

Q	Part	Sub Part	Marking Guidance	Mark	Comments
6	(a)		$\Delta H = \Sigma \Delta H_f(\text{products}) - \Sigma \Delta H_f(\text{reactants})$ $= -201 - 242 - (-394)$ $= -49 \text{ kJ mol}^{-1}$	1 1 1	+49 kJ mol <sup>-1</sup> = 1 mark units not required, wrong units lose 1 mark
6	(b)		$\Delta S = \Sigma S(\text{products}) - \Sigma S(\text{reactants})$ $= 238 + 189 - (214 + 3 \times 131)$ $= -180 \text{ J K}^{-1} \text{ mol}^{-1}$	1 1 1	+180 = 1 mark units not required, wrong units lose 1 mark
6	(c)		$\Delta G = \Delta H - T\Delta S$  $(\Delta S \text{ is negative so) at high temp } -T\Delta S \text{ (is positive and) greater than } \Delta H / \text{large}$  So $\Delta G > 0$  $(\text{Limiting condition } \Delta G = 0 \text{ so) } T = \Delta H / \Delta S$  $= 272 \text{ K}$  Reaction is too slow at this temperature/to speed up the reaction	1 1 1 1 1 1	If use $G$ not $\Delta G$ penalise M1 but not M2 and M3  Do not award M2 or M3 if positive $\Delta S$ value used  Independent mark unless positive $\Delta S$ value used  Allow 297-298 if used given values. Do not award M5 if $T$ -ve or if M4 should give $T$ -ve

6	(d)	<p><math>\text{CH}_3\text{OH} + 3/2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}</math></p> <p>2.5 mol give 3 mol (gases)</p> <p>Therefore <math>\Delta S</math> is positive/entropy increases</p> <p>( combustion exothermic so <math>\Delta H</math> -ve so <math>\Delta H - T\Delta S</math>) and hence <math>\Delta G</math> always negative (less than zero)</p>	1	<p>Allow multiples. Ignore state symbols. Do not allow equation for wrong compound but mark on provided number of moles increases or stays the same. If no equation or equation that gives a decrease in the number of moles, CE = 0</p>
			1	<p>Allow statement 'increase in number of moles/molecules' If numerical values given, they must match the equation in M1 Ignore the effect of incorrect state symbols on the number of moles of particles unless used correctly</p>
			1	<p>If correct deduction from wrong equation is <math>\Delta S = 0</math> or <math>\Delta S</math> very small must say <math>\Delta H</math> -ve</p>
			1	<p>Allow <math>G</math> instead of <math>\Delta G</math> Can score 3 out of 4 marks if equation wrong but leads to increase or no change in number of moles M4 dependent on M3</p> <p>Note, if equation wrong AND there is an incorrect deduction about the change in number of moles, CE = 0</p>
6	(e)	<p><math>\text{CO}_2/\text{CO}/\text{CH}_4</math> may be produced during <math>\text{H}_2</math> manufacture/building the plant/transport/operating the plant</p>	1	