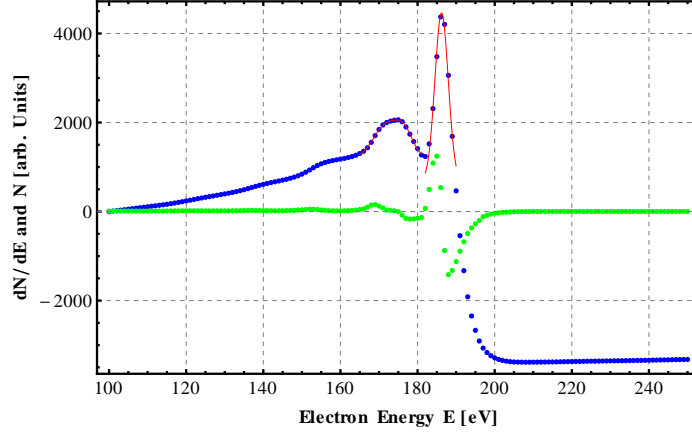
Figure 7: Integrated (blue) and differential (green) electron energy loss spectrum for the surface plasmon measurement. Fits for peak positions are plotted on top of the integrated spectrum in red.

3.3 Surface Plasmons

The differential and integrated spectrum for the surface plasmon EELS can be found in figure 7.

We can see, that the integrated spectrum does not exhibit many features, which is why we were only able to succefully fit two peaks: The well-pronounced major peak on the right, which is due to, more or less, elastically backscattered electrons, and the misshapen peak to its right. Further analysis showed, that this peak is a superposition of two peaks, one describing the excitation of the surface plasmon and one describing the excitation of a bulk plasmon.

All three peaks were fitted with gaussians, the respective peak positions can be found in table 2. The resulting excitation energies are:

$$\Delta E_{bulk} = (14.3 \pm 0.5)\,\text{eV}, \tag{6}$$
$$\Delta E_{surf} = (\;9.6 \pm 0.3)\,\text{eV}. \tag{7}$$

Both numbers are not consistent with the excitation energies of $15.3\,\text{eV}$ and $10.8\,\text{eV}$ for bulk an surface plasmons from [4], but the bulk plasmon excitation energy is consistent with the one obtained earlier and both are in the vicinity of the literature value.

The same possible sources as for the bulk plasmon EELS apply, but surface contamination is naturally a more probable error source in this case.

3.4 Auger Spectrum

The Auger spectrum of the sample was taken with an electron gun energy of $4.8\,\text{eV}$. The interesting part of the differential spectrum can be found in figure 8.

9

Peak Position E [eV]
171.95 ± 0.40
176.67 ± 0.25
186.23 ± 0.12

Table 2: Peak positions of surface plasmons.

We fitted all minima and maxima of the differential spectrum with gaussians to obtain their position in the spectrum. The appropriate fit paramaters can be found in table 3.

The distribution of the minima, agrees qualitatively very well, with the literature values form [5], but they seem to be shifted by a constant additive factor to higher energies. If we artificially put the last minimum on its literature value and use this offset to recalibrate the spectrum, all minima agree remarkably well, with their literature values.

Figure 9 shows the integrated Auger spectrum of our sample, which should be used to obtain the peaks for the KLL transitions in alumnium. Unfortunately this spectrum is unusable, since only the two rightmost peaks are pronounced enough to be discernible.To solve this problem we took the maxima and minima positions in figure 8 and took the average between a minimum and its preceding maximum, which gives a rough estimate for the position of a peak in figure 9. The estimated peak positions can be found in the lower section of table 3.

The estimated peak positions in the integrated Auger spectrum need to be identified with a Auger KLL transition from table 4. Possible identifications, taking possible energy losses due to plasmon excitation into account, can be found in table 5, although for some measure peaks several transitions come into consideration, but cannot be discerned. This may be due to two neighbouring peaks being combined to one peak of average energy.

4 Conclusion

We performed Auger and Electron Energy Loss Spectroscopy on an alumminum sample.

References

[1] G. Dufour, J.M. Mariot, P.-E. Nilsson-Jatko, and R. Karnatak. K-ll auger spectrum of aluminium. *Physica Scripta*, 13:370, 1976.

