

Question	Marking Guidance	Mark	Comments
5(a)	loses electrons / donates electrons	1	penalise donates electron pair
5(b)	Zn	1	can only score M2 if M1 correct do not allow e.m.f instead of $E^\circ$
	(most) negative $E^\circ$ / lowest $E^\circ$ / least positive	1	
5(c)	$E^\circ \text{F}_2 / \text{F}^- > E^\circ \text{O}_2 / \text{H}_2\text{O}$	1	or e.m.f is positive or e.m.f = 1.64 V
	Fluorine reacts to form oxygen (can score from equation in M3 even if equation unbalanced provided no contradiction) or fluorine oxidises water or fluorine is a more powerful oxidising agent than oxygen	1	
	$2\text{F}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{F}^- + 4\text{H}^+ + \text{O}_2$	1	
			allow 4HF in equation balanced equation scores M2 and M3

5(d)(i)	<p>order correct Zn Zn<sup>2+</sup> Ag<sub>2</sub>O Ag or reverse of this order</p> <p>all phase boundaries correct</p> <p>e.g. Zn Zn<sup>2+</sup>  Ag<sub>2</sub>O Ag or Ag Ag<sub>2</sub>O  Zn<sup>2+</sup> Zn scores 2</p>	<p>1</p> <p>1</p>	<p>ignore ss , H<sup>+</sup> and H<sub>2</sub>O, no. of moles</p> <p>allow Zn Zn<sup>2+</sup>  Ag<sub>2</sub>O,Ag or Zn Zn<sup>2+</sup>  Ag<sub>2</sub>O H<sup>+</sup> Ag for M1 &amp; M2</p> <p>M2 cannot be gained unless M1 scored</p> <p>allow H<sup>+</sup> either side of Ag<sub>2</sub>O with comma or   for M2 penalise</p> <ul style="list-style-type: none"> <li>wrong phase boundary (allow dashed lines for salt bridge)</li> <li>Pt</li> <li>use of + (from half equation)</li> <li>water/H<sup>+</sup> outside Ag in Ag electrode</li> </ul>
5(d)(ii)	1.1 (V)	1	<p>Allow no units, penalise wrong units</p> <p>allow correct answer even if no answer to (d)(i) or answer to (d)(i) incorrect</p> <p>allow –1.1 if silver electrode on Left in (d)(i) even if the species are in the wrong order.</p>
5(d)(iii)	<u>Reaction(s)</u> not reversible or H <sub>2</sub> O electrolyses	1	<p>do not allow hard to reverse</p> <p>mention of primary cell is not enough to show that reaction(s) are irreversible</p>
5(e)(i)	–0.46 (V)	1	Allow no units, penalise wrong units

5(e)(ii)	$2\text{PbSO}_4 + 2\text{H}_2\text{O} \rightarrow \text{Pb} + \text{PbO}_2 + 2\text{HSO}_4^- + 2\text{H}^+$ lead species correct on correct sides of equation	1	allow ions / species must be fully cancelled out or combined allow 1/2 for balanced reverse equation
	equation balanced and includes $\text{H}_2\text{O}$ , $\text{HSO}_4^-$ and $\text{H}^+$ (or $\text{H}_2\text{SO}_4$ )	1	
5(f)(i)	reagents / $\text{PbO}_2$ / $\text{H}_2\text{SO}_4$ / acid / ions used up (or concentration decreases)	1	
5(f)(ii)	fuel cell	1	Ignore any other words
5(f)(iii)	reagents / fuel supplied continuously	1	
	concentrations (of reagents) remain constant	1	