

 $P_1 = 1500 \, \text{kPa}$   $P_2 = 750 \, \text{kPa}$   $P_1 = 750 \, \text{kPa}$   $P_2 = 750 \, \text{kPa}$   $P_3 = 750 \, \text{kPa}$   $P_4 = 750 \, \text{kPa}$   $P_5 = 750 \, \text{kPa}$   $P_6 = 750 \, \text{kPa}$ 

Assuming ideal gas 
$$Pv = PT$$

$$P_1 v_1 = P_2 v_2$$

$$T_1 = P_2 v_2$$
and  $P_1 v_1^n = P_2 v_2^n$ 

Combining  $T_{2} = T_{1} \left( \frac{P_{2}}{P_{1}} \right)^{\frac{n-1}{n}} = 750 \left( \frac{750}{1500} \right)^{\frac{0.2}{1.2}}$ 

T2 = 668.2 K.

For a polytropic process

$$W_{12} = \frac{P_2 V_2 - P_1 V_1}{N-1} = mR \frac{(T_2 - T_1)}{N-1}$$

$$\omega_{12} = \frac{W_{12}}{m} = P\left(\frac{T_2 - T_1}{n - 1}\right) = 0.2968 (668.2 - 750)$$

$$q_{12} = (u_2 - u_1) - \omega_{12} = c_v (T_2 - T_1) - \omega_{12}$$
  
= 0.7448 (668.2-750) + 121.4  
 $q_{12} = 60.5$  ET/by

Mars flow 
$$\mathring{m} = \frac{AV}{U}$$
 where  $U = \frac{RT}{P}$ 

The solution  $\mathring{m} = \frac{AV}{U}$  where  $U = \frac{RT}{P}$ 

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