[2] Christian Fräßdorf, Simon Jaekel, and Johannes Mosig. Auger and eel spectroscopy - lab report, 2011.
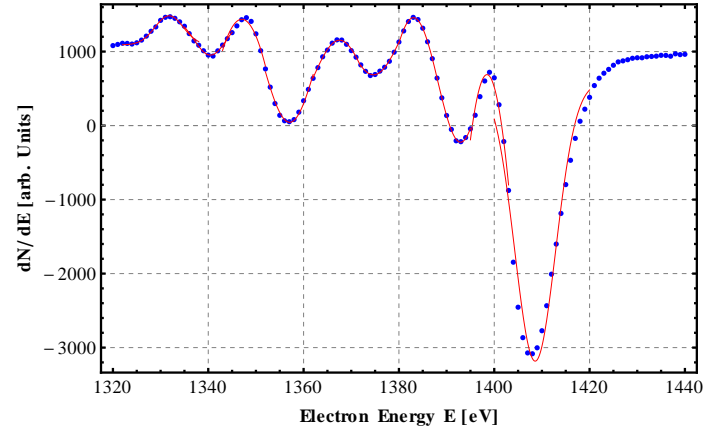
Figure 8: Differential Auger spectrum for aluminum. Fits for the maxima and minima are plotted on top of the spectrum in red.
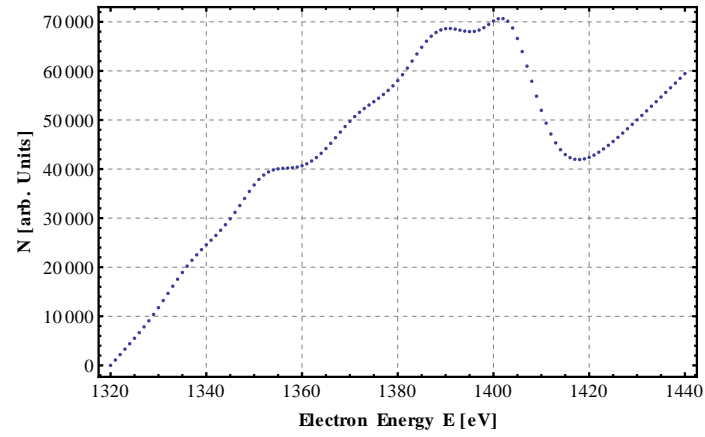

Figure 9: Integrated Auger spectrum for aluminum.

11

peak position [eV]	recalibrated peak position [eV]	literature value [eV]
1340.45 ± 0.07	1327.86 ± 0.07	1329
1357.08 ± 0.07	1344.49 ± 0.07	1345
1374.66 ± 0.08	1362.07 ± 0.08	1364
1392.93 ± 0.04	1380.34 ± 0.04	1380
1408.6 ± 0.2	1396.0 ± 0.2	1396
1331.8 ± 0.1	1319.2 ± 0.1	
1347.0 ± 0.2	1334.5 ± 0.2	
1367.54 ± 0.05	1354.95 ± 0.0	
1383.05 ± 0.03	1370.46 ± 0.0	
1398.5 ± 0.2	1385.89 ± 0.2	
1336.1 ± 0.2	1323.5 ± 0.2	
1352.1 ± 0.2	1339.5 ± 0.2	
1371.1 ± 0.2	1358.5 ± 0.2	
1388.0 ± 0.2	1375.4 ± 0.2	
1403.5 ± 0.2	1390.9 ± 0.2	

Table 3: Upper part: minimum positions of aluminum auger spectrum. Recalibrated peak positions shifted that last peak coincides with literature value. Literature values taken from [5]. Middle part: maximum positions of aluminum auger spectrum. Lower part: means of minimum and preceding maximum as approximation for peak position in integrated spectrum.

Auger Line	Energy E [eV]	$+B_1$	$+B_2$
$KL_1L_{2,3}(^1P)$	1342.3	1328.1	1313.9
$KL_1L_{2,3}(^3P)$	1358.5	1344.3	1330.1
$KL_{2,3}L_{2,3}(^1S)$	1386.6	1372.4	1358.2
$KL_{2,3}L_{2,3}(^1D)$	1393.5	1379.2	1365.0
$KL_{2,3}L_{2,3}(^3P)$	1398.1	1383.9	1369.7

Table 4: Auger peak energies from [1] with corrections for subsequent excitation of up to two bulk plasmons.

Auger Line		Theoretical Energy [eV]	Measured Energy [eV]
$KL_1L_{2,3}:$	$^1P + B_1$	1328.1	1323.5 ± 0.2
$KL_1L_{2,3}:$	^1P	1342.3	1339.5 ± 0.2
$KL_1L_{2,3}:$	^3P	1358.5	1358.5 ± 0.2
$KL_{2,3}L_{2,3}:$	$^1S + B_2$	1358.2	1358.5 ± 0.2
$KL_{2,3}L_{2,3}:$	$^1S + B_1$	1372.4	1375.4 ± 0.2
$KL_{2,3}L_{2,3}:$	$^1D + B_1$	1379.4	1375.4 ± 0.2
$KL_{2,3}L_{2,3}:$	^1S	1386.6	1390.9 ± 0.2
$KL_{2,3}L_{2,3}:$	^1D	1393.5	1390.9 ± 0.2

Table 5: Auger peak energies from [1] with corrections for subsequent excitation of up to two bulk plasmons.

[3] Benjamin Huber and Carolin Wille. Auger electron and electron energy loss spectroscopy - lab report, 2011.

[4] D. Massignon, F. Pellerin, J. M. Fontaine, C. Le Gressus, and T. Ichinokawa. Comparision of the secondary-electron spectrum with the electron-loss spectrum on pure al by low-energy electron-reflection spectroscopy. *J. Appl. Phys.*, 51:808, 1980.

[5] P.W. Palmberg, G.E. Riach, and R.E. Weber. *Handbook of Auger electron spectroscopy data.* Physical Electronics Industries Inc, 1972.

[6] unknown. Ma 10: Auger- and electron energy loss spectroscopy - skript, November 2011.