

Balancing Equations Worksheet

1. ... $\text{H}_3\text{PO}_4 + \dots \text{KOH} \rightarrow \dots \text{K}_3\text{PO}_4 + \dots \text{H}_2\text{O}$
2. ... $\text{K} + \dots \text{B}_2\text{O}_3 \rightarrow \dots \text{K}_2\text{O} + \dots \text{B}$
3. ... $\text{HCl} + \dots \text{NaOH} \rightarrow \dots \text{NaCl} + \dots \text{H}_2\text{O}$
4. ... $\text{Na} + \dots \text{NaNO}_3 \rightarrow \dots \text{Na}_2\text{O} + \dots \text{N}_2$
5. ... $\text{C} + \dots \text{S}_8 \rightarrow \dots \text{CS}_2$
6. ... $\text{Na} + \dots \text{O}_2 \rightarrow \dots \text{Na}_2\text{O}$
7. ... $\text{N}_2 + \dots \text{O}_2 \rightarrow \dots \text{N}_2\text{O}_5$
8. ... $\text{H}_3\text{PO}_4 + \dots \text{Mg}(\text{OH})_2 \rightarrow \dots \text{Mg}_3(\text{PO}_4)_2 + \dots \text{H}_2\text{O}$
9. ... $\text{NaOH} + \dots \text{H}_2\text{CO}_3 \rightarrow \dots \text{Na}_2\text{CO}_3 + \dots \text{H}_2\text{O}$
10. ... $\text{KOH} + \dots \text{HBr} \rightarrow \dots \text{KBr} + \dots \text{H}_2\text{O}$
11. ... $\text{Na} + \dots \text{O}_2 \rightarrow \dots \text{Na}_2\text{O}$
12. ... $\text{Al}(\text{OH})_3 + \dots \text{H}_2\text{CO}_3 \rightarrow \dots \text{Al}_2(\text{CO}_3)_3 + \dots \text{H}_2\text{O}$
13. ... $\text{Al} + \dots \text{S}_8 \rightarrow \dots \text{Al}_2\text{S}_3$
14. ... $\text{Cs} + \dots \text{N}_2 \rightarrow \dots \text{Cs}_3\text{N}$
15. ... $\text{Mg} + \dots \text{Cl}_2 \rightarrow \dots \text{MgCl}_2$
16. ... $\text{Rb} + \dots \text{RbNO}_3 \rightarrow \dots \text{Rb}_2\text{O} + \dots \text{N}_2$
17. ... $\text{C}_6\text{H}_6 + \dots \text{O}_2 \rightarrow \dots \text{CO}_2 + \dots \text{H}_2\text{O}$
18. ... $\text{N}_2 + \dots \text{H}_2 \rightarrow \dots \text{NH}_3$
19. ... $\text{C}_{10}\text{H}_{22} + \dots \text{O}_2 \rightarrow \dots \text{CO}_2 + \dots \text{H}_2\text{O}$
20. ... $\text{Al}(\text{OH})_3 + \dots \text{HBr} \rightarrow \dots \text{AlBr}_3 + \dots \text{H}_2\text{O}$
21. ... $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3 + \dots \text{O}_2 \rightarrow \dots \text{CO}_2 + \dots \text{H}_2\text{O}$
22. ... $\text{C}_3\text{H}_8 + \dots \text{O}_2 \rightarrow \dots \text{CO}_2 + \dots \text{H}_2\text{O}$
23. ... $\text{Li} + \dots \text{AlCl}_3 \rightarrow \dots \text{LiCl} + \dots \text{Al}$
24. ... $\text{C}_2\text{H}_6 + \dots \text{O}_2 \rightarrow \dots \text{CO}_2 + \dots \text{H}_2\text{O}$
25. ... $\text{NH}_4\text{OH} + \dots \text{H}_3\text{PO}_4 \rightarrow \dots (\text{NH}_4)_3\text{PO}_4 + \dots \text{H}_2\text{O}$
26. ... $\text{Rb} + \dots \text{P} \rightarrow \dots \text{Rb}_3\text{P}$
27. ... $\text{CH}_4 + \dots \text{O}_2 \rightarrow \dots \text{CO}_2 + \dots \text{H}_2\text{O}$
28. ... $\text{Al}(\text{OH})_3 + \dots \text{H}_2\text{SO}_4 \rightarrow \dots \text{Al}_2(\text{SO}_4)_3 + \dots \text{H}_2\text{O}$
29. ... $\text{Na} + \dots \text{Cl}_2 \rightarrow \dots \text{NaCl}$
30. ... $\text{Rb} + \dots \text{S}_8 \rightarrow \dots \text{Rb}_2\text{S}$
31. ... $\text{H}_3\text{PO}_4 + \dots \text{Ca}(\text{OH})_2 \rightarrow \dots \text{Ca}_3(\text{PO}_4)_2 + \dots \text{H}_2\text{O}$
32. ... $\text{NH}_3 + \dots \text{HCl} \rightarrow \dots \text{NH}_4\text{Cl}$
33. ... $\text{Li} + \dots \text{H}_2\text{O} \rightarrow \dots \text{LiOH} + \dots \text{H}_2$
34. ... $\text{Ca}_3(\text{PO}_4)_2 + \dots \text{SiO}_2 + \dots \text{C} \rightarrow \dots \text{CaSiO}_3 + \dots \text{CO} + \dots \text{P}$
35. ... $\text{NH}_3 + \dots \text{O}_2 \rightarrow \dots \text{N}_2 + \dots \text{H}_2\text{O}$
36. ... $\text{FeS}_2 + \dots \text{O}_2 \rightarrow \dots \text{Fe}_2\text{O}_3 + \dots \text{SO}_2$
37. ... $\text{C} + \dots \text{SO}_2 \rightarrow \dots \text{CS}_2 + \dots \text{CO}$