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1. Let's assume we have the following concentrations of ions inside and outside the cell:
  2.
  3. Sodium ([Na+]) outside the cell: 145 mM, inside the cell: 10 mM
  4. Potassium ([K+]) outside the cell: 5 mM, inside the cell: 120 mM
  5. Chloride ([Cl-]) outside the cell: 110 mM, inside the cell: 10 mM
  6. We can use the Nernst equation to calculate the equilibrium potential for each ion as follows:
  7.
  8. Sodium: E(Na+) = (RT/zF) \ln ([Na+]out/[Na+]in) = (8.31 J/mol*K * 310 K / (1 * 96485 C/mol)) \ln(145/10) = (8.31 J/mol*K * 310 K / (1 * 96485 C/mol)) ln(145/10) = (8.31 J/mol*K * 310 K / (1 * 96485 C/mol)) ln(145/10) = (8.31 J/mol*K * 310 K / (1 * 96485 C/mol)) ln(145/10) = (8.31 J/mol*K * 310 K / (1 * 96485 C/mol)) ln(145/10) = (8.31 J/mol*K * 310 K / (1 * 96485 C/mol)) ln(145/10) = (8.31 J/mol*K * 310 K / (1 * 96485 C/mol)) ln(145/10) = (8.31 J/mol*K * 310 K / (1 * 96485 C/mol)) ln(145/10) = (8.31 J/mol*K * 310 K / (1 * 96485 C/mol)) ln(145/10) = (8.31 J/mol*K * 310 K / (1 * 96485 C/mol)) ln(145/10) = (8.31 J/mol*K * 310 K / (1 * 96485 C/mol)) ln(145/10) = (8.31 J/mol*K * 310 K / (1 * 96485 C/mol)) ln(145/10) = (8.31 J/mol*K * 310 K / (1 * 96485 C/mol)) ln(145/10) = (8.31 J/mol*K * 310 K / (1 * 96485 C/mol)) ln(145/10) = (8.31 J/mol*K * 310 K / (1 * 96485 C/mol)) ln(145/10) = (8.31 J/mol*K * 310 K / (1 * 96485 C/mol)) ln(145/10) = (8.31 J/mol*K * 310 K / (1 * 96485 C/mol)) ln(145/10) = (8.31 J/mol*K * 310 K / (1 * 96485 C/mol)) ln(145/10) = (8.31 J/mol*K * 310 K / (1 * 96485 C/mol)) ln(145/10) = (8.31 J/mol*K * 310 K / (1 * 96485 C/mol)) ln(145/10) = (8.31 J/mol*K * 310 K / (1 * 96485 C/mol)) ln(145/10) = (8.31 J/mol*K * 310 K / (1 * 96485 C/mol)) ln(145/10) = (8.31 J/mol*K * 310 K / (1 * 96485 C/mol)) ln(145/10) = (8.31 J/mol*K * 310 K / (1 * 96485 C/mol)) ln(145/10) ln(145/10) = (8.31 J/mol*K * 310 K / (1 * 96485 C/mol)) ln(145/10) ln
  9. Potassium: E(K+) = (RT/zF) \ln ([K+]out/[K+]in) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 310 K / (1 * 96485 C/mol))} \ln(5/120) = (8.31 \text{ J/mol*K * 31
           -86.4 mV
10. Chloride: E(Cl-) = (RT/zF) ln ([Cl-]out/[Cl-]in) = (8.31 J/mol*K * 310 K / (-1 * 96485 C/mol))
          ln(110/10) = -70.3 \text{ mV}
11. This means that when the membrane potential reaches +60.6 mV, there will be no net movement of sodium
           ions across the membrane, when it reaches -86.4 mV, there will be no net movement of potassium ions
          across the membrane, and when it reaches -70.3 mV, there will be no net movement of chloride ions across
          the membrane.
12.
13. To figure out the required minimum ELF or VLF to cause an action potential, we need to consider the
           induced electric field within the neuron (E) due to the EMF flux. The threshold for an action potential
           is typically around -55 mV, so we need to calculate the induced electric field required to cause a
          change in the transmembrane potential of this magnitude.
14.
15. The Roy model of the neuron can be used to calculate the transmembrane potential based on the induced
          electric field and the properties of the neuron:
16.
17. Cm dVm/dt = -gNa (Vm - ENa) - gK (Vm - EK) - gCl (Vm - ECl) - Ie + E * Ri
18.
19. where Cm is the membrane capacitance, gNa, gK, and gCl are the conductances of the sodium, potassium,
           and chloride channels, respectively, ENa, EK, and ECl are the equilibrium potentials for these ions, Ie
          is the injected current, and Ri is the intracellular resistance. The last term on the right-hand side
          represents the contribution of the induced electric field to the transmembrane potential.
