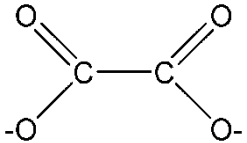
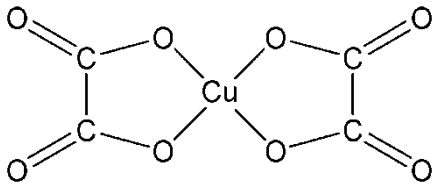


Question	Marking Guidance	Mark	Comments
7(a)(i)	absorbs (certain frequencies of) (white) light / photons	1	not absorbs white / u.v. light
	<u>d</u> electrons excited / promoted	1	or <u>d</u> electrons move between levels / orbitals d electrons can be implied elsewhere in answer
	the colour observed is the light not absorbed / light reflected / light transmitted	1	allow blue light transmitted penalise emission of light in M3
7(a)(ii)	ΔE is the energy gained by the (excited) electrons (of Cu^{2+})	1	allow: <ul style="list-style-type: none"> • energy difference between orbitals / sub-shells • energy of photon / light absorbed • change in energy of the electrons • energy lost by excited electrons • energy of photon / light emitted
	h (Planck's) constant	1	
	ν frequency of light (absorbed by $\text{Cu}^{2+}(\text{aq})$)	1	do not allow wavelength If energy lost / photon lost / light emitted in M1 do not penalised light emitted

7(a)(iii)	$[\text{Cu}(\text{H}_2\text{O})_6]^{2+} + 4\text{Cl}^- \rightarrow [\text{CuCl}_4]^{2-} + 6\text{H}_2\text{O}$ <p>tetrahedral</p> <p>Cl^- / Cl / chlorine too big (to fit more than 4 round Cu)</p>	<p>1</p> <p>1</p> <p>1</p>	<p>note that $[\text{CuCl}_4]^{2-}$ is incorrect</p> <p>penalise charges shown separately on the ligand and overall</p> <p>penalise HCl</p> <p>allow</p> <p>water smaller than Cl^-</p> <p>explanation that change in shape is due to change in <u>co-ordination number</u></p>
7(b)	 <p><u>lone pair(s)</u> on O^- / O</p>	<p>1</p> <p>1</p>	<p>allow:</p> <ul style="list-style-type: none"> ion drawn with any bond angles ion in square brackets with overall / 2- charge shown outside the brackets ion with delocalised $\text{O}=\text{C}-\text{O}$ bonds in carboxylate group(s) <p>allow position of lone pair(s) shown on O in the diagram even if the diagram is incorrect.</p>
7(c)(i)	$[\text{Cu}(\text{H}_2\text{O})_6]^{2+} + 2\text{C}_2\text{O}_4^{2-} \rightarrow [\text{Cu}(\text{C}_2\text{O}_4)_2(\text{H}_2\text{O})_2]^{2-} + 4\text{H}_2\text{O}$ <p>product correct</p> <p>equation balanced</p> <p>6</p> <p>octahedral</p>	<p>1</p> <p>1</p> <p>1</p> <p>1</p>	<p>note can only score M3 and M4 if M1 awarded or if complex in equation has 2 waters and 2 ethanedioates</p> <p>If this condition is satisfied the complex can have the wrong charge(s) to allow access to M3 and M4 but not M1</p>

7(c)(ii)	 <p>90°</p>	1	<p>ignore charges diagram must show both ethanedioates with correct bonding ignore water</p>
		1	<p>allow 180° mark bond angle independently but penalise if angle incorrectly labelled / indicated on diagram</p>