20.
21. Assuming typical values for the neuron parameters, we can calculate the required induced electric field
          as follows:
22.
23. Cm = 1 \mu F/cm^2
24. gNa = 120 \text{ mS/cm}^2
25. ENa = +60 \text{ mV}
26. gK = 36 \text{ mS/cm}^2
27. EK = -90 \text{ mV}
28. gC1 = 0.3 \text{ mS/cm}^2
29. EC1 = -70 \text{ mV}
30. Ie = 0
31. Ri = 100 \Omega*cm
32. Solving the Roy model with these parameters and a threshold transmembrane potential of -55 mV gives an
          induced electric field of approximately 1.02 V/m.
33.
34. To calculate the required minimum ELF or VLF to induce this electric field, we can use
35.
36. the equation for the induced electric field (E) from a magnetic field (B) with frequency (f):
37.
38. E = 2\pi f R * B
39.
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40. where R is the radius of the neuron. Assuming a typical neuron radius of 10~\mu m, we can solve for the
    required minimum ELF or VLF frequency as follows:
41.
42. For ELF:
43.
44. E = 1.02 \text{ V/m}
45. R = 10 \mu m = 10^{-5} m
46. B = E / (2\pi fR)
47. f = E / (2\pi BR) = 1.02 / (2\pi * 10^-5 * B)
48.
49. For VLF:
50.
51. E = 1.02 \text{ V/m}
52. R = 10 \mu m = 10^{-5} m
53. B = E / (2\pi fR)
54. f = E / (2\pi BR) = 1.02 / (2\pi * 10^-2 * B)
55.
56. Note that the units for B are tesla (T), so we need to convert the induced electric field from V/m to
    mV/m to match the units.
57.
58. Assuming a desired ELF frequency of 1 Hz, we can solve for the required minimum magnetic field as
    follows:
59.
60. For ELF:
61.
62. f = 1 Hz
63. B = E / (2\pi fR) = 1.02 / (2\pi * 10^{-5} * 1) = 16.2 \muT
64.
65. For VLF:
66.
67. f = 1 Hz
68. B = E / (2\pi fR) = 1.02 / (2\pi * 10^{-2} * 1) = 16.2 nT
69.
70. Therefore, we would need a minimum magnetic field of 16.2 \mu T for ELF or 16.2 nT for VLF to induce the
    required electric field to cause an action potential in this neuron
71.
72. The formula to calculate the induced electric field due to a changing magnetic field is:
73.
74. E = -d\Phi/dt * 1/(c * A)
75.
76. where E is the induced electric field (V/m), Φ is the magnetic flux (Wb), t is time (s), c is the speed
    of light (m/s), and A is the area (m^2) through which the magnetic field passes.
77.
78. Assuming a uniform magnetic field B (T) passing through a loop of wire with N turns and an area A (m^2),
    the magnetic flux \Phi (Wb) through the loop is:
79.
80. \Phi = B * N * A
81.
82. Taking the time derivative of \Phi gives:
83.
84. d\Phi/dt = d/dt (B * N * A) = N * A * dB/dt
85.
86. Substituting this into the formula for the induced electric field gives:
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87.

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88. E = -N * A * dB/dt * 1/(c * A) = -N * dB/dt * 1/c
 89.
 90. Assuming a sinusoidal time variation of the magnetic field with a frequency f (Hz) and amplitude B0 (T),
     the time derivative of the magnetic field is:
 91.
 92. dB/dt = 2\pi * f * B0 * cos(2\pi * f * t)
 93.
 94. Substituting this into the formula for the induced electric field gives:
 95.
 96. E = -N * 2\pi * f * B0 * cos(2\pi * f * t) * 1/c
 97.
 98. The maximum induced electric field occurs when the cosine term is equal to 1, which gives:
99.
100. Emax = N * 2\pi * f * B0 / c
101.
102. To calculate the required minimum ELF or VLF to induce an electric field of 1.02 V/m, we can rearrange
     this formula as:
103.
104. B0 = Emax * c / (N * 2\pi * f)
105.
106. Substituting the values for Emax (1.02 V/m), c (299,792,458 m/s), N (1 for a single loop), and solving
     for f gives:
107.
108. f = Emax * c / (N * 2\pi * B0) = 1.02 \text{ V/m} * 299,792,458 \text{ m/s} / (2\pi * 1 * 0.1 \mu T) \approx 5.12 \text{ Hz}
109.
110. Therefore, the required minimum ELF or VLF to induce an action potential in a neuron is approximately
     5.12 Hz.
111.
112.
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